



Cite this: *Phys. Chem. Chem. Phys.*, 2023, 25, 4357

Reply to the 'Comment on "Theoretical study of the NO₃ radical reaction with CH₂ClBr, CH₂ICl, CH₂BrI, CHCl₂Br, and CHClBr₂"' by C. J. Nielsen and Y. Tang, *Phys. Chem. Chem. Phys.*, 2022, 24, DOI: 10.1039/D2CP03013F

Ibon Alkorta,^a John M. C. Plane,^b José Elguero,^a Juan Z. Dávalos,^c A. Ulises Acuña^c and Alfonso Saiz-Lopez^{*c}

In this Reply, we answer the main argument raised in the Comment about the energy of the NO₃ radical and its influence in the reaction profiles of the reaction of the NO₃ radical with CH₂ClBr, CH₂ICl, CH₂BrI, CHCl₂Br, and CHClBr₂ by C. J. Nielsen and Y. Tang. The optimized geometry of the NO₃ radical has been obtained using 49 DFT functionals: 26 functionals predict a minimum with D_{3h} symmetry and 23 with C_{2v} symmetry. The former functionals have been used to calculate the thermodynamic values of three reactions (X + HNO₃ → XH + NO₃, X = OH, CH₃ and CCl₃) and compared with experimental data. Those functionals with smaller errors have been used to recalculate the barriers of the reaction of NO₃ with CH₂ClBr, CH₂ICl, CH₂BrI, CHCl₂Br, and CHClBr₂. The results show differences of 10.5 kJ mol⁻¹ when compared to those obtained with the M08HX functional.

Received 13th October 2022,
Accepted 29th November 2022

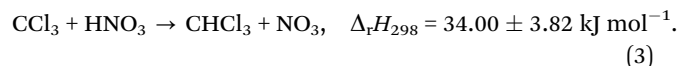
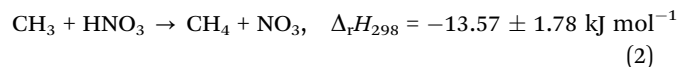
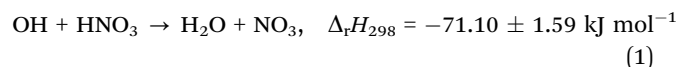
DOI: 10.1039/d2cp04659h

rsc.li/pccp

The molecular symmetry of the ground state of the NO₃ radical has been difficult to determine by experiment¹ and theoretical calculation.^{2–4} However, the latest experimental results indicate that it has D_{3h} symmetry⁵ with an NO distance of 1.240 Å (the bond distance of 1.238 Å mentioned by Nielsen and Tang does not appear in Kawaguchi *et al.*⁷). Our calculations with the M08-HX DFT functional provide a C_{2v} symmetry minimum, while the D_{3h} geometry corresponds to a second-order saddle point with a relative energy of 9.6 kJ mol⁻¹.⁸

We have examined the performance of 49 functionals (using in all cases the 6-311+G(2df,2p) basis set): for 26⁹ of them the D_{3h} geometry is predicted to be the minimum, while for 23¹⁰ the D_{3h} geometry exhibits two degenerate imaginary frequencies. In order to check the reliability for studying proton transfer reactions of those DFT functionals that predict NO₃ to be a minimum with D_{3h} symmetry, we examined the enthalpy of three reactions for which experimental data is available (eqn (1)–(3)). The first reaction was already proposed in the Comment by Nielsen and Tang, but we think that the second and third are probably more appropriate in the present

case since they involve CH groups as the hydrogen donor, as in the reactions studied in our paper.⁸ The heats of formation of all the molecules were obtained from the NIST database,¹¹ apart from the NO₃ (ΔH_f⁰ = 73.72 ± 1.38 kJ mol⁻¹)¹² and OH (ΔH_f⁰ = 37.3 ± 0.7 kJ mol⁻¹)¹³ radicals.



The average calculated errors of the energies obtained in these three reactions, for the 26 functionals under consideration, range between +19 and -69 kJ mol⁻¹. Only five functionals show an average unsigned error below 8 kJ mol⁻¹ (approx. 2 kcal mol⁻¹): B3PW91 (3.8 kJ mol⁻¹), X3LYP (4.6 kJ mol⁻¹), B3LYP (5.1 kJ mol⁻¹), and B971 and B972 (8.0 kJ mol⁻¹).

We therefore used these five functionals to recalculate the proton transfer barrier and the corresponding relative energy of the exit channel, for the five reactions in our study. The average values are listed in Table 1, together with the energies computed using the M08HX functional. The TS of the proton transfer computed at M08HX is between 8.8 and 11.9 kJ mol⁻¹ (average 10.5 kJ mol⁻¹) lower than the average of the five

^a Instituto de Química Médica (CSIC), Juan de la Cierva, 3, E-28006 Madrid, Spain. E-mail: ibon@iqm.csic.es

^b School of Chemistry, University of Leeds, LS2 9TJ Leeds, UK

^c Department of Atmospheric Chemistry and Climate, Institute of Physical Chemistry Rocasolano, CSIC, Madrid E-28006, Spain. E-mail: a.saiz@csic.es



Table 1 Proton transfer barrier and relative energy (kJ mol^{-1}) of the exit channel at M08HX level and average of the selected five DFT functionals, in parenthesis

	$\text{CH}_2\text{ClBr:NO}_3$	$\text{CH}_2\text{ClI:NO}_3$	$\text{CH}_2\text{BrI:NO}_3$	$\text{CHCl}_2\text{Br:NO}_3$	$\text{CHClBr}_2:NO_3$
TS-H transfer	16.7 (28.1 \pm 2.3)	14.7 (23.5 \pm 2.6)	13.3 (22.4 \pm 2.7)	13.7 (25.6 \pm 2.6)	11.7 (23.2 \pm 2.7)
Exit channel	-58.2 (-20.4 \pm 1.6)	-57.1 (-21.9 \pm 1.4)	-53.6 (-19.8 \pm 1.2)	-74.0 (-38.9 \pm 1.6)	-73.4 (-21.9 \pm 1.4)

selected DFT methods. The effect is larger in the evaluation of the proton transfer exit channel, where the differences in the overall reaction energy predicted by M08HX are 37.8, 35.2, 35.1, and 51.5 kJ mol^{-1} for the four reactions (with three of these giving errors that closely match the 37 kJ mol^{-1} difference anticipated in the Comment).

These results partially agree with the comments of Nielsen and Tang, although the difference in the barrier heights between M08HX and the five chosen DFT methods is significantly smaller than the $\sim 37 \text{ kJ mol}^{-1}$ that they anticipated. Importantly, the main conclusion of our paper is unchanged: the oxidations of CH_2ClBr , CH_2ClI , CH_2BrI , CHCl_2Br , and CHClBr_2 by NO_3 are not competitive with other removal processes.

Conflicts of interest

The authors declare that they have no conflicts of interest.

Acknowledgements

This work was carried out with financial support from the Ministerio de Ciencia, Innovación y Universidades (PGC2018-094644-B-C22) and Comunidad de Madrid (P2018/EMT-4329 AIRTEC-CM). Thanks are also given to the CTI (CSIC) for their continued computational support. This study received funding from the European Research Council Executive Agency under the European Union's Horizon 2020 Research and Innovation Programme (Project ERC-2016-COG 726349 CLIMAHAL).

References

- J. F. Stanton, On the vibronic level structure in the NO_3 radical. I. The ground electronic state, *J. Chem. Phys.*, 2007, **126**(13), 134309.
- U. Kaldor, The ground state geometry of the NO_3 radical, *Chem. Phys. Lett.*, 1990, **166**(5), 599–601.
- J. F. Stanton, J. Gauss and R. J. Bartlett, Potential nonrigidity of the NO_3 radical, *J. Chem. Phys.*, 1991, **94**(5), 4084–4087.
- W. Eisfeld and K. Morokuma, A detailed study on the symmetry breaking and its effect on the potential surface of NO_3 , *J. Chem. Phys.*, 2000, **113**(14), 5587–5597.
- K. Kawaguchi, R. Fujimori, J. Tang and T. Ishiwata, Infrared spectroscopy of difference and combination bands of the NO_3 radical and anharmonicity analysis, *J. Mol. Spectrosc.*, 2022, **385**, 111594.
- T. Ishiwata, I. Tanaka, K. Kawaguchi and E. Hirota, Infrared diode laser spectroscopy of the $\text{NO}_3 \nu_3$ band, *J. Chem. Phys.*, 1985, **82**(5), 2196–2205.
- K. Kawaguchi, T. Ishiwata, E. Hirota and I. Tanaka, Infrared spectroscopy of the NO_3 radical, *Chem. Phys.*, 1998, **231**(2), 193–198.
- I. Alkorta, J. M. C. Plane, J. Elguero, J. Z. Dávalos, A. U. Acuña and A. Saiz-Lopez, Theoretical study of the NO_3 radical reaction with CH_2ClBr , CH_2ClI , CH_2BrI , CHCl_2Br , and CHClBr_2 , *Phys. Chem. Chem. Phys.*, 2022, **24**(23), 14365–14374.
- The 26 DFT functionals that predict NO_3 radical have D_{3h} geometry in the minimum are: B3LYP, B3PW91, B971, B972, B97D3, B97D, BLYP, BP86, BPW91, BVP86, LSDA, M06-L, M11L, MPWB95, MPWLYP, MPWPW91, N12, O3LYP, OLYP, OPBE, PBEPBE, PW91PW91, SVWN5, TPSSH, TPSSTPSS and X3LYP.
- The 23 DFT functionals that show two imaginary frequencies for the NO_3 radical with D_{3h} geometry are: APFD, BMK, CAM-B3LYP, HSEH1PBE, LC-BLYP, M05-2X, M05, M06-2X, M06-HF, M06, M08-HX, M11, MN15, MPW1LYP, MPW1PW91, N12SX, PBE1PBE, PW6B95D3, PW6B95, SOGGA11X, wB97, wB97XD, and wB97X.
- NIST Chemistry WebBook*, NIST Standard Reference Database Number 69, DOI: [10.18434/T4D303](https://doi.org/10.18434/T4D303).
- H. F. Davis, B. Kim, H. S. Johnston and Y. T. Lee, Dissociation energy and photochemistry of nitrogen trioxide, *J. Phys. Chem.*, 1993, **97**(10), 2172–2180.
- J. T. Herbon, R. K. Hanson, D. M. Golden and C. T. Bowman, A shock tube study of the enthalpy of formation of OH, *Proc. Combust. Inst.*, 2002, **29**, 1201–1208.

