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Wu, Linglin, Mascal, Mark, Farmer, Thomas James orcid.org/0000-0002-1039-7684 et al. (2 more authors) (2017) *Electrochemical Coupling of Biomass-derived Acids : New C8 Platforms for Renewable Polymers and Fuels*. *ChemSusChem*. 166–170. ISSN 1864-564X

<https://doi.org/10.1002/cssc.201601271>

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2,7-Octanedione and 2,5-Dimethyladipic Acid From the Electrochemical Coupling of Biomass-derived Carboxylic Acids: C₈ Platforms for Renewable Polymers and Fuels

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Abstract: *Electrolysis of biomass-derived carbonyl compounds is an alternative to condensation chemistry for supplying products with chain length >C₆ for biofuel and renewable material production. Kolbe coupling of biomass-derived levulinic acid gives 2,7-octanedione, a new platform molecule only two low process-intensity steps removed from raw biomass. Hydrogenation to 2,7-octanediol provides a chiral secondary diol largely unknown to polymer chemistry, while intramolecular aldol condensation followed by hydrogenation yields branched cycloalkanes suitable for use as high-octane, cellulosic gasoline. Analogous electrolysis of an itaconic acid-derived methylsuccinic monoester gives chiral 2,5-dimethyladipic acid, another underutilized monomer due to lack of availability.*

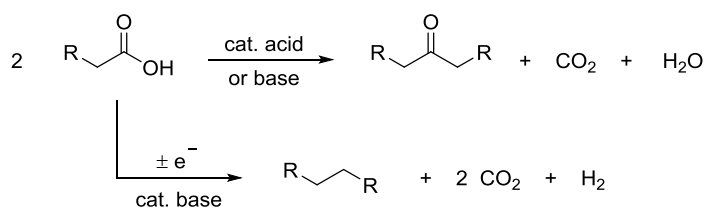
Introduction

Biofuels and renewable polymers play an undisputed role in the green technology movement, and their production from biomass-derived carbohydrates has been investigated by multiple research groups. In the former case, since common monosaccharides are C₆ or less, a C–C coupling step is essential to

eventually achieve the hydrocarbon volatility range required for automotive fuels. In the case of polymers, virtually any platform molecule that can be rendered bifunctional may potentially serve as a monomer, and C–C coupling reactions can also be of value here, likewise offering products not limited to six carbons.

In the majority of cases, the approach to accessing suitable biofuel precursors from biomass-derived carbonyl compounds involves aldol or related condensation reactions, which are catalyzed, thermodynamically driven processes resulting in C–C bond formation.^[1] We however became attracted to the proposition of carrying out electrochemical coupling of such molecules to accomplish the necessary chain elongation for the following reasons: 1) the electrochemical dimerization of sugar derivatives is an inexpensive, non process-intensive method where the driving force for the reaction essentially comes from the power grid, 2) novel structures may be accessed which are not available via condensation chemistry, and 3) coupled products can be obtained in a more advanced state of reduction, thus avoiding extensive hydrodeoxygenation. Comparison of a condensation reaction versus the Kolbe electrolysis serves as an illustration (Scheme 1). Using the same substituted acetic acid as a hypothetical model, Claisen-type condensation/**decarboxylation** gives a ketone, whereas the electrolysis gives a substituted ethane. In both cases, oxygen is carried away in the form of CO₂, although in the latter H₂ is co-generated, reminiscent of the production of hydrogen and CO₂ from hydrocarbons by a combination of steam reforming and the water-gas shift reaction.

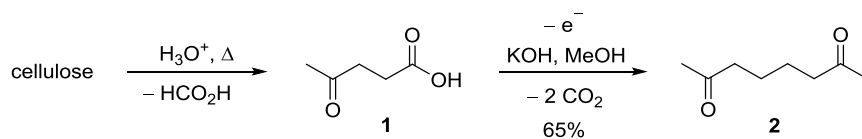
Herein, we will demonstrate that electrolysis of biomass-derived carbonyl compounds has disruptive potential to deliver >C₆ biorefinery outputs of value both in materials and fuels markets.



Scheme 1. Comparison of products derived by a condensation reaction versus electrochemical coupling. Claisen chemistry, while feasible with carboxylic acids,^[2] is routinely carried out with the corresponding esters, in which case the condensate is an alcohol instead of water.

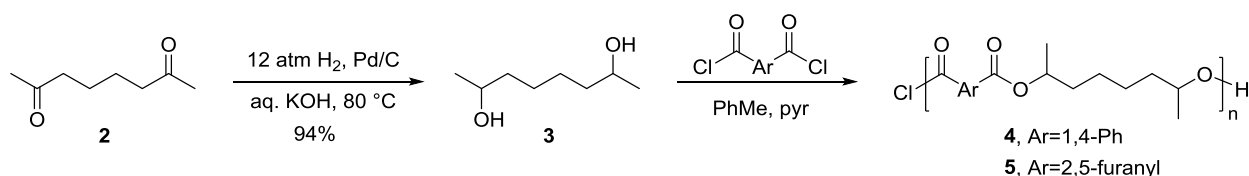
Results and Discussion

Looking to the literature, there has been little activity around the direct coupling of sugars, although it was interesting to find that glucose itself has been made to undergo a cathodic process whereby it was transformed into a dodecitol, presumably via its aldehyde form.^[3] The electrolysis however of sugar derivatives presents a viable alternative, and we were drawn to reports of the dimerization of levulinic acid (LA) **1** to 2,7-octanedione **2** (Scheme 2).^[4-6] The recent upsurge of interest in LA **1** as a renewable, carbohydrate-derived platform molecule, and the continuing development of technologies to produce it on an industrial scale,^[7] advance dione **2** as a potential "second-generation" platform only two steps removed from biomass. The electrolysis reaction is straightforward, providing 65% yield of **2** at 90% conversion using platinum plate electrodes in an undivided cell under constant current conditions (see Supporting Information for details).



Scheme 2. Cellulose to 2,7-octanedione via levulinic acid.

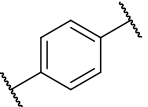
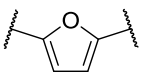
The only substantive account of the use of **2** in materials applications was published by Joshi and Limaye, who in the 1980s reported the conversion of **2** into 2,7-octanediamine via reduction of the corresponding dioxime and subsequent production of a terephthalate polymer.^[8] Another straightforward approach to monomer synthesis would be to simply reduce **2** to corresponding diol **3**. Due to the (historically) limited access to **2**, compound **3** is largely unknown to polymer chemistry, the only reference to our knowledge being its use as one of a series of diols to test the concept of iterative tandem catalysis by polymerization with adipic esters.^[9] In that case, **3** was prepared via the corresponding α,ω -diene. It has also been previously made from **2** using Meerwein-Ponndorf-Verley reduction.^[10] However, we opted for more industrially relevant hydrogenation, which proceeded smoothly and in high (94%) yield (Scheme 3).



Scheme 3. Preparation of 2,7-octanediol **3** and its polyesters.

While a range of bio-derived diols have been investigated for polyester synthesis, both via chemo- and enzyme-catalyzed reactions, most are primary diols, typically 1,3-propanediol, 1,4-butanediol and 1,6-hexanediol.^[11] Polymerization of secondary alcohols and high molecular weight monomers is known to be more challenging as a result of steric issues and high boiling points, respectively. In the standard esterification/transesterification procedure for polyester synthesis, the diol is generally used in excess relative to the diacid or diester component and high temperatures are applied to remove the excess diol and drive the polymer to high chain lengths. Attempts have been made to circumvent both the reactivity and volatility issues by using diacid chlorides in a 1:1 stoichiometric ratio with the diol.^[12] In the first instance we applied this approach to reaction of terephthaloyl chloride with diol **3** to obtain a 63% yield of poly(2,7-octanediol)terephthalate **4**, which was found to have a good chain length ($M_n > 8500$ Da) and low polydispersity ($Pd_i < 1.5$). Another motivation for using diacid chlorides was our facile, two-step preparation of 2,5-furandicarbonyl chloride (FDCC) from 5-(chloromethyl)furfural (CMF), which itself is one step removed from raw biomass.^[13] Since levulinic acid **1**, the precursor to **3** (via **2**) is also a single step from CMF,^[14] this provides us an opportunity to showcase the synthesis of a novel polyester with both monomers ultimately derived from a single platform molecule. The resultant poly(2,7-octanediol)-2,5-furanoate polymer **5** from the reaction between FDCC and diol **3** was also isolated in good yield (54%) and with a reasonable chain length and low Pd_i . The data for these polymers are presented in Table 1.

Table 1. Co-polyesters of diol **3**.

Polymer	Ar	%recovery	M_n (Da) ^a	M_w (Da) ^b	Pd_i ^c
4		63%	8531	12220	1.43
5		54%	3978	5079	1.28

^aNumber average molecular weight. ^bWeight average molecular weight. ^cPolydispersity index.

Thermogravimetric analysis under an N_2 atmosphere was used to determine the stability of **4** and **5** (Table 2), which showed that little mass loss occurs in either prior to 290 °C. The terephthalate polyester **4** was shown to have a slightly higher thermal stability than furandioate **5**, with both its TD10 and TD50 being roughly 20 °C higher than the furan equivalent. Likewise the glass transition temperature (T_g) of **4**

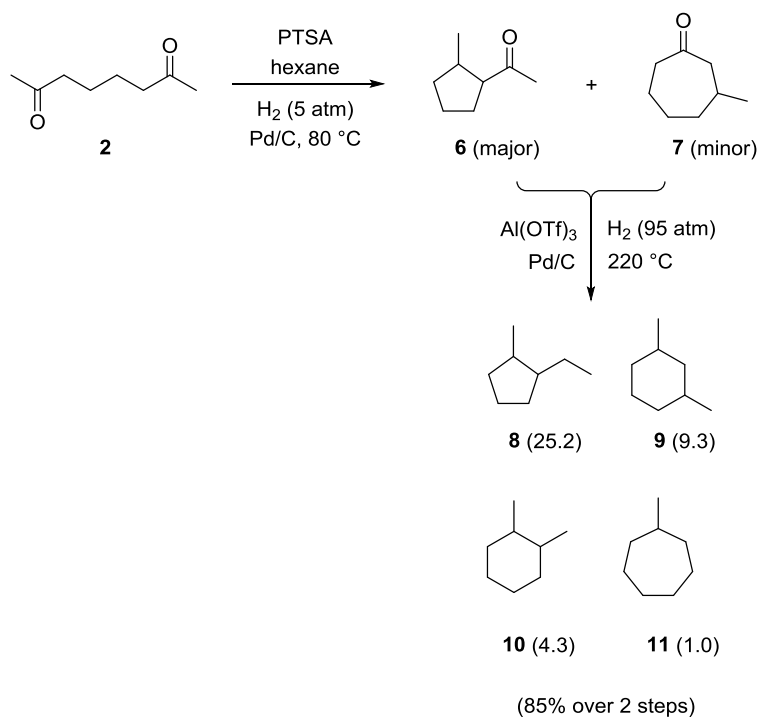
was found to be higher than of **5**. In comparison to literature values for analogous aromatic-aliphatic polyesters (Table S3 in the Supporting Information), it can be seen that both **4** and **5** conform to the previously observed trend whereby extending the length of the diol reduces the T_g while diols of secondary alcohols show increased T_g relative to their primary alcohol regioisomers. This highlights the potential value of **3** as a new bio-based monomer that allows further control of characteristics that have a direct impact on the processability and applications of the final product. For example, the near room temperature T_g of **5** may point to potential thermoresponsive polymer applications where changes from the glassy to rubbery state of a plastic between ambient temperature and body temperature is desired.

Table 2. Thermal analysis of polymers **4** and **5**.

Polymer	TD10 (°C) ^a	TD50 (°C) ^b	T_g (°C) ^c	T_m ^d
4	312.4	330.5	62.6	n.d.
5	289.5	301.6	26.0	n.d.

^aTemperature at 10% mass loss. ^bTemperature at 50% mass loss. ^c T_g = glass transition temperature as determined by differential scanning calorimetry (DSC). ^d T_m = melt point as determined by DSC. n.d. = none detected.

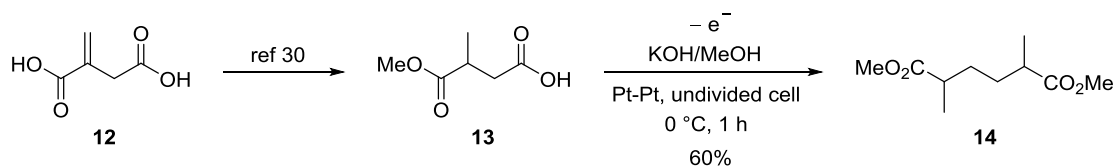
To date, the only application of **2** directed towards biofuels involved complete reduction to n-octane,^[15,16] a compound of no use in gasoline (RON = -19) and too volatile for diesel fuel.^[17] We recognized an opportunity to use dione **2** to much better advantage in the production of fuels by employing intramolecular condensation chemistry, leading ultimately to branched, cyclic alkanes. Thus, **2** could be made to undergo aldol condensation by treatment with either acid or base. The α,β -unsaturated ketone product mixture was however prone to side reactions under these conditions which made it difficult to achieve good selectivity, a result also noted by Bouillon et al.^[18] We therefore performed the aldol reaction in tandem with hydrogenation of the double bond to provide mainly methyl 2-methylcyclopentyl ketone **6** alongside small amounts of the alternative cyclization product 3-methylcycloheptanone **7** (Scheme 4). This product mixture could be isolated and characterized, but the best yield of hydrocarbon was obtained by introducing additional Pd/C and the hydrodeoxygenation catalyst Al(OTf)₃^[19] at this point, then increasing the reaction temperature and H₂ pressure to 220 °C and 50 atm, respectively. Using this approach, the total yield of cycloalkanes **8-11** starting from **2** was 85%, with the relative ratios as determined by GC-MS shown. The observation of dimethylcyclohexanes **9** and **10** is the result of carbocation rearrangements, and since commercial gasoline is a mixture of hundreds of hydrocarbons,^[20] the lack of selectivity to **8** presents no problems.



Scheme 4. Production of cycloalkane biofuels from 2,7-octanedione.

Since we propose the conversion of **2** to **8-11** as a new approach to cellulosic gasoline, we undertook to determine the fuel properties of these products. Existing biomass to biofuel processes generally target linear hydrocarbons appropriate to diesel or jet fuel applications.^[1] Efficient, renewable methods that lead to the branched and cycloalkanes found in gasoline therefore fill a commercial void.^[21] **8-11** are all previously described compounds with boiling points between 118-134 °C, thus well within the volatility range of motor gasoline (*ca.* 40-200 °C).^[20] The key measure of fuel performance in spark ignition engines is the antiknock index, commonly referred to as octane number. For individual molecules, a value for the research octane number (RON) will either have been measured or can be calculated using various algorithms. The experimental RON value for **9** is 69 and for **10** is 81.^[22] A modeled value for **8** = 93 has been reported.^[23] As for the minor component **11**, the RON of cycloheptane itself is 38.8. Methyl branching generally increases octane numbers, and the index value for this molecule has been calculated at RON = 68 using the model of Dahmen and Marquardt.^[24] An estimated RON of the blend of **8-11** in the proportions observed in the hydrogenation reaction would be 86.8. In combination with 10% ethanol as an oxygenate, the calculated RON is 89.

Finally, our success in processing levulinic acid into useful biofuel and polymer components prompted us to consider the electrochemical dimerization of other bio-derived acids. Itaconic acid **12** is another up and coming biorefinery platform molecule,^[25] and its applications to heterocycle and polymer chemistry have recently been reviewed.^[26,27] Like levulinic acid, it appears on the NREL list of the top 12 value added chemicals from biomass.^[28] Although the electrolysis of itaconic acid has been investigated,^[29] no Kolbe coupling of **12** or its monoesters has been reported to date. However, Hancock and Linstead reported that a methylsuccinic acid monoester **13**, derived by the methanolysis of methylsuccinic anhydride, underwent electrolytic dimerization to 2,5-dimethyladipic ester **14**. In that work, a poorly defined mixture of half-esters was used and the reported yield of **14** was 30%. We set out to improve access to **14** by optimization of the electrochemical reaction. Thus, ester **13** was prepared from itaconic acid **12** using a published method, wherein the diester of **12** was selectively hydrolyzed and the double bond asymmetrically hydrogenated.^[30] While a stereodefined form of **13** presents an attractive option for future work, for the purposes of this study, the racemate was used. Subjecting **13** to the same conditions as in the conversion of **1** to **2**, *i.e.* constant current electrolysis at a current density of 180 mA cm⁻² across platinum electrodes in an undivided cell, gave diester **14** in 60% yield at 85% conversion.



Scheme 5. Production of dimethyl 2,5-dimethyladipate by electrolysis of itaconate-derived succinic acid monoester **13**.

Despite being a simple analogue of one of the highest volume monomers used in industry, limited access has meant that very little in the way of polymer chemistry has been described for **14**, as was also the case for **3**. The only systematic study of materials derived from **14** involved the production of stereoisomeric polyamides from hexamethylene diamine and *meso*-, *d*-, and *dl*-**14**, all of which showed considerably less crystallinity than the parent polymer.^[31] The development of new applications for **14** would be stimulated by improved availability. Of particular interest would be a study of the properties of polyesters of **3** with **14**, and we will report on such materials in a separate paper.

Conclusions

The purpose of this work is to showcase the opportunities for accessing novel platform molecules via the electrochemical processing of primary biomass derivatives. The diversity of >C₆ structures attainable using this approach points to future materials and fuels markets comprising a range of novel products. Here, we have presented levulinic acid **1** as a precursor to branched C₈ monomers and cycloalkane components that embody a high-octane, cellulosic gasoline, both via 2,7-octanedione **2**. The monomer 2,7-octanediol **3** was used to prepare novel polyesters with terephthalic and 2,5-furandicarboxylic acids. An analogous renewable monomer synthesis was demonstrated in 2,5-dimethyladipic diester **14**. Novel access to **3** and **14** by means of electrochemical C-C coupling chemistry opens up new opportunities for products that were formerly both limited in their availability and produced from unsustainable feeds.

Experimental Section

Electrolysis of levulinic acid

Levulinic acid (464 mg, 4.00 mmol) was subjected to constant current electrolysis (178 mA/cm²) on platinum plate electrodes (1.5 x 1.5 cm²; distance between electrodes = 12 mm) in 0.075 M methanolic KOH (10 mL) using an undivided cell with magnetic stirring at 22 °C. The reaction was terminated after the consumption of 1.0 F/mol of charge (16 min). The mixture was acidified to pH 3 with 1M HCl and the volatiles were evaporated under vacuum. The conversion (90%) and yield of 2,7-octanedione **2** (65%) were determined by NMR using 1,4-dioxane as an internal standard. To the residue was added 1M NaOH (50 mL) and the mixture was extracted with dichloromethane (50 ml x 3). The combined organic extract was washed with brine and dried over sodium sulfate. The solution was filtered and concentrated under vacuum, and the residue was passed through a short plug of silica using ethyl acetate as eluent. The solvent was evaporated to give 2,7-octanedione **2** as a pale yellow solid (178 mg, 63%). ¹H NMR (600 MHz, CDCl₃) δ 2.44-2.42 (m, 4H), 2.12 (s, 6H), 1.55-1.53 (m, 4H); ¹³C NMR (150 MHz, CDCl₃) δ 208.63, 43.38, 29.90, 23.14.

Hydrogenation of 2,7-octanedione

2,7-Octanedione (994 mg, 6.99 mmol), 5% palladium on carbon (280 mg), potassium hydroxide (140 mg, 2.1 mmol) and water (35 mL) were introduced into a Parr hydrogenator. The vessel was sealed, flushed three times with hydrogen and pressurized to 12 bar H₂. The mixture was heated at 80 °C with stirring for 140 min, then allowed to cool to rt. The pressure was released and the reaction was filtered through

a short pad of Celite. The filtrate was concentrated under vacuum and the residue was filtered through a short plug of silica gel using acetone as eluent. Evaporation of the solvent gave 2,7-octanediol **3** as a colorless oil (962 mg, 94%). ¹H NMR (400 MHz, D₂O) δ 3.78-3.71 (m, 2H), 1.39-1.26 (m, 8H), 1.08 (d, *J* = 6.3 Hz, 3H); ¹³C NMR (100 MHz, D₂O) δ 67.87, 37.78, 24.84, 21.82.

Tandem intramolecular aldol condensation and catalytic hydrogenation of 2,7-octanedione

2,7-Octanedione (284 mg, 2.00 mmol), 10% palladium on carbon (42 mg), *p*-toluenesulfonic acid monohydrate (38 mg) and hexanes (10 ml) were introduced into a Parr hydrogenator. The vessel was sealed, flushed three times with hydrogen and pressurized to 5 bar H₂. The reaction was heated at 80 °C with stirring for 3.5 h and allowed to cool to rt followed by further cooling to 10 °C in an ice-water bath. The reaction mixture was filtered through silica gel and further eluted with an ethyl acetate/hexane (1:5) solvent mixture. Evaporation of the solvent gave a mixture of isomeric ketones **6** (major) and **7** (minor) (227 mg, 90%). The above reaction was repeated but instead of isolating **6** and **7**, the reactor was opened and Al(OTf)₃ (95 mg), additional 10% palladium on carbon (168 mg), and hexanes (30 ml) were added. The vessel was sealed, flushed three times with hydrogen and pressurized to 50 bar H₂. The reaction was heated to 220 °C, which increased the internal pressure to *ca.* 95 bar. After 24 h at this temperature, the reactor was allowed to cool to rt and then further to 0 °C in an ice-water bath. The interior walls of the vessel were washed down with acetone. The catalyst was removed by filtration through Celite and the yields of **8-11** (85% total) were determined by GC-MS analysis with a dodecane internal standard and data matching against the National Institute of Standards and Technology (NIST) mass spectral library.

1-Ethyl-2-methylcyclopentane **8**. MS (EI): *m/z* (% of max intensity) 41.2 (58), 55.1 (100), 70.2 (55), 83.2 (95), 97.2 (11), 112.1 (24). Retention time in GC-MS: 3.42 min (trans) and 3.85 min (cis). Trans/cis ratio is 2.1:1.

1,3-Dimethylcyclohexane **9**. MS (EI): *m/z* (% of max intensity) 41.2 (21), 55.2 (70), 69.2 (19), 97.2 (100), 112.1 (31). Retention time in GC-MS: 3.26 min (trans) and 3.62 min (cis). Trans/cis ratio is 1.8:1.

1,2-Dimethylcyclohexane **10**. MS (EI): *m/z* (% of max intensity) 41.1 (40), 55.2 (91), 70.2 (27), 83.1 (20), 97.2(100), 112.2 (36). Retention time in GC-MS: 3.54 min (trans) and 3.97 min (cis). Trans/cis ratio is 2.3:1.

Methylcycloheptane **11**. MS (EI): *m/z* (% of max intensity) 41.1 (71), 55.1 (98), 69.1 (35), 83.1 (40), 97.2(100), 112.1 (20). Retention time in GC-MS: 4.37 min.

Electrolysis of 2-methylsuccinic acid 1-methyl ester

2-Methylsuccinic acid 1-methyl ester **13** (1.17 g, 8.01 mmol) was subjected to constant current electrolysis (178 mA/cm²) on platinum plate electrodes (1.5 x 1.5 cm²; distance between electrodes = 12 mm) in 0.10 M methanolic KOH (20 mL) using an undivided cell with magnetic stirring at 0 °C. The reaction was terminated after the consumption of 2.0 F/mol of charge (64 min). A small sample of the reaction was acidified to pH=3 with 1M HCl and evaporated under vacuum. The conversion was determined by NMR (85%) using 1,4-dioxane as an internal standard. The reaction mixture was evaporated under vacuum and to the residue was added 0.1M NaOH (70mL). The solution was extracted with dichloromethane and the combined organic extract was washed with saturated brine and dried over sodium sulfate. The solution was filtered and the solvent was evaporated to give dimethyl 2,5-dimethyladipate **14** as a colorless oil (489 mg, 60%). ¹H NMR (400 MHz, CDCl₃) δ 3.64 (s, 6H), 2.43-2.37 (m, 2H), 1.65-1.56 (m, 2H), 1.44-1.32 (m, 2H), 1.12 (d, *J* = 7.0 Hz, 6H); ¹³C NMR (100 MHz, CDCl₃) δ 176.81, 176.75, 51.47, 39.37, 39.20, 31.32, 31.08, 17.06, 16.89.

Experimental details for the preparation of polyesters **4** and **5** are provided in the Supporting Information.

Acknowledgements

MM acknowledges support from NSF-CBET grant 1335646 and Dr. Andrew Sutton at Los Alamos National Laboratory for helpful discussions. TJF and SPA would like to thank the Engineering and Physical Sciences Research Council (EPSRC) UK for funding their involvement in this research through the grant EP/L017393/1.

Access Statement

All data used in the preparation of this manuscript for the sections funded by the EPSRC grant EP/L017393/1 is contained within this document, the electronic supplementary information, or available from DOI: 10.15124/93ff2cd7-9408-4c34-aac6-3baa4f50abff

Keywords: Biomass Conversion • Electrochemistry • Itaconic Acid • Kolbe Coupling • Levulinic Acid

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