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Published paper

Stepwise assembly of an adamantoid Ru$_4$Ag$_6$ cage by control of metal coordination geometry at specific sites†

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The geometrically pure ‘complex ligand’ fac-[Ru(Lph)$_3$]$^{2+}$, in which three pendant bidentate binding sites are located on one face of the complex, reacts with Ag(i) ions to form the adamantoid decanuclear cage [(Ru(Lph)$_3$)$_4$Ag$_6$](PF$_6$)$_{14}$ which contains a 6-coordinate Ru(ii) ion at each vertex of a large tetrahedron and a 4-coordinate Ag(i) ion along each edge.

The self-assembly and host–guest chemistry of metal–ligand coordination cages continue to fascinate. Such high-symmetry cages represent appealing synthetic targets for supramolecular chemists to test their skills at controlling self-assembly so as to generate elaborate, multi-component assemblies from simple starting materials. In addition the ability of cages to bind guest molecules in their central cavity leads to potential applications ranging from catalysis to drug delivery. A few examples of mixed-ligand cages are known in which the same type of metal ion and one type of ligand.

The self-assembly process specifically occurs with selection of two different types of metal ion, resulting in a heteroleptic complex geometric isomerism at specific sites in a cage as a way of directing the self-assembly. In several other types of cage assembly however there is a 3:1 mixture of mer:fac tris-chelate vertices. Therefore, the ability to control the self-assembly of such cages – particularly mixed-metal versions – relies on the ability to prepare kinetically stable, geometrically pure fac or mer tris-chelate subcomponents as starting points to propagate a specific assembly.

We report here the use of this principle – viz. control of geometric isomerism at specific sites in a cage as a way of directing assembly – in the formation of an unusual [Ru$_4$Ag$_6$(Lph)$_{12}$]$_{48}$ mixed-metal cage which combines octahedral fac tris-chelate Ru(n) vertices and pseudo-tetrahedral Ag(i) bis-chelate edges in an adamantane-type cage structure having tetrahedral symmetry. The novelty lies both in the structure of the cage [a combination of three-connected and two-connected metal vertices based on Ru(n) and Ag(i) respectively], and in the use of the pre-formed, kinetically stable [Ru(Lph)$_3$]$^{2+}$ units as purely the fac isomer to direct the course of the self-assembly.

As fac-[ML$_3$]$^{2+}$ units from this family occur at specific sites in many of our cages, we wished to start with fac-[Ru(Lph)$_3$]$^{2+}$ in which the three pendant binding sites, where cage propagation occurs by coordination to additional metal ions, have a fac arrangement. Simple reaction of Ru(dms)$_2$Cl$_2$ with >3 equivalents of Lph afforded [Ru(Lph)$_3$]$^{2+}$ as a 1:3 statistical mixture of fac:mer isomers as shown by the $^1$H NMR spectrum in which every type of proton (e.g. coordinated pyridyl H) occurred in four different environments in a 1:1:1:1 ratio. However column chromatography or HPLC under a range of conditions did not give a good separation of the geometric isomers.
We therefore went back a step in the synthesis, to fac- [Ru(PyPzH)3]2+ [Scheme 1; PyPzH = 3-(2-pyridyl)-pyrazole] which can be readily separated from the mer isomer using a method reported earlier.12 In fac-[Ru(PyPzH)3]3+ the three pyrazole rings, with their acidic NH protons, lie of course on the same face of the complex. Alkylation of these 12 a with the bromomethyl compound A completed the formation of the L ph ligands coordinated to the metal centre at one end, to give fac-[Ru(Lph)3]2+ which was isolated as the hexafluorophosphate salt (see ESI†). The 1HN M R spectrum showed 20 1H environments confirming the threefold symmetry with all three ligands equivalent. Notably the CH2 protons close to the Ru(II) chiral centre are diastereotopic, giving a coupled pair of doublets at 5.5 and 4.8 ppm, whereas the CH2 protons more remote from the Ru(II) centre give a singlet at 5.3 ppm (Fig. S3, ESI†). The crystal structure of the complex cation of fac-[Ru(Lph)3][PF6]2 acetone is shown in Fig. 1; (ESI†) it is clear how the three pendant pyrazolyl-pyridine arms are directed to the same face of the complex. The phenyl group of each pendant arm forms a π-stacking interaction with a coordinated pyrazolyl-pyridine group from another ligand, as we have observed in related complexes.

Ag(I) generally forms four-coordinate bis-chelate complexes with pyrazolyl-pyridine ligands of this type.13 On the basis that three pendant bidentate sites are available for coordination in fac-[Ru(Lph)3][PF6]2, we combined (ESI†) fac-[Ru(Lph)3][PF6]2 with 1.5 equivalents of AgPF6 to maximise the likelihood of a structure forming that conforms to the principle of maximum site occupancy, with all metal ions coordinatively saturated and all ligands fully coordinated.14 If each pendant ligand fragment coordinates to a different Ag(i) ion, as is likely on steric grounds given the distance between the pendant pyrazolyl-pyridine units, we expect a mixed-metal cage in which each fac-[Ru(Lph)3]2+ unit caps a triangular array of Ag(i) ions.

Slow crystallisation of the reaction mixture afforded X-ray quality crystals of what proved to be a decanuclear Ru4Ag6 cage [(RuLph)3Ag6][PF6]14 (Fig. 2–4) (ESI†). The cage has an adamantane-like structure, with a Ru(n) tris-chelate unit at each of the four three-connected vertices which are arranged in an approximate tetrahedron. An Ag(i) bis-chelate unit occupies each of the six two-connected vertices. Thus the structure can be described as a tetrahedral array of Ru(n) ions with an Ag(i) ion lying in the centre of each Ru···Ru edge (Fig. 2), with every adjacent Ru(n)–Ag(i) pair connected by a bis-bidentate bridging ligand Lph.

The molecule lies astride a crystallographic C3 axis such that half of it is unique. This axis passes through Ag(2) and Ag(3) such that these lie on special positions with 50% occupancy in the asymmetric unit, whereas Ag(1) and Ag(4) are in general positions. There is a (non-crystallographic) C3 axis through each Ru(n) tris-chelate vertex, with all four being homochiral; thus the complex belongs to the pure rotation symmetry point group T which is a common consequence of removing mirror planes from high-symmetry polyhedra.
Fig. 3. A view of the complete complex cation of \([(\text{Ru}(\text{Lph})_3)_3\text{Ag}_6]^{14-}\). The two ligands coordinated to each Ag(i) have the same colour. Labels A and B denote the electron-deficient (pyrazolyl-pyridine) and electron-rich (phenyl) units involved in the pairwise π-stacking interactions.

Fig. 4. Space-filling views of the complex cation of \([(\text{Ru}(\text{Lph})_3)_3\text{Ag}_6]^{14-}\). (a) A view down one of the threefold axes, through a Ru(i) tris-chelate centre; (b) a view from the opposite side of the complex looking at one of the Ru3Ag3 faces, with one of the encapsulated [PF6]− anions (F atoms in green) visible through the portal.

The six Ag(i) ions lie on the three C3 axes associated with T symmetry of which one [the Ag(2)–Ag(3) axis, as mentioned above] occurs in the crystal structure; necessarily, all six Ag(i) ions have the same chirality associated with their two non-symmetrical chelating ligands. The nearest-neighbour Ru–Ag separations (i.e. along an edge spanned by a bridging ligand) lie in the range 8.86–9.32 Å, averaging 9.06 Å.

The flexibility of the ligands associated with the CH2 ‘ hinges’ allows them to adopt a conformation which maximises inter-ligand π-stacking – a key driver for assembly of such cages.1c,15 This can be seen in the view shown in Fig. 3, in which the octahedral disposition of the six Ag(i) ions is emphasised with these being placed top/bottom, left/right and front/back with each pair of Ag(i) ions lying on a C2 axis. In this view, Ag(4)/Ag(4A) form the ‘vertical’ C2 axis. The two ligands attached to each Ag(i) ion have the same colour (i.e. the twelve ligands are coloured in six sets of two). The ligands are disposed such that a central phenyl ring of a bridging ligand (denoted ‘B’ in Fig. 3) lies parallel to, and overlapping with, a pyrazolyl-pyridine unit of another ligand coordinated to the adjacent Ru(ii) ion (denoted ‘A’ in Fig. 3), forming a charge-assisted π-stack between electron-rich (phenyl) and electron-deficient (coordinated pyrazolyl-pyridine) ligands. In the view in Fig. 3 we can readily see four such A/B stacked pairs; there are necessarily, therefore, twelve such interactions overall – involving every phenyl group – as the orientations with Ag(1)/Ag(1A) and Ag(2)/Ag(3) as the ‘vertical’ axis are equivalent.

An alternative space-filling view, looking down one of the C3 axes associated with a Ru(ii) tris-chelate centre, is in Fig. 4a. The cage complex has an approximate cavity size of 178 Å3 (calculated assuming that the windows are blocked; Fig. S2, ESI†). The cavity is occupied by a tetrahedral array of four [PF6]− anions (Fig. 2a), each one blocking the window in one of the Ru3Ag3 faces of the cage, as shown in Fig. 4b in which three of the F atoms of the [PF6]− anion in that window can be clearly seen. The P⋯F separations between the four encapsulated anions are in the range of 5.44–5.61 Å, resulting in peripheral F⋯F contacts between anions of ≈ 3 Å which is the sum of the van der Waals’ radii of two F atoms. Each anion is involved in a range of CH⋯F interactions with ligand H atoms. Fig. S1(a) (ESI†) shows one of the anions embedded in the window in one of the Ru3Ag3 faces, with dotted lines indicating some of the short non-bonded C⋯F contacts (≤ 3.15 Å) which are indicative of weak hydrogen-bonding interactions between anion and ligand. This view also nicely shows how the array of six ligands around each Ru3Ag3 face forms a cyclic helicate with every ligand in the cycle having the same sense of ‘under and over’ around the ring. Fig. S1(b) (ESI†) shows how the four anions fill the cavity.

The structural integrity of the complex in solution was confirmed by ES mass spectrometry, which showed peaks corresponding to the species \([(\text{Ru}_3\text{Ag}_3(\text{Lph})_3)_3]^{14-}\) (n = 3, 4, 6), and also by 1H NMR spectroscopy (ESI†). The 1H NMR spectrum at room temperature is very broad, indicative of molecular motions [possibly associated with the highly flexible Ag(i) centres] at a rate comparable to the 1H NMR timescale. However at 75 °C the spectrum sharpened satisfactorily and showed the expected 20 independent 1H signals associated with one environment for Lph with no internal symmetry (Fig. S6, ESI†); this spectrum is considerably different from that of \(\text{fac-[Ru}(\text{Lph})_3]^{14-}[\text{PF}_6]_2\). Significantly, the chirality associated with the \(\text{Ag(}n\text{N})_3\) centres ensures that both independent sets of CH3 protons are now diastereotopic, giving two pairs of coupled doublets in the 4.5–5.5 ppm region (Fig. S5, ESI†). That this species is a large assembly is confirmed by its DOSY spectrum which clearly shows that all of its 1H signals belong to a single species which has a much lower diffusion rate \([\log D (\text{m}^2 \text{s}^{-1}) = -9.2]\) than \(\text{fac-[Ru}(\text{Lph})_3]^{14-}[\text{PF}_6]_2\) \([\log D (\text{m}^2 \text{s}^{-1}) = -8.4]\) (Fig. S7, ESI†).

Assembly of this cage with its adamantane-like structure thus relies on two different types of geometric control at specific metal sites. Firstly it requires the appropriate combination of metal vertices that are three-connected [each tris-chelate, Ru(ii) ion is connected to three Ag(i) ions] and two connected [each bis-chelate Ag(i) ion is connected to two Ru(n) ions]. This is achieved by using metal ions with different stereoelectronic preference, i.e. a combination of...
6-coordinate Ru(II) and 4-coordinate Ag(I) ions at alternating sites. Secondly, the structure relies on exclusive use of pre-formed, kinetically inert fac isomers of the [Ru(L^{3+})_3]^3- unit. We note that there are a few other examples of mixed-metal M_nM'_n(L_{10})_m complexes with an adamantane-like core structure.¹⁶,¹⁷ Many of these arise from one-pot reactions, but some – which use cyanide bridges along the M–M’ edges – are based on a kinetically stable, pre-formed hexacyanometallate unit as one precursor in a manner related to ours.¹⁷ In principle the method we have reported here should have high generality as it could be extended to other kinetically inert octahedral d⁶ metal complexes of the type fac-[M(PyPzH₂)]₃⁺ (M = Os, n = 2; M = Rh, Ir, n = 3 etc.).

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Notes and references


