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# Enhancing Efficiency in a Dye-Sensitized Solar Cell Based on a Metal Oxide Nanocomposite Gel Polymer Electrolyte

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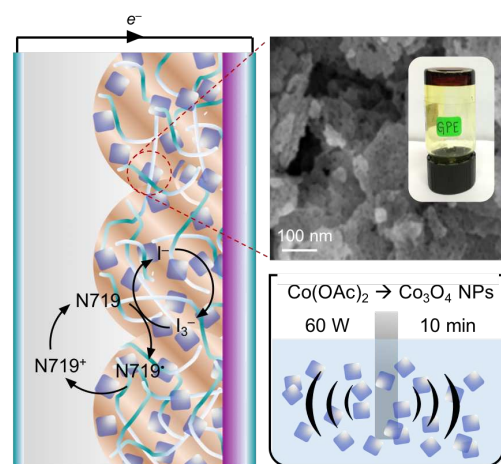
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**KEYWORDS** Gel Polymer Electrolytes, Dye-Sensitized Solar Cells, Metal Oxide Nanoparticles, Current Density, Efficiency

**ABSTRACT:** In order to overcome the critical limitations of liquid electrolyte-based dye-sensitized solar cells, quasi-solid-state electrolytes have been explored as a means of addressing long-term device stability, albeit with comparatively low ionic conductivities and device performances. Although metal oxide additives have been shown to augment ionic conductivity, their propensity to aggregate into large crystalline particles upon high-heat annealing hinders their full potential in quasi-solid-state electrolytes. In this work, sonochemical processing has been successfully applied to generate fine  $\text{Co}_3\text{O}_4$  nanoparticles that are highly dispersible in a PAN:P(VP-co-VAc) polymer blended gel electrolyte, even after calcination. An optimized nanocomposite gel polymer electrolyte containing 3 wt% sonicated  $\text{Co}_3\text{O}_4$  nanoparticles (PVVA-3) delivers the highest ionic conductivity ( $4.62 \times 10^{-3} \text{ S cm}^{-1}$ ) of the series. This property is accompanied by a 51% enhancement in the apparent diffusion coefficient of triiodide versus both unmodified and unsonicated electrolyte samples. The dye-sensitized solar cell based on PVVA-3 displays a power conversion efficiency of 6.46% under AM1.5G,  $100 \text{ mW cm}^{-2}$ . By identifying the optimal loading of sonochemically processed nanoparticles, we are able to generate a homogenous extended particle network that effectively mobilizes redox active species through a highly amorphous host matrix. This effect is manifested in a selective 51% enhancement in photocurrent density ( $J_{\text{sc}} = 16.2 \text{ mA cm}^{-2}$ ) and a lowered barrier to N719 dye regeneration ( $R_{\text{CT}} = 193 \Omega$ ) versus an unmodified solar cell. To the best of our knowledge, this work represents the highest known efficiency to-date for dye-sensitized solar cells based on a sonicated  $\text{Co}_3\text{O}_4$ -modified gel polymer electrolyte. Sonochemical processing, when applied in this manner, has the potential to make meaningful contributions towards the ongoing mission to achieve the widespread exploitation of stable and low-cost dye-sensitized solar cells.



## 1. INTRODUCTION

Ever since the first report of dye-sensitized solar cells (DSSCs) by O'Regan and Grätzel<sup>1</sup> in 1991, energy conversion efficiencies have experienced a steady upward trend, approaching 14% for liquid electrolyte (LE)-based cells.<sup>2</sup> Though they still lag in terms of efficiency compared to more other solid-state photovoltaics, the theoretical price-performance ratio of these simple thin-film devices make DSSCs highly attractive candidates for achieving global grid parity<sup>3</sup> against fossil fuels. Their wider scale implementation has been delayed, however, by a number of critical issues including degradation, cell leakage, dye desorption, thermal instability, and electrode corrosion—

all of which can be traced back to the use of traditional LEs.<sup>4-7</sup> As a result, research efforts have been focused on replacing conventional LE media with other suitable alternatives, such as ionic liquids<sup>8-13</sup>, solid-state perovskites<sup>14-17</sup>, and solid/quasi-solid-state polymer gels and their composites.<sup>8-34</sup>

Gel polymer electrolytes (GPEs) have received considerable attention in recent years due to their straightforward formulation into thermally and mechanically robust solids and quasi-solids that overcome many drawbacks of LEs.<sup>7</sup> To date, optimized DSSCs based on GPEs can achieve efficiencies as high as 9.61%.<sup>33</sup> Poly(acrylonitrile) (PAN) is a popular component of GPEs that typically provides a high

Table 1. Literature Performance Parameters<sup>a</sup> of DSSCs Based on Metal Oxide-Nanocomposite GPEs

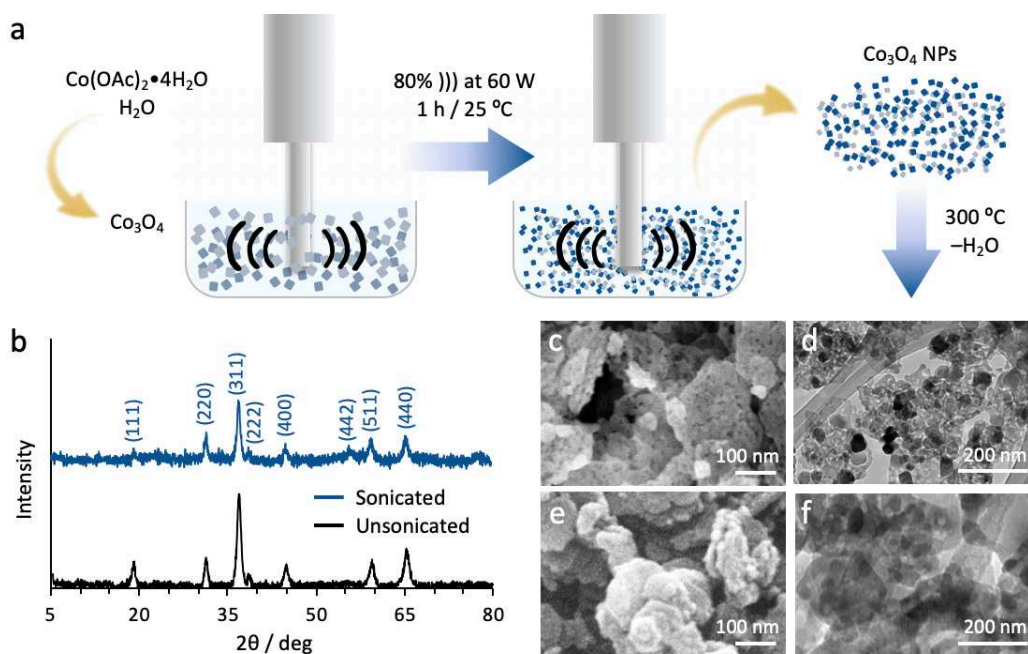
Additive <sup>b</sup> (wt%)	GPE Matrix <sup>c</sup> (Solvent <sup>d</sup> )	$\sigma / 10^{-4} \text{ S cm}^{-1}$	$J_{sc} / \text{ mA cm}^{-2}$	$V_{oc} / \text{ V}$	$\eta / \%$	ref
TiO <sub>2</sub> (–) <sup>e</sup>	PEO/TBP/LiI/I <sub>2</sub> (EC:PC:MeCN)	— <sup>e</sup>	3.77	0.61	5.05	54
TiO <sub>2</sub> (2)	PMMA/KI/I <sub>2</sub> (PC)	1.75	6.08	0.62	2.34	55
TiO <sub>2</sub> (2.5)	Agarose/LiI/I <sub>2</sub> (NMP)	4.40	10.96	0.55	4.74	46
TiO <sub>2</sub> (0.5)	PVDF-HFP (MPN)	21.2	12.82	0.68	5.19	54
TiO <sub>2</sub> (7)	PEO:PVDF-HFP/LiI/I <sub>2</sub> (DMF)	7.24	6.37	0.71	2.80	63
TiO <sub>2</sub> (10)	PEO/KI/I <sub>2</sub> (MeCN)	1.35	3.71	0.76	1.68	45
TiO <sub>2</sub> (15)	PMII/TBP/LiI/I <sub>2</sub> (MeCN)	— <sup>e</sup>	9.00	0.61	3.74	48
SiO <sub>2</sub> (–) <sup>e</sup>	PEO/TBP/LiI/I <sub>2</sub> (EC:PC:MeCN)	— <sup>e</sup>	4.19	0.66	6.85	54
SiO <sub>2</sub> (2.8)	PEGDME-150/TBP/NaI/I <sub>2</sub>	4.45	19.3	0.69	6.30	54
SiO <sub>2</sub> (3)	PMII/TBP/LiI/I <sub>2</sub> (MeCN)	— <sup>e</sup>	8.80	0.62	3.63	48
SiO <sub>2</sub> (7)	PAN:PVDF/HMII/TBP/LiI/I <sub>2</sub> (EC:PC)	696	11.60	0.79	5.61	53
SiO <sub>2</sub> (10)	PU:PPy/NaI/I <sub>2</sub> (PC)	2.68	3.70	0.63	2.49	57
SiO <sub>2</sub> (10)	PU:PANI/NaI/I <sub>2</sub> (PC)	2.77	3.88	0.72	3.10	57
ZnO (–) <sup>e</sup>	PEO/KI/I <sub>2</sub> (MeCN)	— <sup>e</sup>	2.10	0.63	1.70	56
ZnO (3)	PVDF:PEO/NaI/I <sub>2</sub> (EC:PC)	83.6	19.5	0.62	7.33	50
ZnO (3)	PMMA/KI/I <sub>2</sub> (PC)	1.84	6.05	0.61	2.38	55
ZnO (35)	PMII/TBP/LiI/I <sub>2</sub> (MeCN)	— <sup>e</sup>	9.00	0.70	4.17	48
NiO (3)	Agarose/LiI/I <sub>2</sub> (NMP)	33.3	6.20	0.63	2.02	67
Fe <sub>2</sub> O <sub>3</sub> (7)	PAN:PVDF/ HMII/TBP/LiI/I <sub>2</sub> (EC:PC)	681	10.40	0.77	4.90	52
Co <sub>3</sub> O <sub>4</sub> (3)	Agarose/LiI/I <sub>2</sub> (NMP)	42.0	3.82	0.69	1.44	70
Co <sub>3</sub> O <sub>4</sub> (5)	Agarose/LiI/I <sub>2</sub> (NMP)	36.6	3.01	0.64	1.00	70

<sup>a</sup>Best performance parameters reported for TiO<sub>2</sub>/Ru dye-based DSSCs using nanocomposite GPEs measured at room temperature.

<sup>b</sup>Metal oxide additives commercially obtained or prepared *via* conventional hydro/solvothermal syntheses, followed by high-temperature activation. Values for wt% reflect the nanoparticle solid content loading of the GPE. <sup>c</sup>TBP: *tert*-butylpyridine; PVDF-HFP: poly(vinylidene fluoride-*co*-hexafluoropropylene); PMII: 1-propyl-3-methylimidazolium iodide; PEGDME: poly(ethylene glycol) dimethyl ether; PPy: polypyrrole; PANI: polyaniline. <sup>d</sup>NMP: *N*-methyl-1-pyrrolidinone; MPN: 3-methoxypropionitrile; EC: ethylene carbonate; PC: propylene carbonate; MeCN: acetonitrile; DMF: *N*-dimethylformamide. <sup>e</sup>Data unavailable.

structural and thermochemical stability.<sup>35–38</sup> Yet, it is practically inactive towards ionic conductivity as a result of –CN-induced crystallization that tends to increase overall GPE resistance to ion transport.<sup>39</sup> Copolymerization and/or physical blending of PAN<sup>40–43</sup> with other polymer chains such as poly(ethylene oxide) (PEO), poly(methylmethacrylate) (PMMA), and poly(vinylacetate) (PVAc) help to overcome its shortcomings by promoting an amorphous quasi-solid state. Indeed, the spontaneous assembly of blended and block co-polymers into a variety of nanoscale morphologies can naturally facilitate ionic conduction in PAN-based GPEs.<sup>44</sup> Weak binding interactions of salt additives with functional groups on the ion-conducting polymer component of these GPEs can also improve device performance by maintaining a low conduction band edge in TiO<sub>2</sub>-based DSSCs.<sup>18,22,26</sup> Likewise, the introduction of nanomaterial additives such as metal

oxides<sup>45–57</sup>, mesoporous particles<sup>58,59</sup>, nanoclays<sup>60–62</sup> and carbon nanomaterials<sup>63–68</sup> at low (e.g., <5 wt%) concentrations has been proven to augment ionic conductivity in nanocomposite GPEs, primarily through the formation of a conductive network than spans the polymer matrix. Metal oxide nanoparticles (e.g., TiO<sub>2</sub>, ZnO, NiO, Co<sub>3</sub>O<sub>4</sub>, Fe<sub>2</sub>O<sub>3</sub>, etc.) are particularly effective at discouraging the formation of polymer crystalline domains, whilst accelerating charge transport dynamics along their surfaces. Recently, it was also shown that the addition of magnetic nanoparticle additives like Co<sub>3</sub>O<sub>4</sub>, Fe<sub>3</sub>O<sub>4</sub> or NiO can even enable external control over DSSC properties, i.e., *via* magnetic field stimuli.<sup>69,70</sup> In these examples, DSSC efficiencies nearly double as a result of metal oxide alignment within GPE matrices with the magnetic field axis. See Table 1 for a summary of performance parameters for DSSCs based on GPE-oxide nanocomposites.



**Figure 1.** (a) Sonochemical processing followed by high temperature calcination at 300 °C yields a finer crop of microcrystalline  $\text{Co}_3\text{O}_4$  nanoparticles that readily index by (b) powder XRD ( $\lambda = 1.5418 \text{ \AA}$ ) to the (111), (220), (311), (222), (400), (442), (511) and (440) reflections anticipated for the cubic spinel structure of  $\text{Co}_3\text{O}_4$  (in agreement with JCPDS Card No. 76-1802). (c) FESEM imaging of sonochemically processed and heat-treated  $\text{Co}_3\text{O}_4$  nanoparticles reveals a flake-like porous structure caused by Ostwald ripening of individual 40 nm-sized nanocuboids. (d) HR-TEM image of sonochemically processed  $\text{Co}_3\text{O}_4$ . By comparison, (e) FESEM and (f) HR-TEM images of nanoparticles prepared without sonication form compact flakes (45–136 nm) with no apparent pore structure.

Unfortunately, such affordable metal oxide additives still come with their own limitations. Thermal activation at high temperatures ( $\geq 300 \text{ }^\circ\text{C}$ ) is required to remove water (i.e., from hydrothermal syntheses), minimise hydroxide intermediates, and create good contact between fused nanoparticles for improved ionic conduction. However, this treatment also encourages particle agglomeration, thus making their uniform dispersion more difficult to achieve in GPEs. Relatively large particle agglomerates adversely affect DSSC performance<sup>71–73</sup> by trapping redox species and ultimately limiting the rate at which the sensitized dye can be regenerated. Resolving the aggregation of thermally activated nanoparticles is, therefore, a highly relevant task in achieving the overall mission of extending the long-term stability and performance of GPE-based DSSCs.

We identified sonochemical processing as a simple and straightforward technique that could potentially mitigate the negative effects of high temperature calcination on metal oxide nanoparticles. Acoustic cavitation has already proven to be useful for controlling particle morphologies, achieving high surface areas and inherent porosity in metal oxide materials.<sup>74–78</sup> This approach has also afforded carbon nanocomposite materials with high dispersity in various media.<sup>79–81</sup> Despite these advances, however, the technique has not been applied towards metal oxide additives in GPEs for DSSCs. We hypothesized that the use of high-power ultrasound during the solvothermal synthesis of metal oxide nanoparticles would result in finer colloids that consequently yield smaller, more easily dispersed particle agglomerates upon calcination. As a result, these sonochemically processed additives would boost the performance of GPE nanocomposite-based DSSCs.

In this work, we apply the principle of sonochemical synthesis to generate metal oxide nanoparticles that, upon annealing at 300 °C, undergo natural Ostwald ripening to form relatively smaller aggregates compared to conventional (typically unsonicated) bulk metal oxides. In our experience, we have found that an annealing temperature of 300 °C is sufficiently high enough to efficiently and effectively drive off residual moisture (i.e.,  $\text{H}_2\text{O}$ ) trapped in and between metal oxide nanoparticles during the hydrothermal synthesis and washing steps. This condition is not dissimilar to temperatures (ranging from as low as 250 °C up to 1000 °C) applied by others to eliminate  $\text{H}_2\text{O}$  impurities.<sup>82–85</sup> Different wt% loadings of sonicated nanoparticles were incorporated into a PAN host matrix blended with an ion conductive poly(vinylpyrrolidone-*co*-vinylacetate)<sup>29,81</sup> (P(VP-*co*-VAc);  $10^{-3} \text{ S cm}^{-1}$ ) and their effects on ionic conductivity and DSSC performance were measured. We have shown that when optimized, our  $\text{Co}_3\text{O}_4$ -modified GPE (i) displays an enhancement of the apparent diffusion coefficient of triiodide ( $\text{I}_3^-$ ) against an unmodified control GPE, (ii) raises photoconversion efficiency of DSSC devices by selectively enhancing photocurrent density, and (iii) extends their performance stability. Our results imply that a sonochemical processing strategy can broadly impact the development of composite GPEs for enhanced energy device performance.

## 2. EXPERIMENTAL SECTION

**2.1. Materials.** P(VP-*co*-VAc) (weight-average molecular weight ( $M_w$ )  $\sim 50,000 \text{ g mol}^{-1}$  by GPC vs. PEO), vinylpyrrolidone/vinyl acetate = 1.3/1 mole ratio), PAN ( $M_w \sim 150,000 \text{ g mol}^{-1}$  by GPC vs. PEO), cobalt(II) acetate tetrahydrate ( $\text{Co}(\text{OAc})_2 \cdot 4\text{H}_2\text{O}$ ; 99.9% trace metal basis), NaOH,

PC (99.7%, anhydrous), EC (98%, anhydrous), poly(ethylene glycol *t*-octylphenyl ether) (Triton™ X-100, non-ionic), TiO<sub>2</sub> (Aeroxide® P25 (21 nm; ≥99.5%) and Aeroxide® P90 (14 nm; ≥99.5%)), and *cis*-diisothiocyanatobis(2,2'-bipyridyl-4,4'-dicarboxylato)ruthenium(II) bis-(tetrabutylammonium) (N719; 95%) dye were used as received from Sigma-Aldrich, USA. NaI, I<sub>2</sub> and HNO<sub>3</sub> (65% v/v aqueous solution) were used as received from Friendmann Schmidt, Australia. Fluorine-doped tin oxide (FTO) conducting glass plates (sheet resistance = 8 Ω sq<sup>-1</sup>; Solaronix, Switzerland) were cut into 2 × 2 cm sheets and used as substrates for the fabrication of photoanodes and platinum (Pt) counter electrodes. Tetramethylsilane (TMS) and absolute ethanol (EtOH) was used as received from Fluka, USA.

**2.2. Preparation of Co<sub>3</sub>O<sub>4</sub> Nanoparticles.** This study took advantage of sonochemical processing, followed by calcination, to produce uniformly sized nanoparticles of cobalt oxide (Co<sub>3</sub>O<sub>4</sub>). A 2 M aqueous solution of NaOH (20 mL) was added dropwise to a stirring solution of Co(OAc)<sub>2</sub>•4H<sub>2</sub>O (249 mg, 1.0 mmol) in deionized water (30 mL) under simultaneous probe ultrasonication (120 W, 60% maximum amplitude; Fisher Scientific, USA). Afterwards, the resulting blue-green solution was probe sonicated for 1 h, during which time a maximum internal temperature of 60 ± 5 °C was recorded and a fine black precipitate was formed. The colloidal mixture was spun down by centrifugation (6000 rpm for 15 min; Hettich, Germany), the supernatant decanted, and the particles re-suspended in deionized water (12 mL). This washing protocol was repeated several times (*ca.* 5–6 times) until the supernatant was observed to be colorless. After removal of the supernatant, the solid was dried in an oven at 90 °C for 8 h before being ground into a fine powder by mortar and pestle and calcined in an electric muffle furnace (JS Research Inc., Korea) under air at 300 °C for 3 h to afford activated Co<sub>3</sub>O<sub>4</sub> nanoparticles with an average particle size of 40.9 ± 0.3 nm. See *Supporting Information* for details on the preparation unsonicated Co<sub>3</sub>O<sub>4</sub> nanoparticles. The surface morphology of as-prepared Co<sub>3</sub>O<sub>4</sub> nanoparticles were examined using field emission scanning electron microscopy (FESEM; JEOL JSM-7600F, Japan) and high-resolution transmission electron microscopy (HR-TEM; Hitachi-7100, Japan). Structural crystallinity and phase purity were confirmed by powder X-ray diffraction (XRD) analysis using a Empyrean X-ray diffractometer (PANalytical, USA) at 25 °C with Cu Kα radiation (45 kV, 40 mA; λ = 1.5418 Å) in the 2θ range from 5° to 80° at 0.1° step intervals and a 10 s count time.

**2.3. Preparation of Co<sub>3</sub>O<sub>4</sub>-Modified GPEs.** The GPEs investigated in this study were prepared using a 'PVVA' electrolyte formulation previously reported by us<sup>81,86</sup> that has been optimized with respect to the NaI/I<sub>2</sub> redox couple (see the *Supporting Information* for further details.) Varying amounts of either sonicated or unsonicated Co<sub>3</sub>O<sub>4</sub> nanoparticles (*i.e.*, 0, 1, 3, or 5 wt% Co<sub>3</sub>O<sub>4</sub> with respect to total solid contents) were added to a 1:1 w/w (2 g) PC:EC solvent mixture containing a 10:1 w/w ratio (200 mg) of the NaI:I<sub>2</sub> redox couple. The black suspension was sonicated while stirring for 30 min to ensure the complete and uniform dispersion of Co<sub>3</sub>O<sub>4</sub> before finally adding a

1:1 w/w blend (300 mg) of PVP-*co*-PVAc and PAN. The resulting mixture was stirred at 100 °C until the polymer components were fully dissolved, then cooled back down to room temperature, triggering gelation. The resulting GPEs were used directly in electrochemical studies and DSSC fabrication without additional treatment.

The extent of GPE crystallinity was determined using powder XRD analysis, while complexation of the polymer matrix was investigated by absorbance-mode Fourier transform infrared (FTIR) spectroscopy on a Nicolet™ iS10 FTIR spectrometer (Thermo Fisher Scientific, USA) equipped with a Smart iTR™ accessory. Solid-state cross-polarized magic angle spinning (CPMAS) nuclear magnetic resonance (NMR) spectra were recorded using a 400 MHz Bruker Avance III HD SS-NMR spectrometer (Germany) at 25 °C. GPE films were cut into approximately 1.0 mm<sup>2</sup> pieces using a clean razor blade and packed into a 4 mm (∅) CPMAS rotor. These NMR samples were then spun at 54.71° (the 'magic angle') at a spin rate of 10 Hz. <sup>13</sup>C CPMAS NMR spectra were obtained at frequencies of 100.63 MHz, respectively, using a 1 ms contact time and 2 s recycle delay and referenced with respect to neat TMS (δ<sub>c</sub> = 0 ppm) by setting the high-frequency signal of adamantane to a δ<sub>c</sub> of 38.5 ppm. <sup>23</sup>Na CPMAS NMR spectra were recorded at a frequency of 105.85 MHz using direct excitation with a 0.1 s recycle delay and referenced with respect to a 1 M aqueous solution of NaCl (δ<sub>Na</sub> = 0 ppm).

Ionic conductivities of the GPEs were determined by electrochemical impedance spectroscopy (EIS; HIOKI LCR Hi-Tester Model 3532-50, Japan) at 25 °C at an AC voltage of 10 mA within a frequency range of 50 Hz to 5 MHz. GPE samples were sandwiched between two stainless steel blocking electrodes (hole area, *S* = 2.0 cm<sup>2</sup>) to a thickness, *l*, of 2.9 mm. Temperature-dependent EIS measurements were carried out on GPE samples from 30 to 100 °C in 10 °C intervals. The effective diffusion coefficient of triiodide ions (*D*<sub>app</sub><sup>I<sub>3</sub></sup>) in GPEs was determined by carrying out steady-state linear-sweep voltammetry (LSV) experiments with a potentiostat (PGSTAT-128N; Metrohm Autolab, The Netherlands) at a scan rate of 10 mV s<sup>-1</sup> between ±700 mV at 25 °C. The dummy cell used in LSV studies consisted of GPE samples sandwiched between two Pt counter electrodes (hole area, *S* = 0.2 cm<sup>2</sup>) in a Pt|GPE|Pt configuration with a spacer thickness, *l*, of 48 μm.

**2.4. DSSC Fabrication and Characterization.** See *Supporting Information* for details on preparing the TiO<sub>2</sub> photoanode. DSSCs were fabricated by sandwiching the desired GPE sample between a 12 μm-thick TiO<sub>2</sub> photoanode (0.10 cm<sup>2</sup> active area) and a Pt counter electrode. The height of the GPE layer was controlled using a spacer of 48 μm thickness. Photovoltaic DSSC performance and current density-voltage (*J*-*V*) curves were measured using a 'Class A' quality solar simulator (AM1.5G, Xe lamp; Newport LCS-100 Series, USA) coupled to a potentiostat. EIS data of fabricated DSSCs were recorded between 0.1 Hz and 100 kHz at an open-circuit AC potential of 10 mV using a frequency response analyzer (Metrohm Autolab, The Netherlands) and fitted to an equivalent circuit model using the Metrohm Autolab Nova software. EIS data was recorded under illumination at an applied bias of 0.55 V and 10 mV s<sup>-1</sup> scan rate. The illumination intensity

(100 mW cm<sup>-2</sup>) used in all DSSC measurements was calibrated with a solar power meter (TENMARS TM-207; Taiwan) using a standard Si cell. The active area of DSSC devices was fixed to 0.10 cm<sup>2</sup>.

### 3. RESULTS AND DISCUSSION

#### 3.1. Sonochemically Synthesized Co<sub>3</sub>O<sub>4</sub> Nanoparticles.

We originally envisioned that the addition of well-defined porous Co<sub>3</sub>O<sub>4</sub> nanoparticles would encourage fast ion diffusion and conductivity in GPE-based DSSCs by enhancing interactions in the GPE matrix and yielding lower-energy charge-transport pathways. Sonochemical processing, followed by calcination under air at 300 °C for three hours, proved to be an excellent method for achieving the reliable synthesis of crystalline nanoporous Co<sub>3</sub>O<sub>4</sub> nanocubes (Figure 1a). Powder XRD analysis of sonicated and unsonicated Co<sub>3</sub>O<sub>4</sub> nanoparticles post-calcination display similar sets of diffraction peaks located at 19.1, 31.5, 36.9, 38.6, 44.9, 55.8, 59.6 and 65.5° 2θ that clearly index (Figure 1b) to the reflections of a face-centered cubic spinel structure for bulk crystalline Co<sub>3</sub>O<sub>4</sub>, in agreement with standard (JCPDS No. 76-1802) values. Considering the FWHM of the powder diffraction peaks, the average crystallite size of the sonicated and unsonicated Co<sub>3</sub>O<sub>4</sub> nanoparticles can be calculated using Scherrer's formula<sup>87,88</sup>:

$$B_{2\theta\max} = K\lambda / \beta \cos\theta \quad (1)$$

where  $K$  is unitless Scherrer or shape constant,  $\lambda$  is the diffraction wavelength (= 1.5418 Å), and  $\beta$  is the FWHM at angle 2θ in radians.  $K$  is assumed to be 0.94 for the FWHM of a crystal with cubic symmetry. By this method, sonicated Co<sub>3</sub>O<sub>4</sub> nanoparticles in this study are found to possess smaller crystallite sizes (ca. 10.2 nm) compared to Co<sub>3</sub>O<sub>4</sub> nanoparticles prepared without sonication (ca. 11.7 nm). These calculations imply that the applied acoustic cavitation was effective at limiting the size of metal oxide nanoparticles and agglomerates post-calcination. Comparing the powder XRD pattern of bulk Co<sub>3</sub>O<sub>4</sub> with that of the sonicated Co<sub>3</sub>O<sub>4</sub> nanoparticles, we can see that the latter appears slightly broadened, which can be attributed to the mechanical tension associated with the high-heat extraction of solvent from the particle surface. Even so, the effect of thermal annealing on the overall crystallinity of Co<sub>3</sub>O<sub>4</sub> nanoparticles appears to be negligible.

Electron micrographs of sonicated Co<sub>3</sub>O<sub>4</sub> (Figures 1c–f) reveal that the calcination step is critical for fusing sonicated Co<sub>3</sub>O<sub>4</sub> nanoparticles into a porous network morphology. FESEM images (Figure 1c of drop-cast aliquots show the porous flake-like structure of annealed Co<sub>3</sub>O<sub>4</sub> nanoparticles that appear to be composed of smaller Co<sub>3</sub>O<sub>4</sub> nanocuboids (40.9 ± 0.3 nm). HR-TEM images (Figure 1d) of the Co<sub>3</sub>O<sub>4</sub> nanoparticles further illustrate the cuboidal shapes observed by FESEM. Compared to other sonochemically prepared metal oxides reported in the literature,<sup>72–76</sup> the nanocuboids obtained *via* sonochemical processing occupy a rather narrow size distribution. In contrast, FESEM and HR-TEM images of the unsonicated Co<sub>3</sub>O<sub>4</sub> (Figures 1e,f) show densely compacted nanoparticles with no apparent porous structure. Indeed, Brunauer–Emmett–Teller (BET) surface area analysis of the two samples (see Supporting Information, Figure S1) reveals

N<sub>2</sub> adsorption–desorption isotherms with a Type IV hysteresis between P/P<sub>0</sub> = 0.9 and 1, indicating their mesoporous natures. We determined BET surface areas of 80.909 and 65.168 m<sup>2</sup>/g for the sonicated and unsonicated Co<sub>3</sub>O<sub>4</sub> nanoparticles, respectively, thus providing evidence of enhanced surface area and total pore volume as a result of sonochemical processing. We surmise that the higher surface area nanoparticles will display excellent dispersion and contact with the polymer matrix to facilitate ion transport within GPEs.

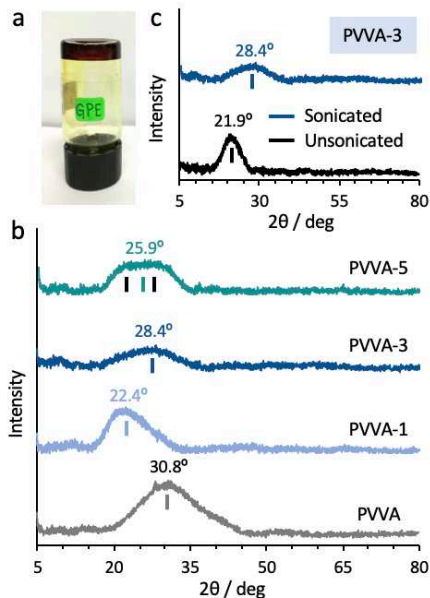
**3.2. Co<sub>3</sub>O<sub>4</sub>-Modified GPEs.** We prepared GPE samples PVVA, PVVA-1, PVVA-3, and PVVA-5 consisting of 0, 1, 3, 5 wt% of sonicated Co<sub>3</sub>O<sub>4</sub> solid content, respectively. Please see the *Experimental Section* and *Supporting Information* for further details on GPE optimization. GPE samples consisted of Co<sub>3</sub>O<sub>4</sub> nanoparticles dispersed in a matrix containing a 50:50 ratio of a 1:1 w/w PAN:P(VP-co-VAc) blend and the NaI/I<sub>2</sub> redox couple along with PC and EC as small molecule plasticizers. Room temperature gelation of the viscous mixtures resulted afforded GPEs (Figure 2a) that were pliable yet mechanically robust thanks to the structural stability offered by the PAN component. Increasing the sonicated Co<sub>3</sub>O<sub>4</sub> content beyond 5 wt% was found to impede gelation, presumably due to a higher concentration of nanoparticle aggregates interfering with the association of polymer chains needed to achieve a stable quasi-solid state.

Brief inspection of the FTIR spectra (Figures S2) reveals particularly strong stretching frequencies associated with the PC and EC plasticizers for all GPE samples. A notable shift (Figure S3) in the diagnostic lactam C=O stretch of the P(VP-co-VAc) component from 1680 cm<sup>-1</sup> in the pure polymer to between 1657 and 1669 cm<sup>-1</sup> for GPEs is indicative of their increased *lp* interactions within the host polymer matrix. Indeed, *N*-tertiary pyrrolidinones such as P(VP-co-VAc) have a propensity to form diverse multivalent complexes with alkali metal cations, even resulting in linear coordination polymers.<sup>89</sup> Relative to PVVA, bathochromic shifts of the GPE lactam C=O stretches suggest that metal complexation is taking place, i.e., through an amide –O···Na<sup>+</sup> interaction. Minor variations in the position of the C=O stretch may indicate the relative strength of cation–*lp* interaction taking place as a function of Co<sub>3</sub>O<sub>4</sub> nanoparticle loading. Line broadening effects observed by solid-state <sup>23</sup>Na CPMAS NMR spectroscopy performed on GPE samples (Figures S4 and S5) also support the notion weak complexes are formed between Na<sup>+</sup> and pendant lactam/ester carbonyl groups of the P(VP-co-VAc) component. These dynamic interactions are anticipated to lower the activation barrier for ion transport to occur.

Powder XRD analysis (Figure 2b) of Co<sub>3</sub>O<sub>4</sub>-modified GPEs more clearly reveals the impact of sonicated nanoparticles on polymer chain interactions in the quasi-solid state. The PVVA matrix is characterized by a single broad diffraction peak spanning 25° and centered at 2θ = 30.8°, confirming its highly amorphous nature. Powder XRD spectra of the sonicated Co<sub>3</sub>O<sub>4</sub>-modified GPEs PVVA-1, PVVA-3 and PVVA-5 display relatively narrowed diffraction signals situated at lower angles of 2θ (maxima centered around 22.4°, 28.4° and 25.9°, respectively) relative to PVVA. This trend towards lower angles of 2θ results from an increase in the interlayer spacing  $d$  (i.e., of Bragg's relation)

between (micro)ordered polymer chains within the matrix, which would be consistent with the intercalation of nanoparticle additives. Increasing the solid content of sonicated  $\text{Co}_3\text{O}_4$  nanoparticles from 1 wt% to 3 wt% (i.e., PVVA-3) reduces the degree of crystallinity in GPEs, as evidenced by broadening of the overall peak area centered at  $2\theta = 28.4^\circ$  and a reduction main peak intensity. Conversely, raising the solid content of sonicated  $\text{Co}_3\text{O}_4$  to 5 wt% in PVVA-5 seems to re-establish the main peak area and intensity, suggesting that nanoparticles coalesce to form more crystalline domains within the GPE matrix at (and above) this concentration threshold. On the whole, these data reveal that the introduction of 3 wt% sonicated  $\text{Co}_3\text{O}_4$  to PVVA-based GPEs achieves a ‘sweet spot’ with respect to (i) retaining the most amorphous GPE phase whilst (ii) reducing the distance between polymer chains (i.e.,  $d = 3.14 \text{ \AA}$ ) to overall enhance ion transport and mobility in the matrix. Comparison with the homologous GPE consisting of 3 wt% *unsonicated*  $\text{Co}_3\text{O}_4$  nanoparticles (Figure 2c) reveals a relatively sharper peak at  $21.5^\circ 2\theta$ , highlighting the role of sonochemical processing to avoid the formation of higher crystalline matrixes.

FESEM micrographs taken of the homologous GPE films (Figure S6) reveal that sonochemical processing leads to finer  $\text{Co}_3\text{O}_4$  nanoparticles dispersions and a more uniform polymer matrix in PVVA-3, whereas crystallite domains can be clearly identified in the FESEM image of the film containing Unsonicated nanoparticles. Such bridging interactions between nanoparticles and the host matrix are thought to encourage cooperative ion complexation along the polymer backbones in order to achieve more efficient charge transport in the solid state. From these initial data,



**Figure 2.** (a) A typical GPE used in this study. (b) Powder X-ray diffraction analysis of GPEs consisting of 0, 1, 3, 5 wt% of sonicated  $\text{Co}_3\text{O}_4$  (i.e., PVVA, PVVA-1, PVVA-3, and PVVA-5, respectively) reveals that a 3 wt% loading provides an optimum balance of low matrix crystalline phase and close interlayer spacing ( $d = 3.14 \text{ \AA}$ ) between polymer chains. On the other hand, (c) a homologous control GPE of PVVA-3 bearing 3 wt% of *unsonicated*  $\text{Co}_3\text{O}_4$  nanoparticles displays relatively larger distances within a more crystalline polymer matrix.

it is clear that balancing nanoparticle loading against the amorphousness of the solid-state is a non-trivial but critical task for optimizing GPEs for DSSC development.

**3.3. Electrochemical Performance of GPEs.** AC electrochemical impedance analyses (Figure 3) was used to investigate the ionic conductivity of GPEs as a function of  $\text{Co}_3\text{O}_4$  content and processing.  $\text{Co}_3\text{O}_4$ -modified GPEs were sandwiched between two stainless steel blocking electrodes (hole area,  $S = 2.0 \text{ cm}^2$ ; spacer thickness,  $l = 2.9 \text{ mm}$ ) and evaluated against the control GPE (PVVA). The ionic conductivities ( $\sigma$ ) of the GPEs were determined by the following equation:

$$\sigma = R_l^{-1}S^{-1}l = R_B^{-1}S^{-1} \quad (2)$$

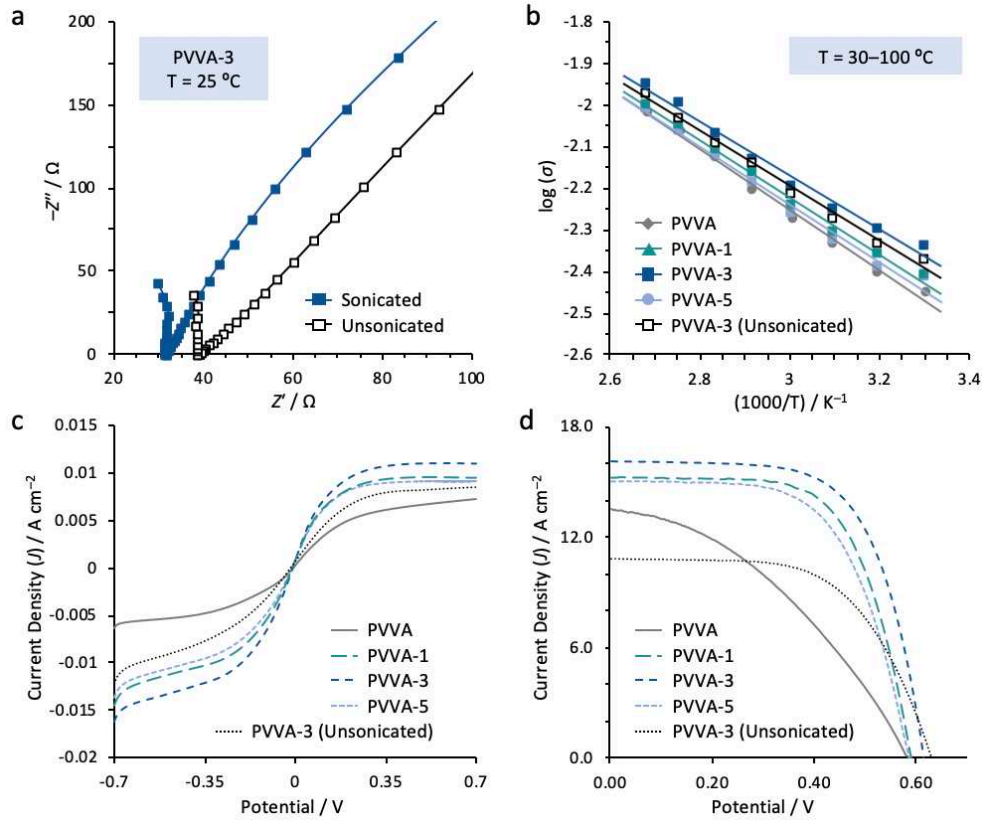
where  $R_l$  is the real-axis intercept of impedance spectra from the Nyquist plot,  $S$  is the geometric surface area of the GPE-electrode interface, and  $l$  is the distance between the two electrodes.

Room temperature impedance spectra as well as variable temperature data between 30 and  $100^\circ\text{C}$  (Figures 3a, S7 and S8) show all three sonicated  $\text{Co}_3\text{O}_4$ -modified GPEs delivering real-component impedance ( $Z'$ ) values smaller than PVVA. These data confirm that  $\text{Co}_3\text{O}_4$ -modified GPEs must be more efficient in facilitating ion transport through the host polymer matrix, where the following values for  $\sigma$  were obtained at  $25^\circ\text{C}$ :  $3.89 \times 10^{-3}$  (PVVA-1),  $4.62 \times 10^{-3}$  (PVVA-3),  $3.93 \times 10^{-3}$  (PVVA-5),  $3.97 \times 10^{-3}$  (PVVA-3 unsonicated) and  $3.57 \times 10^{-3} \text{ S cm}^{-1}$  (PVVA). A summary of these ionic conductivities and related values obtained from EIS measurements can also be found in Table 2. Indeed, the phase-optimized PVVA-3 displayed the lowest bulk resistance ( $R_B = R_l^{-1}d = 31.7 \text{ }\Omega$ ) and highest ion conductivity overall. The Nyquist data satisfyingly follow the trend in electrochemical performance (i.e., PVVA-3 > PVVA-5 > PVVA-1) anticipated from powder XRD analyses (*vide supra*). Beyond a certain critical wt% threshold,<sup>71</sup> the formation tendency of larger crystalline aggregates begins to hinder charge transport efficiencies by raising the percolation threshold of redox species between ion-conducting polymer chains. Here, it is evident that a solid content loading of  $\text{Co}_3\text{O}_4$  beyond 3 wt% does not lead to more favourable charge mobilities.

Temperature-dependent ionic conductivity measurements taken between 30 and  $100^\circ\text{C}$  provided us with further insight into the ion transport mechanism taking place in  $\text{Co}_3\text{O}_4$ -modified GPEs. Arrhenius plots (Figure 3b) show the  $\log(\sigma)$  of GPEs with respect to inverse temperature plotted as  $1000/T$ . Here, changes in ionic conductivity are observed to be linearly dependent. Their ionic conductivity behavior can thus be expressed by the Arrhenius<sup>90</sup> equation:

$$\sigma = \sigma_0 \exp(-E_a/k_B T) \quad (3)$$

in which  $\sigma_0$  is the pre-exponential frequency factor,  $E_a$  is the activation energy,  $k_B$  is Boltzmann's constant, and  $T$  is absolute temperature. Such a profile suggests that  $\text{Co}_3\text{O}_4$ -modified GPEs are homogenous (as far as ion mobility phase is concerned) with no large polymer chain rearrangements taking place as the temperature is increased. In other words, general diffusion and conductive behaviour of charge carriers in our GPE design can be described



**Figure 3.** (a) Room temperature impedance spectra reveal a marked increase in resistance towards ion mobility for the sonochemically enhanced GPE PVVA-3 versus its homologue containing 3 wt% unsonicated  $\text{Co}_3\text{O}_4$  nanoparticles, evidenced by a lower  $Z'$  intercept. The linear relationship between  $\log(\sigma)$  and inverse temperature in the (b) Arrhenius plots are indicative of a charge-hopping ion transport mechanism within the GPEs. (c) Linear sweep voltammograms taken at room temperature reveal an 51% enhancement in the apparent diffusion coefficient of triiodide species ( $D_{app}^{I_3^-}$ ) within the optimized PVVA-3 sample versus PVVA and its unsonicated homologue. (d)  $J$ - $V$  curves of  $\text{Co}_3\text{O}_4$ -modified DSSCs show enhanced photovoltaic performance over an unmodified cell, where the device containing PVVA-3 outperforms all other devices in the study. DSSCs were operated under AM1.5G at an illumination intensity of  $100 \text{ mW cm}^{-2}$ .

Table 2. Electrochemical Impedance-Derived Properties of GPEs

GPE	wt% $\text{Co}_3\text{O}_4$	$R_B / \Omega$	$\sigma^a / 10^{-3} \text{ S cm}^{-1}$	$E_a^b / \text{ eV}$	$D_{app}^{I_3^-} / 10^{-8} \text{ cm}^2 \text{ s}^{-1}$
PVVA	0	40.5	3.57	0.144	2.19
PVVA-1	1	37.2	3.89	0.137	2.71
PVVA-3	3	31.3	4.62	0.128	3.30
PVVA-5	5	36.8	3.93	0.136	2.81
PVVA-3 Unsonicated	3	36.5	3.97	0.131	2.51

<sup>a</sup>Ionic conductivity at 25 °C. <sup>b</sup>Activation energy determined from the Arrhenius plot (equation 3)

as hopping to nearest and energetically favorable vacant (e.g., empty, interstitial, and (sub)lattice) sites<sup>91</sup> throughout the host polymer framework.

This model is unlike other common ion transport mechanisms reported in the literature, such as that described by the Vogel–Thammann–Fulcher<sup>92–94</sup> model where charge carrier diffusion instead arises from the molecular motion of the polymer chains at greater temperatures to generate free volume within the host matrix. Even at higher temperatures, the Arrhenius behaviour of our GPEs is preserved, demonstrating the structural contributions of a PAN blend to limit segmental motion<sup>91,95,96</sup> of polymer chains within the host matrix. With respect to  $E_a$ ,  $\text{Co}_3\text{O}_4$ -

modified GPEs display lowered activation energies in comparison to unmodified PVVA. These variations suggest that a homogenous dispersion of fine  $\text{Co}_3\text{O}_4$  nanoparticles can aid the formation of interstitial lattice sites between polymer chains which in turn facilitate ion hopping through the quasi-solid-state. The rise in  $E_a$  from PVVA-3 to PVVA-5 can be attributed to the fact that higher weight percentages of  $\text{Co}_3\text{O}_4$  nanoparticles can lead to the formation of larger nanocrystalline aggregates, which ultimately can impede ion diffusion by disrupting access of ions to the vacant sites of the host lattice.

The results of linear sweep voltammetry (LSV) experiments (Figure 3c; Table 2) carried out on  $\text{Co}_3\text{O}_4$ -modified

GPEs to determine the effective diffusion coefficient of triiodide ions ( $D_{app}^{I_3^-}$ ) stand in agreement with our EIS analyses. Based on the known mechanism of dye-sensitized photovoltaics, the  $D_{app}^{I_3^-}$  of this key rate-determining species could be determined by the following equation:

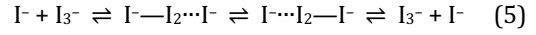
$$D_{app}^{I_3^-} = J_{lim} \left( \frac{l}{2nFC} \right) \quad (4)$$

where  $J_{lim}$  is the steady-state photocurrent density,  $l$  is the GPE thickness,  $n$  is the number of electrons (i.e.,  $n = 2$ ),  $F$  is the Faraday constant and  $C$  is the concentration of  $I_3^-$  ions. Here,  $J_{lim}$  is maximized at 0.011 mA cm<sup>-2</sup> in PVVA-3. Moreover, PVVA-3 achieves the greatest enhancement in  $D_{app}^{I_3^-}$ , going from  $2.19 \times 10^{-8}$  cm<sup>2</sup> s<sup>-1</sup> for PVVA to a value of  $3.30 \times 10^{-8}$  cm<sup>2</sup> s<sup>-1</sup>. This 51% overall enhancement in effective ion diffusion is consistent with our observation of a highly amorphous matrix offered by an optimized 3 wt% loading of sonicated Co<sub>3</sub>O<sub>4</sub>. Notably, a literature survey of diffusion coefficients for GPEs augmented with other conventionally obtained and/or untreated metal oxides reveals  $D_{app}^{I_3^-}$  values to typically span a 10<sup>-6</sup> to 10<sup>-10</sup> cm<sup>2</sup> s<sup>-1</sup> range,<sup>48,55,63</sup> calling attention to the general advantage provided by sonochemical processing for improving the performance of GPE nanocomposites.

**3.4. GPE Performance in DSSCs.** Lab-scale DSSC devices were fabricated by sandwiching an N719-sensitized TiO<sub>2</sub> mesoporous anode with the prepared GPEs and a Pt counter electrode between two transparent FTO substrates.  $J$ - $V$  curves of the as-prepared DSSCs were measured and are provided in Figure 3d above. From Table 3 below—which summarizes the short-circuit photocurrent ( $J_{sc}$ ), open circuit potential ( $V_{oc}$ ), fill factor ( $FF$ ) and photovoltaic conversion efficiency (PCE;  $\eta$ )—it is clear that device  $J_{sc}$  values are sensitive and positively influenced by sonochemically processed Co<sub>3</sub>O<sub>4</sub> nanoparticles in the PVVA host matrix—up to 16.2 mA cm<sup>-2</sup> in PVVA-3. On the other hand, the presence of 3 wt% unsonicated nanoparticles appears to have a detrimental effect on the  $J_{sc}$  value (11.5 mA cm<sup>-2</sup>), though the device efficiency is still better than the one containing unmodified PVVA electrolyte.

Given that our DSSCs were constructed and measured under consistent temperatures and device components, increases in  $J_{sc}$  cannot be attributed to either thermal acceleration of ions in the GPE or differences in photoanode

characteristics. On the other hand, improved dye regeneration kinetics, which are ultimately influenced by the efficiency of ion diffusion and charge transport within the electrolyte layer, can lead to increased  $J_{sc}$  values. In this vein, the apparent rise in  $J_{sc}$  here can be reasonably attributed to ion diffusion of the redox-active species across a bridged nanoparticle network, made more effective as a result of sonochemical processing, *via* a Grotthus-type electron hopping-ion exchange mechanism, i.e.,



Previous reports<sup>95,96</sup> have shown that iodide exchange processes occur in viscous/gel electrolyte media to afford the adsorption of redox-active species at the nanoparticle surface. The formation of dense, highly oriented regions enables rapid  $I_3^-$  transport with little to no physical species transfer observed. This mechanism bears similarity to the chemical bond exchange processes that enables proton transport in aqueous media.<sup>97,98</sup> We surmise that sonochemically prepared nanoparticles may act in a similar capacity, i.e., as potential redox mediators for dye regeneration. By lessening the degree of aggregation that is inherently caused by thermal annealing with sonochemical processing, the relatively high surface area of the nanoparticles can be retained. This, as well as the internal pore structure generated by acoustic cavitation, enable the alignment of a local density of  $I^-/I_3^-$  to accelerate redox couple diffusion along a bridged particle network<sup>99</sup> in GPEs. It is this re-sensitization of the TiO<sub>2</sub> photoanode by mobilization of the charge carriers that ultimately accounts for the observed increases in  $J_{sc}$ ,  $FF$ , and PCE.

We were pleased in the end to find that the device based on PVVA-3—our best performing GPE—also delivers the best PCE ( $\eta$ ) of 6.46%. This result represents a more than two-fold improvement over the DSSC lacking any additive as well as a more than four-fold efficiency enhancement over a recent previous report.<sup>70</sup> To the best of our knowledge, our work represents the highest PCE reported for Co<sub>3</sub>O<sub>4</sub>-modified GPEs of either homo-, co- or blended polymer host compositions (Table 1.)

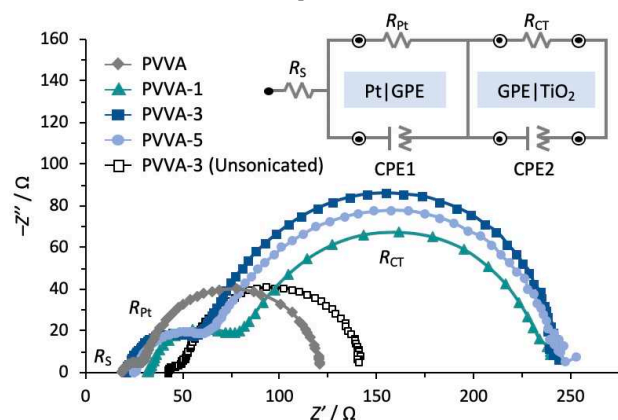
EIS analysis via Nyquist plots (Figure 4) of our devices helped us gain a more complete understanding of which interfacial processes may be influenced by the presence of sonochemically processed Co<sub>3</sub>O<sub>4</sub> nanoparticles. The equivalent circuit model of our devices is given in the inset

Table 3. Photovoltaic Performance and Equivalent Circuit Parameters Fitted to EIS Data of DSSCs

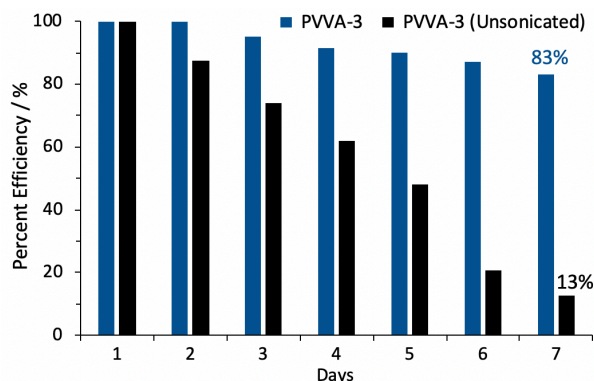
DSSC <sup>a</sup>	$J_{sc}$ / mA cm <sup>-2</sup>	$V_{oc}$ / V	$FF$ / %	$\eta$ / %	$R_s$ / $\Omega$	$R_{PT}$ / $\Omega$	$R_{CT}$ / $\Omega$
PVVA	13.5 ± 0.3	0.590 ± 0.01	0.383 ± 0.01	2.96 ± 0.09	40.4 ± 0.10	10.50 ± 0.05	92.2 ± 0.50
PVVA-1	15.1 ± 0.2	0.580 ± 0.01	0.616 ± 0.01	5.33 ± 0.10	32.0 ± 0.20	49.4 ± 0.10	157.0 ± 0.50
PVVA-3	16.2 ± 0.1	0.610 ± 0.01	0.657 ± 0.01	6.46 ± 0.10	20.70 ± 0.05	41.60 ± 0.09	193.0 ± 0.10
PVVA-5	15.3 ± 0.2	0.580 ± 0.01	0.640 ± 0.01	5.78 ± 0.09	25.20 ± 0.06	41.30 ± 0.04	180 ± 0.20
PVVA-3 unsonicated	11.5 ± 0.4	0.620 ± 0.01	0.58 ± 0.01	3.99 ± 0.04	18.80 ± 0.05	10.90 ± 0.05	95 ± 0.10

<sup>a</sup>Data is measured and averaged for three devices.

of Figure 4. Two distinct semi-circle profiles are detected at low (0.1–10 Hz) and high (10–10<sup>5</sup> Hz) frequencies, representing the occurrence of charge recombination phenomena ( $R_{CT}$ ) and charge transfer processes at the Pt|GPE interface ( $R_{Pt}$ ), respectively. The intersection at the x-axis is interpreted as ohmic serial resistance<sup>100</sup> ( $R_s$ ) of the as-prepared FTO substrate. We found that the total impedance of the sonochemically enhanced DSSCs is dominated by a high charge-transfer resistance ( $R_{CT}$  ranging between 60 and 250  $\Omega$ ), meaning that the tendency of electrons to recombine with  $I^-/I_3^-$  ions in GPEs has been significantly reduced in the presence of processed nanoparticles. Interaction of the oxidized redox species with the surface and internal pore structure of the  $Co_3O_4$  nanoparticles are likely to mitigate the back reaction of electrons within the  $TiO_2$  conduction band, therefore extending the effective lifetime and overall performance of the present DSSC devices. This argument stands in agreement with EIS data collected for the DSSC comprised of unmodified PVVA: a



**Figure 4.** Electrochemical impedance data collected for GPE-based DSSCs show that sonochemically processed  $Co_3O_4$  additives confer higher resistance ( $R_{CT}$ ) towards charge recombination events (i.e., of  $I_3^-$ ) in GPEs, thus enabling fast N719 dye regeneration for enhanced DSSC performance. Inset: An equivalent circuit model has been fitted to the EIS data.



**Figure 5.** Photovoltaic performance stability of DSSCs employing  $Co_3O_4$ -modified GPEs, plotted as percent efficiency over time. The sonochemically enhanced DSSC consisting of the GPE PVVA-3 retains 83% of its initial efficiency over the course of seven days, whereas the untreated cell is found to experience an 87% efficiency loss in the same period.

resistance to iodide charge. From these results, the apparent increases in  $J_{sc}$  and PCE can be clearly attributed to the ability of sonochemically processed  $Co_3O_4$  additives to substantially mediate the charge-transfer dynamics of redox couples in the PVVA gel state.

### 3.5. Stability of DSSCs Bearing $Co_3O_4$ -modified GPEs.

Finally, given the importance of addressing device stability in additional to enhanced efficiencies, we sought out to interrogate the relationship between our  $Co_3O_4$  processing method and DSSC performance over time. The performance of our best DSSC (PVVA-3) was monitored over seven days and compared to the homologous device containing 3 wt% unsonicated  $Co_3O_4$  nanoparticles. Both devices were stored at room temperature and under ambient conditions between each test. Plotting the normalized efficiencies over time (Figure 5) revealed a dramatic loss in DSSC performance for the device employing unsonicated nanoparticles. Compared to that of the DSSC based on PVVA-3, which retained at least 83% of its initial efficiency, the unsonicated cell only retained 13% efficiency after seven days. We suspect this remarkable retention in DSSC performance for sonochemically enhanced PVVA-3 has to do with the finer porous structure and higher surface area of the nanoparticle network slowing down cell degradation and GPE collapse caused by the evaporation of carbonate solvent molecules from the matrix. Thus, on top of improving overall device performance parameters, we find rather serendipitously that the sonochemical process has the potential to extend the performance integrity and stability of quasi-solid-state devices.

## 4. SUMMARY AND CONCLUSIONS.

We have successfully applied a sonochemical processing method to access nanoporous  $Co_3O_4$  particles that display relatively high dispersibilities and surface areas to promote faster ion transport within GPEs for enhanced quasi-solid-state DSSC performance. FESEM and HR-TEM analysis of the annealed metal oxide reveal the formation of porous flakes composed of individual nanocuboids, in agreement with the powder XRD data. The high dispersibility of sonicated  $Co_3O_4$  nanoparticles (1, 3, and 5 wt% solid content) in a PVVA polymer-blended host matrix resulted in homogenous quasi-solid-state GPEs with highly amorphous phases. A GPE modified with 3 wt% sonicated  $Co_3O_4$  nanoparticles (PVVA-3) was found to be the best performing GPE in our series, providing an ionic conductivity and apparent diffusion coefficient value of  $4.62 \times 10^{-3} \text{ S cm}^{-1}$  and  $3.30 \times 10^{-8} \text{ cm}^2 \text{ s}^{-1}$ , respectively. The latter value represents a 51% enhancement in triiodide diffusion kinetics over an unmodified PVVA-derived GPE lacking any nanoparticle content, which we attribute to an interconnected particle network for contiguous ion transport. These attributes translated to a maximum photovoltaic conversion efficiency,  $\eta$ , of 6.46% in quasi-solid-state DSSCs—the highest known efficiency to-date for a  $Co_3O_4$ -modified GPE-based cell. Electrochemical impedance analysis of DSSCs have revealed that the addition of  $Co_3O_4$  nanoparticles selectively enhance the photocurrent density ( $J_{sc}$ ) and

rate of dye regeneration *via* a high recombination charge transfer resistance ( $R_{CT}$ ). The GPE containing a higher weight percent content of sonicated  $\text{Co}_3\text{O}_4$  (i.e., PVVA-5) led to reduced DSSC efficiency ( $\eta = 5.85\%$ ) on account of more crystalline particle deposits present within the GPE matrix. Likewise, the homologous GPE of PVVA-3 containing unsonicated  $\text{Co}_3\text{O}_4$  nanoparticles showed a lowered photovoltaic conversion efficiency of 3.99%, further attributing the improvement in DSSC performance to the sonochemical process. Finally, the use of sonochemically processed nanoparticles was found to extend the performance stability of DSSCs with an overall 83% retention in device efficiency (versus 13% in the homologous untreated cell) after one week of monitoring.

Our studies demonstrate the importance in balancing optimizing metal oxide content with device performance. By comparing our sonochemically enhanced photovoltaic device to similar devices reported prior, especially that employing equivalent quantities of conventionally obtained  $\text{Co}_3\text{O}_4$ , it appears that including a sonochemical processing step can lead to a greater than two-fold increase in DSSC performance. Moreover, sonochemically prepared  $\text{Co}_3\text{O}_4$  nanoparticles may yet reveal further enhanced properties in DSSCs when subjected to a magnetic field versus their bulk analogues. We are hopeful that such a straightforward and generally applicable technique will allow stable quasi-solid-state DSSCs based on low-cost metal oxides to reach performance levels that are meaningful for real-world exploitation.

## ASSOCIATED CONTENT

### Supporting Information

Details on the preparation of unsonicated  $\text{Co}_3\text{O}_4$  nanoparticles; GPE optimization notes; details of photoanode fabrication; FTIR data; solid-state CPMAS NMR spectra; electrochemical impedance data of PVVA, PVVA-1, PVVA-3 and PVVA-5 at varying temperatures. The Supporting Information is available free of charge on the ACS Publications website at <http://pubs.acs.org>.

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## AUTHOR CONTRIBUTIONS

Nanoparticle synthesis, gel polymer electrolyte processing and electrochemical experiments were performed by N.M.S., F.S.O., and A.N. Solid-state NMR analysis was provided by D.C.A. Materials imaging analysis was completed by M.M.A. The research was designed by R.T.S. and A.-J.A. with intellectual contributions from R.K. The manuscript was written through the contributions of all authors, with final figures and editing performed by A.-J.A. and R.T.S. All authors have given approval to the final version of the manuscript.

## Notes

The authors declare no competing financial interest.

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