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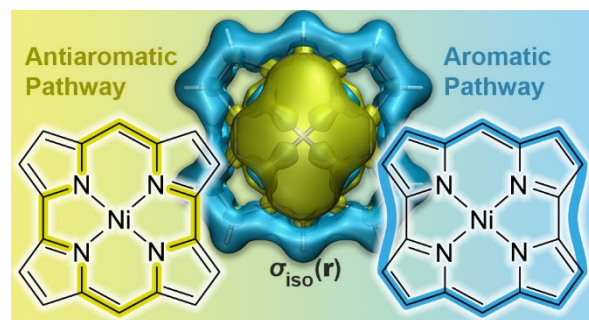
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Norcorrole: Aromaticity and Antiaromaticity in Contest

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ABSTRACT: Magnetic shielding studies demonstrate that nickel norcorrole (NiNc) and norcorrole (H₂Nc) provide unusual examples of stable molecules with high antagonistic levels of antiaromaticity and aromaticity: Both incorporate an antiaromatic “core”, a 14-membered cyclic conjugated subsystem with 16 π electrons, surrounded by an aromatic “halo” in the form of a ring of either 14 atoms and 14 π electrons with a new type of homoconjugation (NiNc), or 18 atoms with 18 π electrons (H₂Nc).

Aromaticity and antiaromaticity, two essential concepts deeply ingrained in chemical thinking, are usually associated with Hückel’s familiar $4n + 2$ and $4n$ rules. Having both aromatic and antiaromatic behavior within the same molecule is uncommon; examples are limited to fused rings with $4n + 2$ and $4n$ π electrons such as benzocyclobutadiene¹⁻⁵ and biphenylene.^{3,4,6,7} As a rule, fused $4n + 2$ and $4n$ rings both distort, leading to significant decreases of aromatic and antiaromatic character. The analysis of the isotropic magnetic shielding^{8,9} enveloping Ni^{II} norcorrole (NiNc) and norcorrole (H₂Nc) shows that these molecules provide unusual examples of conjugated systems with high antagonistic levels of antiaromaticity and aromaticity. Norcorrole, the smallest cyclic tetrapyrrole porphyrin analogue synthesized to date,^{10,11} was initially proposed on the basis of DFT calculations.¹² The contraction of porphyrin through removal of one meso-methine unit results in corrole; further contraction, through removal of a second meso-methine unit, produces norcorrole (H₂Nc). Although NiNc has been found to be reasonably stable,¹¹ both experiment and theory indicate that this complex has certain antiaromatic features. Experimental evidence of antiaromatic behavior comes from the NiNc ¹H NMR and MCD spectra;¹¹ on the theory side, antiaromaticity is suggested by DFT-level NICS, HOMA values, ACID plots, out-of-plane bond magnetizabilities and magnetically induced current densities.^{11,13-16} NiNc units have been used in the construction of “an antiaromatic-walled nanospace”.¹⁷ In contrast to other antiaromatic species, such as the classical example of antiaromaticity, cyclobutadiene, C₄H₄, NiNc does not undergo a 2nd-order Jahn-Teller transition lowering the energy, reducing the symmetry and introducing bond length alternation (in the case of C₄H₄, from square D_{4h} to rectangular D_{2h}). Instead, DFT results¹⁵ show that the highest-symmetry geometry of NiNc, planar D_{2h}, corresponds to the transition state (TS) for a bowl-to-bowl inversion connecting two equivalent shallow bowl-shaped geome-

tries of C_{2v} symmetry with not much in the way of bond length alternation, which are just 0.7 kcal mol⁻¹ lower in energy. The reason why NiNc behaves so differently from other antiaromatic molecules has been attributed to the high stability of the dipyrin fragments, assumed to retain their electronic and structural integrity within the norcorrole ring system.¹⁵ While this model accounts for some of the features of bonding in norcorrole, it does not explain what holds the two dipyrin fragments together and why the C _{α} -C _{α} bonds between these fragments are long and readily oxidized.¹¹ The variations in isotropic magnetic shielding around NiNc and H₂Nc shows that bonding in these systems is even more unusual and intriguing than previously thought and involves an antiaromatic “core” surrounded by an aromatic outer ring which, in NiNc, features a new type of homoconjugation over the C _{α} -C _{α} bonds.

The isotropic magnetic shielding, $\sigma_{\text{iso}}(\mathbf{r}) = \frac{1}{3}[\sigma_{xx}(\mathbf{r}) + \sigma_{yy}(\mathbf{r}) + \sigma_{zz}(\mathbf{r})]$, can be calculated at any point \mathbf{r} in the space surrounding a molecule; differences between $\sigma_{\text{iso}}(\mathbf{r})$ values at nuclei correspond to experimentally measurable NMR shifts. 3D grids of $\sigma_{\text{iso}}(\mathbf{r})$ values computed at the B3LYP-GIAO/6-311++G(d,p) level were employed to construct NiNc and H₂Nc $\sigma_{\text{iso}}(\mathbf{r})$ isosurfaces (Figure 1). For the purposes of this study, the ground state potential energy surfaces (PES) of NiNc and H₂Nc in the gas phase were investigated at the B3LYP-D3BJ/def2-TZVP level. In addition to the C_{2v} local minimum and planar D_{2h} TS geometries of NiNc, and the C_i local minimum geometry of H₂Nc, for which the current results agree very well with those reported previously,¹⁵ two further stationary points on the H₂Nc PES were located, a bowl-shaped C₂ local minimum and a planar C_{2h} 2nd-order saddle point; these last have not been reported in the literature. The C_{2h} 2nd-order saddle point acts as TS for both the bowl-to-bowl inversion between two equivalent C₂ local minima and the flip between two equivalent C_i local minima. The C₂ and C_i geome-

tries of H₂Nc were found to be 2.10 and 2.04 kcal mol⁻¹, respectively, lower in energy than the C_{2h} 2nd-order saddle point. The H₂Nc PES analysis was repeated with a second DFT method, PBE0-D3BJ/def2-TZVP. With this method, the C₂ and C_i geometries of H₂Nc were found to be 1.37 and 1.57 kcal mol⁻¹, respectively, lower in energy than the C_{2h} 2nd-order saddle point. These small energy differences suggest that the C₂ and C_i geometries of H₂Nc can easily interconvert in the gas phase through the C_{2h} 2nd-order saddle point; thorough examination of the H₂Nc PES indicates that there is no TS linking these geometries directly. The flatter C_i geometry is preferred in the solid state and in solution.¹⁴ A better estimate of the experimentally observed Ni–C distance in NiNc (1.771 Å to 1.789 Å¹⁵) was obtained at the TPSSh-D3BJ/def2-TZVP level, 1.796 Å against 1.810 Å at the B3LYP-D3BJ/def2-TZVP level.

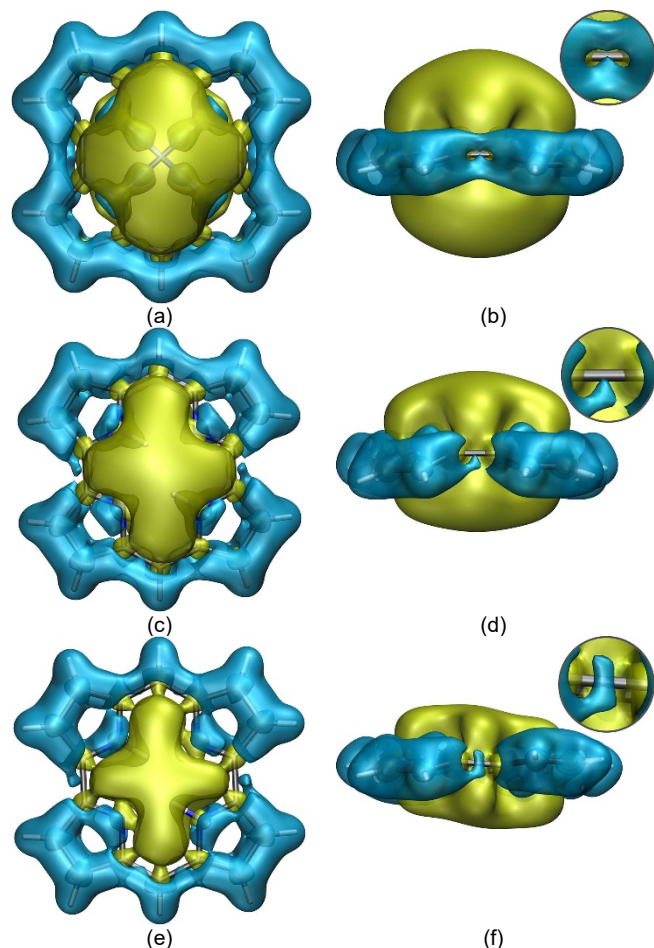


Figure 1. Shielding around NiNc and H₂Nc. Isosurfaces at $\sigma_{\text{iso}}(\mathbf{r}) = +12$ ppm (aromatic regions, blue) and $\sigma_{\text{iso}}(\mathbf{r}) = -12$ ppm (antiaromatic regions, yellow). (a, b) NiNc top and side views; H₂Nc: C₂ (c, d) and C_i (e, f) geometries, top and side views. Magnified sections in (b, d, f) show C_α–C_α bonds.

Before discussing the NiNc and H₂Nc $\sigma_{\text{iso}}(\mathbf{r})$ isosurfaces (Figure 1), it is instructive to examine shielding around benzene and cyclobutadiene (C₆H₆ and C₄H₄, Figure 2, at D_{6h} and D_{2h} optimized geometries, same levels of theory as for NiNc and H₂Nc). The benzene ring is enclosed within a thick shielded “doughnut” which demonstrates strong bonding interactions and aromatic stability. Antiaromatic destabilization in C₄H₄ follows from the central deshielded region which eliminates most of the shielding over C–C “single” bonds and pushes “double” bonds towards the exterior of the ring, weakening both types of bond, albeit to somewhat different extents. While the isosurfaces in Figure 2 are qualitatively similar to those obtained previously at other levels of theory,^{8,9} it should

be noted that rectangular (D_{2h}) C₄H₄ is considerably less antiaromatic than square (D_{4h}) C₄H₄.^{18,19} As shown previously,^{20,21} $\sigma_{\text{iso}}(\mathbf{r})$ is capable of exposing the differences between bonds of different strengths in much greater detail than is achievable using the total electronic density. The isovalues $\sigma_{\text{iso}}(\mathbf{r}) = +12$ ppm were chosen so as to show optimal levels of detail in Figures 1 and 2. Isosurfaces for other isovalues can be inspected using the Gaussian cube files provided in the Supporting Information.

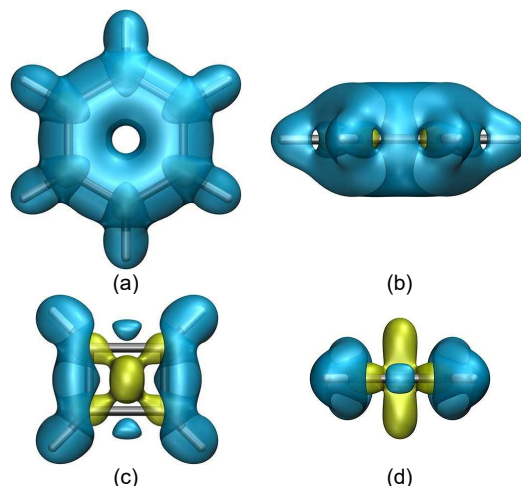


Figure 2. Shielding around C₆H₆ and C₄H₄. Top and side views: (a, b) C₆H₆, (c, d) C₄H₄. Isosurface details as in Figure 1.

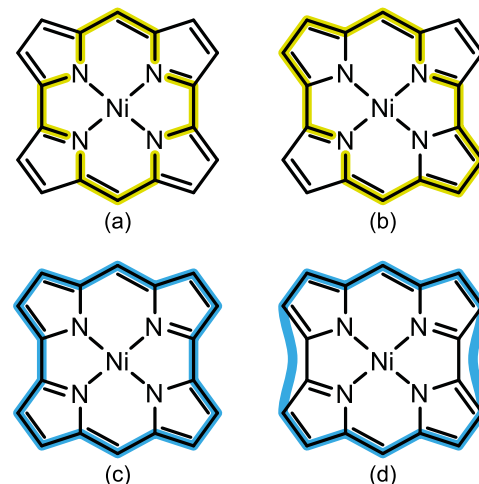


Figure 3. Conjugation pathways in NiNc. (a) antiaromatic, 14 atoms and 16 π electrons (this work); (b) antiaromatic, 16 atoms and 16 π electrons^{11,12–14} (c, d) aromatic, 18 atoms and 18 π electrons, or 14 atoms and 14 π electrons (this work).

Both NiNc and H₂Nc feature central deshielded regions resembling that in C₄H₄ but considerably larger in size (Figure 1). These regions extend over a 14-membered cyclic conjugated subsystem with 16 π electrons (Figure 3(a)) which is more compact than the 16-membered conjugated circuit with the same number of π electrons (Figure 3(b)) hitherto thought to be the source of antiaromatic character in norcorrole.^{11,12,14} The view that the antiaromaticity of NiNc follows from the pathway in Figure 3(b) has not been challenged previously, although evidence that the pathway in Figure 3(a) provides the more realistic option has been hiding in plain sight. Large magnetically induced current densities over the pathway in Figure 3(a) were observed in previous studies of NiNc¹⁵ and Ni^{II} cyanonorcorrole¹⁶; a NICS contour plot for a NiNc derivative²² akin to a planar slice through the surface in Figure 1(a, b) shows deshielding beneath the area of this pathway; it has been argued that in ACID plots for substituted NiNc the main

ring current “flows along the inner periphery involving four nitrogen atoms and fourteen bonds”^{23,24} although inspection of these ACID plots shows significant paratropic currents all over the NiNc unit. The antiaromatic pathway in NiNc (Figure 3(a)) resembles an antiaromatic pathway with 16 atoms and 16 π electrons in the octaethylporphyrin Zn^{II} dication,²⁵ established through DFT GIMIC²⁶ calculations; the absence of two meso-methine units in neutral NiNc produces the same 16 π electron count as that in the octaethylporphyrin Zn^{II} dication.

Nickel enhances the antiaromaticity of the central region in NiNc, most likely by creating additional conjugation pathways. To check if Ni contributes electrons to the π system, natural population analyses (NPA) were carried out at the level of theory used in the geometry optimizations. The NPA charge on Ni was obtained as 0.470e and the NPA charge distributions in NiNc and H₂Nc turned out to be reasonably similar (Figure S2). Thus, the Ni contribution to the NiNc π system is minor and does not need to be taken into account in π electron counts. The shielded ovoids between Ni and N atoms (Figure 1(a)) provide evidence of Ni–N bonding; despite their relatively small sizes, shielding within these ovoids reaches values of over 72 ppm. The effect of the central deshielded regions in NiNc and H₂Nc on C–N and C–C bonds is similar to that observed in C₄H₄ (Figure 2(c, d)) but, surprisingly, none of the bonds along the NiNc perimeter are affected as much as the C₄H₄ “single” bonds. In fact, the antiaromatic “core” in NiNc is surrounded by a shielded 18-membered ring with 18 π electrons which, despite the two “weaker” links across C $_{\alpha}$ –C $_{\alpha}$ bonds, suggests well-established bonding interactions. While a conjugated ring with 18 π electrons (Figure 3(c)) can be considered as Hückel-aromatic, it is apparent that the level of aromatic stabilization of the ring surrounding NiNc is lower than that of C₆H₆ (Figure 2(a, b)). A closer look at the shielding around the C $_{\alpha}$ –C $_{\alpha}$ bonds reveals additional details: The distancing of the $\sigma_{\text{iso}}(\mathbf{r}) = 12$ ppm isosurface from these bonds (Figure 1(a)) and the apertures in the isosurface (Figure 1(b)) suggest that the conjugated “halo” around NiNc is bypassing the α carbons, preferring through-space conjugation similar to homoconjugation^{27–29} between the β carbons instead. With homoconjugation, the conjugated ring around the perimeter of NiNc is reduced to 14 atoms and 14 π electrons (Figure 3(d)) and remains Hückel-aromatic. In this way, the shielding around the C $_{\alpha}$ –C $_{\alpha}$ bonds identifies these bonds as the weakest and most reactive along the NiNc outer perimeter.¹¹ Similarly to NiNc, the deshielded central region in H₂Nc involves a 14-membered circuit with 16 π electrons but its smaller size and different shape indicate weaker antiaromatic character. In this case, shielding along the outer frame outlines two interacting dipyrin fragments with much less if any homoconjugation across the links between these fragments. Of course, a lower isovalue for H₂Nc, say $\sigma_{\text{iso}}(\mathbf{r}) = \pm 10$ ppm, would produce a positive $\sigma_{\text{iso}}(\mathbf{r})$ isosurface enveloping the C $_{\alpha}$ –C $_{\alpha}$ bonds to a greater extent. The central antiaromatic region in the bowl-shaped H₂Nc geometry of C₂ symmetry is larger and closer in size and shape to that in NiNc.

NICS(0)³ and NICS(1),^{30,31} which measure $-\sigma_{\text{iso}}(\text{at ring center})$ and $-\sigma_{\text{iso}}(\text{at } 1 \text{ \AA above ring center})$, respectively, provide a more conventional way of estimating aromaticity in NiNc and H₂Nc. A complication arises from the fact that both molecules are non-planar. Calculating NICS(0) is less of a problem because the center of a non-planar ring can be located by averaging the coordinates of its atoms; NICS(0) values have been reported for some norcorrole derivatives.^{11,14,22–24} One way of defining NICS(1) positions for non-planar rings is to fit a plane to the ring atoms and ring center, and then use the points 1 \AA above and below along the normal passing through the ring centre.³² In general, these two points are not equivalent by

symmetry and give different NICS(± 1) values. B3LYP-GIAO and PBE0-GIAO NICS(0), NICS(1) and nuclear isotropic shielding values for C₆H₆ and C₄H₄ are shown in Figure S3. Analogous information for NiNc and H₂Nc, with the addition of NICS(–1) values, is shown in Figure S4. Whereas the choice of DFT method affects Ni, N and C nuclear shieldings, related changes in proton and off-nucleus shieldings, such as NICS and shielding isosurfaces (Figures 1 and 2) are much more minor. The NICS(0) and NICS(± 1) values calculated within the central 14-membered circuits with 16 π electrons in NiNc and H₂Nc confirm that the interior of NiNc is significantly more antiaromatic than that of H₂Nc; the level of antiaromaticity in the interior of H₂Nc at the C_i geometry is close to that in C₄H₄ (Figures S3 and S4). The NICS(0) and NICS(± 1) values for the pyrrole rings in NiNc are negative but of relatively low magnitudes; the corresponding values for the pyrrole rings in H₂Nc are a mixture of positive and negative numbers of even lower magnitudes. These observations suggest that in both NiNc and H₂Nc the individual pyrrole rings should be considered as non-aromatic. Clearly, while NICS can be used to identify the antiaromatic features of NiNc and H₂Nc, these indices are unable to spot the presence of the aromatic “halos” surrounding these molecules. A detail worth drawing attention to is the very substantial deshielding of the two hydrogens in the interior of H₂Nc (Figure S4(c–f)), which is a consequence of the positioning of these hydrogens inside the central deshielded region of this molecule (Figure 1(c, e)) and has been observed experimentally.¹⁴ Not coincidentally, these “inner” hydrogens have high positive NPA charges (Figure S2(b, c)).

As shown in this communication, through a visual approach in which bonding and aromaticity can be inspected using 3D isotropic magnetic shielding isosurfaces, norcorrole combines a strongly antiaromatic interior with an aromatic “halo” along its outer perimeter, providing a unusual example of a stable molecule in which aromaticity and antiaromaticity coexist in balance. This description of norcorrole provides an exception to the notion that the stability of porphyrinoids follows from “appended 6 π electronic sextets”:³³ As shown in Figure 1, the large deshielded region in the center of norcorrole disrupts the shielding within pyrrole rings and makes them non-aromatic; this is also supported by the NICS(0) and NICS(± 1) values for these rings. It has been emphasized,^{18,19,34} that the proper description of the magnetic properties of antiaromatic molecules requires the use of CASSCF wavefunctions as such molecules often exhibit significant open-shell singlet character. Computational evidence indicates that the antiaromaticity of rectangular (*D*_{2h}) cyclobutadiene is overestimated by methods based on closed-shell wavefunctions such as HF, MP2 and B3LYP.^{18,19} However, a CASSCF-GIAO calculation on NiNc in an active space large enough to take into account all conjugation pathways is not feasible at present, as the size of this active space is well beyond the capabilities of current codes. Were it possible to do such a calculation, the central antiaromatic region would become smaller while the aromatic “halo” would expand, enhancing the aromatic stabilization of this molecule and lending further support to the description suggested in this communication.

ASSOCIATED CONTENT

Supporting Information.

The Supporting Information is available free of charge at <http://pubs.acs.org>.

Figures showing bond lengths in the local minimum geometries of NiNc and H₂Nc; NPA charges for NiNc and H₂Nc; NICS values and isotropic shieldings for benzene, cyclobutadiene, NiNc and H₂Nc; computational details including optimized geometries; additional references (PDF). Zip archive of Gauss-

ian cube files with isotropic shielding values for benzene, cyclobutadiene, NiNc and H₂Nc (ZIP).

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