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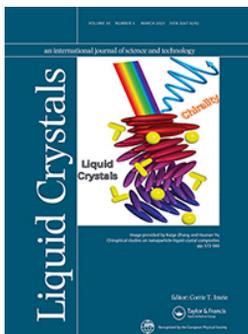
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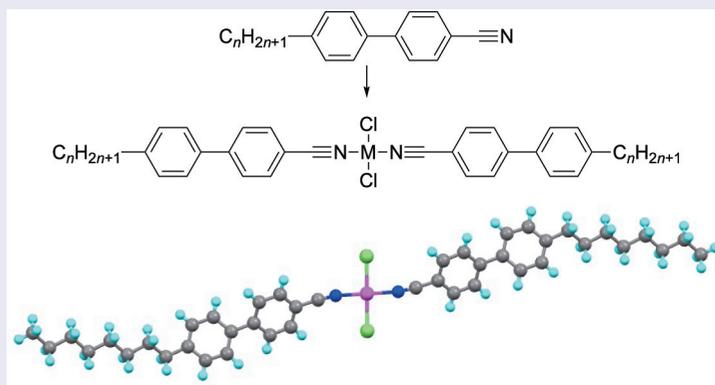
Cyanobiphenyls and metallomesogens – where it started and where it went

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ABSTRACT

The author's work in metallomesogens stemmed from initial preparations of κ^1 -complexes of cyanobiphenyls of the Group 10 metals, palladium and platinum. In this personalised account, these initial studies are described, put in the context of where the field was at that time and contextualised for the way the author's own contributions to the field then developed.



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isonitriles

Introduction

Our story begins in Sheffield in 1984 with a conversation between the author (DWB) and the late Peter Maitlis, who had just appointed him to a temporary lectureship in Chemistry. Having come from a PhD in homogeneous photocatalysis using soluble metal-phosphine complexes with David Cole-Hamilton, DWB had been keen to pursue his interests in catalysis further and had indeed written a research proposal based on heterobinuclear metal complexes as part of his application. But Peter had seen the way the chemistry world was going, had understood that organometallics was at that time a crowded field and had therefore suggested that a visible career might be better followed in newer and less-developed areas of the subject. The conversation, held over several meetings, had ranged far and wide until Peter drew on the board a cyanobiphenyl and mentioned liquid crystals. This was a new and foreign field to both, although DWB does remember attending a lecture on the subject by George Gray while a postgraduate student at Liverpool and even being exposed to some of it even earlier on

a visit to Hull as a prospective undergraduate, although ironically little of the detail had stuck from either exposure.

The cyanobiphenyl idea struck some sort of chord and while, as recalled earlier, as an organometallic chemist of many years standing, Peter saw η^6 -coordination as the way forward. However, as someone versed in coordination chemistry, DWB saw coordination through nitrogen and analogues of the well-known bis(benzonitrile)dichlorometal(II) complexes of palladium and platinum as the way to go [1].

An exploration of the literature began and DWB recalls clearly that the first paper he read was from Kazuo Ohta [2] on some copper(II) β -diketonates and then, papers from the other names that became familiar from what was at that time a relatively limited literature – Anne-Marie Giroud with her Group 10 dithiolene complexes [3], Mauro Ghedini with *ortho*-palladated azobenzenes [4], Yury Galyametdinov with salicylaldimato complexes [5], Jacques Simon with phthalocyanines [6] and, of course, Daniel Vorländer with mercury complexes [7,8]. All were using bi- or tetra-dentate

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lasted some 10 years and generated 22 papers. Thus, when Elena had prepared *trans*-[PdCl₂(5CB)₂] for the first time, we gave it to David to characterise and a short while later he came back with the news that we had our first liquid crystal!

The early work

It was undergraduate student Bill Snooks who, in undertaking a 10-week, final-year project, made the first studies in the group. Motivated by the work with β -diketones mentioned above, he reacted diketene with long-chain alcohols to produce alkylacetoacetates, which were then complexed with copper(II) (Figure 2). None of the complexes was mesomorphic and the observed melting points showed a steady decrease with alkyl chain length [10]. There was simply insufficient anisotropy in these complexes to stabilise a mesophase and it seemed as though trying to elaborate these further would lead to what would become a crowded field.

However, the summer that Elena arrived, progress became rapid. She prepared several examples of first palladium(II) complexes of cyanobiphenyls and then extended this to analogous complexes of platinum(II) – we had our foot in the door [11]. As will be outlined below, Elena made other contributions, too, but having put some momentum behind the cyanobiphenyls, some of this work was then taken up by new PhD student Peter Styring (now Professor of Chemical Engineering in Sheffield) alongside undergraduate project students Mick Wragg, Lesley Green and Andy Maggs. Between them they made palladium and platinum complexes of a very wide range of organonitrile liquid crystals, and while several of them were published in *Chemical Communications* in 1986 [11] and then in this journal in 1987 [12], data for the vast majority never saw the light of day. Although a good deal of the basic data for these complexes still exists in old project reports and lab books, the Supplementary Information would be of a poor standard in trying to report them as we do now (e.g. note that in the *Chemical Communications* paper [11], it was sufficient to say ‘All new materials have been characterised by i.r. and ¹H n.m.r. spectroscopy and by C, H, N, and Cl microanalysis for which satisfactory results were obtained’, with none of those data provided for referees! [11]). This is not meant to imply that the data are poor, rather that we do not have all of them readily to hand and so rather we give a flavour of what was done with the structure of some of the nitriles used given in Figure 3.

Single crystal structures were published for one palladium and one platinum complex (Figure 4), confirming the *trans* nature of the complexes, with the

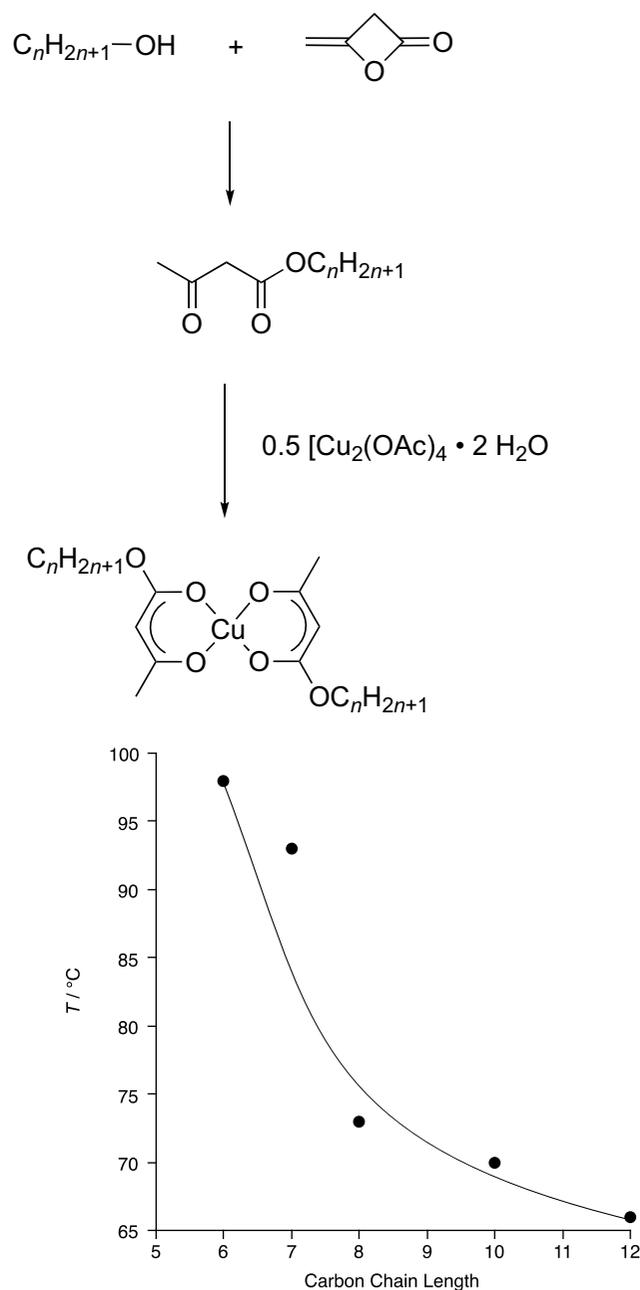


Figure 2. Preparation and plot of the melting points of a series of bis(alkylacetoacetato)copper(II) complexes.

palladium complex crystallising in the triclinic *P*-1 space group, while the Pt complex crystallised in the monoclinic *P*2₁/*a* [12].

As indicated, while reliable, tabulated transition temperatures do not exist for most of these complexes, Styring's thesis [13] contains phase diagrams for [PdCl₂(*n*-CB)₂] and [PdCl₂(*n*-OCB)₂], which are now reproduced in Figure 5.

Thus, the palladium CB complexes all show monotropic nematic phases, with clearing points in the approximate range 70–95°C, while the complexes of the OCBs all show enantiotropic phases, which are

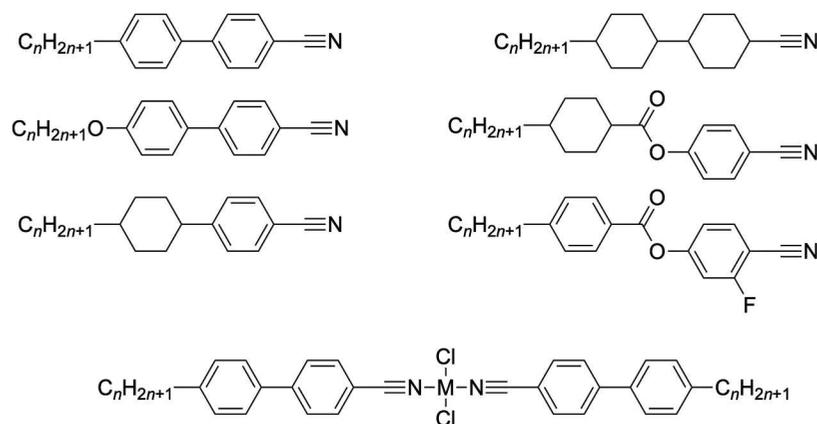


Figure 3. Structures of the come of the organonitriles used to prepared metal complexes and the structure of the complexes (M = Pd or Pt).

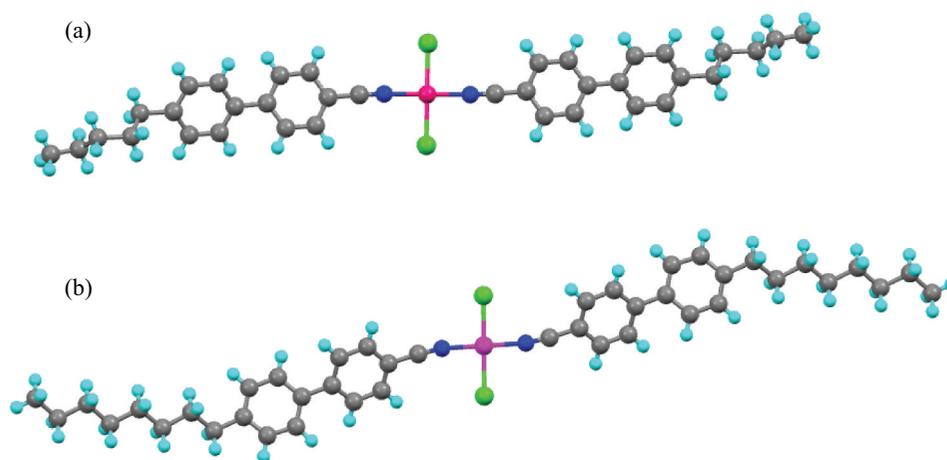


Figure 4. (Colour online) Molecular structure of (a) $[\text{PdCl}_2(5\text{CB})_2]$ and (b) $[\text{PtCl}_2(8\text{CB})_2]$.

nematic for $4 \leq n \leq 8$, while for $n = 9$ a SmC and a SmA phase are seen. As ever in liquid crystals, enantiotropic and monotropic behaviour are a function of *both* the melting point *and* the clearing point and, in these complexes, a comparison of each is really rather instructive. Thus, [Figure 6\(a\)](#) shows the melting points of the two series plotted on the same axes *as a function of the total chain length*, while [Figure 6\(b\)](#) likewise shows the clearing points plotted using Kelvin as the temperature axis (see below).

The melting points are not so dissimilar as a function of total chain length, although for the shorter-chain ligands the complexes of the OCBs have a generally higher melting point, which may be an effect of the dipole moment introduced on addition of the alkoxy oxygen to the ligand. However, for the ligands 5OCB to 7OCB, the melting points of the complexes fall below those of the complexes with ligands 6CB to 8CB, which is not expected and which may indicate that they crystallise in different space groups.

Then in considering the clearing points ([Figure 6\(b\)](#)), it can be seen that while not unexpectedly the temperatures found for the complexes are higher than those of the ligands, nonetheless they track one another rather well and the observed odd-even effect also tracks the total number of atoms in the terminal chain. Interestingly, if the ratio is taken between the clearing point of the complex and the clearing point of the related ligand in Kelvin, it comes out in effect as 1.2 for each pair considered. Thus, the higher clearing points for the complexes compared to the ligands would appear to have a common origin related to complexation, with those of the complexes of n -OCB being highest on account of the dipole induced from introduction of the ether oxygen.

There are many fewer data available for platinum complexes that can be relied upon and so for the purposes of this article, the two published examples will be considered, namely $[\text{PtCl}_2(5\text{CB})_2]$ and $[\text{PtCl}_2(8\text{CB})_2]$, whose melting points (189°C and 166°C , respectively)

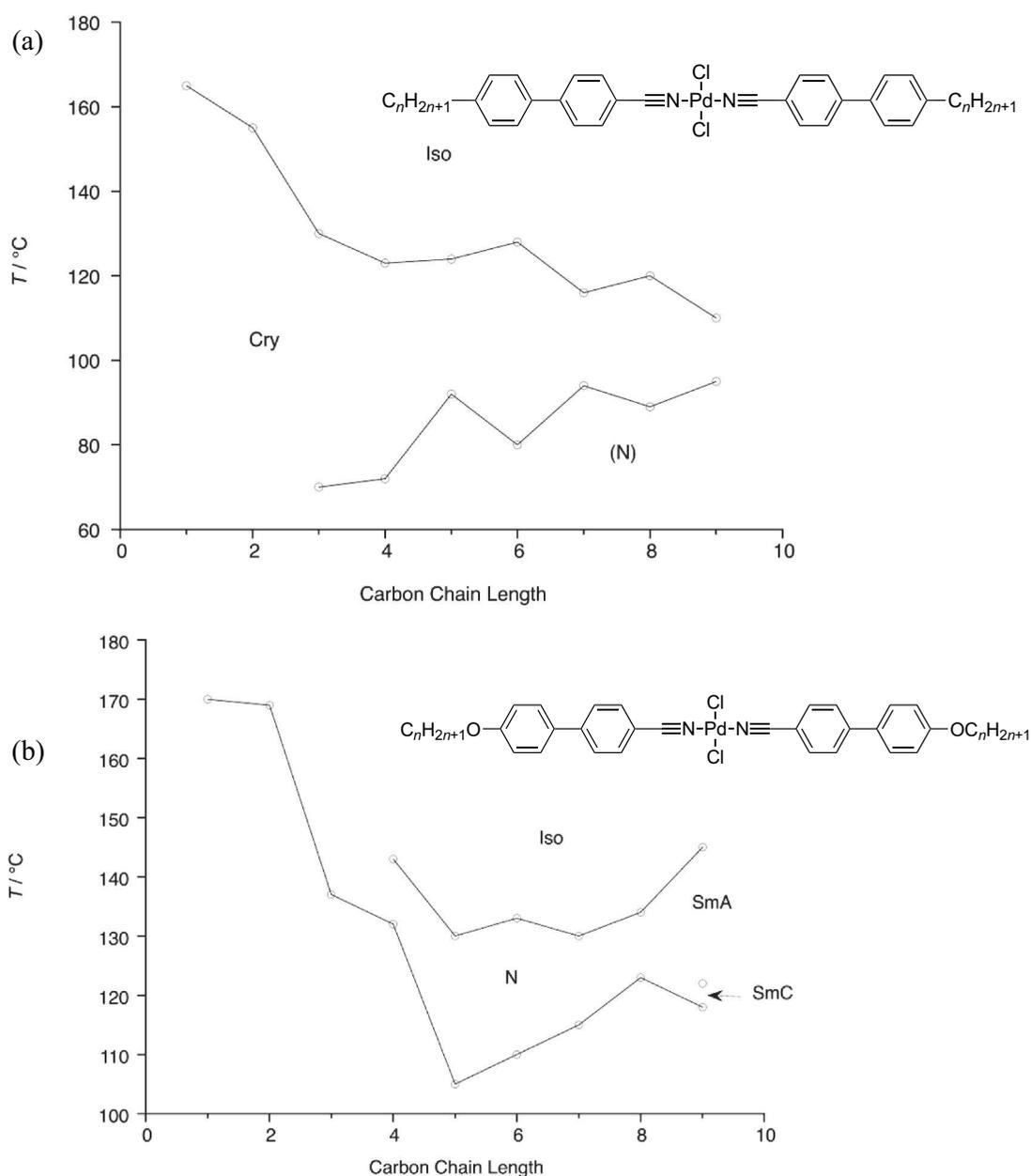


Figure 5. Phase diagram for (a) 4-alkyl-4-cyanobiphenyl complexes of (b) 4-alkoxy-4-cyanobiphenyl complexes of palladium. Temperatures were estimated visually from a plot in reference [12].

and clearing points ($T_{N-Iso} = 204$ °C and 176 °C, respectively) are very significantly higher than those of the analogous palladium complexes ($\Delta T_{N-Iso} = 112$ °C and 87 °C, respectively) [11]. Further, the ratio (in K) of the clearing point of the complexes to the ligands is between 1.5 for $[PtCl_2(5CB)_2]$ and 1.4 for $[PtCl_2(8CB)_2]$. How might the differences between Pd and Pt be understood?

There is a simple, linear correlation between the length of a mesogen and its clearing point as set out in a molecular field theory of Kloczkowski and Luckhurst [14] and, on this basis, it is quite reasonable to expect that the clearing points of the palladium and platinum complexes ought to be rather close (there is a difference

in molecular weight of *ca* 89 g mol^{-1} ($\approx 12\%$) on account of the larger value of *Z* for Pt, but this should not have a significant effect).

In the two crystal structures, the metal atoms sit on a crystallographic inversion centre and so both the N-M-N and Cl-M-Cl angles are formally 180°. However, as monodentate ligands and in the case of the nitriles quite extended ligands, then there is the possibility for bending about the M – N bond in particular. Further, the result of the lanthanide contraction means that while the nuclear charge at Pt is some 32 greater than at Pd, the radii of the two metal(II) centres will be all but identical, giving a much greater charge at

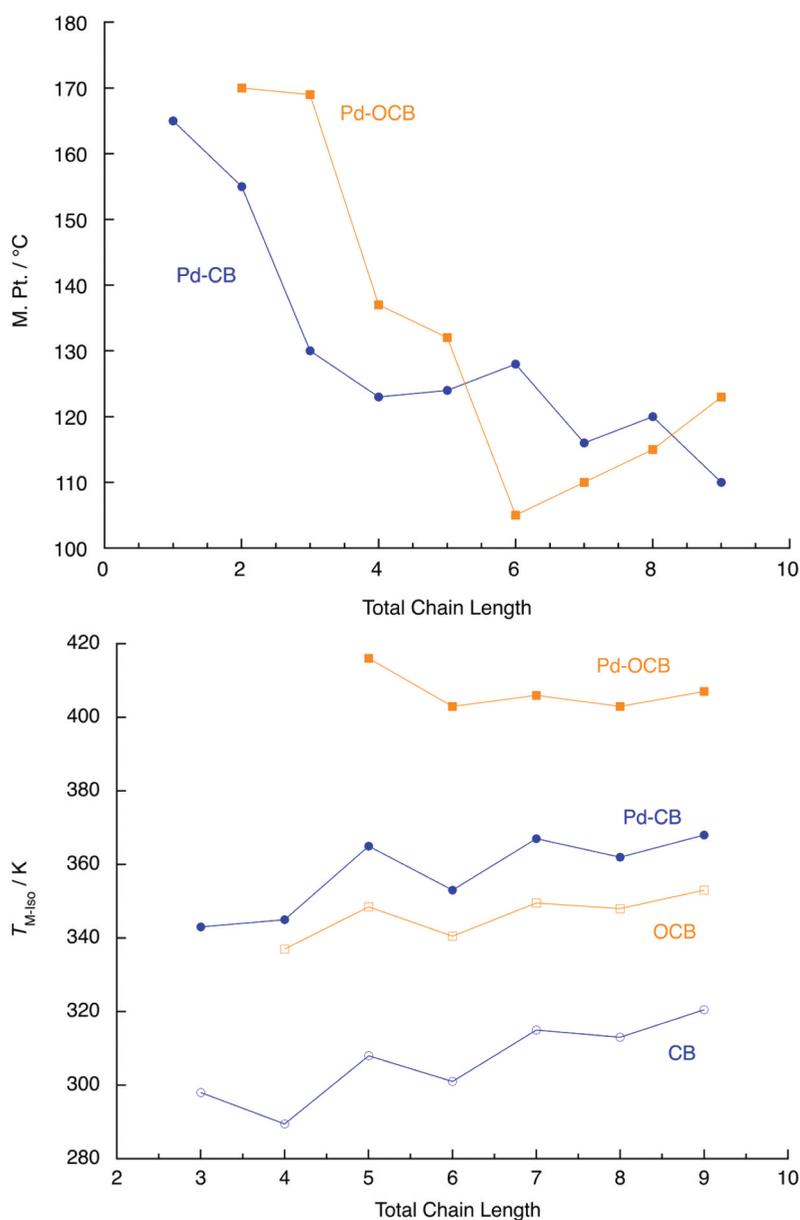


Figure 6. (Colour online) (a) Melting points (in °C) and (b) clearing points (in K) for the complexes $[\text{PdCl}_2(n\text{-CB})_2]$ and $[\text{PdCl}_2(n\text{-OCB})_2]$ plotted as function of the total terminal chain length. All of the clearing points are for N-Iso except for $[\text{PdCl}_2(9\text{-OCB})_2]$, where it is SmA-Iso. Note that the two y-axis scales are different.

Pt and resulting in tighter binding of ligands. On this basis, it is entirely reasonable to assume that there is greater motional flexibility in the palladium complexes compared to those of platinum and, this being the case, then there is a potential explanation for the lower mesophase stability of the palladium complexes.

In fact, there is another interesting issue when comparing the clearing point of 5CB and its complexes, for the same molecular field theory would predict that the ratio of their clearing points (in K) should be close to 1:2 as the complex is almost exactly twice the length of the ligand, when in fact it is highest in platinum complexes where it is 1.5. In order to probe this observation,

^2H NMR spectra were recorded of 5CB and its platinum complex dissolved in the liquid crystal I35 (Figure 7) using a d^2 -labelled cyanobiphenyl.¹ The work was done with Geoffrey Luckhurst and Shimei Fan [15]. The NMR experiment allowed values of the order parameter P_2 to be obtained from the observed quadrupolar splitting and, as might have been expected, at the clearing point of the solvent, P_2 was greater for the complex (0.48) compared to the ligand (0.33), although the two converged to a reasonable degree at lower reduced temperatures. An explanation was proposed in terms of what are now the fashionable cybotactic clusters, although the fact that the complexes have two flexible

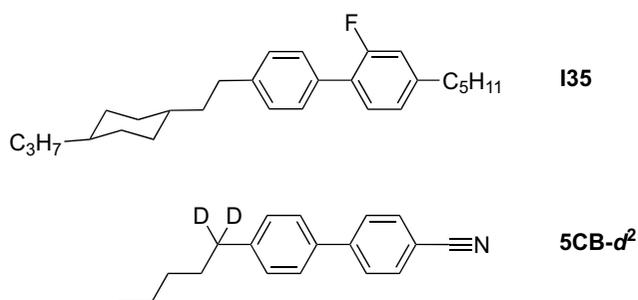


Figure 7. The nematic solvent known as I35 and the d^2 -labelled 5CB.

chains (as opposed to one for the ligand) and the model assumes rigid molecules is likely a pertinent factor.

In addition to these NMR studies, the enhanced anisotropy of the complexes was studied by optical means and solutions of $[\text{PdCl}_2(5\text{CB})_2]$ in 5CB showed slightly enhanced clearing points and birefringence (Figure 8) [12]. This work was followed up by a more complete series of measurements for the complexes $[\text{PdCl}_2(n\text{-OCB})_2]$ ($n = 5, 6, 7$) and, for example, at a reduced temperature of 0.985, Δn for

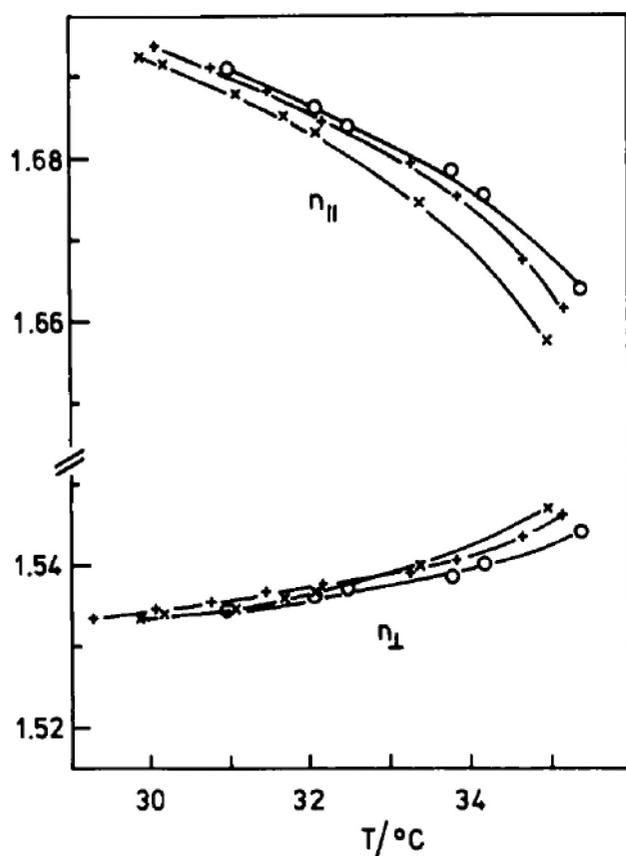


Figure 8. Plot of refractive indices versus temperature for a solution of $[\text{PdCl}_2(5\text{CB})_2]$ in 5CB at the following compositions: \bullet 1.27 wt per cent Pd complex; $+$ 0.77 wt per cent Pd complex; \times pure 5CB (reproduced from reference [12]).

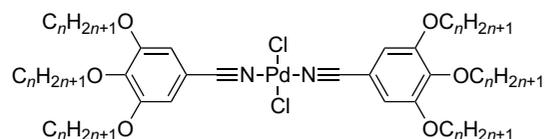


Figure 9. 3,4,5-trialkoxybenzonitrile complexes of palladium(II).

$[\text{PdCl}_2(6\text{-OCB})_2]$ was 0.1636 compared to 0.1433 for the free ligand [16].

After this, there are few reports regarding cyanobiphenyl complexes, one being the report by Hanabusa *et al.* of the use of PtCl_2 to cross-link side-chain cyanobiphenyl polymers thus creating elastomeric materials [17]. However, a particularly elegant extension was the report by Lee *et al.* of the palladium complexes of 3,4,5-tridodecyloxy- and 3,4,5-trioctadecyloxy-benzonitrile (Figure 9), which showed a columnar mesophase between 73.3 and 92.1°C and between 58.4 and 80.4°C, respectively [18]. Elegant simplicity!

Beyond cyanobiphenyls to isonitriles

Takahashi and Espinet later mixed things up by preparing a many and interesting series of complexes of palladium(II), platinum(II) [19], silver(I), gold(I) [20,21] and iron(0) [22] based upon the isomeric isonitriles, representative examples of which are shown in Figure 10. In fact, one might also regard an acetylide anion as isoelectronic with a nitrile and of course these have been used on their own and in combination with organoisonitriles to generate many more metallomesogens, with an example shown also in Figure 10 [23,24].

Further, Coco *et al.* prepared isonitrile analogues of the trialkoxybenzonitriles shown in Figure 9, which they complexed to both palladium and platinum (Figure 11). Interestingly, where the halide ligand was iodide, then the complexes of both metals showed a *trans* geometry in the solid state and (insofar as they could tell) the mesophase. However, things were not so clear cut for $X = \text{Cl}$ and Br , so that for $M = \text{Pt}$, both showed exclusively a *cis*-geometry, whereas for $M = \text{Pd}$, both isomers were found at room temperature (95% *trans* where $X = \text{Br}$ and 65% *cis* where $X = \text{Cl}$). Irrespective of the geometry, the complexes were mesomorphic forming columnar phases for $4 \leq n \leq 10$ for all *cis* complexes and for $n = 8, 10$ for the *trans* complexes [25].

What happened next?

During that eventful stay that Lalinde made to Sheffield in the summer of 1985, I recall having been away for 2 weeks on holiday and coming back to find a new ligand

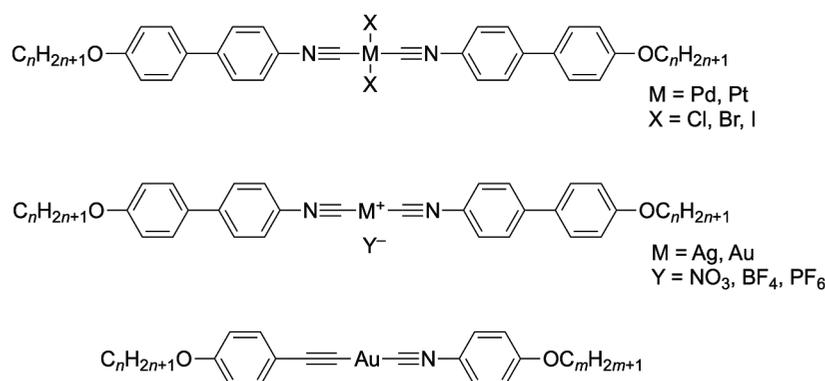


Figure 10. Examples of metallomesogens containing organoisonitrile ligands.

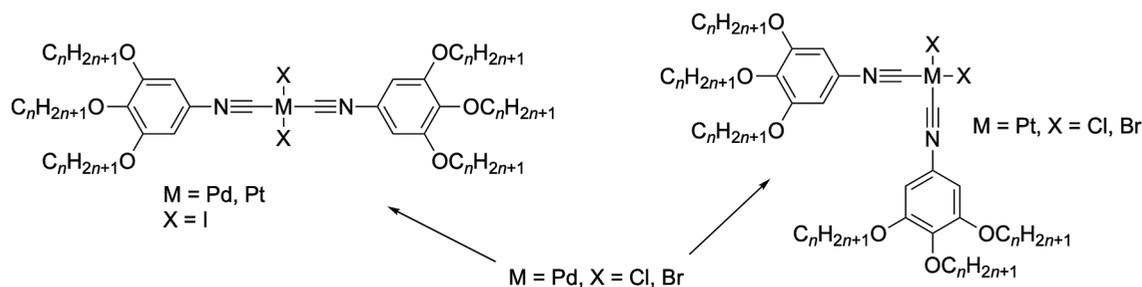


Figure 11. Trialkoxyphenyl isonitrile complexes of palladium(II) and platinum(II).

system, which had been proposed to her by Maitlis. These were the alkoxy stilbazoles (Figure 12), which were made at that time by the reaction of an alkoxybenzaldehyde with 4-picoline in acetic anhydride giving a recovered yield of <5% [26]. Despite the awful yields (not that long after we began to prepare them in much greater quantity and in much higher yields using a Heck coupling between 4-alkoxyiodobenzene and 4-vinylpyridine) these were attractive. For, while organonitriles are in general rather poor ligands, pyridines are rather better and so there was

room to hope for more stable/inert complexes that might result. The simple 4-alkoxy stilbazoles never gave anything terribly interesting with palladium and platinum as the complexes were very high melting/clearing and rather difficult to study, but complexes of related dialkoxy stilbazoles did open up quite a field of work [27,28]. However, while nitrile complexes of rhodium(I) had been really thermally labile, the stilbazoles formed a whole series of complexes (Figure 10) with both rhodium and iridium of the formula *cis*-[MCl(CO)₂(stilbazole)]

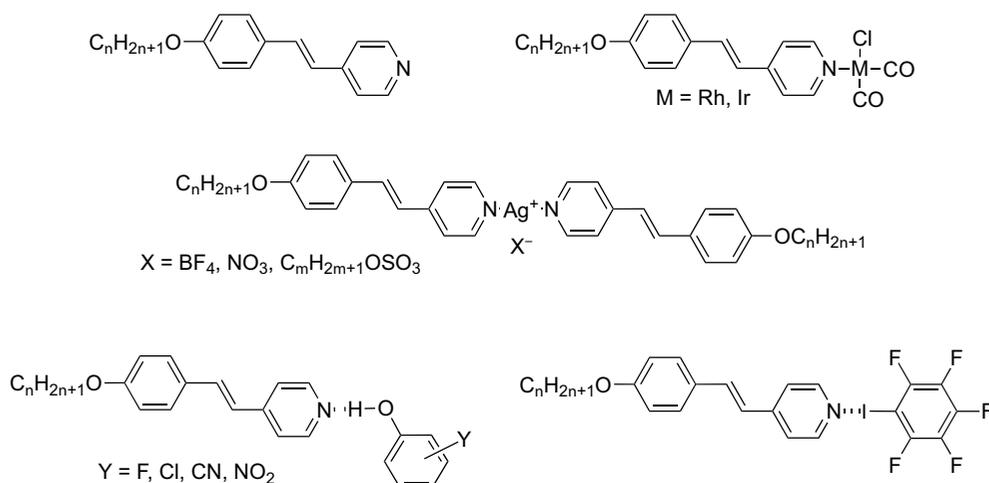


Figure 12. 4-alkoxy stilbazoles. Complexes with Rh, Ir and Ag and both hydrogen- and halogen-bonded mesogens.

[29], and developing these ideas led us to work in both Langmuir-Blodgett films (with the late Tim Richardson) [30] and second-order non-linear optics [31].

Stilbazoles [32] also opened up our work on hydrogen-bonded liquid crystals [33] following inspiration from Takashi Kato and Jean Fréchet [34] and, in turn, this led us to halogen-bonded liquid crystals (Figure 10) where we reported the first examples [35] – an area we have reviewed recently [36].

However, stilbazoles also got us into silver chemistry [37] and while we published our first paper on silver in 1986 [38], the most recent one was published (albeit using related dialkoxyphenylpyridine ligands) in 2022 [39]. Through silver (Figure 10) and with major contributions from Peter Styring, Sarah Hudson and Bertrand Donnio, we discovered ionic nematic phases [37,40] (now more common through some lovely work by former postdoc Viorel Cîrcu and co-workers [41]), in 1986/1987 we doubled the number of materials showing thermotropic cubic phases (from around 6 to 12) [42], we studied the cubic-columnar transition in detail [43] and obtained freeze-fracture electron microscopy images [44], we developed a material with a direct N-Cub transition which, working with Anne-Marie Levelut (sadly no longer with us), Michéle Veber and Bertrand Donnio gave us a monodomain X-ray determination of a cubic phase ($Ia\bar{3}d$) [45] and, having been able to stabilise the S_4 phase that often occurred alongside the cubic phase, once more with Anne-Marie and Bertrand we were able to use X-ray to determine that it has tetragonal symmetry [46]. Furthermore, using a wider range of stilbazole complexes, we were able to propose some very strong structure-activity correlations on the formation of cubic phases in calamitic mesogens [28,43,44,47,48], which then informed studies with Antonina Smirnova on solvent-modified mesomorphism [49].

Epilogue

To pull this article together has been a great deal of fun, not least because it made me dig out so many old reports and in so doing recall so many previous students and co-workers. I have named many, but by no means all, of the key players and I think those not named are otherwise recognised as co-authors in the references. But I did want to reflect once more on how it all started. From sage advice to change direction and find something ‘new and different’ to do, from eschewing the organometallic route and following an instinct for κ^1 rather than η^6 coordination, from the support of being given the chance to work over two summers with a hugely talented visitor (Elena did come back in summer 1986, but that is another

story) who not only got the whole thing moving but also opened up the stilbazoles, to the fortune in having a colleague in the department to show that the ideas actually had something to them. Cyanobiphenyls are 50-years old, and this story has almost a 40-year history attached to it. George Gray was also a very early supporter of the work we started to do and that feels apposite given that, after all, it was he and his research group that showed what the cyanobiphenyls could do. It’s a privilege to have been part of that story.

Note

1. DWB went to Southampton to prepare the complex with Geoffrey Luckhurst, staying on a day after giving a seminar there to the inorganic chemists. Whereas previously the platinum complexes had been made under reflux in toluene leading to only quite moderate yields [12], the route had recently been modified to carry it out in a small excess of molten ligand above 100°C to which $PtCl_2$ was added [50]. These regioselectively deuteriated CBs take a bit of making and are, therefore, valuable materials. DWB remembers well the look of discomfit he encountered on his collaborator’s face when this beautiful, pure and colourless material has heated to the fluid state only to have ‘dirty’ $PtCl_2$ powder dumped into it. Nonetheless, coming back after coffee, a slightly dirty yellow product was recovered that cleaned up just fine.

Acknowledgments

It has been a real pleasure to be able to recall all of those with whom I was able to work in this enterprise over many years and who are named either explicitly in the text or through the references cited. We were funded generously by the EPSRC (and its predecessor the SERC), the EU, the Royal Society, the British Council and by the Universities of Sheffield, Exeter and now York, while palladium and platinum came from the wonderful loans scheme run by Johnson Matthey. All are thanked warmly.

Disclosure statement

No potential conflict of interest was reported by the author.

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