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Atmospheric Chemistry and Physics Discussions

Heterogeneous reaction of HO_2 with airborne TiO_2 particles and its implication for climate change mitigation strategies

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Abstract.

One geoengineering mitigation strategy for global temperature rises resulting from the increased concentrations of greenhouse gases is to inject particles into the stratosphere to scatter solar radiation back to space, with TiO_2 particles emerging as a possible candidate. Uptake coefficients of HO_2, $\gamma(HO_2)$, onto sub-micrometre TiO_2 particles were measured at room temperature and different relative humidities (RH) using an atmospheric pressure aerosol flow tube coupled to a sensitive HO_2 detector. Values of $\gamma(HO_2)$ increased from 0.021 ± 0.001 to 0.036 ± 0.007 as the RH was increased from 11% to 66%, and the increase in $\gamma(HO_2)$ correlated with the number of monolayers of water surrounding the TiO_2 particles. The impact of the uptake of HO_2 onto TiO_2 particles on stratospheric concentrations of HO_2 and O_3 was simulated using the TOMCAT three-dimensional chemical transport model. The model showed that by injecting the amount of TiO_2 required to achieve the same cooling effect as the Mt. Pinatubo eruption, heterogeneous reactions between HO_2 and TiO_2 would have a negligible effect on stratospheric concentrations of HO_2 and O_3.

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1. Introduction

It has been suggested that injection of aerosols into the stratosphere in order to scatter solar radiation back to space could be a possible solar radiation management scheme (Shepherd and Working Group on Geoengineering the Climate, 2009). Such a scheme would have the effect of cooling the Earth's surface and serve as a measure to mitigate enhanced global warming. A possible candidate for stratospheric injection is sulphuric acid aerosols as these aerosols occur naturally in the stratosphere. These aerosols are formed by transport of precursors, e.g. SO_2 and carbonyl sulphide (OCS), from the troposphere to the stratosphere (Holloway and Wayne, 2010), which form sulphuric acid vapour that condenses onto particles. However, injection of these particles could have important negative effects on the stratosphere through enhanced ozone depletion. Concentrations of sulphate aerosols can increase dramatically as the result of volcanic eruptions. Following the eruption of Mt. Pinatubo in 1991, it was estimated that around 30 Tg of H_2SO_4 was injected in the stratosphere, dramatically increasing stratospheric aerosol loading and hence the available surface area for heterogeneous chemistry to occur. Following this volcanic event, the average global lower tropospheric temperature decreased by 0.5 K (Dutton and Christy, 1992), however, stratospheric ozone concentrations reached a record low in northern mid-latitudes (Dutton and Christy, 1992;McCormick et al., 1995) showing sulphate aerosols to be unsuitable for solar radiation management. Other candidates for particle types, such as TiO_2, have been put forward due to their large refractive indices (Pope et al., 2012), meaning that less stratospheric aerosol loading would be necessary to achieve the same level of cooling. The refractive index of TiO_2 at 550 nm is 2.5 compared to a value of 1.5 for naturally occurring sulphate aerosols (Tang et al., 2014). Assuming that the size of TiO_2 particles can be optimised, it has been reported that to achieve the same cooling effect that sulphate aerosols had during the Mt. Pinatubo event, approximately three times less in mass, and seven times less in volume of TiO_2 would be required compared with sulphuric acid (Pope et al., 2012). However, the impacts of the presence of TiO_2 particles on stratospheric chemistry have to be determined before this kind of geoengineering solution can be considered. Mineral dust particles are commonly found in the troposphere and contribute the largest fraction to tropospheric aerosol loading in terms of mass (Textor et al., 2006;Huneeus et al., 2011). Typically TiO_2 (which is classified as a mineral) constitutes from 0.1% to 10% of overall atmospheric mineral dust loading depending on the location of sources (Usher et al., 2003;Karagulian et al., 2006).

The heterogeneous chemistry of sulphate aerosols in the stratosphere is fairly well understood (Ammann et al., 2013); for example the conversion of NO_x to nitric acid in the aerosol via N_2O_5 adsorption and reaction, and also the activation of chlorine via the reaction of $ClONO_2$ with HCl to form Cl_2 and nitric acid within cold aerosols. However, the heterogeneous reactivity of mineral particles, in particular TiO_2, is not as well understood. Removal and production of trace gases in the stratosphere may significantly perturb concentrations of O_3, therefore it is important when assessing the potential impact of such a solar radiation management scheme to evaluate the kinetics of likely heterogeneous chemistry.

The hydroperoxyl radical, HO_2, is an important species within the stratosphere, being present at about 5 parts per trillion per unit volume (pptv) around the tropopause, and is involved in a HO_x catalytic cycle responsible for about 40% of O_3 depletion in the lower stratosphere via to the following reactions (Wennberg, 1994):

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$$OH + O_3 \rightarrow HO_2 + O_2 \quad (R1)$$

$$HO_2 + O_3 \rightarrow OH + 2O_2 \quad (R2)$$

$$\text{Net: } 2O_3 \rightarrow 3O_2$$

Moreover, HO_2 can also react with stratospheric ClO and BrO to produce HOCl and HOBr respectively, which can be photolysed to produce further OH and atomic halogen species that can contribute to O_3 loss. HO_2 can undergo self-reaction upon surfaces of mineral dust, which is thought to result in the generation of H_2O_2 (Matthews et al., 2014), whose two predominant removal pathways are photolysis and reaction with OH (Versick et al., 2012):

$$H_2O_2 + h\nu\ (\lambda \leq 360\ \text{nm}) \rightarrow 2OH \quad (R3)$$

$$H_2O_2 + OH \rightarrow HO_2 + H_2O \quad (R4)$$

Although the kinetics of the uptake of HO_2 onto Arizona Test Dust (ATD), a proxy of mineral dust, has been previously investigated (Matthews et al., 2014;Bedjanian et al., 2013) the heterogeneous reaction of HO_2 with TiO_2 has not been studied. However, the kinetics of N_2O_5 uptake (Tang et al., 2014) and $ClONO_2$ (Tang et al., 2016) onto TiO_2 have been studied. The heterogeneous reaction of N_2O_5 results in the conversion to reactive nitrogen oxides (NO and NO_2) involved in a catalytic cycle that leads to significant O_3 depletion and non-reactive HNO_3. The reactive uptake coefficient, γ, which is the probability that a species will collide with an aerosol and be removed by reaction, was measured to be more than an order of magnitude larger for HO_2 onto ATD than for N_2O_5 and $ClONO_2$ onto TiO_2 sub-micron particles, and contrasting dependences of γ with relative humidities (RH) were observed. Therefore, by analogy it is expected that HO_2 uptake onto TiO_2 may be faster than N_2O_5 uptake. $ClONO_2$ uptake by TiO_2 particles resulted in similar values of γ, however, no dependency of γ with RH between 7 – 33% was observed.

In this investigation an aerosol flow tube coupled with a sensitive HO_2 detector based on chemical conversion followed by laser-induced fluorescence detection of OH (George et al., 2013) has been used to study the kinetics of the heterogeneous reaction of HO_2 with airborne TiO_2 nanoparticles at different RH. While it has been possible here to study such kinetics over a range of RH representative to those typically found in the lower stratosphere (< 40%) (Wennberg et al., 1994), experimental limitations meant that experiments were only conducted at room temperature (~293 K). The TOMCAT off-line three-dimensional (3D) chemical transport model (Chipperfield; 1999) has also been used to predict the likely impact of HO_2 uptake by TiO_2 particles on the stratospheric concentrations of HO_2 and O_3.

2. Methods.

2.1. Overview of experimental apparatus

The experimental setup deployed for this investigation is similar to other investigations of HO_2 uptake by aerosols undertaken at the University of Leeds (George et al., 2013;Matthews et al., 2014;Lakey et al., 2016) therefore a

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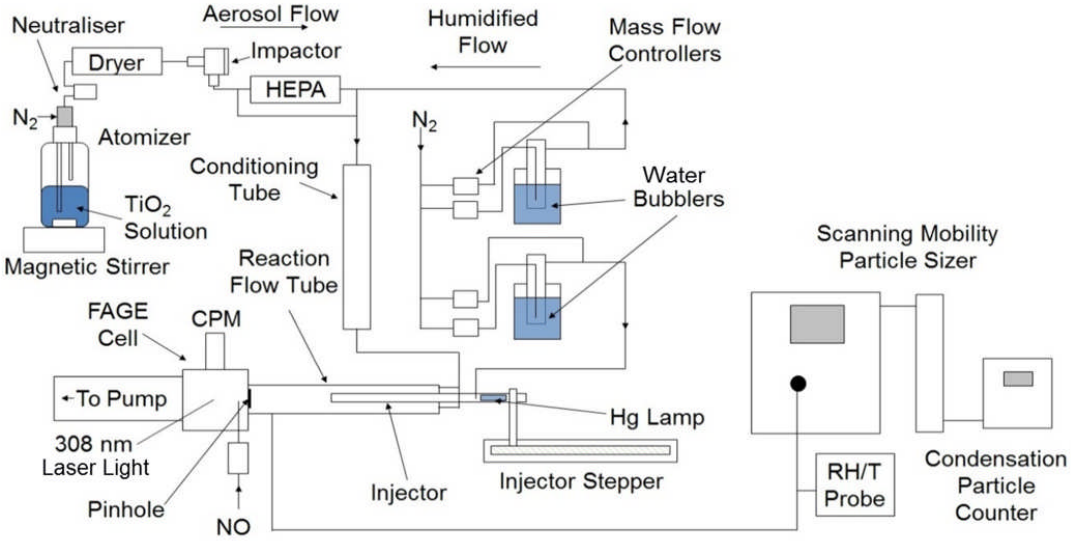
95 detailed description of the components of the experiment is not given. A schematic diagram of the experiment is
96 shown in Figure 1, and all experiments were undertaken at room temperature (~293 K) and under normal laboratory
97 levels of illumination. For some experiments the flow tube was covered with a black shield to eliminate light and no
98 differences in the results were observed.



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100 **Figure 1**. Schematic diagram of the aerosol flow tube experiment. CPM: Channel Photomultiplier, HEPA: high-
101 efficiency particulate air filter, FAGE: fluorescence assay by gas expansion, RH/T: Relative Humidity/Temperature.

102

103 Compressed nitrogen, which had been passed through a gas purification system (TSI 3074B) consisting of particle
104 filters, a dryer and a carbon filter, was used as the carrier gas for the experiments. A humidified flow of TiO_2 particles
105 was introduced through two inlets located at the rear of the aerosol flow tube and the flow of HO_2 radicals enters the
106 flow tube via the movable injector. The total flow through the flow tube (107 cm length, 5.9 cm I.D.) was 5.4 L
107 min^{-1} which resulted in a Reynolds number of 130, and therefore is considered laminar as confirmed by radial
108 concentration gradient measurements of gases exiting the injector (George et al., 2013). Experiments consisted of
109 moving the injector using a linear drive (BSL Engineering 15 KR4610A) to different fixed positions along the flow
110 tube (30-70 cm from injector tip to HO_2 detector inlet in steps of 5 cm) corresponding to reaction times between ~ 8
111 and 20 s from the injector, with detection of HO_2 at the end of the flow tube. All gas flows within the experiment
112 were controlled using mass flow controllers (Brookes and MKS). The RH of the flow was measured using a probe
113 (Rotronics) in the exhaust of the flow tube, which itself was calibrated against a dew point hygrometer (Buck
114 Research Instruments CR-4).

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116 **2.2. Aerosol Generation and Detection**

117 A solution of TiO_2 (Aldrich Chemistry 718467, 99.5% Degussa, 5 g in 500 ml of Milli-Q water) was placed in a
118 commercial atomizer (TSI 3076) in order to produce a 1 L min^{-1} flow entrained with TiO_2 particles, referred to as the
119 aerosol flow. The aerosol flow was then passed through a neutraliser (Grimm 5522) to reduce static wall losses, a
120 diffusion drier (TSI 3062) and an impactor (TSI 1035900) to ensure larger aerosols, beyond the detection range of
121 the SMPS (~ 750 nm diameter), do not enter the aerosol flow tube. A high-efficiency particulate air (HEPA, PALL
122 Life Sciences) filter situated within a by-pass loop was used to control the number concentration of particles entering
123 the aerosol flow tube. The aerosol flow was then mixed with a humidified flow of nitrogen (3 L min^{-1}) to control the
124 RH within the system. The RH of the humidified flow was altered by changing the ratio of dry nitrogen and nitrogen
125 passed through a water bubbler. This flow was then passed through a conditioning tube (residence time ~ 6 s) before
126 entering the aerosol flow tube to allow time for water adsorption onto the surface of the TiO_2 particles to equilibrate
127 at the given RH.

128 The total surface area of TiO_2 particles available for heterogeneous reaction with HO_2 was measured with a SMPS
129 instrument from the flow exiting the aerosol flow tube. Previous experiments showed that there is a negligible loss
130 of aerosols during the transit of the flow tube (George et al., 2013). The SMPS consisted of a Differential Mobility
131 Analyzer (DMA, TSI 3080, 3081) that creates a monodisperse flow of aerosols based on their electrical mobility
132 which is related to their size. A condensation particle counter (CPC, TSI 3775) connected in parallel to the DMA
133 quantified particle number concentrations. These two instruments connected in parallel can be used to create an
134 aerosol size distribution from which the total surface area and average radius of particles can be calculated by making
135 the assumption that particles are spherical (as demonstrated experimentally in section 3.2 below). The average
136 diameter of the particles is 136 nm and 173 nm at RH = 11% and 37%, respectively.

137 **2.3 HO_2 Generation and Detection**

138 HO_2 radicals were produced within the movable injector (110 cm length, 1.9 cm O.D., 1.6 cm I.D.) by passing a 1.3
139 L min^{-1} humidified flow of nitrogen (consisting of a mixture of 0.9 L min^{-1} of dry N_2 and 0.4 L min^{-1} N_2 passed
140 through a water bubbler) containing trace amounts of oxygen over a mercury lamp (L.O.T.-Oriel 6035) via the
141 following reactions:

142 $$H_2O + h\nu\ (\lambda = 185\ nm) \rightarrow OH + H \quad (R5)$$

143 $$H + O_2 + M \rightarrow HO_2 + M \quad (R6)$$

144 OH is also created by the photolysis of water vapour in R5, but no OH was observed exiting the injector, presumably
145 owing to rapid losses at the walls of the injector. HO_2 was sampled by a 0.7 mm diameter pinhole at the end of the
146 flow tube, and after chemical conversion to OH by addition of excess NO (50 sccm, BOC, 99.5 %) just inside the
147 pinhole, laser induced fluorescence (LIF) at low-pressure (the fluorescence assay by gas expansion (FAGE) technique
148 (Heard and Pilling, 2003)) was used to measure OH. The relative LIF signal from converted HO_2 was calibrated
149 using an established method (Winiberg et al., 2015) developed for field measurements of OH and HO_2 radicals.
150 Hence the experiment is able to measure the absolute concentration of HO_2 during passage from the injector to the

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sampling inlet. The $Q_1(2)$ line of the OH ($A^2\Sigma^+$ $v' = 0 - X^2\Pi_i$ $v'' = 0$) transition at ~ 308 nm was used to detect OH. A Nd:YAG pumped dye laser (JDSU Q201-HD Q-series, Sirah Cobra Stretch) was used to produce the required 308 nm radiation (line width ~ 0.1 cm^{-1}) at a pulse repetition rate of 5 kHz. As the flows through the movable injector (1.3 L min^{-1}) and mercury lamp current (20 mA) were kept constant, it is assumed that the initial HO_2 concentration, $[HO_2]_0$ (defined in this investigation as $[HO_2]$ at the first injector position, i.e. 30 cm downstream of the injector), determined by calibration to be 1.6×10^9 molecule cm^{-3}, was the same for all experiments. These HO_2 concentrations are ~ 50 times higher than typical levels in the sunlit stratosphere (Wennberg et al., 1994). A reference fluorescence cell, in which a large concentration of OH was generated and detected by LIF, was used to facilitate the identification of OH lines and tune the laser wavelength. The FAGE cell was continuously evacuated using a combination of a rotary pump (Edwards, model E1M80) and a roots blower (EH1200), and was kept at 0.8–0.9 Torr, which was monitored using a capacitance monitor (Tylan General, CDC 11).

2.4. Experimental Procedure and Data Analysis

The HO_2 signal was measured at 8 positions as the moveable injector was drawn back from 30 to 70 cm using the linear stepper drive and again as the moveable injector was pushed forwards back to its initial position. The HO_2 signal was averaged over 20 s (average of twenty 1 s measurement points, each corresponding to 5000 laser shots) at each injector position with a 22 s delay between measurements at each injector position in order to allow time for mechanical vibrations to subside, and to ensure a full flush of the aerosol flow tube so that the LIF signal corresponds to HO_2 emitted from the injector position being measured. The laser power was recorded for each injector position and used to normalise the HO_2 signal to correct for any fluctuations in laser power (< 5%). The HO_2 signals with the injector moving forwards and backwards were then averaged, and this procedure repeated six times with varying concentrations of aerosols present in the aerosol flow tube. The wall loss rate of HO_2 (k_{wall}) was determined by recording the HO_2 decay in the absence of aerosols, but at the same RH, and was repeated four times for each experiment.

The HO_2 concentration as a function of time along the flow tube can be expressed as:

$$\ln[HO_2]_t = \ln[HO_2]_0 - k_{obs}\, t \qquad \text{(Eqn. 1)}$$

where $[HO_2]_t$ and $[HO_2]_0$ are concentrations of HO_2 at time t and $t = 0$ respectively, and k_{obs} is the observed pseudo-first-order rate coefficient for HO_2 uptake. As the HO_2 signal is directly proportional to the concentration of HO_2, the gradient of a plot of ln(HO_2 Signal) against time (calculated from the injector position and measured flow rate) yields k_{obs}, as shown in Figure 2.

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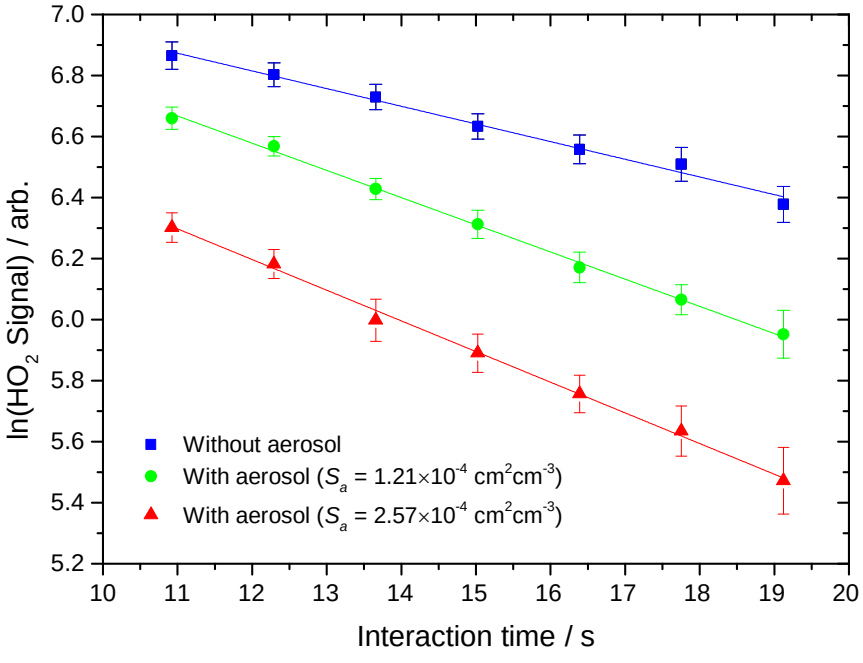

Figure 2. Measured HO_2 signal at different interaction times, in the presence of TiO_2 particles (surface area density 1.21×10^{-4} cm^2 cm^{-3} and 2.57×10^{-4} cm^2 cm^{-3}, green circles and red triangles respectively) and in their absence (blue squares). The RH in the aerosol flow tube was 11%, the lowest used in this study. The lines represent linear-least squares fits to the data yielding $k_{obs} = 0.079 \pm 0.005$ s^{-1} and $k_{obs} = 0.093 \pm 0.003$ s^{-1} (aerosols present, green circles and red triangles respectively) and 0.049 ± 0.003 s^{-1} (no aerosols, k_{wall}).

The flow tube was coated with halocarbon wax (Halocarbon Products Corporation, Halocarbon Wax Series 600) to reduce HO_2 wall loss rate (k_{wall}) and an average of k_{wall} from several determinations in the absence of aerosols was subtracted from k_{obs}. A correction (typically about 30%) to account for non-plug flow conditions in the aerosol flow tube using the procedure outlined by Brown (Brown, 1978) was applied to yield, k', the pseudo-first-order loss of HO_2 by heterogeneous reaction with TiO_2 particles. The relationship between k' and total surface area of TiO_2 particles (S_a) can be expressed as (George et al., 2013):

$$k' = \frac{w_{HO_2} \gamma_{obs} S_a}{4} \quad \text{(Eqn. 2)}$$

where w_{HO2} is the mean velocity of HO_2 (~ 435 ms^{-1} at 293 K) and γ_{obs} is the observed reactive uptake coefficient, obtained from a plot of k' versus S_a, an example of which is given in Figure 3. γ_{obs} was in turn corrected (typically ~1%) to account for the gas diffusion limitation (Fuchs and Sutugin, 1970), to yield $\gamma(HO_2)$.

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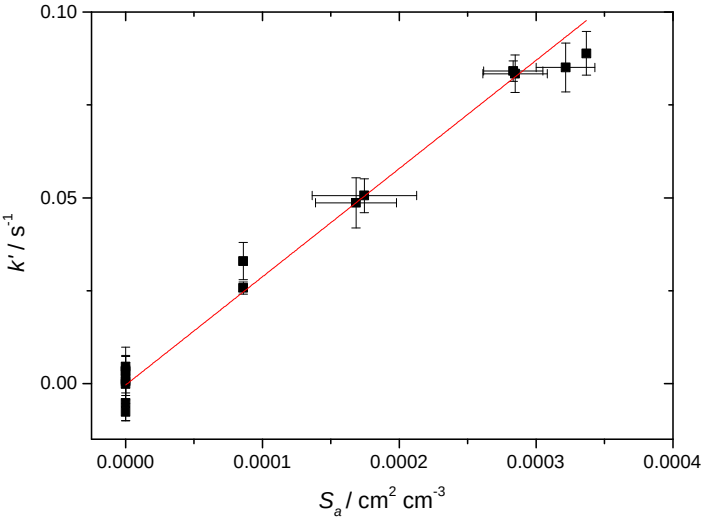

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197 **Figure 3**. First order rate coefficient for loss of HO_2 due to heterogeneous reaction with TiO_2 particles at different
198 total surface areas for T = 296 K and RH = 37%. The gradient yielded a value of $\gamma(HO_2) = (2.68 \pm 0.01) \times 10^{-2}$, with
199 the uncertainty representing 2σ random errors from the fit (95% confidence limits).

200 **2.5. TOMCAT Model Description**

201 The TOMCAT off-line three-dimensional (3D) chemical transport model (CTM) (Chipperfield, 1999, 2006) has been
202 used to predict the impact of the heterogeneous reaction of TiO_2 with HO_2 to stratospheric concentrations of O_3 and
203 HO_2. The model has been widely used in previous studies of stratospheric chemistry and performs well in reproducing
204 stratospheric ozone and the trace species which control is distribution (Chipperfield et al., 2015). The model includes
205 a detailed treatment of stratospheric chemistry of O_x, HO_x, NO_y, Cl_y and Br_y species along with the main source
206 gases. The model has a comprehensive gas-phase chemistry scheme and includes a number of heterogeneous
207 reactions on stratospheric sulphate aerosols and polar stratospheric clouds (Chipperfield, 1999).

208 The loss rate of HO_2 due to heterogeneous reaction with TiO_2 was included in the model as:

209 $$k = 0.25\, S_a\, w_{HO2}\, \gamma(HO_2) \qquad \text{(Eqn. 3)}$$

210 where S_a is the surface area density of TiO_2, and w_{HO2} and $\gamma(HO_2)$ are defined above. Three TOMCAT simulations
211 were performed at a horizontal resolution of 5.6° × 5.6° and 32 levels from the surface to ~ 60 km. The model was
212 forced with wind and temperature fields from the European Centre for Medium-Range Weather Forecasts (ECMWF)
213 ERA-Interim reanalyses and integrated for 2 years from January 2007 until December 2008, initialised with the
214 output from a standard TOMCAT run which had spun-up from 1977. More information on the model experiments is
215 given in Section 3.3.

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3. Results and Discussion

3.1. The reactive uptake coefficient dependence with relative humidity

The reactive uptake coefficient for HO_2 radicals, $\gamma(HO_2)$, onto TiO_2 particles was determined at eight different RH (11 - 66%), as shown in Figure 4 and summarised in Table 1. The number of monolayers of water adsorbed onto the surface of TiO_2 particles has been previously determined experimentally by transmission FTIR spectroscopy (Goodman et al., 2001), and is also shown as a function of RH in Figure 4.

Table 1. Reactive HO_2 uptake coefficients, $\gamma(HO_2)$, for TiO_2 particles at different RH.

RH (%) ± 1.0	$\gamma(HO_2) \times 10^{-2}$
11.1	2.08 ± 0.11
11.8	2.11 ± 0.13
24.9	2.48 ± 0.33
29.1	2.54 ± 0.18
37.5	2.68 ± 0.09
45.1	2.90 ± 0.53
45.3	2.92 ± 0.37
63.1	3.43 ± 0.27
66.4	3.65 ± 0.70

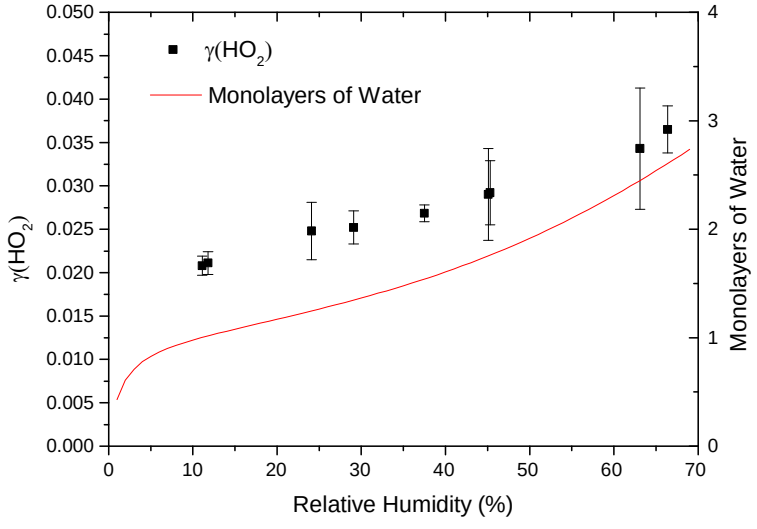

Figure 4. Reactive uptake coefficients of HO_2, $\gamma(HO_2)$, onto airborne TiO_2 particles (black squares, left *y* axis) at different RH for $T = 295 \pm 2$ K. The number of monolayers of the adsorbed water on TiO_2 particles (red curve, right *y* axis) at 296 K is also plotted as a function of RH, determined using FTIR spectroscopy (Goodman et al., 2001).

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The results clearly show a positive dependence of $\gamma(HO_2)$ across the range of RH investigated, and as shown in Figure 5, $\gamma(HO_2)$ correlates well with the number of monolayers of water adsorbed onto the TiO_2 particles, V/V_m, determined by Goodman *et al.* (2001), and which was parameterised by Eqn. 4:

$$\frac{V}{V_m} = \left[\frac{c\left(\frac{P}{P_0}\right)}{1-\left(\frac{P}{P_0}\right)}\right]\left[\frac{1-(n+1)\left(\frac{P}{P_0}\right)^n+n\left(\frac{P}{P_0}\right)^{n+1}}{1+(c-1)\left(\frac{P}{P_0}\right)-c\left(\frac{P}{P_0}\right)^{n+1}}\right] \quad \text{(Eqn. 4)}$$

where V is the volume of gas (water vapour) adsorbed at equilibrium pressure P, V_m is volume of gas necessary to cover the surface of the adsorbent TiO_2 particles with a complete monolayer, P is the equilibrium pressure of the adsorbing gas, P_0 is the saturation vapour pressure of the adsorbing gas at that temperature, c is a temperature-dependent constant related to the enthalpies of adsorption of the first and higher layers and n is the asymptotic limit of monolayers (~8).

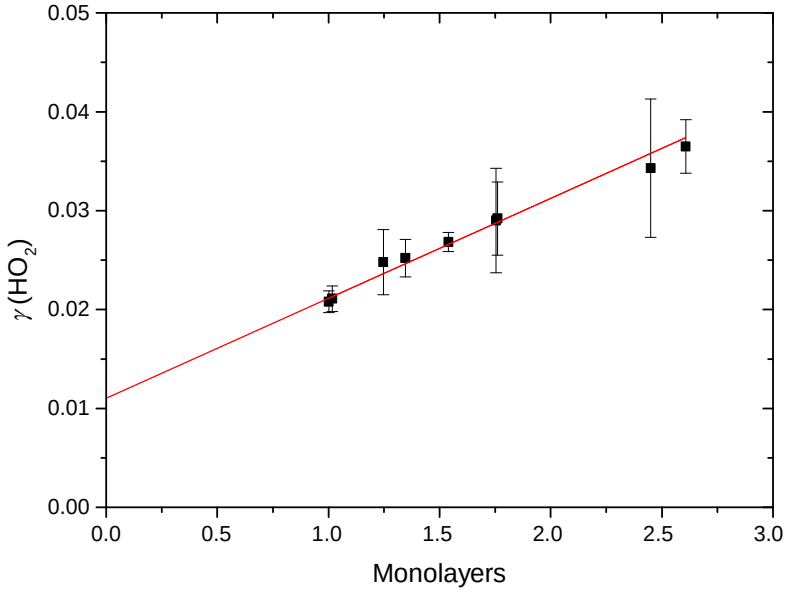

Figure 5. Variation of the reactive uptake coefficient, $\gamma(HO_2)$, with the number of monolayers of water surrounding TiO_2 particles (as determined by Goodman et al., 2001) for T = 295 ± 2 K. The red line represents a linear least-squares fit to the data (r^2 = 0.987).

Wall losses of HO_2 also increase as RH is increased within the aerosol flow tube, and so in the absence of aerosols, the $[HO_2]$ for a given distance from the injector will decrease with RH. Previous work in this laboratory has shown that $\gamma(HO_2)$ for uptake on ATD aerosols increases as the $[HO_2]$ is lowered (Matthews et al., 2014), and hence some of the positive dependence of $\gamma(HO_2)$ with RH shown in Figure 4 might be expected simply owing to the $[HO_2]$ impinging on the aerosol for a given injector position decreasing with RH. To investigate this further, uptake onto

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TiO_2 at RH = 11% was measured as a function of $[HO_2]_o$ exiting the injector, and $\gamma(HO_2)$ increased from 2.08×10^{-2} to 2.72×10^{-2} as $[HO_2]_0$ was decreased from 1.6×10^9 molecule cm^{-3} to 8.9×10^8 molecule cm^{-3}. However, the wall loss rate for HO_2 only increased from 0.049 to 0.079 s^{-1} across the entire range of RH (11% to 66%) resulting in only a small change in $[HO_2]$, decreases of ~ 2.6×10^8 molecule cm^{-3} and ~ 2.7×10^8 molecule cm^{-3} at the first and last injector position, respectively. In previous studies of HO_2 uptake onto ATD (Matthews et al., 2014) it was shown that the increase of $\gamma(HO_2)$ with decreasing $[HO_2]$ is linear. Therefore, assuming the same behaviour for uptake onto TiO_2 particles, the expected change in $\gamma(HO_2)$ as a result of RH increasing from 11-66% due only to a change in initial HO_2 concentration is only $\Delta\gamma(HO_2)$=0.0023 (or ~6% of the averaged measured $\gamma(HO_2)$ across this range). Hence the 175 % increase in $\gamma(HO_2)$ observed in Figure 4 across this range of RH is due to another reason.

Figure 5 shows that, for the range of RH studied, $\gamma(HO_2)$ is a linear function of the number of monolayers of water surrounding a TiO_2 particle, suggesting that water plays a role in the reactive HO_2 uptake process on TiO_2 particles. Studies using ambient pressure photoelectron spectroscopy (APPES) (Yamamoto et al., 2008; Ketteler et al., 2007) have shown that water adsorption on rutile, a polymorph of TiO_2 (110), occurs in distinct steps, as illustrated in Figure 6.

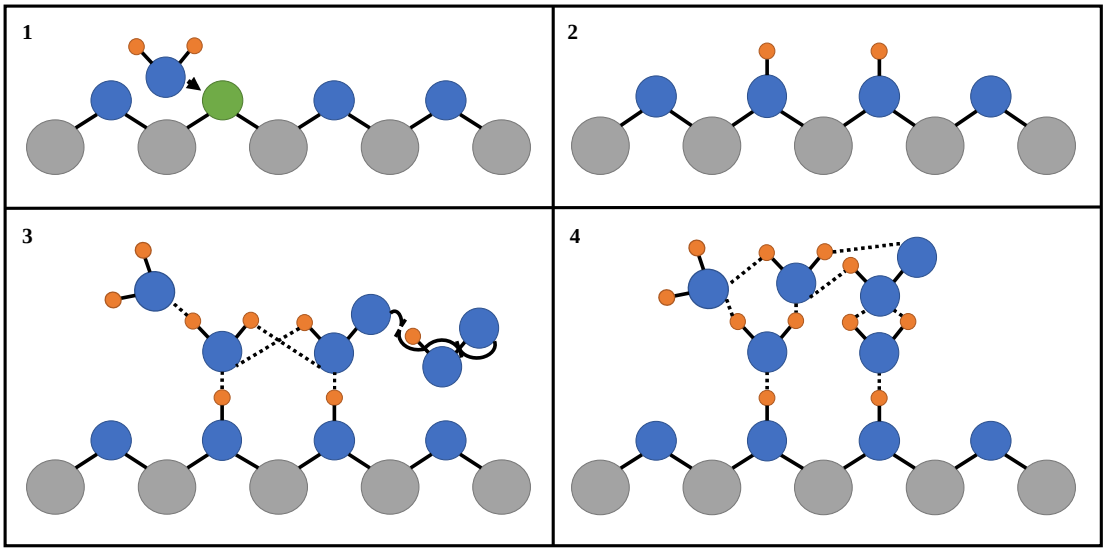

Figure 6. Simplified diagram of the important steps of HO_2 reactive uptake onto the surface of TiO_2. Grey circles = Ti, blue circles = O, orange circles = H, green circles = V_{bridge}, solid black lines = chemical bond and dashed lines = hydrogen bond. Panel 1 shows the diffusion of water molecule towards a bridging O vacancy. Panel 2 shows the resultant formation of two neighbouring bridging OH groups at the original site of the O vacancy. Panel 3 shows the bridging OH groups acting as anchoring sites for water and HO_2 adsorption via hydrogen bonding leading to multilayer water adsorption and HO_2 self-reaction via an Eley-Rideal type mechanism. Panel 4 shows the build up of a more extensive hydrogen bonded network as more water molecules adsorb onto the particle, which stabilises HO_2 increasing its desorption lifetime and hence probability that it will react.

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First, water molecules dissociate at O-vacancies (V_{bridge}) in bridge sites, producing a stoichiometric amount of adjacent bridging OH groups (OH_{bridge}) equal to twice the initial vacancy concentration upon initial exposure of rutile to water vapour:

$$H_2O + V_{bridge} + O_{bridge} = 2OH_{bridge} \quad (R7)$$

This step takes place even at very low RH, with these bridging OH groups acting as nucleation sites for subsequent water adsorption by anchoring water molecules to form strongly bound OH–H_2O complexes. The OH–H_2O complexes continue to act as nucleation centres for further water adsorption. The wetting properties of TiO_2 (110) are thus driven by moderate amounts of strongly attractive OH sites that nucleate water molecules. These OH-H_2O complexes have a relatively high enthalpy of adsorption for water of -72 kJ mol^{-1} (Ketteler et al., 2007), whereas additional adsorption of water beyond a monolayer of water coverage is more characterised by enthalpies associated with the bulk enthalpy of water condensation (-45 kJ mol^{-1}) (Chen et al., 2012), explaining the variation of the number of monolayers of water with RH shown in Figure 4. Computational studies (Aloisio and Francisco, 1998) showed that in the gas-phase HO_2 can also readily form complexes with water through hydrogen bonding with a binding energy of 28.9 kJ mol^{-1}. The observed correlation of $\gamma(HO_2)$ with the number of monolayers of water surrounding TiO_2 particles could be explained by two effects. An increase in the network of hydrogen bonding would increase the stability of a molecular system, ~ 20.9 kJ mol^{-1} for each hydrogen bond (Joshi and Ghanty, 2013), or that simply more HO_2 can adsorb onto the particle surface as the number of available sites for hydrogen bonding increases. An increase in the adsorption lifetime of HO_2 owing to the more extended H-bonding network (Joshi and Ghanty, 2013) will result in an increased probability of HO_2 reacting with another HO_2^-, increasing the value of $\gamma(HO_2)$:

$$HO_2 + HO_2 \rightarrow H_2O_2 + O_2 \quad (R8)$$

For up to ~2 monolayers of H_2O, an Eley-Rideal (ER) mechanism of a gas phase HO_2 molecule reacting an adsorbed HO_2 is more likely than for two adsorbed HO_2 molecules diffusing together to react via a Langmuir-Hinshelwood (LH) mechanism. The coverage of V_{bridge} across the surface of TiO_2 (110) is 0.125 monolayers (i.e. one vacancy for every eight unit cells) (Ketteler et al., 2007) meaning water will initially adsorb onto the surface of the particle in clusters. As the binding energy of the HO_2 water complex is fairly high, the rate of HO_2 diffusion across the surface after the initial adsorption at a bridging site will be slow, making a LH type mechanism seem unlikely. Above ~2 monolayers, the adsorbed water clusters will begin to interact with each other via an extended H-bonded network more characteristic of bulk liquid water, allowing HO_2 to diffuse around the thin surface film of water with thermodynamic properties similar to liquid water (Ketteler et al., 2007).

The superoxide ion, O_2^-, which is the conjugate base of HO_2, reacts with HO_2 over a hundred times quicker than with HO_2 via:

$$HO_2 + O_2^- + H_2O \rightarrow H_2O_2 + O_2 + OH^- \quad (R9)$$

It is unlikely that O_2^- will form readily on the surface of the particle at low coverages of water. O_2^- can form on surfaces via direct surface–oxygen electron transfer, photo-induced electron transfer, surface intermolecular electron

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transfer or decomposition of hydrogen peroxide (H_2O_2) (Anpo et al., 1999). The dissociation of HO_2 to O_2^- occurs in bulk liquid (pK_a = 4.7, (Thornton and Abbatt, 2005)), however, water surrounding the particles only begins to acquire liquid like properties once the coverage of water is greater than ~2 monolayers. If significant formation of O_2^- does occur above ~2 monolayers of H_2O, then owing to its much higher reactivity, some deviation from linearity of $\gamma(HO_2)$ versus monolayers of water (Figure 5) might be expected. However, this was not observed, although the number of data points is very limited.

3.2. Comparison of $\gamma(HO_2)$ with literature values

There are previous studies of HO_2 uptake onto ATD (Matthews et al., 2014), a proxy for mineral dust, and both N_2O_5 (Tang et al., 2014) and $ClONO_2$ (Tang et al., 2016) uptake onto TiO_2 particles. Values of $\gamma(HO_2)$ measured for ATD are comparable to those measured onto TiO_2 particles at similar concentrations of HO_2 in the aerosol flow tube ($\gamma(HO_2)_{ATD}$ ~ 0.008 – 0.030). For ATD, $\gamma(HO_2)$ was only determined at 4 values of RH, but a general increase with RH was observed (although there was a dip around RH = 50% before a further increase), consistent with HO_2 uptake being driven by the number of defects in the crystal lattice of mineral dust surfaces which provides bridging OH groups upon exposure to water and subsequently by the number of monolayers of water adsorbed onto the surface of such aerosols. The number of monolayers of water on the ATD surface at different RH has been determined (Gustafsson et al., 2005), and also showed a general increase with RH but with a shoulder around RH = 50% where the observed $\gamma(HO_2)$ also contained a small dip.

Values of $\gamma(N_2O_5)$ (Tang et al., 2014a; and corrigendum Tang et al. 2014b) and $\gamma(ClONO_2)$ (Tang et al., 2016) onto TiO_2 particles were an order of magnitude lower than $\gamma(HO_2)$ and may be associated with the lower polarity of these molecules compared to HO_2. It is likely that, as with uptake of HO_2, both N_2O_5 and $ClONO_2$ react with TiO_2 via complexing with bridging OH groups and adsorbed H_2O, therefore less polar molecules will be less bound to the surface and more likely to desorb back into the gas phase, leading to a smaller γ. The dependence of $\gamma(N_2O_5)$ for TiO_2 with RH is also different to that for HO_2 observed here, with a small decrease of $\gamma(N_2O_5)$ observed as the RH is increased from ~ 5 – 23% where a minimum is reached, then beyond 23% $\gamma(N_2O_5)$ increases as RH is increased. (Tang et al., 2014). Competition between water and N_2O_5 for surface active OH groups was suggested for the initial observed decrease in $\gamma(N_2O_5)$ with RH, whereas for RH > ~ 23% heterogeneous hydrolysis of N_2O_5 to form HNO_3 starts to drive reactive uptake resulting in an increase of $\gamma(N_2O_5)$ (Tang et al., 2014). Measurements of $\gamma(ClONO_2)$ onto TiO_2 particles were only made at two values of RH, insufficient to determine any systematic dependence, although Tang *et al.* (2016) expected $\gamma(ClONO_2)$ to increase as more water adsorbs onto the surface of TiO_2 particles.

George et al. (2013) previously measured $\gamma(HO_2)$ onto dry salt aerosols. Values of $\gamma(HO_2)$ for NaCl and $(NH_4)_2SO_4$ at RH values (33 – 54%) below their deliquescence point were below the limit of detection ($\gamma(HO_2)$ <0.004). Values of $\gamma(HO_2)$ for TiO_2 reported here (Figure 4 and Table 1) are more than an order of magnitude greater than that of solid salts, even for RH = 11%. A possible explanation for the difference in $\gamma(HO_2)$ values could be that even though the sizes of the aerosols determined by the SMPS are similar, dry salt aerosols are more spherical in shape than TiO_2 particles which may be more fractal in nature. As the SMPS indirectly measures the surface area of aerosols by

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measuring their mobility through an electric field, an assumption that the aerosol is spherical has to be made. If this is not the case, this may lead to a significant under-prediction of the surface area of non-spherical aerosols and therefore an over-prediction of $\gamma(HO_2)$. In order to measure the geometry of the TiO_2 particles, a scanning electron microscope (SEM, FEI Nova NanoSEM 450) operating at 3 kV was used to image the TiO_2 nanoparticles used within these experiments. Samples for the SEM were prepared by dispersing the nanoparticles in ethanol and allowing a drop of this solution to dry on a silicon wafer. The wafer with nanoparticles was then mounted on SEM stubs using conductive carbon tapes and coated with 2 nm of Ir, using a Cressington 208HR high resolution sputter coater. Figure 7 shows an example of an SEM image providing evidence that the TiO_2 particles are spherical and therefore any error associated with SMPS measurements of surface area is minimal. A more likely explanation for the higher $\gamma(HO_2)$ for TiO_2 particles is that dry salt aerosols do not adsorb a significant amount of water onto their surface until the deliquescence point reached, whereas at RH = 11% Figure 4 shows that the TiO_2 particles already have a monolayer coverage which can form relatively strongly bound complexes with HO_2. Measurements of $\gamma(HO_2)$ onto aqueous salt aerosols show that $\gamma(HO_2)$ significantly increases above the deliquescence point (George et al., 2013) and is comparable to $\gamma(HO_2)$ measured for TiO_2 ($\gamma(HO_2) \sim 0.01$).

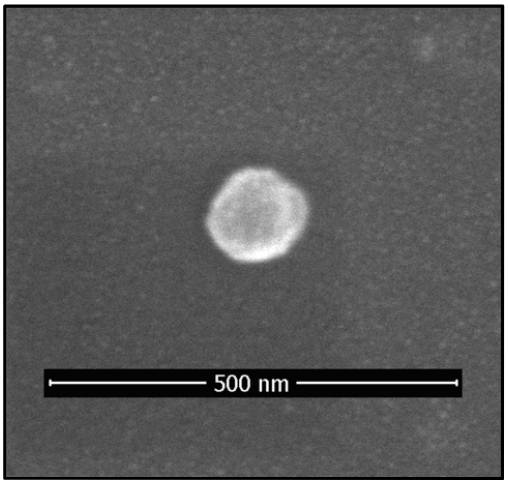

Figure7. SEM image of a single TiO_2 particle magnified 80,000× used within this study.

3.3 Implications of HO_2 uptake by TiO_2 for stratospheric chemistry

The effect of HO_2 uptake onto TiO_2 particles upon the stratospheric concentrations of HO_2 and O_3 was assessed using the TOMCAT model (Chipperfield, 1999, 2006). At RH relevant to the lower stratosphere (< 40%) the measurements showed that $\gamma(HO_2)$ is in the range 0.020 - 0.028 at 295 K. An inverse temperature dependence of $\gamma(HO_2)$ onto dry sea-salt aerosols has previously been observed (Remorov et al., 2002), and although there have been no systematic experimental studies of the temperature dependence of $\gamma(HO_2)$, parameterisations have developed (Thornton et al., 2008;Macintyre and Evans, 2011). At stratospherically relevant temperatures (T = 200 - 220 K), $\gamma(HO_2)$ is likely to be considerably larger than observed at 295 K, however it is not possible to cool the aerosol flow tube/SMPS system to verify this experimentally. Therefore $\gamma(HO_2)$ =1 was used in the model simulations representing an upper limit,

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with three TOMCAT simulations performed as follows. A control simulation, similar to that presented in Chipperfield et al. (2015) did not include TiO_2 particles. A specified latitude-height distribution of TiO_2 particles was then included in two simulations with an effective aerosol surface area density equal to that of sulphate aerosols in 1992 the year after the eruption of Mt. Pinatubo. This is an assumption which allows for the fact that less TiO_2 mass is needed in order to produce the same radiative impact as sulphate aerosol from Mt. Pinatubo, but the TiO_2 particle size is smaller. Hence these effects largely cancel (Tang et al., 2014). Stratospheric injection via a geo-engineering solution will result initially in a different distribution of TiO_2 particles compared with after the Mt. Pinatubo eruption, but it is assumed that following mixing and transport the distributions would resemble one another and not lead to any significant difference in model behaviour and the conclusions drawn. Support from this assumption comes from a model run in which a globally uniform distribution of TiO_2 was assumed initially and which yielded very similar results.

The first of these simulations included only the loss of N_2O_5 on TiO_2 particles with $\gamma(N_2O_5) = 0.005$, the upper limit used in the modelling of Tang et al. (2014), which allows us to compare our results with their study. The second TiO_2 simulation also included the loss of HO_2 on TiO_2 particles with a $\gamma(HO_2) = 1$, as discussed above. Figure 8 shows the impact of including heterogeneous loss of N_2O_5 on TiO_2 particles in the model.

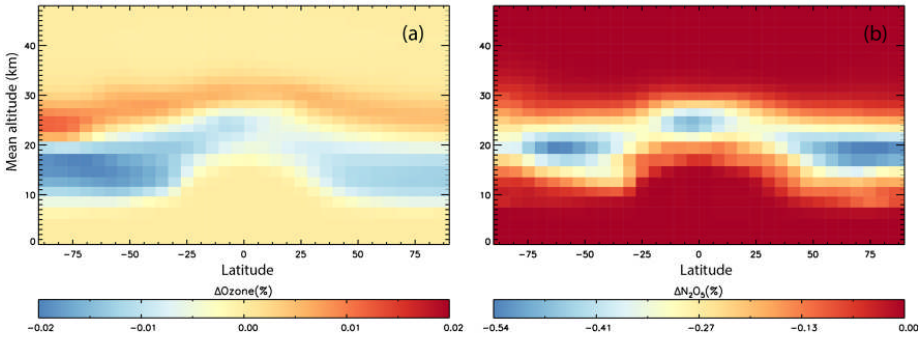

Figure 8. Annual mean change (%) in atmospheric O_3 (a) and N_2O_5 (b) calculated using the TOMCAT 3-D model from inclusion of heterogeneous loss of N_2O_5 on TiO_2 particles for 2008 and using $\gamma(N_2O_5) = 0.005$.

N_2O_5 is decreased by up to 0.5% in the region of TiO_2 particles, which is assumed to follow the distribution of sulphate particles after the Mt. Pinatubo eruption, in order follow the approach of previous modelling studies. Assuming a globally uniform distribution initially yielded results that were very similar. Inclusion of uptake is only a minor effect and considerably smaller than the impact of around -20% modelled by Tang et al. (2004) for the same assumed $\gamma(N_2O_5) = 0.005$. The reasons for this are not clear, although it is noted that the effect modelled in our off-line chemical transport model, with specified meteorology, is clearly confined to regions of high aerosol loading.

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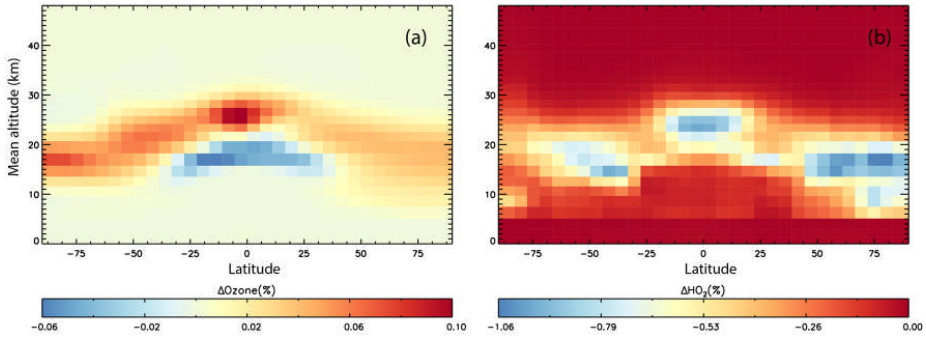
393 The impacts modelled in the nudged chemistry-climate model study of Tang et al. (2004) are not confined to the
394 region of high aerosol and even extend to the upper stratosphere. It is possible that their simulations, although nudged,
395 also include some dynamical feedback which enhances an otherwise small signal. Figure 8 also shows that the
396 resulting impact on O_3 is small with changes less than 0.02%. The model produces a region of slight decrease in the
397 very low stratosphere, with a region of slight increase above.

398 Figure 9 shows results from the simulation which also included the loss of HO_2 on stratospheric TiO_2 particles, and
399 using $\gamma(HO_2) = 1$ (a simulation using the measured value of $\gamma(HO_2) = 0.021$ at 295 K led to no impact on O_3 or HO_2).

400



401

402 **Figure 9**. Annual mean change (%) in atmospheric O_3 (a) and HO_2 (b) calculated using the TOMCAT 3-D model
403 from inclusion of heterogeneous loss of HO_2 on TiO_2 particles for 2008, and using $\gamma(HO_2) = 1$.

404

405 It is evident that HO_2 loss due to heterogeneous reaction between HO_2 and TiO_2 particles in 2008 is < 1% and is
406 confined to the lower stratosphere where it is assumed TiO_2 particles are located. Figure 9 also shows that the
407 subsequent effect of the TiO_2 particles on the O_3 concentrations through the effects of this reaction is also small
408 (< 0.1%), with a small decrease in the tropical upper troposphere/lower stratosphere and a small increase at all
409 latitudes in the lower stratosphere. This small effect of TiO_2 particles on stratospheric HO_2 and O_3 concentrations is
410 due to the reactive nature and short lifetime of HO_2. The species readily reacts with other gas phase species (e.g. O_3)
411 and so loss on TiO_2 surfaces does not compete significantly.

412

413 **4. Conclusions and further work**

414 The reactive uptake coefficients for the heterogeneous reaction of HO_2 onto TiO_2 particles were measured at different
415 RH and at room temperature for the first time using an aerosol flow tube reactor coupled with a sensitive FAGE HO_2
416 detection system. A range of HO_2 uptake coefficients on TiO_2 particles were measured varying from $\gamma(HO_2) = 0.021$
417 ± 0.001 to 0.036 ± 0.007 for RH over the range 11% to 66%, respectively. The HO_2 uptake coefficient showed a

418 positive dependence on RH which correlated well with the number of monolayers of water adsorbed onto the TiO_2
419 particle. These results suggest a mechanism by which HO_2 adsorbs to the surface if the TiO_2 particle by forming
420 complexes with water molecules bound to bridging OH groups. As the number of water layers increases so does the
421 network of hydrogen bonds that stabilises HO_2 leading to a longer adsorption lifetime and increased $\gamma(HO_2)$. The
422 TOMCAT chemical transport model was used to evaluate the possible effects of HO_2 uptake (using an upper limit of
423 $\gamma(HO_2) = 1$) onto the surface of TiO_2 particles on the stratospheric concentrations of HO_2 and O_3. The amount of
424 TiO_2 used was chosen to achieve a similar cooling to that following the Mt. Pinatubo eruption, but the model predicted
425 a very small loss of both stratospheric HO_2 and O_3. TiO_2 possesses photocatalytic properties and water adsorbed
426 onto its surface may dissociate under stratospheric illumination providing a source of radicals (Chen et al.,
427 2012;Romanias et al., 2012). Production of OH and HO_2 from irradiated TiO_2 surfaces should be evaluated in future
428 studies to fully understand the consequences of injection of TiO_2 particles into the stratosphere.

429

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436

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