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Unveiling the multi-step solubilization mechanism of sub-micron size vesicles by detergents

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Abstract

The solubilization of membranes by detergents is critical for many technological applications and has become widely used in biochemistry research to induce cell rupture, extract cell constituents, and to purify, reconstitute and crystallize membrane proteins. The thermodynamic details of solubilization have been extensively investigated, but the kinetic aspects remain poorly understood. Here we used a combination of single-vesicle Förster resonance energy transfer (svFRET), fluorescence correlation spectroscopy and quartz-crystal microbalance with dissipation monitoring to access the real-time kinetics and elementary solubilization steps of sub-micron sized vesicles, which are inaccessible by conventional diffraction-limited optical methods. Real-time injection of a non-ionic detergent, Triton X, induced biphasic solubilization kinetics of surface-immobilized vesicles labelled with the DiI/DiD FRET pair. The nanoscale sensitivity accessible by svFRET allowed us to unambiguously assign each kinetic step to distortions of the vesicle structure comprising an initial fast vesicle-swelling event followed by slow lipid loss and micellization. We expect the svFRET platform to be applicable beyond the sub-micron sizes studied here and become a unique tool to unravel the complex kinetics of detergent-lipid interactions.

Introduction

Detergent-induced membrane solubilization is critical for applications including membrane-protein purification^{1, 2}, and targeted drug delivery, where vesicle rupture enables release of encapsulated therapeutics³. Despite decades of widespread use, the complexity of membrane solubilization, coupled with limitations in current technology, have made characterizing its mechanism extremely challenging.⁴

Initial biochemical experiments indicated that the rate of membrane solubilization depends on the lipid phase, and the type and concentration of detergent⁵. The non-ionic detergent Triton X 100 (TX-100), for example, solubilizes phosphocholine (PC) rich membranes relatively slowly below the gel-to-liquid transition temperature but speeds up rapidly in the fluid phase^{6, 7}. In the gel state, however, the concentration of detergent required to achieve complete solubilization is strongly dependent on the lipid chain length⁸. For most biochemical applications, TX-100 is the solubilizer of choice, and is used as a reference for measuring the activity of other surfactants^{9, 10}. Turbidity measurements also reported the TX-100: lipid ratios required to solubilize lipid vesicles as a function of phase^{11, 12, 13} and lipid¹⁴, and isothermal titration calorimetry has probed the initial TX-100-membrane interaction¹⁵. These experiments suggest an interplay between surfactant monomers and lipids at the detergent's critical micellar concentration (CMC) in which lipid re-arrangement leads to heat transfer and mixed-micelle formation within the intact membrane¹⁶.

Importantly, the solubilization activity of TX-100 is inhibited by membrane cholesterol, though the precise mechanism is still unclear. Cholesterol may initiate liquid ordered, detergent-resistant regions across the membrane^{12, 17}, but whether these are microdomains¹⁸, detergent-resistant rafts¹⁹ or a combination of both²⁰ requires confirmation. Alternatively, TX-100 may promote liquid-ordered phases via interaction with order-preferring cholesterol-rich regions rather than initiating lipid reorganization²¹.

Despite its complexity, the most widely adopted model to describe membrane solubilization is the three state mechanism²². In State 1, detergent monomers partition the bilayer until a saturation value (R_{sat}) is reached^{23, 24}. This results in an increase in mass and an increase in turbidity due to vesicle swelling, and potentially, fusion between vesicles. In State 2, the membrane starts to disintegrate, and this phase involves the formation of mixed detergent-lipid micelles coexisting with the bilayer. State 3 corresponds to breakdown of the membrane into mixed-micelles in solution. There are, however, many unanswered questions, particularly regarding the timescales of these processes mostly due to a lack of methods that can unambiguously dissect each stage of the solubilization process. While Cryo-TEM, NMR and conventional dynamic light scattering all provide snapshots of the membrane conformation^{25, 26, 27}, they cannot provide dynamic insight. Conversely, ITC probes the thermodynamics and turbidity measurements reveal solubilization conditions, but neither reveal structure^{28, 29}. Molecular dynamics simulations have attempted to bridge this gap^{17, 30, 31} and coarse-grained simulations reveal a sequence of events in broad agreement with the three-state mechanism¹⁰. Additionally, phase contrast and fluorescence microscopy have also been used to study the solubilization mechanism of giant (10-20 μm) unilamellar vesicles (GUVs) and the influence of cholesterol¹¹. These studies demonstrated that alterations in vesicle shape following the injection of non-ionic TX-100 or anionic sodium dodecyl sulfate (SDS) are regulated by the very different flip-flop rate of both detergents. For TX-100, an almost instantaneous flip-flop rate (<0.5 s) ensures equilibration of detergent molecules across both leaflets^{5, 13, 32} that results in swelling and an increase in surface area of the GUVs. In these studies, the formation of pores in the POPC bilayer leads to complete solubilization into micelle-like structures for TX-100 concentrations of ~0.18 mM, well below the detergent CMC (~0.28 mM). In contrast, SDS exhibits a flip-flop rate in the range of minutes to hours at room temperature, and a concentration of SDS (~30 mM) much higher than its CMC (~8 mM) was needed to induce solubilization and no increase in surface area was observed. These

studies confirmed a stepwise solubilization mechanism of GUVs by both detergents and demonstrated that the structures adopted by the lipid-detergent complex are strongly influenced by the ability of the detergent to rapidly equilibrate between both leaflets. Unfortunately, conventional techniques such as phase contrast and fluorescence microscopy used in these studies can only resolve changes in shape for large objects with diameters usually above $\sim 5\text{-}10\text{ }\mu\text{m}$, which constitutes only one end of the curvature space of lipid membranes. Furthermore, traditional optical microscopy quantifies macroscopic changes in size and packing density but provides little structural information at the molecular level. Therefore, there is a need to develop complementary methods that can monitor the solubilization process of individual vesicles smaller than the diffraction limit ($\sim 250\text{ nm}$) that constitute the other end of the membrane curvature space, and which are commonly used in biotechnological applications but cannot be studied by conventional microscopy.

In this work we demonstrate the combination of single-molecule fluorescence and Förster resonance energy transfer (FRET) to monitor, in real-time, the detergent-induced solubilization of **large** unilamellar vesicles (**LUVs**) with sizes smaller than the diffraction limit. FRET is sensitive to 1-10 nm distances between two small organic dyes termed donor (D) and acceptor (A)³³ and in svFRET, lipophilic D and A fluorophores are incorporated directly within the membrane to act as reporters of molecular interactions. Although svFRET has been applied to investigate the kinetics of membrane fusion^{34, 35} and pore formation³⁶, its application for characterizing solubilization kinetics has not been reported. Importantly, by immobilizing individual vesicles on the surface of a microscope slide via biotin-streptavidin interactions, structural changes in the lipid vesicle can be monitored without interference from vesicle fusion. We demonstrate that by monitoring the time-dependent variations in FRET efficiency (E_{FRET}) and total emission intensity ($I_{\text{D}}+I_{\text{A}}$) following the addition of detergent provides a means to differentiate, for the first time, each structural step along the solubilization process and unambiguously extract the rates of swelling and lysis events as a function of detergent concentration and cholesterol content.

Solubilization profiles of LUVs of sizes between 100 nm and 200 nm at concentrations of TX-100 near the critical micellar concentration ($\text{CMC}_{\text{TX-100}} \sim 0.28 \text{ mM}$) were characterized by a rapid increase ($\sim 5 \text{ s}$) in vesicle surface area, reported by a pronounced decrease in FRET efficiency with no change in total intensity. This swelling step is followed by a slow lysis phase ($\sim 40 \text{ s}$) involving loss in lipid content that results in a remarkable decrease in total intensity without significant variation in FRET efficiency. By measuring the diffusion coefficient of labelled LUVs using fluorescence correlation spectroscopy (FCS), we estimated a 34% increase in vesicle size induced by TX-100 and confirmed that the observed decrease in E_{FRET} on immobilized LUVs reflects vesicle swelling. Quartz-crystal microbalance with dissipation (QCM-D) revealed a 5% mass gain during the first few seconds after TX-injection followed by a 63% mass loss at the final stages of solubilization with timescales comparable to those observed by svFRET, thus confirming the assignation of each event in the svFRET trajectories. When the same experiments were carried out in the presence of 20% cholesterol, we observed a remarkable decrease in the rate of swelling with much smaller impact on the lysis rate, suggesting that the previously reported resistance to solubilization conferred by cholesterol to POPC vesicles might arise from cholesterol inhibiting the initial step of detergent insertion in the lipid bilayer. In summary, we demonstrate the use of a multi-disciplinary approach combining the novel application of svFRET to dissect the elementary stages of the solubilization process and extract kinetic information with FCS and QCM-D to quantify changes in size and mass of sub-micron size vesicles with high membrane curvature. Given that membrane curvature is emerging as an important mechanism regulating the recruitment of numerous proteins and peptides³⁷, the svFRET technique should become an exceptional tool that complements current optical microscopy and phase contrast methods when targeting the entire curvature space of lipid bilayers.

Results and Discussion

PC- and phosphoserine (PS)-rich model-membrane vesicles incorporating 20% cholesterol were prepared as detailed in the **Methods** and are schematically shown in **Figure S1**. The amounts of donor (Dil) and acceptor (DiD) per vesicle were optimized (1:1, 0.1 % of each dye) such that the average FRET efficiency (E_{FRET}) per vesicle was initially close to 0.5, enabling nanometer length scale changes to be quantified by an observable change in E_{FRET} in either direction. The production of homogeneously distributed unilamellar vesicles ($d \sim 200$ nm) was confirmed by dynamic light scattering (**Figure S2**). Steady-state fluorescence measurements were carried out as an initial step to characterize the interaction between TX-100 and labelled vesicles. As the concentration of TX-100 was progressively increased, we observed a decrease in E_{FRET} (**Figure 1a**), from a value of 0.43 ± 0.05 in the absence of TX-100, to 0.13 ± 0.02 in the presence of 4.4 mM with a half-maximal concentration constant of 0.39 ± 0.07 mM. These data suggest that the addition of TX-100 induces changes in vesicle structure, or composition, that results in a high distance separation between the dyes. The decrease in E_{FRET} was further confirmed by time-correlated single photon counting, where the amplitude weighted average lifetime of Dil progressively increased as a function of TX-100 (**Figure 1b, Table S1**).

Having established FRET as a sensor of fluorophore separation in the ensemble, fluorescence correlation spectroscopy (FCS) was used to probe the diffusion of single vesicles. The high sensitivity of FCS to the size of diffusing vesicles makes it an attractive technique for accessing their diameter under solubilizing conditions. Interestingly, the use of FCS techniques has mainly focused on the formation of micelle-like structures, the understanding of vesicle fusion and protein-lipid interactions, but their application in the context of lipid-detergent interactions remains under-explored^{38, 39, 40}.

The translational diffusion times of vesicles prepared with 0.2 % (1-palmitoyl-2-{6-[(7-nitro-2-1, 3-benzoxadiazol-4-yl) amino] hexanoyl}-*sn*-glycero-3-phosphocholine) (NBD-PC) (**Figure S3**), a fluorescent analog of PC, were recorded. NBD-PC replaced Dil and DiD for compatibility with the apparatus. Normalized cross-correlation functions obtained from freely-diffusing vesicles progressively shifted towards longer diffusion times as a function of TX-100 (**Figure 1c**). In the absence of detergent, vesicles of diameters ~100 nm and ~200 nm displayed diffusion times, τ_D , across the confocal volume of 6.1 ± 0.03 ms and 8.0 ± 0.3 ms, respectively (**Figure 1d**). τ_D associated with the diffusion of the smaller vesicles increased by ~13 % in 0.1 mM TX-100 and by ~34 % in 0.2 mM TX-100. The larger 200 nm-diameter vesicles also displayed a similar trend, representing an ~8 % and ~18 % increase in hydrodynamic diameter at 0.1 mM and 0.2 mM TX-100, respectively. These data point toward an increase in mean vesicle diameter when incubated with TX-100, and was attributed to vesicle expansion, fusion or a combination of both in solution. It is interesting to note that the relative increase in τ_D and the resulting increase in size (**Table S2**) are slightly higher for smaller LUVs and that this trend is maintained at the two concentrations of TX-100 investigated. This results in a 25% higher slope for the 100 nm LUVs $(9.40 \pm 1.4) \times 10^3 \mu\text{s mM}^{-1}$ compared to $(7.0 \pm 1.4) \times 10^3 \mu\text{s mM}^{-1}$ for the 200 nm LUVs (**Figure 1d**). We interpreted this as evidence that the degree of swelling depends on vesicle size; an observation that emphasizes the importance of membrane curvature modulating the initial stages of TX-100 insertion in the lipid bilayer. As shown in recent molecular dynamics studies, the formation of lipid packing defects is intimately linked to membrane curvature in addition to lipid composition⁴¹ and we hypothesize that the higher slope observed for the smaller vesicles might reflect the formation of stress-induced defects that facilitate the insertion of detergents. This is an interesting finding to consider for further studies because many natural and synthetic compounds must partition in the lipid bilayer to reach their targets.

In the next step, to rule out the possibility of fusion and investigate in more detail each step of the solubilization process, DiI/DiD labelled vesicles containing a low percentage of biotinylated lipids and 20% cholesterol were immobilized onto a NeutrAvidin-coated surface and imaged via total internal reflection fluorescence microscopy. As illustrated in **Figure 2a**, biotinylated vesicles were anchored to NeutrAvidin tethered to the surface via biotinylated polyethylene glycol (PEG). In the absence of TX-100 the vesicles were stable with no variation in the svFRET efficiency observed (**Figure 2b**). Perturbation of single vesicles by TX-100 was then reported as observable changes in the svFRET efficiency in real-time with 50 ms time integration (**Figure 2b, 2c**). To suppress photobleaching and optimize conditions for svFRET, the fluorescence response of single vesicles labelled with DiD were investigated as a function of excitation intensity and percentage of dye-loading content. As demonstrated in **Supplementary Text I, Figure S4** and **Table S3**, excitation intensities $< 0.04 \text{ mW/cm}^2$ with 0.25 % dye were necessary for long-term (180 s) stability of the incorporated dyes. In the absence of TX-100, the FRET efficiency from single vesicles remained largely invariant with a value of $E_{\text{FRET}} \sim 0.5$. Injection of 0.16 mM TX-100 then induced variations in E_{FRET} and total intensity on remarkably different time scales. Immediately after injection of TX-100, the FRET efficiency decreased from a value of ~ 0.46 to ~ 0.22 in a 250 sec time window (**Figures 2d, S5**). Most of this change happened in the first 100 seconds after injection, pointing towards a $\sim 20\%$ increase in the separation distance between FRET pairs. Assuming spherical vesicles, this distance scales directly with the vesicle radius and thus agrees well with the FCS data. Importantly, within this time window, the total intensity decreased only by $\sim 7\%$. At time scales longer than 250 seconds, the FRET efficiency did not change further, whereas the total intensity decreased progressively to $\sim 50\%$ of its initial value. The different timescales and responses of both signals suggest that they represent different distortions of the vesicle structure. The rapid decrease in FRET efficiency without significant variation in total intensity indicates a structural change involving no loss of lipid content and we assigned it as arising from vesicle swelling induced by TX-100 molecules inserting into the lipid

bilayer and increasing the average inter-dye distance. The time window where the FRET efficiency does not change but the total intensity is strongly decreased suggests a structural distortion involving the diffusion of lipids into solution and it was assigned to a lysis step resulting in the formation of micelles. The FRET efficiency plateau value observed in this time window ($E \sim 0.22$) represents an inter-dye distance ~ 6.5 nm within these micelles. These observations point towards a fast vesicle expansion event with half-life t_E , followed by a slower lysis event (t_L). The expansion step was observed to occur on average 12 times faster than lysis, and at 0.08 mM TX-100, the half-lives associated with each event increased by ~ 75 % (**Figure 2e**). The ability to unambiguously discriminate between expansion and lysis events and extract individual kinetic rates for each stage during LUV solubilization is a completely new finding only afforded by the development of svFRET for surface-immobilized vesicles.

The observation of a stepwise solubilization mechanism in LUVs qualitatively agrees with optical microscopy and phase contrast measurements reported in GUVs composed of POPC and sizes of ~ 20 - 30 $\mu\text{m}^{11, 12}$. In these studies, it was found that the addition of sub-CMC concentrations of TX-100, in the range of those used in our work, resulted in an increase in vesicle diameter a few seconds after injection. Increasing further the concentration of TX-100 to near CMC values resulted in dynamic perforation of the bilayer leading to complete solubilization in a timescale of ~ 20 - 30 seconds. Importantly, when SDS was used as the detergent, no increase in surface area was detected and the solubilization process was slow and only efficient at concentrations well above the CMC (8 mM). It was suggested that the different rates of detergent equilibration between the two leaflets of the bilayer may govern the solubilization process and dictate the morphological changes taking place in the vesicle during the initial stages. A direct quantitative comparison between our findings on LUVs and the previous work on GUVs is not straightforward for several reasons. First, as discussed in those studies, optical microscopy and phase contrast have access only to the cross-section of the focal plane of the vesicle and thus

quantification of surface area in vesicles that might adopt non-trivial shapes becomes unfeasible. Secondly, given the size of the GUVs and the injection of detergent using a micropipette, the concentration of detergent around the vesicle is not homogeneous, the vesicle deforms more rapidly on the side facing the micropipette, and this might influence the overall kinetics of the process. In contrast, for the immobilized LUVs used in our study, filling the entire microchannel takes place in ~ 1 second and the detergent flow wave front should achieve a steady-state concentration much faster than any of the processes investigated. Nevertheless, it is clear that LUVs and GUVs shared some common solubilization features by TX-100. This includes similar concentration requirements and the presence of an initial and rapid expansion step followed by a slower step where lipids are released to the solution to form micelle-like structures. Thus, our svFRET studies on highly curved sub-micron size vesicles are complementary to optical microscopy on giant vesicles and both suggest a common general mechanism of solubilization.

Since cholesterol alters the membrane structure⁴², we next employed svFRET to assess the influence of cholesterol on the stability and kinetic mechanism of vesicle solubilization. 200 nm diameter vesicles were prepared in the absence of cholesterol and were induced to solubilize by 0.08 mM TX-100. Here, the expansion and lysis half-lives reduced by ~ 76 % and ~ 21 % respectively, compared with vesicles loaded with 20 % cholesterol. When the TX-100 concentration was doubled, t_E and t_L reduced further to 1.2 ± 0.4 and 17.6 ± 1.6 seconds, respectively, as shown via a comparative bar plot summarizing the relative variations in t_E and t_L as a function of cholesterol and TX-100 concentration (**Figure 2f, Figure S6**). To the best of our knowledge, these values constitute the first direct measurement on the effect of cholesterol on each individual step of the solubilization process. The relatively small decrease observed for the kinetic rate of the lysis step in the presence of 20% cholesterol compared to that with no cholesterol added agrees with previous observations on binary POPC/Cholesterol mixtures where no significant effect on the solubilization rate was found at any temperature using turbidity measurements⁴³. In contrast, the remarkable decrease in the timescale for vesicle swelling in the presence of

cholesterol was not detected in previous turbidity measurements but agrees with recent phase contrast and light scattering observations on a similar lipid bilayer composition^{11, 44}. In these studies, a significant decrease in the surface area variation during the swelling step was reported from a value of 48% with no cholesterol to ~32% in the presence of 30% cholesterol. Because the swelling event reflects mostly the initial interaction between TX-100 and the lipid membrane, the slower rate of swelling in the presence of cholesterol can be interpreted as a decrease in the association rate of TX-100 to the lipid bilayer. A decrease in the association rate of TX-100 aligns well with the reported decrease in binding constant induced by cholesterol in POPC vesicles¹¹ and potentially reinforces the curvature-mismatch hypothesis²⁷ whereby the incorporation and packing of cholesterol is incompatible with the association of TX-100 molecules due to their opposite curvature.

The svFRET and FCS experiments discussed so far have provided new structural and kinetic information regarding the solubilization mechanism of LUVs by TX-100. However, given that the main consequence of swelling and lysis events is the transfer of mass from the vesicle to solution and *vice versa*, a complete description of the mechanism will benefit from the development of methods that can evaluate such mass gain or loss at each stage of the process. To quantify the transfer of mass during solubilization, we employed a label-free quartz-crystal microbalance with dissipation (QCM-D) monitoring approach (**Supplementary Text II**). QCM-D has recently emerged as a very useful method to monitor vesicle fusion⁴⁵, vesicle adsorption to surfaces⁴⁶ and protein-vesicle interactions⁴⁷ but its application in the context of detergent-induced solubilization has not been reported. Here, PC/PS vesicles containing 20 % cholesterol were immobilized onto a quartz sensor surface and interactions with 0.16 mM TX-100 were followed by changes in oscillation frequency and dissipation, reflecting mass and viscoelasticity on the sensor surface, respectively (**Figure 3a**). Interactions between immobilized vesicles and TX-100 were observed via changes in both frequency and dissipation traces immediately after TX-100 injection corresponding to a ~5% mass gain at the sensor which we attributed to TX-

100 incorporation into vesicles. This was followed by an interaction that leads to a conformational change in intact vesicles, with no mass loss, over the first 35 s. As the local TX-100 concentration then increased, the deposited mass accumulated on the surface leading to a decrease in resonance frequency. A substantial mass loss of ~63% was then observed via an increase in resonance frequency, indicating material immobilized to the surface was released into solution (**Figure 3b**). These processes occurred on similar timescales to those obtained under the same conditions using svFRET and control experiments performed simultaneously indicated no interaction with the PEG-coated sensor surface and TX-100 (**Figure 3a**). These findings support a mechanism through which TX-100 accumulates on the curved membrane surface, preceding a rapid expansion of the vesicle structure that, in turn, precedes a slower lysis event (**Figure 4**).

In summary, we have directly monitored the solubilization of sub-micron size lipid vesicles that cannot be resolved using conventional optical techniques in response to TX-100 using svFRET, FCS and QCM-D. We show that the combination of svFRET and surface-immobilization of LUVs is a unique method to discriminate between the swelling and lysis steps and unambiguously quantify the kinetic rates of each process without interference from vesicle fusion. Our data demonstrate that detergent-induced swelling is a relatively fast process that is strongly slowed down by the presence of cholesterol whereas the subsequent lysis step is only marginally affected. The increase in vesicle size during swelling was calculated using FCS and suggested that increasing the membrane curvature facilitates the insertion of detergent molecules. Our assignment and characterization of the timescales for the swelling and lysis steps by svFRET and FCS was further supported by quantifying the mass gain (swelling) or loss (lysis) using QCM-D measurements. Understanding the structural stability and dynamics of small vesicles with high curvature is crucial not only because they are commonly employed in biotechno-

logical applications but also because important traffic pathways between the endoplasmic reticulum (ER) rely on the formation of LUVs and a large fraction of the ER consists of tubules of ~ 30 nm radius⁴⁸. In the future, the experimental approach presented here may be useful in several directions: to quantify the effect of membrane curvature on each step of the solubilization process, to determine the role of lipid composition on the solubilization of LUVs and to quantify the interaction with small molecules that need to target and cross the lipid membrane.

Methods

Materials

1-palmitoyl-2-oleoyl-sn-glycero-3-phospho-L-serine (sodium salt) (PS), 1-palmitoyl-2-oleoyl-glycero-3-phosphocholine (PC) and 1,2-dioleoyl-sn-glycero-3-phosphoethanolamine-N-(cap biotinyl) (sodium salt) (biotinylated lipid) phospholipids were purchased from Avanti Polar Lipids Inc. 1,1'-Diocadecyl-3,3,3',3'-Tetramethylindocarbocyanine Perchlorate (DiI) and 1,1'-Diocadecyl-3,3,3',3'-Tetramethylindodicarbocyanine, 4-Chlorobenzenesulfonate Salt) (DiD) membrane stains were purchased from ThermoFisher Scientific. 1-palmitoyl-2-{6-[(7-nitro-2-1,3-benzoxadiazol-4-yl)amino]hexanoyl}-sn glycerophosphocholine (NBD-PC) was purchased from Sigma Aldrich. All phospholipid samples were used without additional purification and stored in chloroform at -20°C prior to use. DiI, DiD and cholesterol stocks were stored at 4°C in chloroform prior to use. Triton X-100 was purchased from Sigma Aldrich and freshly suspended in 50 mM Tris (pH 8) prior to use.

Preparation of large unilamellar vesicles

Mixtures of lipids and lipophilic dyes were homogeneously dispersed in chloroform, dried by nitrogen flow and stored under continuous vacuum pumping at room temperature for 5 hours.

Phospholipid mixtures were subsequently re-suspended in buffer solution (50 mM Tris, pH 8) and mixed well by vortex. **Large** unilamellar vesicles were prepared by the extrusion method at room temperature, in which they were passed through a polycarbonate membrane filter of defined pore size. A molar ratio of 65: 35 PC: PS was used throughout. Vesicles were labelled with dyes (0-0.5 %), cholesterol (0-20 %) and biotin (1 %) at the molar percentages specified in the text. The mean size of the prepared vesicles in solution was evaluated by dynamic light scattering using a Zetasizer μ V molecular size detector (Malvern Instruments Ltd., UK).

Steady-state fluorescence spectroscopy

Fluorescence emission spectra were acquired using a Varian Eclipse fluorescence spectrophotometer. Spectra from Dil and DiD were recorded using an excitation wavelength of 532 nm at magic angle. FRET efficiencies were approximated by the apparent FRET efficiency, $E_{FRET} = (I_{665}/[I_{665} + I_{565}])$, where I_{665} and I_{565} represent the fluorescence intensities of the acceptor at 665 nm, and donor at 565 nm, respectively. The FRET efficiency data shown in Figure 1a was fitted to a Hill model of the form, $E_{FRET} = A + B \frac{[TX-100]^n}{k^n + [TX-100]^n}$, where A and B are the measured FRET efficiencies at the start and end of the titration, k is the half-maximal concentration constant and n is the Hill coefficient. The parameters of the fit shown in **Figure 1a** are $A = 0.43 \pm 0.04$, $B = 0.12 \pm 0.01$, $k = 0.39 \pm 0.07$ and $n = 2.0 \pm 0.3$ ($\chi^2 = 0.99$). Error bars represent the standard error of the mean from 3 individual experimental runs.

Time-resolved fluorescence spectroscopy

Fluorescence lifetime measurements were performed with a Hamamatsu C6860 Synchroscan streak camera. The 80 MHz, 100 fs (full width half maximum) 800 nm output of a Ti: Sapphire oscillator was frequency doubled with a beta barium borate crystal, giving 400 nm excitation

pulses. The 400 nm light, with an average power of less than 1 mW, was subsequently focused through the optical path length (1cm) of the solution cuvette. Fluorescence from the sample was then collected and collimated with a lens before being focused onto the entrance slit of a Chromex 250is imaging spectrograph. Excitation light was removed with a yellow schott glass filter that cuts all light below 420 nm. A spectral window of 585–607 nm corresponding to Dil fluorescence emission was selected with the spectrograph before the light was directed into the streak camera. Time-resolved fluorescence dynamics were then recorded enabling time constants of approximately 10 ps to be resolved with instrument response deconvolution.

Fluorescence Correlation Spectroscopy

Samples were deposited on glass-bottomed well plates (Whatman) and excited by the linearly polarized light of a 488 nm continuous wave laser (Becker & Hickl) which was spectrally cleaned (Semrock, US, FF01-482/18), redirected by a dichroic mirror (Semrock, US, DI01-R488) and focused into the sample by a 60 x water immersion microscope objective (Olympus, UPLSAPO60xW/1.2) mounted in an inverted microscope (Olympus, IX-71). The fluorescence was focused onto a $\phi = 50 \mu\text{m}$ pinhole (Thorlabs) before being split by a 50:50 nonpolarizing beamsplitter cube (Thorlabs). Each beam was then focused onto an avalanche photodiode (MPD50CTC APD, $\phi = 50 \mu\text{m}$, Micro Photon Devices). An emission filter (Semrock, 525/45) placed in front of the beamsplitter was used to discriminate fluorescence from scattered light. The detector signals were processed and stored by two time-correlated single photon counting (TCSPC) modules (Becker & Hickl, SPC 132). Typically 20 million photons were collected for each correlation curve with count rates between 5 and 20 kHz. All measurements were made at a stabilized temperature of $25.0 \pm 0.5 \text{ }^{\circ}\text{C}$. The excitation power as measured in the focus of the

microscope objective by a power meter (Thorlabs) was 0.02 mW corresponding to a mean irradiance of 7.15 kW/cm² assuming a Gaussian intensity distribution along the optical axis. The focal area and the detection volume were calibrated with Rhodamine 123 in aqueous solutions at low irradiance using an estimated diffusion coefficient of $4.6 \pm 0.4 \times 10^{-10} \text{ m}^2 \text{ s}^{-1}$, yielding a radial 1/e² radius of $\omega_{xy} = 0.27 \text{ } \mu\text{m}$ and volume of focus of $V = 0.53 \text{ } \mu\text{m}^3$. Correlation functions were calculated according to $G = \frac{\langle I(t)I(t+\tau) \rangle}{\langle I(t)^2 \rangle}$ where $I(t)$ is the intensity at time t and fitted according to $G = b_0 + \frac{1}{N} \left(1 + \frac{t}{\tau_D} \right)^{-1} \sqrt{\left(1 + \frac{t}{\Omega^2 \tau_D} \right)^{-1}} (1 + A_T e^{-t/\tau_T})$ where N is the number of molecules, t is the correlation time, A_T is the amplitude of the triplet and τ_T is the triplet time. Ω defines the ratio between the axial and radial 1/e² radii, ω_z and ω_{xy} , respectively: $\Omega = \frac{\omega_z}{\omega_{xy}}$. Translational diffusion coefficients, D , were determined by $D = \frac{\omega_{xy}^2}{4\tau_D}$ where τ_D is the diffusion time. All diffusion coefficients were corrected for temperature and viscosity effects and are reported for 25°C. Hydrodynamic radii, R_h , were estimated according to $R_h = \frac{kT}{6\pi\eta D}$ where k is Boltzmann's constant, T is the system temperature and η is the solution viscosity. Power series were performed in order to determine the photobleaching limits. A triplet-state contribution of 1 μs with the expected irradiance-dependent amplitude was observed in all cases. All measurements were repeated at least 20 times and curves distorted due to occasional transits of big aggregates were excluded. The surfactant was added to the diluted vesicle samples immediately before the FCS measurement. Error bars indicate the standard error of the mean.

Single-vesicle TIRF Spectroscopy

Fluorescence emission at the donor and acceptor wavelengths were acquired from single vesicles by using a prism-type total internal reflection fluorescence microscope equipped with green

(532nm) and red (635nm) lasers (Crystalaser, USA). Microscope slides were successively treated with biotinylated poly-ethyleneglycol (PEG) and NeutrAvidin, before pM concentrations of fluorescently-labelled vesicles were added. Fluorescence trajectories were acquired with an integration time of 50 ms. The base buffer used for imaging was 50 mM Tris (pH 8), 6 % (w/v) glucose, 165 U/mL glucose oxidase, 2170 U/mL catalase and 2 mM trolox. Specified concentrations of TX-100 were included in imaging buffer prior to being injected into the sample. Spatially-separated fluorescence images of donor and acceptor emission were collected in custom built relay optics with a 550 nm long-pass filter and imaged in parallel using an EMCCD camera (iXON, Andor Technology). All measurements were performed at room temperature. SvFRET efficiency after background correction was approximated by the apparent FRET efficiency, $E_{\text{FRET}} = (I_A/[I_A + I_D]) \sim R_0^6/([R_0^6 + R^6])$, where I_A and I_D are the fluorescence intensities of the acceptor and donor, respectively, R_0 is the Forster radius and R is the separation distance between the probes. Since the quantum yields of DiI and DiD are similar, E_{FRET} closely matches the true efficiency of energy transfer. Half-lives were calculated by applying double exponential fits consisting of a rise ($I = Ae^{t/E}$) and decay ($I = Be^{-t/L}$) component to the donor trajectories. Data analysis was carried out using laboratory-written analysis routines developed in MATLAB 7. Ensemble information from svFRET measurements was obtained by assembling single-vesicle FRET trajectories into population FRET contour plots.

Quartz Crystal Microbalance with Dissipation (QCM-D) monitoring.

Quartz crystal microbalance with dissipation monitoring (QCM-D) experiments were performed using a Q-sense E4 system (Biolin Scientific). SiO₂-coated AT-cut quartz sensors (QSX 303, Biolin Scientific) were used, for which the fundamental frequency was 4.95 ± 0.05 MHz. The sensors were initially subjected to a 10 minute cleaning step by UV-ozone, prior to being soni-

cated in solutions of 2 % Hellmanex III and 2 x ultrapure Milli-Q water for 10 minutes. The sensors were then dried with N₂ and placed under UV–ozone for a further 30 minutes. Each sensor was then immersed in 100% ethanol for 30 minutes and dried with N₂ before installation in the flow modules. The QCM-D flow chambers were first flushed with ultrapure Milli-Q water for 1 hour, and then with 50 mM Tris buffer (pH 8) for 20-30 minutes before each measurement until a stable baseline was established (<0.5 Hz shift over 10 min). The flow rate was kept constant at 20 µL/min. The sensor surfaces were then functionalized with biotinylated polyethyleneglycol (Iris Biotech) which acts as a biocompatible support for specific immobilization of Avidin. The sensor surfaces were then rinsed with 50 mM Tris buffer (pH 8.0) for 15 min to remove unadsorbed molecules. Thereafter, Avidin was immobilized on the sensor surfaces by incubating a 0.1 mg/mL Avidin solution in 50 mM Tris buffer (pH 8.0) for 20 min, following a rinse step with 50 mM Tris buffer (pH 8.0) for 20 min to wash unbound Avidin molecules. Subsequently, vesicles coated with 1 % biotinylated lipids were immobilized on the sensor surfaces by incubation with a 33 µg/mL vesicle solution for 70 min. Triton X-100 (TX-100) detergent solutions at specified concentrations were then introduced into the QCM-D flow chambers. Changes in mass (Δm) were related to changes in frequency (Δf) via the Sauerbrey equation $\Delta m = - (C \cdot \Delta f)/n$ where n is the overtone number and C is a constant related to the properties of the quartz (17.7 ng Hz⁻¹cm⁻²).

References

1. Brown, D. A. & London, E. Structure and function of sphingolipid- and cholesterol-rich membrane rafts. *J. Biol. Chem.*, **275**, 17221–17224 (2000).
2. Sotgia, F.; Razani, B.; Bonuccelli, G.; Schubert, W.; Battista, M.; Lee, H.; Capozza, F.; Schubert, A. L.; Minetti, C.; Buckley, J. T. & Lisanti, M. P. Intracellular retention of glycosylphosphatidyl inositol-linked proteins in caveolin-deficient cells. *Mol. Cell. Biol.* **22**, 3905–3926 (2002).
3. Hamada, T.; Sugimoto, R.; Vestergaard, M. D. C.; Nagasaki, T. & Takagi M. Membrane disk and sphere: controllable mesoscopic structures for the capture and release of a targeted object. *J. Am. Chem. Soc.* **132**, 132, 10528–10532 (2010).

4. Lichtenberg, D.; Ahlyayauch H.; Alonso, A. & Goñi, F. M. Detergent solubilization of lipid bilayers: a balance of driving forces. *Trends. Biochem. Sci.*, **38**, 85-93 (2013).
5. Lichtenberg, D.; Ahlyayauch, H. & Goñi, F. M. The mechanism of detergent solubilization of lipid bilayers. *Biophys. J.*, **105**, 289–299 (2013).
6. De la Maza, A. & Parra, J. L. Vesicle-micelle structural transition of phosphatidylcholine bilayers and Triton X-100. *Biochem. J.*, **303**, 907–914 (1994).
7. Schnitzer, E.I Lichtenberg D. & Kozlov M. M. Temperature-dependence of the solubilization of dipalmitoylphosphatidylcholine (DPPC) by the non-ionic surfactant Triton X-100, kinetic and structural aspects. *Chem. Phys. Lipids*. **126**, 55–76 (2003).
8. Patra, S. K., Alonso, A. & Goñi, F. M. Detergent solubilisation of phospholipid bilayers in the gel state: the role of polar and hydrophobic forces. *Biochim. Biophys. Acta*. **1373**, 112-118 (1998).
9. Urbaneja, M. A.; Alonso, A.; Gonzalez-Mañas, J. M.; Goñi, F. M.; Partearroyo, M. A.; Tribout M. & Paredes, S. Detergent solubilization of phospholipid vesicle. Effect of electric charge. *Biochem. J.*, **270**, 305–308 (1990).
10. Goñi, F. M.; Urbaneja, M. A.; Arrondo, J. L.; Alonso, A.; Durrani, A. A. & Chapman, D. The interaction of phosphatidylcholine bilayers with Triton X-100. *Eur. J. Biochem.*, **160**, 659–665 (1986).
11. Sudbrack, T. P. Archilha, N. L.; Itri, R. & Riske, K. A. Observing the solubilization of lipid bilayers by detergents with optical microscopy of GUVs. *J. Phys. Chem. B*, **115**, 269–277 (2011).
12. Mattei, B.; França A. D. C. & Riske, K. A. Solubilization of binary lipid mixtures by the detergent Triton X-100: the role of cholesterol. *Langmuir*, **31**, 378–386 (2015).
13. Fan, H. Y.; Das, D. & Heerklotz, H. “Staying out” rather than “cracking in”: assymetric membrane insertion of 12:0 lysophosphocholine. *Langmuir*, **32**, 11655–11663 (2016).
14. Ahlyayauch, H.; Larijani, B.; Alonso, A. & Goñi, F. M. Biophysics of sphingolipids I. Detergent solubilization of phosphatidylcholine bilayers in the fluid state: Influence of the acyl chain structure. *Biochim. Biophys. Acta. Biomembr.* **1758**, 190–196 (2006).
15. Heerklotz, H.; Tsamaloukas A. D. & Keller, S. Monitoring detergent-mediated solubilization and reconstitution of lipid membranes by isothermal titration calorimetry. *Nat. Protoc.* **4**, 686-697 (2009).
16. Heerklotz, H. & Seelq, J. Correlation of membrane/water partition coefficients of detergents with the critical micelle concentration. *Biophys. J.*, **78**, 2435–2440 (2000).
17. Muddana, S. H.; Chiang, H. H. & Butler, P. J. Tuning membrane phase separation using nonlipid amphiphiles. *Biophys. J.*, **102**, 489–497 (2012).

18. Brown, J. K. & Rose J. K. Sorting of GPI-anchored proteins to glycolipid-enriched membrane subdomains during transport to the apical cell surface. *Cell*. **68**, 533-544 (1992).
19. Simons, K. & Vaz W. L. Model systems, lipid rafts and cell membranes. *Annu. Rev. Biophys. Biomol. Struct.* **33**, 269-295 (2004).
20. Lichtenberg, D.; Goñi, F. & Heerklotz, H. Detergent-resistant membranes should not be identified with membrane rafts. *Trends Biochem. Sci.* **30**, 430-436 (2005).
21. Nazari, M.; Kurdi, M. & Heerklotz, H. Classifying surfactants with respect to their effect on lipid membrane order. *Biophys. J.* **102**, 498–506 (2012).
22. Helenius, A. & Simons, K. Solubilization of membranes by detergents. *Biochim. Biophys. Acta.* **415**, 29–79 (1975).
23. Stuart, M. C. A. & Boekema, E. J. Two distinct mechanisms of vesicle-to-micelle and micelle-to-vesicle transition are mediated by the packing parameter of phospholipid-detergent systems. *Biochim. Biophys. Acta. Biomembr.* **1768**, 2681–2689 (2007).
24. Ahyayauch, H.; Bennouna, M.; Alonso, A. & Goñi, F. M. Detergent effects on membranes at subsolubilizing concentrations: transmembrane lipid motion, bilayer permeabilization, and vesicle lysis/reassembly are independent phenomena. *Langmuir*, **26**, 7307–7313 (2010).
25. Almgren, M. Mixed micelles and other structures in the solubilization of bilayer lipid membranes by surfactants. *Biochim. Biophys. Acta. Biomembr.* **1508**, 146–163 (2000).
26. Edwards, K. & Almgren, M. J. Solubilization of lecithin vesicles by C₁₂E₈: structural transitions and temperature effects. *Colloid Interface Sci.*, **147**, 1–21 (1991).
27. Ahyayauch, H. Collado, M. I.; Goñi, F. M. & Lichtenberg, D. Cholesterol reverts Triton X-100 preferential solubilization of sphingomyelin over phosphatidylcholine: a ³¹P-NMR study. *FEBS Lett.* **583**, 2859–2864 (2009).
28. Arnulphi, C.; Sot, J.; García-Pacios, M.; Arrondo, J.-L. R.; Alonso, A. & Goñi, F. M. Triton X-100 Partitioning into Sphingomyelin Bilayers at Subsolvubilizing Detergent Concentrations: Effect of Lipid Phase and a Comparison with Dipalmitoylphosphatidylcholine. *Biophys. J.* **93**, 3504–3514 (2007).
29. Ahyayauch, H.; Collado, M. I.; Alonso, A. & Goñi, F. M. Lipid Bilayers in the Gel Phase Become Saturated by Triton X-100 at Lower Surfactant Concentrations Than Those in the Fluid Phase. *Biophys. J.* **102**, 2510–2516 (2012).
30. Pizzirusso, A.; De Nicola, A. & Milano, G. J. MARTINI coarse-grained model of Triton TX-100 in pure DPPC monolayer and bilayer interfaces. *Phys. Chem. B*, **120**, 3821–3832 (2016).
31. Bandyopadhyay, S.; Shelley, J. C. & Klein, M. L. Molecular dynamics study of the effect of surfactant on a biomembrane. *J. Phys. Chem. B.* **105**, 5979–5986 (2001).
32. Pizzirusso, A.; De Nicola, A.; Agur Sevink, G. J.; Correa, A.; Cascella, M.; Kawakatsu, T.;

- Rocco, M.; Zhao, Y.; Celino, M. & Milano, G. Biomembrane solubilization mechanism by Triton X-100: a computational study of the three stage model. *Phys. Chem. Chem. Phys.* **19**, 29780-29794 (2017).
33. Stryer, L. Fluorescence resonance energy transfer as a spectroscopic ruler. *Annu. Rev. Biochem.* **47**, 819-846 (1978).
 34. Diao, J.; Su, Z.; Ishitsuka, Y.; Lu, B.; Lee, K. S.; Lai, Y.; Shin, Y. & Ha, T. A single-vesicle content mixing assay for SNARE-mediated membrane fusion. *Nat. Comm.* **1**, 54 (2010).
 35. Yoon, T.; Okumus, B.; Zhang, F.; Shin, Y. & Ha, T. Multiple intermediates in SNARE-induced membrane fusion. *Proc. Natl. Acad. Sci. USA.* **103**, 19731-19736 (2006).
 36. Wang, B. Y.; Liu, Y.; DeBerg, H. A.; Nomura, T.; Hoffman, M. T.; Rohde, P. R.; Schulten, K.; Martinac, B. & Selvin, P. R. Single molecule FRET reveals pore size and opening mechanism of a mechano-sensitive ion channel. *eLife*. **18**, e01834 (2014).
 37. Antonny, B. Mechanisms of membrane curvature sensing. *Annu. Rev. Biochem.* **80**, 101-123 (2011).
 38. Zetti, H.; Portnoy, Y.; Gottlieb, M. & Krausch, G. J. Investigation of micelle formation by fluorescence correlation spectroscopy. *Phys. Chem. B.* **109**, 13397-13401 (2005).
 39. Piñeiro, L.; Freire, S.; Bordello, J.; Novo, M. & Al-Soufi, M. Dye exchange in micellar solutions. Quantitative analysis of bulk and single molecule fluorescence titrations. *Soft Matter*. **9**, 10779-10790 (2013).
 40. Al-Soufi, W.; Reija, W. B.; Felekyan, S.; Seidel, C. A. & Novo, M. Dynamics of supramolecular association monitored by fluorescence correlation spectroscopy. *ChemPhysChem*. **9**, 1819-1827 (2008).
 41. Vanni S.; Hirose, H.; Barelli H.; Antonny, B. & Gautier, R. A sub-nanometre view of how membrane curvature and composition modulate lipid packing and protein recruitment. *Nat. Commun.* **5**, 4916 (2014).
 42. Yang, S.; Kreutzberger, A. J. B.; Lee, J.; Kiessling, V. & Tamm, L. K. The role of cholesterol in membrane fusion. *Chem. Phys. Lipids*. **199**, 136-143 (2016).
 43. Schnitzer, E.; Kozlov, M. M. & Lichtenberg, D. The effect of cholesterol on the solubilization of phosphatidylcholine bilayers by the non-ionic surfactant Triton X-100. *Chem. Phys. Lipids*. **135**, 69-82 (2005).
 44. Carita, A. C.; Mettei, B.; Domingues, C. C.; de Paula, E. & Riske, K. A. Effect of Triton X-100 on raft-like lipid mixtures: phase separation and selective solubilization. *Langmuir*. **33**, 7312-7321 (2017).
 45. Morigaki, K. & Tawa, K. Vesicle fusion studied by surface plasmon resonance and surface plasmon fluorescence spectroscopy. *Biophys. J.* **91**, 1380-1387 (2006).
 46. Lind, T.; Cárdenas, M. & Wacklin, H. P. Formation of supported lipid bilayers by vesicle fusion: effect of deposition temperature. *Langmuir*. **30**, 7259-7263 (2014).

47. Heider, S.; Reimhult, E. & Metzner, C. Real-time analysis of protein and protein mixture interaction with lipid bilayers. **1860**, 319-328 (2018).
48. Shibata, Y.; Hu, J.; Kozlov, M. M. & Rapoport, T. A. Mechanisms of shaping the membranes of cellular organelles. *Annu. Rev. Cell Dev. Biol.* **25**, 329-354 (2009).

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Data Availability Statement

The datasets generated during the current study are available from the corresponding author on reasonable request.

Author Contributions

S. D. Q. and J. C. P. designed the research. P. A. D., S. D. Q., J. J. –C., G. J. H., L. P., and C. P. G. performed the research. S. D. Q., J. J. –C., G. J. H., M. N., I. D. W. S., M. C. L., S. J., W. A. and J. C. P. analyzed the data. S. D. Q. and J. C. P. wrote the manuscript.

Additional Information

Figures S1-S6, Tables S1-S3 and text are found in the Supporting Information.

The authors declare no competing interests.

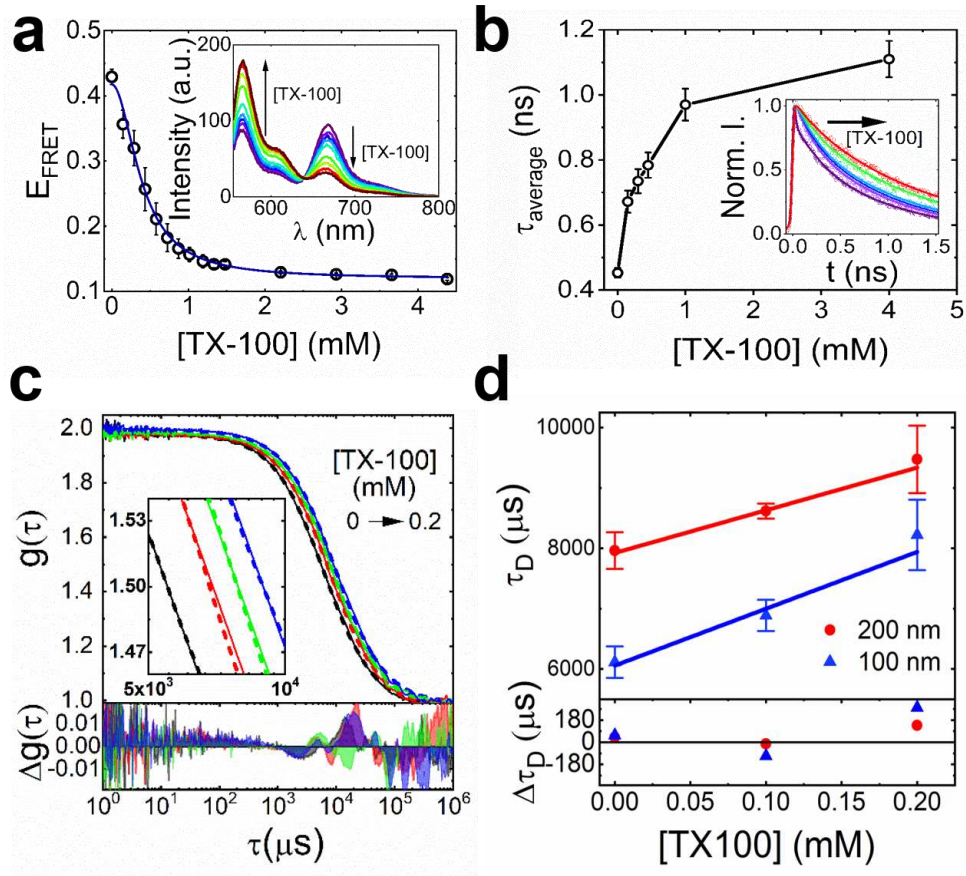


Figure 1. TX-100 vesicle interactions reported using ensemble FRET and FCS. (a) FRET efficiency of Dil/DiD labeled vesicles versus TX-100. The solid line represents a Hill model fit. Inset: corresponding variation in fluorescence spectra. (b) Average lifetime of Dil versus TX-100. Inset: corresponding time-resolved fluorescence decays. Solid lines represent bi-exponential fits. (c) Top: FCS cross-correlation curves (solid lines), fits (dashed lines) (inset: zoomed in) associated with 100 nm- (black) and 200 nm-sized (red) vesicles. Also shown are FCS curves for 200 nm-sized vesicles in the presence of 0.1 mM (green) and 0.2 mM (blue) TX-100. Bottom: residuals of the fits. (d) Top: Diffusion times of NBD-PC labeled vesicles as a function of TX-100. A PC: PS lipid ratio of 65: 35 was used under all conditions. Solid lines represent linear fits. Bottom: corresponding residuals.

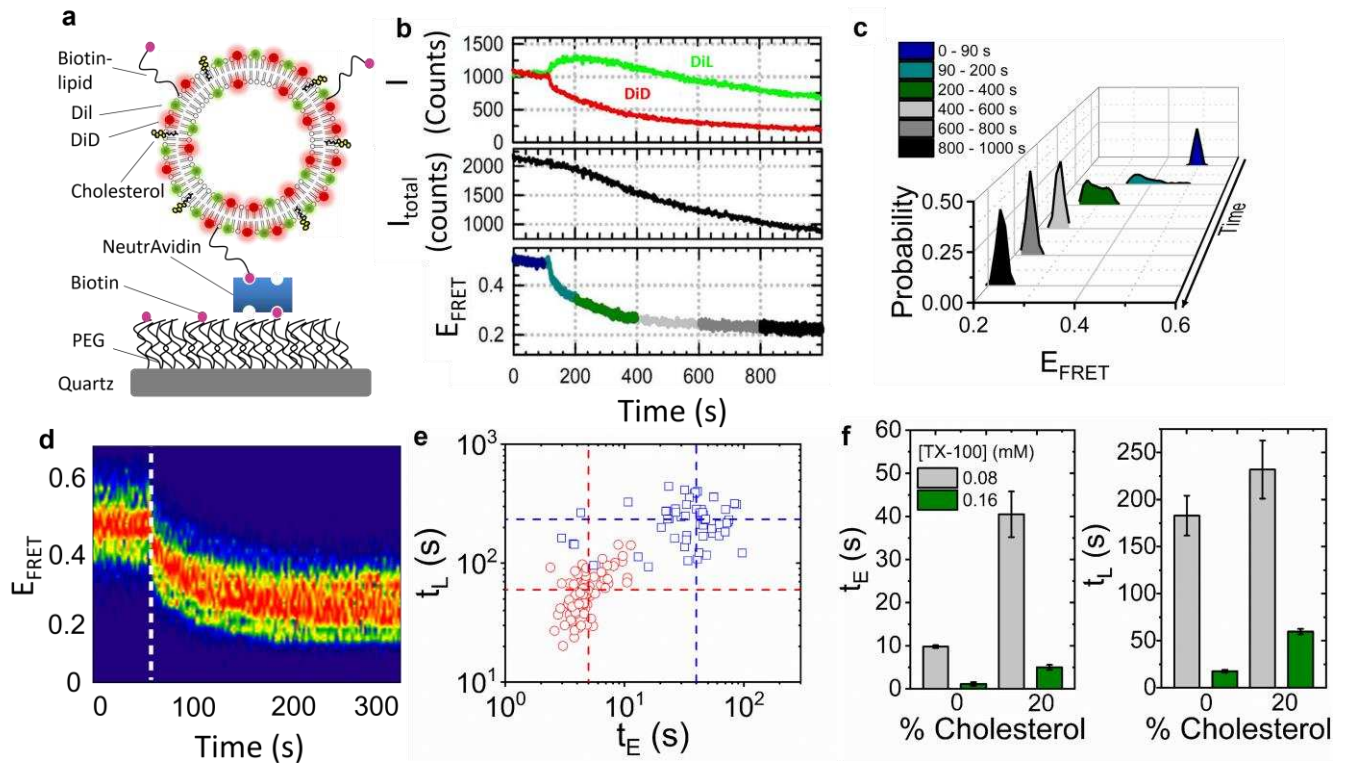


Figure 2. Real-time visualization of solubilization kinetics by svFRET. (a) Schematic of the immobilization scheme. The schematic is for illustration purposes only and is not to scale. (b) Representative variation in the fluorescence emission of Dil and DiD (top panel), the sum of their intensities (middle panel) and the corresponding variation in FRET efficiency obtained before (< 90 s) and after (>90 s) injection of 0.16 mM TX-100. (c) Relative FRET state occupancies observed over 1000 s. (d) FRET contour plot showing the variation in E_{FRET} before and after TX-100 injection (dashed white line) ($N = 105$). (e) Corresponding scatter plot of expansion half-life, t_E versus that of lysis, t_L obtained after injection of 0.08 mM (blue) and 0.16 mM (red) TX-100. Dashed lines represent the center of each distribution. (f) Comparative bar plots summarizing the variation in t_E and t_L as a function of TX-100 and percentage of cholesterol incorporated within the vesicle bilayer. Error bars indicate the standard error of the mean.

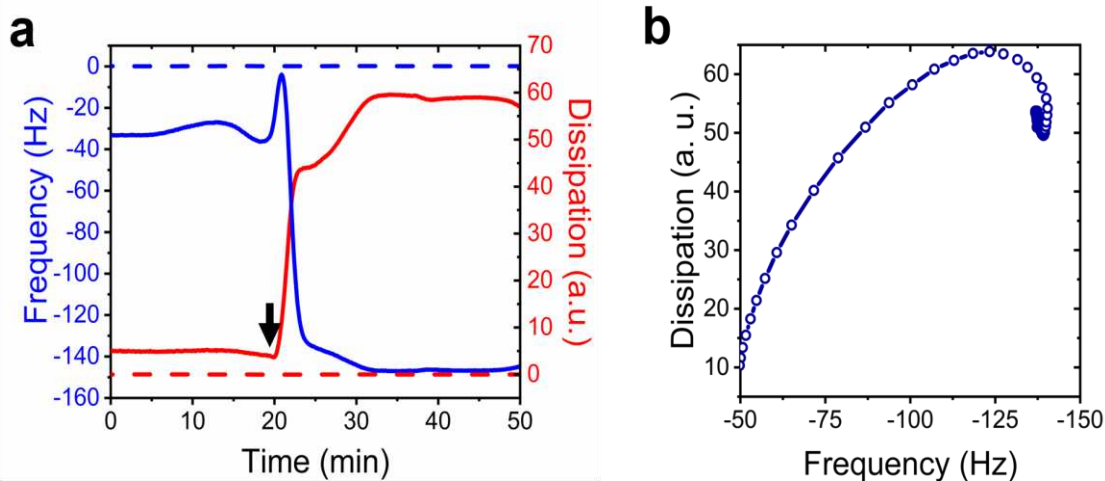


Figure 3. TX-100 induced vesicle solubilization monitored by QCM-D. (a) Variation in frequency (blue) and dissipation (red) of the 7th overtone associated with surface immobilized vesicles in the presence of TX-100. The dashed lines represent data collected from a control sensor pre-treated with PEG and NeutrAvidin, but lacks vesicles. The arrow indicates the time-point of the solubilization. (b) Frequency versus dissipation observed during the interaction between surface immobilized vesicles and TX-100.

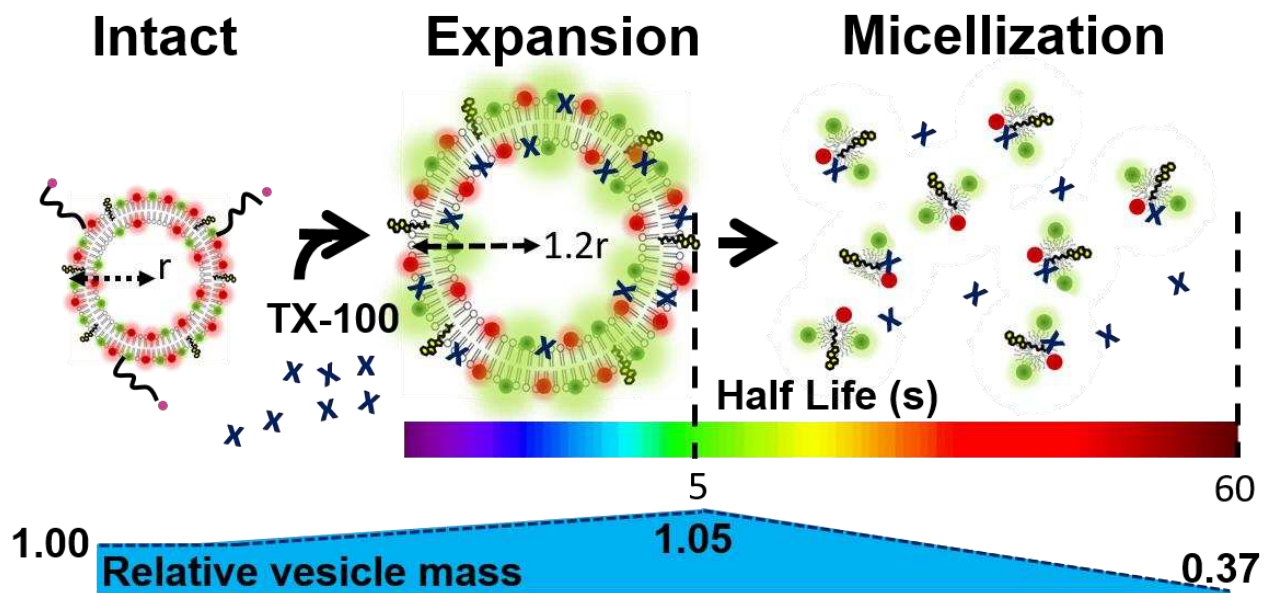


Figure 4. Mechanism of TX-100 induced vesicle solubilization. Detergent molecules approach lipid vesicles inducing a fast conformational expansion prior to lysis and the release of mixed-micelles into solution.