

This is a repository copy of *Encapsulation of Crabtree's Catalyst in Sulfonated MIL-101(Cr): Enhancement of Stability and Selectivity between Competing Reaction Pathways by the MOF Chemical Microenvironment*.

White Rose Research Online URL for this paper:

<https://eprints.whiterose.ac.uk/155089/>

Version: Published Version

Article:

Grigoropoulos, Alexios, McKay, Alasdair I., Katsoulidis, Alexandros P. et al. (6 more authors) (2018) Encapsulation of Crabtree's Catalyst in Sulfonated MIL-101(Cr): Enhancement of Stability and Selectivity between Competing Reaction Pathways by the MOF Chemical Microenvironment. *Angewandte Chemie - International Edition*. pp. 4532-4537. ISSN 1433-7851

<https://doi.org/10.1002/anie.201710091>

Reuse

This article is distributed under the terms of the Creative Commons Attribution (CC BY) licence. This licence allows you to distribute, remix, tweak, and build upon the work, even commercially, as long as you credit the authors for the original work. More information and the full terms of the licence here:

<https://creativecommons.org/licenses/>

Takedown

If you consider content in White Rose Research Online to be in breach of UK law, please notify us by emailing eprints@whiterose.ac.uk including the URL of the record and the reason for the withdrawal request.

Metal–Organic Frameworks

International Edition: DOI: 10.1002/anie.201710091
German Edition: DOI: 10.1002/ange.201710091**Encapsulation of Crabtree's Catalyst in Sulfonated MIL-101(Cr): Enhancement of Stability and Selectivity between Competing Reaction Pathways by the MOF Chemical Microenvironment**

Alexios Grigoropoulos, Alasdair I. McKay, Alexandros P. Katsoulidis, Robert P. Davies, Anthony Haynes, Lee Brammer, Jianliang Xiao, Andrew S. Weller,* and Matthew J. Rosseinsky*

In memory of Gérard Férey

Abstract: Crabtree's catalyst was encapsulated inside the pores of the sulfonated MIL-101(Cr) metal–organic framework (MOF) by cation exchange. This hybrid catalyst is active for the heterogeneous hydrogenation of non-functionalized alkenes either in solution or in the gas phase. Moreover, encapsulation inside a well-defined hydrophilic microenvironment enhances catalyst stability and selectivity to hydrogenation over isomerization for substrates bearing ligating functionalities. Accordingly, the encapsulated catalyst significantly outperforms its homogeneous counterpart in the hydrogenation of olefinic alcohols in terms of overall conversion and selectivity, with the chemical microenvironment of the MOF host favouring one out of two competing reaction pathways.

Metal–organic frameworks (MOFs)^[1] are crystalline and permanently porous materials that have emerged as promising hosts for the immobilization of organometallic catalysts,^[2] since they allow control of the steric and chemical microenvironment around the encapsulated catalytically active species. This in turn could promote catalytic activity and selectivity through extended coordination sphere interactions. These concepts lie behind the exceptional reactivity and

selectivity of metalloenzymes,^[3] however their transfer to the design and synthesis of artificial catalysts is challenging.^[4] Several examples of MOF-supported catalysts showing exceptional overall catalytic activity have been reported.^[5,6] Enhancement of selectivity between products of a single reaction pathway by control of the steric^[7] or the chemical^[8] microenvironment has also been demonstrated.

Crabtree's catalyst is one of the best commercially available homogeneous catalysts for hydrogenation of alkenes.^[9] However, it is deactivated in solution under hydrogenation conditions, forming catalytically inactive poly-metallic hydride clusters.^[10] This self-association reaction can be attenuated via modification of the coordination sphere of Ir^[11] or employment of larger weakly coordinating anions.^[12] Substrates bearing ligating functionalities such as olefinic alcohols show a more complicated behavior with Crabtree's catalyst since isomerization^[13] can also take place in parallel with hydrogenation.^[14]

Here we use the Na⁺ salt of sulfonated MIL-101(Cr) MOF (**1-SO₃Na**) to provide the anionic framework host for encapsulation of the cationic component of Crabtree's catalyst [Ir(cod)(PCy₃)(py)][PF₆] (**2-PF₆**) by cation exchange,^[15] forming **2@1-SO₃Na** (Scheme 1). Encapsulation of cation **2** inside a well-defined, anionic and hydrophilic microenvironment forms an efficient heterogeneous catalyst for the hydrogenation of non-functionalized alkenes in solution, enables hydrogenation in the gas phase, and most importantly enhances the catalyst's activity and selectivity for the hydrogenation of olefinic alcohols by suppressing the

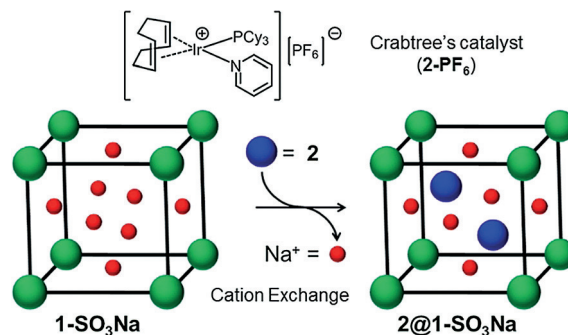
[*] Dr. A. Grigoropoulos, Dr. A. P. Katsoulidis, Prof. J. Xiao, Prof. M. J. Rosseinsky
Department of Chemistry, University of Liverpool
Liverpool L69 7ZD (UK)
E-mail: M.J.Rosseinsky@liverpool.ac.uk
Dr. A. I. McKay, Prof. A. S. Weller
Department of Chemistry, University of Oxford
Chemistry Research Laboratories
Oxford OX1 3TA (UK)
E-mail: andrew.weller@chem.ox.ac.uk

Dr. R. P. Davies
Department of Chemistry, Imperial College London
South Kensington, London SW7 2AZ (UK)

Dr. A. Haynes, Prof. L. Brammer
Department of Chemistry, University of Sheffield
Brook Hill, Sheffield S3 7HF (UK)

Supporting information and the ORCID identification number(s) for the author(s) of this article can be found under:
<https://doi.org/10.1002/anie.201710091>.

© 2018 The Authors. Published by Wiley-VCH Verlag GmbH & Co. KGaA. This is an open access article under the terms of the Creative Commons Attribution License, which permits use, distribution and reproduction in any medium, provided the original work is properly cited.



Scheme 1. Encapsulation of the cationic component of Crabtree's catalyst (**2**, blue spheres) in sulfonated MIL-101(Cr) (**1-SO₃Na**, cube) by exchange of the charge-balancing Na⁺ cations (red spheres).

competing isomerization reaction. The MOF chemical micro-environment directs substrates along one of two distinct reaction pathways.

The sulfonated analogue of MIL-101(Cr) (**1-SO₃H**)^[16] is a robust, readily synthesized anionic MOF. It is isostructural with pristine MIL-101(Cr)^[17] with two charge-balancing cations per formula unit, [H_xNa_{2-x}][Cr₃(μ₃-O)(BDC-SO₃)₃] ($x = 1.8 \pm 0.1$, Figure S1, H₂BDC-SO₃Na = 2-sulfoterephthalic acid sodium salt). Each cubic unit cell ($a = 87.63(3)$ Å) contains 8 bigger and 16 smaller mesopores, large enough to accommodate **2** (Figures S2 and S3). The cations within **1-SO₃H** can be partially exchanged with Ag⁺^[18] or [Rh(cod)-(dppe)]⁺ [dppe = 1,2-bis(diphenylphosphino)ethane].^[19] To increase the number of exchangeable Na⁺ cations, **1-SO₃H** was treated with AcONa/AcOH buffer solution (pH 4.7), forming [H_yNa_{2-y}][Cr₃(μ₃-O)(BDC-SO₃)₃] (**1-SO₃Na**, $y = 0.2 \pm 0.1$, Table S1).

Compound **1-SO₃Na** remains crystalline and mesoporous (Figures 1 a, b) with only a small change in the cubic unit cell parameter ($a = 87.99(4)$ Å) and a slight increase in the measured porosity (BET surface area = 2005 m² g⁻¹, $V_p = 0.91$ cm³ g⁻¹) and the pore size distribution, compared to **1-SO₃H** (Figure S9).

After establishing an appropriate cation exchange protocol using [Cp*₂Co]⁺ as a cationic probe (Table S2 and Figures S6, S9–S12), as we have shown previously,^[15b] **2-PF₆** was used as a cationic guest precursor. Since water can poison the catalytically active species,^[12] cation exchange was carried out using desolvated **1-SO₃Na** as the anionic host in dry and degassed acetone, producing **2@1-SO₃Na**. Crystallinity and particle morphology were retained after cation exchange with only a minor change in the cubic unit cell parameter ($a = 87.74(3)$ Å, Figure 1 a, see Le Bail fit in Figure S7 and SEM images in Figures S11 and S12), whereas BET surface area (1570 m² g⁻¹) and pore volume (0.70 cm³ g⁻¹) were reduced, compared to **1-SO₃Na** (Figure 1 b).

ICP-OES after digestion of **2@1-SO₃Na** gave an Ir content of 2.28 wt %, indicating that 7 % of the Na⁺ cations have been exchanged with **2** (Table S3), which is close to the upper limit of about 9 % calculated by accounting for the guest-accessible

space of the host MOF and the size of the cationic guest (Figures S1–S3). ICP-OES also showed an equimolar Ir/P ratio, and only one broad peak was observed ($\delta_p = 15.65$, fwhm ≈ 15 ppm) in the ³¹P{¹H} MAS NMR spectrum of **2@1-SO₃Na**, assigned to the PCy₃ ligand (Figure 1 c). Signals arising from the [PF₆]⁻ anion were not observed either in the ³¹P{¹H} MAS or the ¹⁹F{¹H} solution NMR spectra of **2@1-SO₃Na** after digestion, in contrast with the respective spectra of **2-PF₆** (Figures 1 c and S13). The down-field chemical shift and peak broadening observed for the signal due to the PCy₃ ligand in the ³¹P{¹H} MAS NMR spectrum of **2@1-SO₃Na**, compared to **2-PF₆**, likely originate from the different anionic environment surrounding **2**.^[20]

The ¹H solution NMR spectrum of **2@1-SO₃Na** after digestion showed three low intensity peaks at $\delta = 8.22$, 7.58 and 7.16 ppm, assigned to pyridine (Figure S14). Treatment of **2@1-SO₃Na** with D₂ gas resulted in deuteration of the cod ligand and formation of [D₄]-cyclooctane, as detected by ²H MAS NMR spectroscopy (Figure S15). These analytical and spectroscopic data are consistent with cation **2** being encapsulated intact inside the mesopores of **1-SO₃Na** by a simple cation exchange process.

To explore the possible interaction of the sulfonate groups decorating the pore walls of **1-SO₃Na** with the Ir center of **2** after encapsulation, the tosylate anion [OTs]⁻ was selected to model the BDC-SO₃ linker. Two new complexes were synthesized, [Ir(cod)(PCy₃)(py)][OTs] (**2-OTs**) and [Ir(cod)-(PCy₃)(OTs)] (**3**), in which OTs⁻ acts as a counter anion or as a ligand to Ir, respectively (Figures 2 a, b, Figures S16, S17, Table S4). ³¹P{¹H} and ¹H EXSY NMR spectroscopy in CD₂Cl₂ (Figures S18–S20) revealed that a dynamic reversible ligand exchange takes place between complexes **2-OTs** and **3**, with OTs replacing pyridine in the coordination sphere of Ir (Figure 2 c). This suggests that the sulfonate groups in **2@1-SO₃Na** may also play a non-spectator role, with potential implications in catalysis, as discussed next.

The catalytic performance of **2@1-SO₃Na** was benchmarked against **2-PF₆** in the hydrogenation of non-functionalized alkenes in CH₂Cl₂ under mild conditions (Table 1). Control experiments verified that **1-SO₃Na** does not catalyze

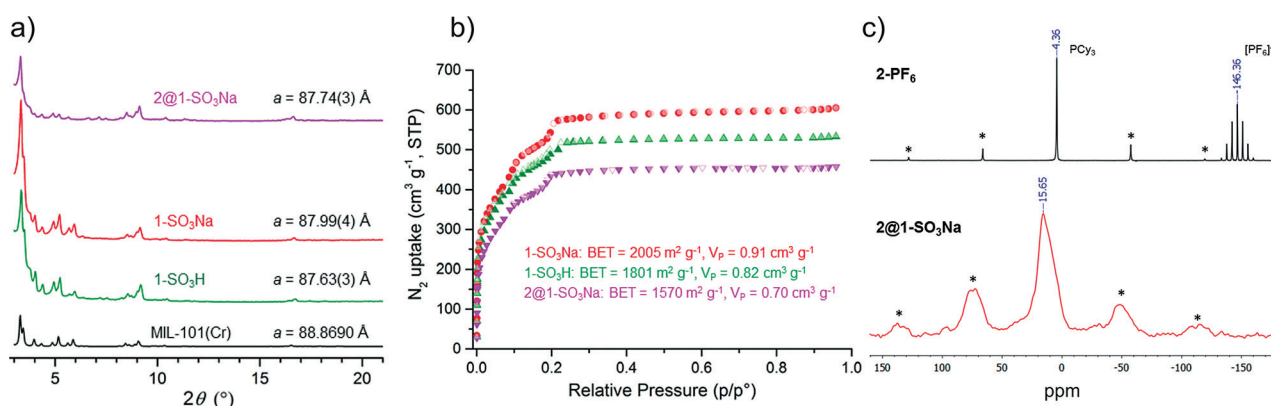


Figure 1. a) Comparison of PXR D patterns and unit cell parameter ($Fd\bar{3}m$ space group) for **2@1-SO₃Na** (magenta), **1-SO₃Na** (red), **1-SO₃H** (green) and MIL-101(Cr) (calculated, black).^[17] Le Bail fits are included in the supporting information. b) N₂ uptake of the desolvated materials at 77 K (BET = surface area, V_p = pore volume). c) ³¹P{¹H} MAS NMR spectrum of **2-PF₆** (black) and **2@1-SO₃Na** (red). Spinning side bands are marked with an asterisk.

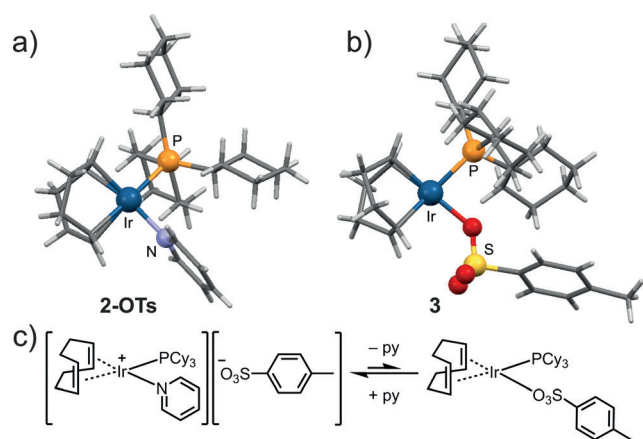


Figure 2. a) Single crystal structure of **2-OTs** (OTs⁻ counter anion is not shown for clarity). b) Single crystal structure of **3**. c) Reversible ligand exchange between **2-OTs** and **3** in CD₂Cl₂.

Table 1: Hydrogenation of non-functionalized alkenes with heterogeneous **2@1-SO₃Na** and homogeneous **2-PF₆** catalysts.^[a]

Entry	Substrate	Loading [ppm]	t [h]	2@1-SO₃Na		2-PF₆	
				Conv ^[b] [%]	TON	Conv ^[b] [%]	TON
1		1000 ^[c]	3	> 99	> 990	100	1000
2		100 ^[d]	20	100	10000	100	10000
3		50 ^[e]	24	100	20000	–	–
4		10 ^[f]	24	83	83000	94	94000
5		1000 ^[c]	3	> 99	> 990	100	1000
6		20	100	1000	–	–	
7		1000 ^[c]	3	10	100	12	120
8		20	26	260	37	370	
9		1000 ^[c]	3	69	690	100	1000
10		20	81	810	–	–	

[a] CH₂Cl₂ solvent, T = 20 °C. [b] Conversion (%) based on GC. [c] [alkene] = 0.5 M, V = 1 mL, 8 mmol of H₂. [d] [alkene] = 1.0 M, V = 4 mL, 16 mmol of H₂. [e] [alkene] = 1.0 M, V = 10 mL, 48 mmol of H₂. [f] [alkene] = 1.5 M, V = 12 mL, 48 mmol of H₂.

the hydrogenation of oct-1-ene (**4**). Introduction of **2@1-SO₃Na** as the catalyst afforded complete hydrogenation of **4** to *n*-octane, at loadings as low as 50 ppm (entries 1–3). When the loading was reduced to 10 ppm (entry 4), conversion of **4** to *n*-octane reached 83% (TON = 8.3 × 10⁴). Homogeneous catalyst **2-PF₆** under identical conditions produced comparable results, demonstrating that encapsulation is not detrimental to catalytic activity.

The branched, but unhindered, aliphatic alkene, 3-methylhex-1-ene (**5**) was also completely hydrogenated using **2@1-SO₃Na** at 1000 ppm loading (entries 5 and 6). The hindered aliphatic alkene, 2-methylhex-1-ene (**6**) was only partially hydrogenated with either catalyst after 20 h (entries 7 and 8). Conversion did not increase any further after 72 h in either

system, reflecting catalyst deactivation. When cyclohexene (**7**) was employed as a substrate, conversion reached 69% in 3 h with **2@1-SO₃Na** as the catalyst but increased only to 81% after 20 h. On the contrary, 100% conversion was observed with **2-PF₆** in 3 h (entries 9 and 10).

The different response observed for this bulkier substrate is consistent with hydrogenation taking place within the pores and not on the surface of **2@1-SO₃Na**. The heterogeneity of the reaction was further established by carrying out a leaching test (Figure S21). Recycling of **2@1-SO₃Na** was also possible with a small decrease in activity (82% conversion) during the third cycle (Figure S22).

Compound **2@1-SO₃Na** is a versatile catalyst which can also be employed in a gas/solid reaction,^[21] as demonstrated by the complete hydrogenation of but-1-ene over **2@1-SO₃Na** in 2.5 h (4000 μmol of but-1-ene hydrogenated per 1 mg of Ir). Although finely ground solid **2-PF₆** was also active, dispersion of **2** in the porous anionic solid-state support increases the number of accessible catalytic sites in **2@1-SO₃Na**, resulting in a sixfold increase in activity compared to the non-porous solid **2-PF₆** (Figure 3). Recycling of **2@1-SO₃Na** was also successful upon exposure to fresh but-1-ene (Figure S23).

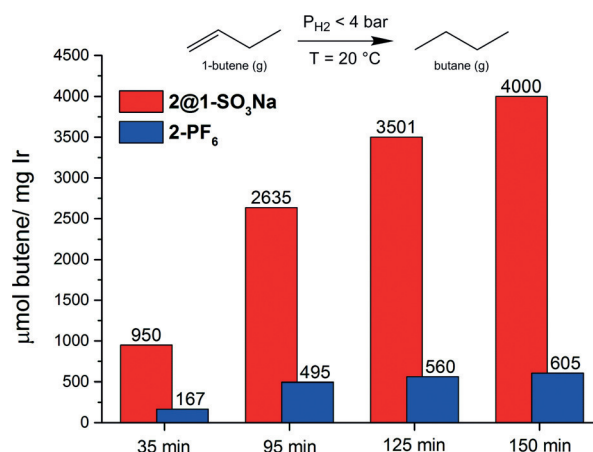


Figure 3. Conversion of but-1-ene into *n*-butane in a gas/solid hydrogenation reaction over **2@1-SO₃Na** (red) and **2-PF₆** (blue). Conditions: T = 20 °C, P_{H₂} < 4 bar, 0.5 mg of solid catalyst used.

The mesopores of **2@1-SO₃Na** are hydrophilic due to the presence of H-bond accepting sulfonate groups as well as Lewis acidic Cr^{III} sites and Na⁺ cations. Therefore, the reactivity of Crabtree's catalyst with substrates bearing functional groups that can interact with such an environment could significantly change due to encapsulation. We chose to explore this by using olefinic alcohols as substrates, whose fundamental characteristic is the competition between hydrogenation and isomerization upon turnover.^[22] Hydrogenation of a series of olefinic alcohols was carried out under a ≈ 20-fold excess of H₂ (Table 2).

Complete hydrogenation of pent-4-en-1-ol (**8a**), pent-4-en-2-ol (**9a**), and 2-methylbut-3-en-1-ol (**10a**) to the respective alcohols **8b–10b** was observed with **2-PF₆** in 3 h. Isomerization products were not detected (Figure S24), as reported for **2-PF₆** using similar substrates.^[23] Complete hydrogenation

Table 2: Substrate conversion^[a] and product selectivity^[a,b] for hydrogenation of olefinic alcohols with heterogeneous **2@1-SO₃Na** and homogeneous **2-PF₆** catalysts.^[c]

$$\text{R-CH=CH-(CH}_2\text{)}_n\text{-OH} \xrightarrow[\text{T = 20 }^\circ\text{C}]{\text{P}_{\text{H}_2} < 4 \text{ bar, 0.1 mol \%}} \text{R-CH}_2\text{-CH}_2\text{-(CH}_2\text{)}_n\text{-OH} + \text{R-CH=CH-(CH}_2\text{)}_n\text{-OH} + \text{R-CH}_2\text{-(CH}_2\text{)}_{n+2}\text{-CHO}$$

$\text{R = H; } n = 1, 2$
 $\text{R = Me, Et; } n = 0$

Entry	Substrate	t [h]	2@1-SO₃Na				2-PF₆			
			Conv [%]	b	c	d	Conv [%]	b	c	d
1		3	34	100	n.d. ^[d]	n.d.	100	100	n.d.	n.d.
2		24	100	100	n.d.	n.d.	–	–	–	–
3		3	22	100	n.d.	n.d.	100	100	n.d.	n.d.
4		24	100	100	n.d.	n.d.	–	–	–	–
5		3	10	100	n.d.	n.d.	100	100	n.d.	n.d.
6		24	100	100	n.d.	n.d.	–	–	–	–
7		3	33	95	5	n.d.	56	85	13	2
8		24	100	100	n.d.	n.d.	57	86	12	2
9		3	41	93	n.d.	7	69 ^[e]	61	n.d.	35
10		24	96	92	n.d.	8	62 ^[e]	55	n.d.	19
11		3	26	92	n.d.	8	54 ^[e]	31	n.d.	54
12		24	82	90	n.d.	10	53 ^[e]	28	n.d.	26

[a] Based on ¹H NMR using mesitylene as standard for verifying mass-balance. [b] Yield of each product over total conversion. [c] 0.1 mol% loading, [substrate] = 0.5 M in CH₂Cl₂, V = 0.7 mL, ≈ 8 mmol of H₂. [d] Not detected. [e] Formation of ill-defined condensation products was also observed, especially in 24 h.

of **8a–10a** to **8b–10b** was also achieved with **2@1-SO₃Na**, albeit in 24 h (Table 2, entries 1–6). Isomerization products were again not detected. Conversion in 3 h correlates well with the steric hindrance around the double bond of the substrate: 10% for **10a** (more hindered), increasing to 22% for **9a** (less hindered), and reaching 34% for **8a** (linear). Olefinic alcohols **8a–10a** were hydrogenated considerably slower with **2@1-SO₃Na**, compared to the sterically comparable non-functionalized alkenes **4** and **5** (Table 1). This is consistent with a strong interaction between the hydroxyl group of the olefinic alcohols and the chemical microenvironment of **2@1-SO₃Na**.

Substrates which are intrinsically more susceptible to isomerization, such as the homoallylic (**11a**) and allylic (**12a**, **13a**) alcohols,^[23,24] revealed a significant enhancement of reactivity and selectivity to hydrogenation with **2@1-SO₃Na**, compared to its homogeneous counterpart. The homogeneous catalyst **2-PF₆** afforded 56% conversion of **11a** in 3 h and 57% in 24 h, indicative of catalyst deactivation (Table 2, entries 7 and 8, Figure S25). Moreover, isomerization of **11a** was also observed, producing a non-negligible amount of the internal olefinic alcohol **11c** and traces of the aldehyde **11d**. As a result, selectivity to hydrogenation and formation of *n*-butanol (**11b**) was only 86% for the homogeneous system.

By contrast, the heterogeneous catalyst **2@1-SO₃Na** afforded complete conversion and 100% selectivity to hydrogenation and formation of **11b** (Table 2, entries 7 and 8, Figure S26). Monitoring conversion over time for both systems (Figure S27) verified that **2-PF₆** is deactivated after 3 h, whereas **2@1-SO₃Na** remained productive, affording full

conversion in 6 h. Although traces of the internal olefin **11c** were detected in short reaction times, **11c** was subsequently also hydrogenated to **11b**. The encapsulated catalyst is thus more stable, more active with respect to overall conversion, and more selective.

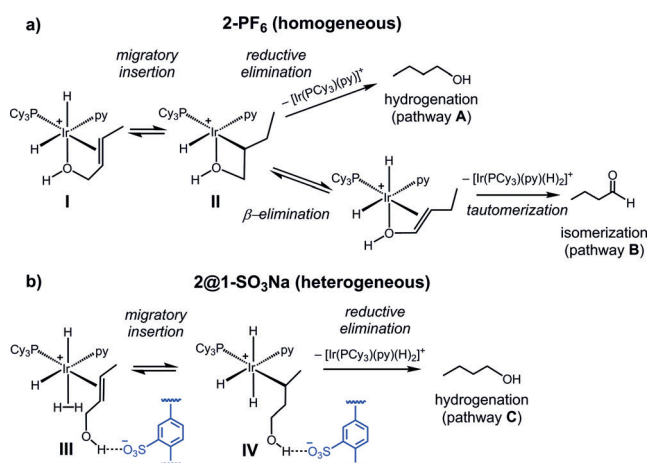
The superior performance of **2@1-SO₃Na** was even more pronounced in the hydrogenation of allylic alcohols that can isomerize directly to the respective aldehydes. Conversion under hydrogenation conditions for *trans*-pent-2-en-1-ol (**12a**, entries 9 and 10) and *trans*-crotyl alcohol (**13a**, entries 11 and 12) in 3 h with **2-PF₆** was 69% and 54%, respectively (Figure S28). Conversion did not increase after 24 h, indicating catalyst deactivation. Selectivity to hydrogenation was poor: 61% for alcohol **12b** in 3 h with a substantial amount of the aldehyde **12d** formed (35% selectivity), and 31% for alcohol **13b** in 3 h with the aldehyde **13d** now being the main product (54% selectivity). By

contrast, overall conversion with **2@1-SO₃Na** as the catalyst reached 96% for **12a** and 82% for **13a** in 24 h (Figure S29). Isomerization to the aldehydes **12d** and **13d** was significantly suppressed, resulting in ≥ 90% selectivity for the alcohols **12b** and **13b**.

To probe the effect of the sulfonate group on stability and selectivity, we also investigated the homogeneous hydrogenation of crotyl alcohol using **2-OTs** and **3** as catalysts (Figure S30). Higher conversions were observed compared to **2-PF₆** (77% for **2-OTs** and 83% for **3** in 24 h) in accordance with OTs[−] being a more strongly coordinating anion, hence prolonging the catalyst's lifetime.^[25] By contrast, selectivity to hydrogenation did not significantly improve (39% for **2-OTs** and 53% for **3**), remaining considerably lower than that of **2@1-SO₃Na** (≥ 90%).

The reaction pathways for the hydrogenation or isomerization of olefinic alcohols with the homogeneous catalyst **2-PF₆** likely share the same starting point, the formation of a cationic Ir^{III}-dihydride complex in which the hydroxyl group is also coordinated to Ir (Scheme 2, intermediate **I**), followed by migratory insertion (intermediate **II**).^[13,14] Bifurcation into separate, competitive pathways then occurs: i) hydrogenation to the respective alcohol via reductive elimination (pathway **A**) or ii) isomerization to the internal olefin via β-elimination, which requires an appropriately orientated vacant coordination site, followed by off-cycle tautomerization to the aldehyde (pathway **B**).

The significantly improved selectivity to hydrogenation observed with **2@1-SO₃Na** suggests that isomerization is suppressed. We propose that this could take place due to



Scheme 2. a) Competing pathways for hydrogenation (A) and isomerization (B) of olefinic alcohols with **2-PF₆**. b) Proposed pathway for hydrogenation (C) and suppression of isomerization with **2@1-SO₃Na**.

extended coordination sphere interactions between the hydroxyl group of the olefinic alcohols and the chemical microenvironment around **2**, such as H-bonding to the sulfonate groups. This disfavors coordination of the hydroxyl group to Ir and enables formation of the dihydrogen complex **III**, in preference to **I** (pathway C). Productive hydrogenation occurs via an octahedral Ir^V-trihydride species (**IV**), as proposed for non-functionalized alkenes with Crabtree-type catalysts^[26] and β -elimination is suppressed since Ir is coordinatively saturated throughout.

Catalyst **2@1-SO₃Na** also resulted in higher overall conversions for the hydrogenation of olefinic alcohols, compared to **2-PF₆**. A series of selective poisoning experiments revealed that the isomerization products are not responsible for catalyst deactivation (Table S7). We thus suggest that **2@1-SO₃Na** has a longer lifetime due to: i) spatial isolation of the positively charged catalytically active species inside the pores of the anionic MOF which hinders the formation of catalytically inactive clusters and/or ii) reversible coordination of the sulfonate anion, as shown with **2-OTs** and **3**.

In summary, we demonstrate that the hybrid catalyst **2@1-SO₃Na** is capable of hydrogenating non-functionalized alkenes at low loadings in solution and in the gas phase under mild conditions. It outperforms its homogeneous counterpart in the hydrogenation of olefinic alcohols, showing significantly higher conversions under otherwise identical conditions. In addition, encapsulation results in a pronounced selectivity enhancement in favor of hydrogenation by suppressing the competing isomerization reaction due to extended coordination sphere interactions of the catalytic center with the chemically functionalized internal surface of the MOF. Capitalizing on such stability and selectivity enhancements is likely to be important in catalytic applications in continuous flow.^[27] In metalloenzymes, it is well-established that well-positioned amino acid residues around the active site control reactivity and selectivity.^[3] Here, the well-defined, readily engineered MOF chemical microenvi-

ronment controls reactivity and selectivity of the encapsulated catalyst, allowing discrimination between two distinct reaction pathways.

Acknowledgements

The EPSRC is thanked for support for this work through the UK Catalysis Hub Consortium under grants EP/K014706/1, EP/K014668/1, EP/K014854/1, EP/K014714/1, and EP/M013219/1. A.S.W. also acknowledges EPSRC for funding (EP/M024210/1). M.J.R. is a Royal Society Research Professor.

Conflict of interest

The authors declare no conflict of interest.

Keywords: allylic alcohols · Crabtree's catalyst · encapsulation · hydrogenation · metal–organic frameworks

How to cite: *Angew. Chem. Int. Ed.* **2018**, *57*, 4532–4537
Angew. Chem. **2018**, *130*, 4622–4627

- [1] a) H. Furukawa, K. E. Cordova, M. O'Keeffe, O. M. Yaghi, *Science* **2013**, *341*, 1230444; b) B. Li, H.-M. Wen, Y. Cui, W. Zhou, G. Qian, B. Chen, *Adv. Mater.* **2016**, *28*, 8819–8860; c) H.-C. Zhou, S. Kitagawa, *Chem. Soc. Rev.* **2014**, *43*, 5415–5418, and references therein.
- [2] a) J. Gascon, A. Corma, F. Kapteijn, F. X. Llabrés i Xamena, *ACS Catal.* **2014**, *4*, 361–378; b) A. H. Chughtai, N. Ahmad, H. A. Younus, A. Laypkov, F. Verpoort, *Chem. Soc. Rev.* **2015**, *44*, 6804–6849; c) A. Grigoropoulos in *Modern Developments in Catalysis*, World Scientific (Europe), London, **2016**, pp. 123–158; d) S. M. J. Rogge, A. Bavykina, J. Hajek, H. Garcia, A. I. Olivos-Suarez, A. Sepulveda-Escribano, A. Vimont, G. Clet, P. Bazin, F. Kapteijn, M. Daturi, E. V. Ramos-Fernandez, F. X. Llabrés i Xamena, V. Van Speybroeck, J. Gascon, *Chem. Soc. Rev.* **2017**, *46*, 3134–3184.
- [3] a) S. W. Ragsdale, *Chem. Rev.* **2006**, *106*, 3317–3337; b) M. Zhao, H.-B. Wang, L.-N. Ji, Z.-W. Mao, *Chem. Soc. Rev.* **2013**, *42*, 8360–8375; c) R. H. Holm, E. I. Solomon, *Chem. Rev.* **2014**, *114*, 3367–3368, and references therein.
- [4] a) C. W. Jones, *Top. Catal.* **2010**, *53*, 942–952; b) M. Raynal, P. Ballester, A. Vidal-Ferran, P. W. N. M. van Leeuwen, *Chem. Soc. Rev.* **2014**, *43*, 1660–1733 and 1734–1787; c) C. J. Brown, F. D. Toste, R. G. Bergman, K. N. Raymond, *Chem. Rev.* **2015**, *115*, 3012–3035; d) S. Hübner, J. G. de Vries, V. Farina, *Adv. Synth. Catal.* **2016**, *358*, 3–25.
- [5] a) A. Fateeva, P. A. Chater, C. P. Ireland, A. A. Tahir, Y. Z. Khimyak, P. V. Wiper, J. R. Darwent, M. J. Rosseinsky, *Angew. Chem. Int. Ed.* **2012**, *51*, 7440–7444; *Angew. Chem.* **2012**, *124*, 7558–7562; b) L. Mitchell, P. Williamson, B. Ehrlichová, A. E. Anderson, V. R. Seymour, S. E. Ashbrook, N. Acerbi, L. M. Daniels, R. I. Walton, M. L. Clarke, P. A. Wright, *Chem. Eur. J.* **2014**, *20*, 17185–17197; c) K. Manna, T. Zhang, M. Carboni, C. W. Abney, W. Lin, *J. Am. Chem. Soc.* **2014**, *136*, 13182–13185; d) H. Fei, S. M. Cohen, *J. Am. Chem. Soc.* **2015**, *137*, 2191–2194; e) Z. Li, J.-D. Xiao, H.-L. Jiang, *ACS Catal.* **2016**, *6*, 5359–5365; f) A. M. Fracaroli, P. Siman, D. A. Nagib, M. Suzuki, H. Furukawa, F. D. Toste, O. M. Yaghi, *J. Am. Chem. Soc.* **2016**, *138*, 8352–8355; g) N. C. Thacker, Z. Lin, T. Zhang, J. C.

- Gilhula, C. W. Abney, W. Lin, *J. Am. Chem. Soc.* **2016**, *138*, 3501–3509.
- [6] a) J. Canivet, S. Aguado, Y. Schuurman, D. Farrusseng, *J. Am. Chem. Soc.* **2013**, *135*, 4195–4198; b) A. M. Rasero-Almansa, A. Corma, M. Iglesias, F. Sanchez, *Green Chem.* **2014**, *16*, 3522–3527; c) S. A. Burgess, A. Kassie, S. A. Baranowski, K. J. Fritzsche, K. Schmidt-Rohr, C. M. Brown, C. R. Wade, *J. Am. Chem. Soc.* **2016**, *138*, 1780–1783; d) M. Rimoldi, A. Nakamura, N. A. Vermeulen, J. J. Henkelis, A. K. Blackburn, J. T. Hupp, J. F. Stoddart, O. K. Farha, *Chem. Sci.* **2016**, *7*, 4980–4984; e) A. Burgun, C. J. Coghlan, D. M. Huang, W. Chen, S. Horike, S. Kitagawa, J. F. Alvino, G. F. Metha, C. J. Sumbly, C. J. Doonan, *Angew. Chem. Int. Ed.* **2017**, *56*, 8412–8416; *Angew. Chem.* **2017**, *129*, 8532–8536.
- [7] a) Z. Guo, C. Xiao, R. V. Maligal-Ganesh, L. Zhou, T. W. Goh, X. Li, D. Tesfagaber, A. Thiel, W. Huang, *ACS Catal.* **2014**, *4*, 1340–1348; b) S. Yuan, Y. P. Chen, J. S. Qin, W. Lu, L. Zou, Q. Zhang, X. Wang, X. Sun, H.-C. Zhou, *J. Am. Chem. Soc.* **2016**, *138*, 8912–8919; c) L. Liu, T.-Y. Zhou, S. G. Telfer, *J. Am. Chem. Soc.* **2017**, *139*, 13936–13943.
- [8] a) D. J. Xiao, J. Oktawiec, P. J. Milner, J. R. Long, *J. Am. Chem. Soc.* **2016**, *138*, 14371–14379; b) W. Shi, L. Cao, H. Zhang, X. Zhou, B. An, Z. Lin, R. Dai, J. Li, C. Wang, W. Lin, *Angew. Chem. Int. Ed.* **2017**, *56*, 9704–9709; *Angew. Chem.* **2017**, *129*, 9836–9841; c) L. Li, Q. Yang, S. Chen, X. Hou, B. Liu, J. Lu, H.-L. Jiang, *Chem. Commun.* **2017**, *53*, 10026–10029.
- [9] R. Crabtree, *Acc. Chem. Res.* **1979**, *12*, 331–337.
- [10] a) D. F. Chodosh, R. H. Crabtree, H. Felkin, S. Morehouse, G. E. Morris, *Inorg. Chem.* **1982**, *21*, 1307–1311; b) Y. Xu, M. A. Celik, A. L. Thompson, H. Cai, M. Yurtsever, B. Odell, J. C. Green, D. M. P. Mingos, J. M. Brown, *Angew. Chem. Int. Ed.* **2009**, *48*, 582–585; *Angew. Chem.* **2009**, *121*, 590–593.
- [11] a) H. M. Lee, T. Jiang, E. D. Stevens, S. P. Nolan, *Organometallics* **2001**, *20*, 1255–1258; b) L. D. Vazquez-Serrano, B. T. Owens, J. M. Buriak, *Inorg. Chim. Acta* **2006**, *359*, 2786–2797; c) E. L. Kolychev, S. Kronig, K. Brandhorst, M. Freytag, P. G. Jones, M. Tamm, *J. Am. Chem. Soc.* **2013**, *135*, 12448–12459.
- [12] a) S. P. Smidt, N. Zimmermann, M. Studer, A. Pfaltz, *Chem. Eur. J.* **2004**, *10*, 4685–4693; b) G. L. Moxham, T. M. Douglas, S. K. Brayshaw, G. Kociok-Kohn, J. P. Lowe, A. S. Weller, *Dalton Trans.* **2006**, 5492–5505.
- [13] H. Li, C. Mazet, *Acc. Chem. Res.* **2016**, *49*, 1232–1241.
- [14] a) G. Stork, D. E. Kahne, *J. Am. Chem. Soc.* **1983**, *105*, 1072–1073; b) R. H. Crabtree, M. W. Davis, *J. Org. Chem.* **1986**, *51*, 2655–2661.
- [15] a) D. T. Genna, A. G. Wong-Foy, A. J. Matzger, M. S. Sanford, *J. Am. Chem. Soc.* **2013**, *135*, 10586–10589; b) A. Grigoropoulos, G. F. S. Whitehead, N. Perret, A. P. Katsoulidis, F. M. Chadwick, R. P. Davies, A. Haynes, L. Brammer, A. S. Weller, J. Xiao, M. J. Rosseinsky, *Chem. Sci.* **2016**, *7*, 2037–2050.
- [16] a) G. Akiyama, R. Matsuda, H. Sato, M. Takata, S. Kitagawa, *Adv. Mater.* **2011**, *23*, 3294–3297; b) Y.-X. Zhou, Y.-Z. Chen, Y. Hu, G. Huang, S.-H. Yu, H.-L. Jiang, *Chem. Eur. J.* **2014**, *20*, 14976–14980; c) J. Juan-Alcañiz, R. Gielisse, A. B. Lago, E. V. Ramos-Fernandez, P. Serra-Crespo, T. Devic, N. Guillou, C. Serre, F. Kapteijn, J. Gascon, *Catal. Sci. Technol.* **2013**, *3*, 2311–2318.
- [17] G. Férey, C. Mellot-Draznieks, C. Serre, F. Millange, J. Dutour, S. Surlblé, I. Margiolaki, *Science* **2005**, *309*, 2040–2042.
- [18] Y. Zhang, B. Li, R. Krishna, Z. Wu, D. Ma, Z. Shi, T. Pham, K. Forrest, B. Space, S. Ma, *Chem. Commun.* **2015**, *51*, 2714–2717.
- [19] D. T. Genna, L. Y. Pfund, D. C. Samblanet, A. G. Wong-Foy, A. J. Matzger, M. S. Sanford, *ACS Catal.* **2016**, *6*, 3569–3574.
- [20] Note: Although we cannot discount chemical shift changes in the ^{31}P SS NMR spectrum due to the presence of paramagnetic Cr^{III} centers, no significant changes were reported when molecular species were incorporated into pristine MIL-101(Cr). See reference [17].
- [21] a) D. Yang, S. O. Odoh, J. Borycz, T. C. Wang, O. K. Farha, J. T. Hupp, C. J. Cramer, L. Gagliardi, B. C. Gates, *ACS Catal.* **2016**, *6*, 235–247; b) F. M. Chadwick, A. I. McKay, A. J. Martinez-Martinez, N. H. Rees, T. Kramer, S. A. Macgregor, A. S. Weller, *Chem. Sci.* **2017**, *8*, 6014–6029.
- [22] a) M. Moreno, L. N. Kissell, J. B. Jasinski, F. P. Zamborini, *ACS Catal.* **2012**, *2*, 2602–2613; b) E. Sadeghmoghaddam, H. Gu, Y.-S. Shon, *ACS Catal.* **2012**, *2*, 1838–1845.
- [23] a) L. Mantilli, C. Mazet, *Tetrahedron Lett.* **2009**, *50*, 4141–4144; b) L. Mantilli, D. Gérard, S. Torche, C. Besnard, C. Mazet, *Angew. Chem. Int. Ed.* **2009**, *48*, 5143–5147; *Angew. Chem.* **2009**, *121*, 5245–5249.
- [24] R. Uma, C. Crevisy, R. Gree, *Chem. Rev.* **2003**, *103*, 27–51.
- [25] a) A. Rifat, G. Kociok-Köhn, J. W. Steed, A. S. Weller, *Organometallics* **2004**, *23*, 428–432; b) E. Piras, F. Läng, H. Rügger, D. Stein, M. Würle, H. Grützmacher, *Chem. Eur. J.* **2006**, *12*, 5849–5858.
- [26] J. J. Verendel, O. Pàmies, M. Diéguez, P. G. Andersson, *Chem. Rev.* **2014**, *114*, 2130–2169.
- [27] U. Hintermair, G. Francio, W. Leitner, *Chem. Commun.* **2011**, *47*, 3691–3701.

Manuscript received: October 20, 2017

Accepted manuscript online: January 28, 2018

Version of record online: March 22, 2018