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Bracing copper for the catalytic oxidation of C–H bonds

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Abstract

A structural unit found in the active site of some copper proteins, the histidine brace is comprised of an N-terminal histidine which chelates a single copper ion through its amino terminus NH₂ and the π -N of its imidazole side chain, and coordination by the τ -N of a further histidine side chain, to give an overall N₃ T-shaped coordination at the copper. The histidine brace appears in several proteins, including *lytic polysaccharide monooxygenases* (L)PMOs and *particulate methane monooxygenases* pMMOs, both of which catalyse the oxidation of substrates with strong C–H bonds (bond dissociation enthalpies \sim 100 kcal/mol). As such, the copper histidine brace is the focus of research aimed at understanding how Nature catalyses the oxidation of unactivated C–H bonds. In this Perspective, we evaluate these studies, which further give bioinspired direction to coordination chemists in the design and preparation of small molecule copper oxidation catalysts.

Introduction

Discovered in 2011, the histidine brace¹, or His-brace (“brace” from the old English noun meaning “a pair” and the verb “to clasp”) is the name given to a structural unit found in the active site of copper enzymes. It is comprised of an N-terminal histidine which chelates to a single copper ion through its NH₂ group and the π -N (also defined as δ -N) of its imidazole side chain. Coordination is completed by the τ -N (also, ϵ -N) of a further histidine side chain to give an overall T-shaped N₃ coordination geometry at the copper ion (Figure 1). The histidine brace appears in the proteins *lytic polysaccharide monooxygenases* (LPMOs; also termed PMOs²), the copper transport protein *CopC*,³ and *particulate methane monooxygenases* (pMMOs). Both LPMOs and pMMOs have roles in the terrestrial biological carbon cycle and have major industrial uses/potential. LPMOs catalyse the O₂ and co-factor-dependent oxidation of polysaccharides, leading to cleavage of a glycosidic bond,^{1,4} and attract attention because of their utility in commercial biofuel production.^{5,6,7,8} pMMOs catalyse the oxidation of CH₄ by O₂,⁹ and are enzymes of importance because of their potential to convert methane into a liquid fuel (e.g. MeOH).

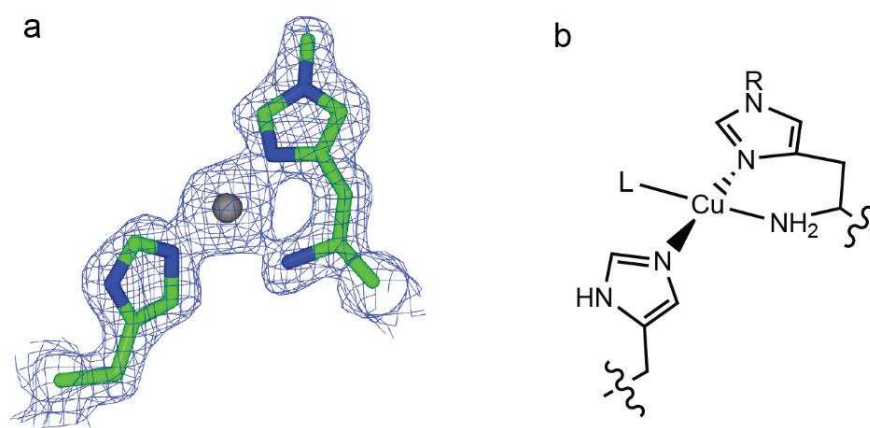


Figure 1. The copper histidine brace. a) Electron density contour map of histidine brace active site from an exemplar Cu(I)-AA9 LPMO (PDB code 5ACH), b) the copper histidine brace (R = Me or H).

Beyond their biological and industrial importance, both pMMO and LPMOs are of interest at a molecular level since they catalyse the oxidation of a strong (bond dissociation enthalpy ~95-103 kcal/mol) C–H bond,^{7, 10} at which the carbon atom is saturated. Thus, any oxidative mechanism employed by the enzymes requires the generation of highly oxidising intermediates. Accordingly, central questions surround the histidine brace, including how it facilitates the generation of such highly oxidising intermediates and what are their structural and electronic features. We address these and related questions in this Perspective by interweaving discussions of enzymatic structure/function relationships with those of synthetic model complexes. Our objective is to highlight the tremendous catalytic potential of enzymes or molecules that contain the histidine brace and the importance of this ligating unit on their possible mechanisms of action.¹¹

Discovery of the copper histidine brace

The first reports^{12, 13} of a new type of “non-catalytic” protein that bound to polysaccharides appeared in 2005. It was evident from crystal structures of these proteins that they had a metal-binding site in which a conserved N-terminal histidine chelated a metal ion. Later studies then demonstrated that these proteins were enzymes (LPMOs) that catalysed the oxidative degradation of polysaccharides and that the metal was essential for any activity.¹⁴ During this time the metal ion was variously assigned as Ca, Na, Zn and Mg,^{4, 13, 14, 15} incommensurate with the oxidoreductase properties of the enzyme. Adding to the confusion and highlighting the need for a small molecule understanding of metal-containing active sites, other studies had also proposed that the redox-active transition metals Co, Ni and Mn could also bind at the active site.¹⁴ The true identity of the metal as copper was determined in 2011 through the use of careful metal-ion binding studies, EPR spectroscopy¹ and activity studies.² This work¹ was the also first to name the active site as the histidine brace, a moniker now in wide use. These studies¹ further revealed that the copper histidine brace in LPMOs isolated from fungal sources features an N-methylation on the N-terminal histidine, a feature which had also been missed in earlier crystallographic studies.^{14, 15} LPMOs are now known to be widespread in Nature,¹⁶ with over 3500 examples already identified in the CAZy database for carbohydrate-active proteins.¹⁷ The enzymes have been classified on the basis of their amino-acid sequences into six distinct classes and given the codes AA9, AA10, AA11,¹⁸ AA13,^{19, 20} AA14²¹ and AA15²² (Figure 2).

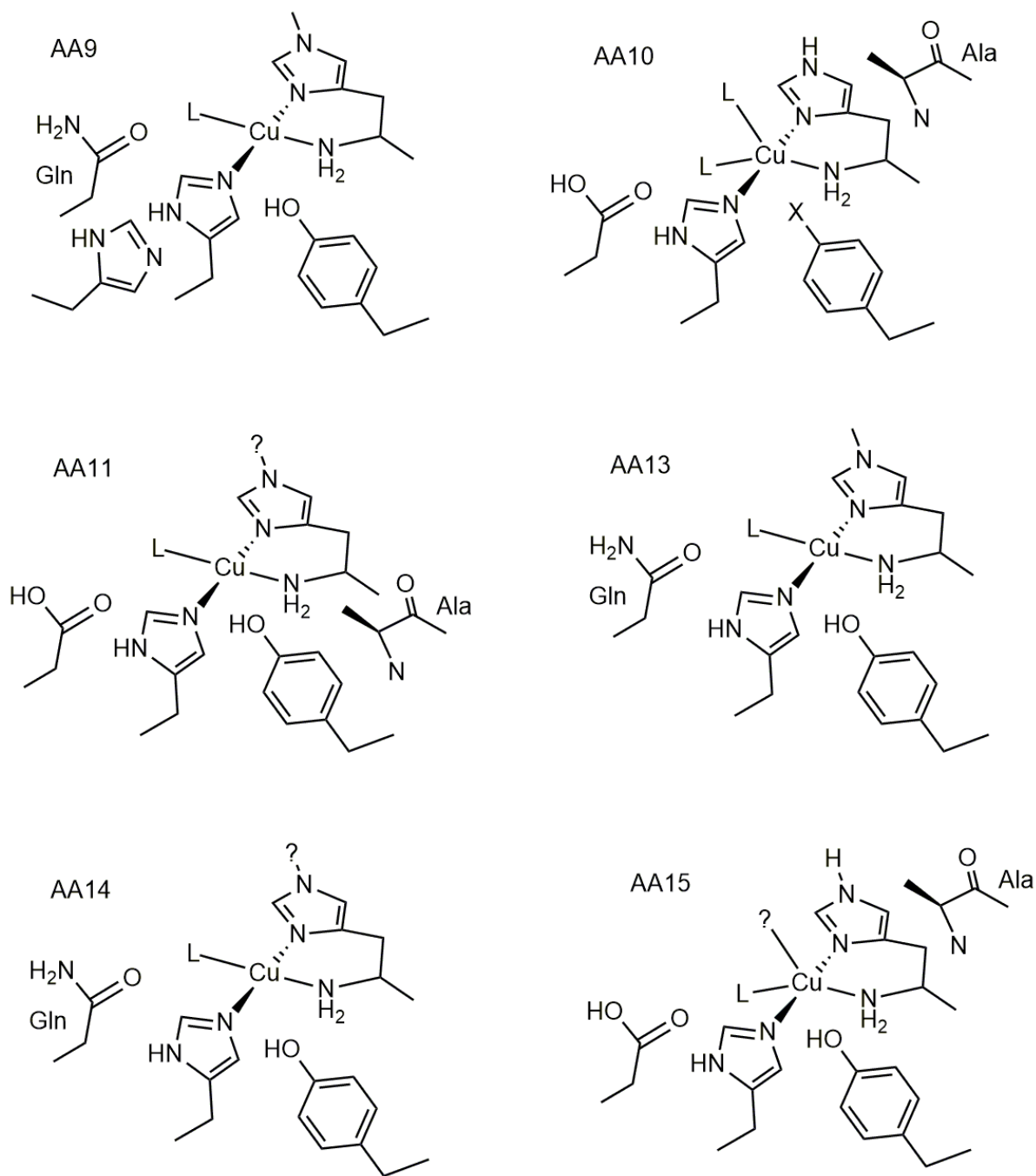


Figure 2. LPMO active sites. Cu(II) histidine brace structures from the active sites of LPMOs (X = H or OH in AA10s, L = H₂O or OH⁻), depicting some of the conserved secondary coordination sphere amino acid side chains. The wild type AA11 and AA14 LPMOs may contain a histidine side chain which is methylated, but this is unknown since the expression system (*Pichia pastoris*) used to express these genes appears to lack the necessary methylation apparatus.

The structure of the copper histidine brace in LPMOs

Various reviews have compiled the biochemical and catalytic aspects of LPMOs.^{6, 16, 23} These commentaries highlight the structural range of the active sites found within the enzymes, as shown not only from the 45 known crystal structures, but also the various EPR spectroscopy studies.^{23, 24} This diversity is evident in the identity and relative positions of amino acids in the secondary coordination spheres of the copper ion which vary to a significant degree from one LPMO class to

another.²⁵ The only common structural aspect of these LPMOs is the T-shaped coordination geometry of the histidine brace—a feature recognised in recent work on artificial copper-metalloenzymes.²⁶ In fact, there is so little structural variation in the primary coordination sphere of the copper (Supplementary Table 1) that the copper histidine brace may be viewed as a geometrically-fixed structural unit, the reactivity of which is finessed through secondary coordination sphere interactions.^{24, 27} Of these secondary interactions, those amino acid side chains which are capable of forming hydrogen bonds to any exogenous substrate bound to the copper ion are of particular importance insofar as highly reactive oxygen-based ligands (e.g superoxide, hydroperoxide, oxyl) are potentially affected through such interactions. Indeed, site-directed mutagenesis studies have already implicated the importance of the secondary glutamine and histidine residues in AA9 LPMOs.²⁵ Other residues, for example the tyrosine in AA9 LPMOs (Figure 2), also have significant effects on catalytic rates/turnover numbers, while others near the active site (e.g. alanine in AA10 LPMOs) may have roles in controlling access of exogenous ligands to the axial coordination site on the copper,^{14, 27} although these residues do not appear to affect the regioselectivity of polysaccharide oxidation.²⁸ In terms of secondary interactions, a recent crystal structure of an AA9 LPMO in contact with its saccharide substrate²⁹ revealed the details of substrate-active site interactions which modulate the electronics and therefore reactivity of the Cu histidine brace.⁶ Included in these changes are an increased covalency in the metal-ligand interactions and a change in hydrogen bonding between the amino terminus and a nearby water molecule. In this context,³⁰ an on-going debate within the field is whether substrate binding is coupled to the activity of the enzyme, such that the generation of highly oxidising copper species that carry out site-specific oxidation on the substrate only takes place in the presence of substrate (as is the case of cytochrome P450, for example).³¹

The primary coordination geometry around the copper in the histidine brace might be as expected from a straightforward consideration of the lone pair orbitals of the coordinating nitrogen atoms. Two *trans* sp^2 -hybridised nitrogen atoms coordinate with average distances of Cu(II)–N(1) and Cu(II)–N(2) ~ 2.0 Å (Supplementary Tables 1 and 2). The angle of twist between the two best fit planes associated with each of the imidazole rings of the two histidines is around 65° and appears to be dependent on the presence of the metal ion, as the angle is around 30° in most of the *apo* structures available (AA10s). This twist angle is significantly different from the analogous angle found in many *trans*-N-heterocycle Cu(II) complexes, where it is usually less than 30° , a result of the conformational demands of the chelate rings of ligands (Supplementary Table 3). The significance of this difference is unclear. If, however, such a twist angle influences the reactivity of the copper histidine brace, particularly in copper-oxygen intermediates, then some role for the π -interaction capacities of the N-heterocycles seems to be worth consideration in the design of ligands aimed at emulating the active site. The coordination sphere of the copper in the histidine brace is completed from coordination in the ‘bridgehead’ position by the NH₂ amino terminus of the protein. In accord with the sp^3 nature of the lone pair on this group, the average Cu–N(amino) bond length from X-ray structures is ~ 2.2 Å, although some care must be exercised in reading too much into this distance since the structures of LPMOs are bedevilled by photo-reduction of the Cu(II) ion to Cu(I) during X-ray diffraction experiments.^{27, 32} For instance, the published DFT calculations of the histidine brace consistently calculate the Cu–NH₂ distance to be 2.1 Å, some 0.1 Å shorter than those usually measured in crystal structures.³³

The histidine brace N₃ coordination is such that the N–Cu–N angles all fall within the range 87° to 106° , averaging around 94° . From this feature it is reasonable to assume that the lone pairs of the nitrogen atoms are maximally overlapped with the $d(x^2-y^2)$ orbital of the copper, necessarily requiring that the redox active molecular orbital (RAMO) on the copper is at high energy compared

to the rest of the d-orbital manifold. Such a high energy redox-active orbital has implications for the reactivity of the copper ion, the first of which is that the strength of the copper(II)-superoxide bond [formed upon initial reaction of O₂ with Cu(I)] is maximised through a lowered difference in energy between the frontier orbital on the copper d(x²-y²) and the oxygen (anti-bonding π_g). The high strength of the bond offsets the energetic penalty associated with the one-electron reduction of O₂ by Cu(I).³⁴ A second feature of the d-orbital configuration of the histidine brace is that any d-orbital mixing can be expected to be low, lending the RAMO near pure d(x²-y²) orbital character, as supported by the values of hyperfine coupling constants observed in the EPR spectra of LPMOs in the presence of substrates.²⁹ We therefore speculate that any inner-sphere redox chemistry on the exogenous ligands associated with the copper is most likely to occur with ligands in the equatorial plane of the copper coordination sphere, thus hindering what might be potentially deleterious generation of reactive oxidation species through electron transfer mechanisms at the axial position of the copper ion.³⁵

An ongoing debate about the mechanism of LPMOs has included discussion about the protonation state of the NH₂ binding unit. There is some circumstantial evidence from studies of synthetic complexes that deprotonation of this group could occur. For instance, Cu(III) oligopeptide species exhibit pK_as for Cu(III) -NH₂ amino terminus moieties of ca. 8.8, a value approaching biological viability.³⁶ Also, [CuOH]²⁺ complexes stabilised by anionic amidato type ligands are competent for hydrogen atom abstraction from C-H bonds with strengths of >90 kcal mol⁻¹ (see below). Accordingly, deprotonation of the NH₂ in LPMOs to give an imido group gains credence, and such a possibility was raised in the 2011 paper which announced the copper histidine brace active site in LPMOs.¹ The notion gained further momentum from neutron diffraction experiments performed on AA10 LPMOs, which appeared to show a lack of density around the amino terminus nitrogen atom³⁷ and a notable average Cu-NH₂-Cα angle of ~117°; confirmation by spectroscopy is nonetheless needed before a definitive assignment can be made. If true, however, the enhanced electron donation by the deprotonated moiety would raise the d(x²-y²) orbital high enough in energy such that a formal Cu(III) species becomes realistic.³⁶ On the other hand, recent DFT calculations on an LPMO in the presence of its oligosaccharide substrate suggested that deprotonation of the amino terminus in LPMOs potentially reduces the capacity of a [CuOH]²⁺ species to perform hydrogen atom abstraction.³³ If a deprotonated amino group is an essential feature of the reactivity of the copper histidine brace, then its presence presents challenges for ligand design for small molecule mimics, since the NH₂ 'bridgehead' in ligand design is commonly used as an anchor point for the nitrogen heterocycles (usually linked to the nitrogen atom of the bridgehead through a methylene or ethylene linker group).

The structure of the copper histidine brace in pMMO

The exact nature of the active site species responsible for attacking the strong C-H bond in methane in pMMO is controversial.^{38, 39, 40} Early EXAFS, XRD and EPR work on this enzyme had concluded that the active site contains not a single copper ion, but two in close proximity to each other (Cu...Cu, ca 2.6 Å).^{38, 41, 42, 43, 44} An alternative proposal suggested that the active site was a tricopper cluster.⁴⁰ These hypotheses, supported by DFT investigations,^{45, 46, 47} spurred many scientists to investigate multicopper structures and their analogous multicopper-oxygen complexes as potential model complexes, catalytic or not, of the active site of pMMO.^{39, 48, 49, 9, 38, 43, 50} However, notwithstanding the extensive work into multi-copper centers and in the light of the discovery of the histidine brace, there is now growing evidence in support of an alternative proposal in which pMMO contains a mononuclear Cu active site. Indeed, the 2011 discovery of the histidine brace in LPMOs prompted a specific proposal from the authors that the active site of pMMO could be mononuclear in copper

and that such a possibility ‘warranted investigation’.^{1, 5} Further work on LPMOs by some of the same authors showed that the EXAFS spectrum of Cu(II)-LPMO was similar to that reported for pMMO, lending further weight to the possibility of a mononuclear copper active site in pMMOs.³⁴ Since then, other commentators have drawn attention to the unusual, if not unreasonable, coordination geometry of the active site histidines of pMMOs if two copper ions are to be accommodated therein, as illustrated in the original X-ray structure by the angle made by the ‘coordinating’ histidine side chains to the two coppers (Figure 3).³⁹ Most revealingly, and perhaps stimulated by the discovery of LPMOs, a reappraisal of the original X-ray data of pMMO (3RGB) by some of the original presenters of the dicopper model using quantum crystallographic methods, led to the conclusion that the electron density at the active does not support a two copper assignment, and that the data are best interpreted as a mononuclear copper active site.⁵¹ This study was coupled to a DFT calculation of methane oxidation by a mononuclear copper site (a $[\text{CuO}]^+$ species, as discussed below) coordinated by the histidine brace, also arriving at the conclusion that oxidation of methane could occur without significant activation barriers. Furthermore, a very recent EXAFS and EPR study has backed-up the findings from the quantum crystallographic work.⁵² With the weight of evidence now swinging away from a dinuclear active site within pMMO and with the discovery of the histidine brace in LPMOs, there are major repercussions for our understanding of the mechanism of O_2 activation and substrate oxidation at monomeric copper sites. Moreover, given that the C–H bonds which both LPMOs and pMMOs oxidise are strong and unactivated, the postulate of a histidine brace coordinating a single copper is inspiring coordination chemists interested in preparing synthetic copper complexes which themselves could then be used as powerful oxidation catalysts.⁵³

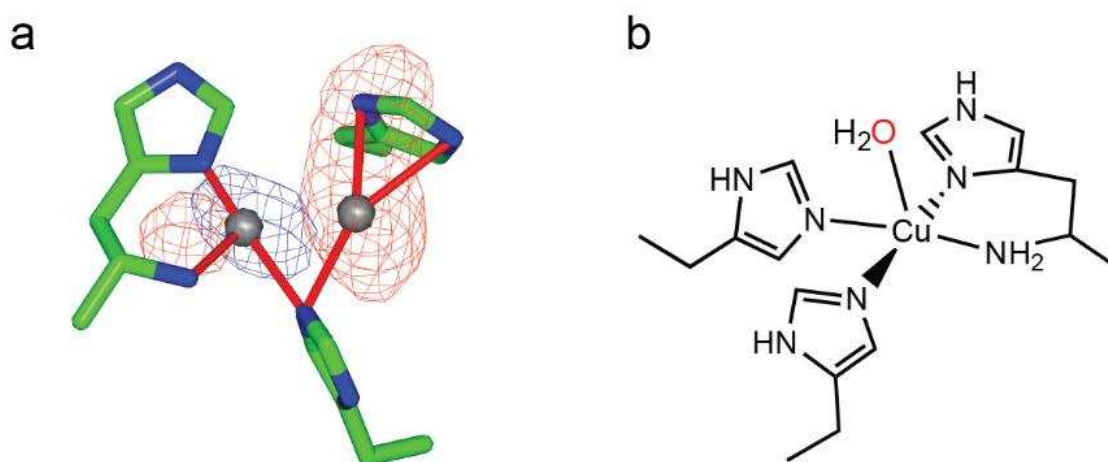


Figure 3. The changing structure of the active site of pMMO. a) Original electron density X-ray maps from one pMMO structure (3RGB),⁵⁰ showing non-canonical bonding arrangements of histidines (Cu-His bonds shown as red cylinders), a single $2F_o-F_c$ peak for a mononuclear Cu and large regions of negative F_o-F_c electron density associated with the modelling of two Cu positions, b) proposed structure of the pMMO site as now re-evaluated by quantum crystallography.⁵¹

Mechanistic Considerations

Proposals exist for the generation of free hydroxyl radicals by LPMOs, but—while certainly viable from the perspective of oxidative cleavage of a recalcitrant substrate—these notions appear incommensurate with the observed site-specific attack.⁵⁴ Attention, therefore, has focussed on reactive copper-oxygen species as the key oxidative intermediate(s). There is currently no single consensus on the mechanism by which LPMOs and pMMOs activate O_2 (or H_2O_2)⁵⁴ to hydroxylate

strong C–H bonds. Contributing to the lack of agreement for LPMOs is the fact that different LPMOs classes may employ different oxidative mechanisms during catalytic turnover.^{6, 23} There are also different opinions as to whether O₂ or H₂O₂ is the natural substrate for LPMOs.^{54, 55} The absence of a consensus may also reflect the inherent difficulty of studying what are naturally highly reactive intermediates, further highlighting the importance of small molecule studies that are capable of providing precedent for structural proposals. For example, the reaction of H₂O₂ with mono-copper complexes is known from small molecule studies to be complicated by unintentional reaction pathways.⁵⁶ Such pathways confound what might otherwise be simple mechanistic interpretations of how oxidised products are generated from the reaction with H₂O₂⁵⁴—a caution also made by other authors.^{8, 56} In this regard, studies of synthetic copper-oxygen complexes provide fundamental structural, spectroscopic, and mechanistic information relevant to their viability as intermediates in LPMOs and pMMOs.^{49, 57, 58, 59, 60} Yet, compared to the plethora of multicopper-oxygen complexes that have been characterised, relatively few mononuclear compounds have been isolated, in large part due to the propensity of such small molecule species to form complexes with oxo, hydroxo, or peroxy bridges. A key synthetic strategy thus has been to use sterically hindered supported ligands to inhibit di- or multicopper complex formation.^{61, 62, 63, 64}

(a) Intermediates with intact O-O bond

Several commentators have proposed what the key oxidative species might be in LPMOs.^{6, 39} A Cu(II)-superoxide ([CuO₂]⁺) intermediate is proposed to form upon reaction of a Cu(I) center with O₂ and surely merits further investigation in enzyme studies.^{34, 65} Extensive synthetic studies of the reactions of appropriate Cu(I) precursors with O₂ led to the identification of [CuO₂]⁺ species that may adopt multiple binding modes (η^1 and η^2) and extents of electron transfer to the bound O₂, bracketed by the resonance structure extremes of a Cu(II)-superoxide and a Cu(III)-peroxide.^{60, 62, 63, 64, 66} Studies of the reactivity of such complexes with C–H bonds generally have indicated poor capabilities to perform hydrogen atom transfers (HATs), with most examples where HAT was observed involving attack at C–H bonds significantly weaker than those of the substrates of LPMO and pMMO.⁵⁹ While there are exceptions, the reactivity observed to date for [CuO₂]⁺ compounds with exogenous substrates, along with results from theory,^{67, 68} suggest that this core may not be responsible for attacking substrate C–H bonds in the enzymes. For this and other reasons, attention has focused on species further activated by the addition of protons and electrons.

Addition of one proton and one electron to the [CuO₂]⁺ unit give a feasible intermediate with the core [CuOOH]⁺. This concept is informed by the numerous examples of synthetic [CuOOR]⁺ complexes which have been reported,^{49, 69} with most prepared by the reaction of ROOH and base with suitable Cu(I) or Cu(II) precursors and characterised via spectroscopy; a few X-ray structures are available.^{69, 70, 71} In a more biomimetic route, a [CuOOH]⁺ intermediate was proposed to be formed by a proton-coupled electron transfer (PCET) of a [CuO₂]⁺ complex with a 1-electron reductant (Fc or Fc*) in the presence of protic acid (trifluoroacetic acid).⁷² The recent identification of a [CuOOH]⁺ unit stabilised by hydrogen bonds in an engineered protein is also a notable advance in this arena.²⁶ Examples of [CuOOR]⁺ complexes reacting with C–H bonds are numerous, with reports of hydroxylations of strong *sp*³ C–H bonds being particularly relevant here.⁷³ Most mechanistic proposals invoke O–O bond homolysis to yield a [CuO]⁺ (“Cu(II)-oxyl”) unit that is responsible for attacking the C–H bond.⁷⁴ Direct evidence for this pathway is scant, however, and differentiating this pathway from one involving direct hydrogen atom abstraction by the [CuOOR]⁺ unit presents experimental challenges (in one study, a case for the latter pathway has been made).⁷⁵ Other work has shown that Cu–O bond homolysis can yield an alkylperoxy radical that attacks substrate.⁵⁶ While more reactivity and mechanistic studies of [CuOOR]⁺ complexes are needed, the synthetic work

accomplished so far supports their feasibility as intermediates along O₂ activation pathways, but it is unclear whether they are sufficiently potent reagents to be responsible for the enzymatic C–H bond cleavage events.

(b) Intermediates derived from O–O bond scission

A key reactive intermediate in LPMOs proposed primarily on the basis of DFT calculations (of varying merit, see reference²³ for detailed comment) is the Cu(II)-oxyl, [CuO]⁺.^{63, 67, 76} Beyond LPMOs, such a species has long been proposed to be involved in monomeric copper oxidation catalysts/enzymes.^{49, 77} The proposal of [CuO]⁺ as a reaction intermediate is particularly appealing in view of the results of computational work that support it being highly reactive with strong C–H bonds^{78, 79, 80} and experimental gas phase studies of the [CuO]⁺ ion focused on its ability to hydroxylate methane.⁸¹ Yet, no examples of complexes with this unit have been identified conclusively in condensed phase. It thus represents a synthetic ‘holy grail’, the attainment of which would provide important precedent for the involvement of this unit in oxidation catalysis.

In terms of LPMO mechanisms, recent DFT calculations have raised the possibility that the [CuO]⁺ species in LPMOs can be formed following O–O bond scission of hydrogen peroxide by the Cu(I) form of the enzyme leading to the formation of the hydroxyl radical and a Cu(II)–OH unit, “caged” together within a cavity formed between the active site and substrate. Within the cage the hydroxyl rebounds to react with the [Cu–OH]⁺ giving a [CuO]⁺ and a water molecule.⁶⁸ While formation of the [CuO]⁺ or hydroxyl radical requires the scission of the O–O bond in oxygen (or peroxide), a step which would be expected to have a high kinetic barrier due to the reorganisation energy associated with 1-electron reductive cleavage of a O–O bond, the calculated barrier for O–O scission is only 5.8 kcal/mol.¹¹ This normally prohibitive barrier is ostensibly reduced by the presence of a secondary coordination sphere glutamine in LPMOs (see above), which forms hydrogen bonds to the peroxide before O–O bond cleavage, thereby orientating O–O σ* orbital towards the redox active d(x²–y²) orbital on the copper.

From a small molecule perspective, well-characterised monocopper-oxygen intermediates lacking an O–O bond are limited to a protonated version of [CuO]⁺, formulated as [CuOH]²⁺. Examples of complexes with the latter have been prepared by 1-electron oxidation of [CuOH]⁺ precursors (Figure 4).^{82, 83, 84, 85, 86} Importantly as described above, the [CuOH]²⁺ complexes (described formally as Cu(III)-OH species) are capable of attacking strong C–H bonds, and thermodynamic and kinetic studies have shed important light on the basis for this high reactivity. Particularly notable in this regard is the formation of a strong O–H bond (BDE ~90 kcal/mol) in the [Cu(OH₂)]²⁺ product of hydrogen atom abstraction from substrate. This thermodynamic driving force may be directly related to the strong electron-donating character of the supporting ligand, which while mitigating the oxidising power of the complex enhances the basicity of the hydroxide. In essence, the ligand stabilises the high oxidation state of the [CuOH]²⁺ core (enabling it to be observed and characterised) while at the same time increasing the thermodynamic driving force for hydrogen atom abstraction. These findings, as well as relevant theory,^{67, 84} suggest that (a) the [CuOH]²⁺ unit could be considered to be a feasible oxidant in LPMOs and pMMOs,⁷⁶ and (b) the role of the anionic ligand in stabilising the complexes while enhancing their reactivity with C–H bonds may be similarly filled by a deprotonated histidine brace in the enzymes. On the other hand, arguing against the [CuOH]²⁺ unit as the active oxidant in the enzymes is the likely inability of the initial product of hydrogen atom abstraction, the [Cu(OH₂)]²⁺ core, to undergo ‘rebound’ to yield the final hydroxylated organic product.

Other possible roles of the deprotonated histidine brace may be to enhance the basicity of (a) a [CuO]⁺ species, potentially enhancing its HAT reactivity, or (b) the initially formed 1:1 adduct [CuO₂]⁺.

Protonation of this adduct would yield a novel $[\text{CuOOH}]^{2+}$ group. The viability of this unit was recently supported by the preparation of complexes with $[\text{CuOOR}]^{2+}$ cores ($\text{R} = \text{cumyl}, \text{t-Bu}$) via 1-electron oxidation of $[\text{CuOOR}]^+$ precursors at low temperatures, using the same supporting ligand shown in Figure 4a (left, $\text{X} = \text{H}$).⁶⁹ While most $[\text{CuO}_2]^+$ complexes behave as electrophiles,⁶² $[\text{CuO}_2]^+$ core supported by the electron donating dicarboxamide ligand exhibits nucleophilic character, presumably derived from the overall -1 charge of the complex.^{87, 88} Taken together, the synthetic work provides some support for the notion that the deprotonated form of the histidine brace may facilitate O_2 activation and stabilise high-valent intermediates. These ideas are illustrated in Figure 4b, which features structural postulates for the initial 1:1 Cu/O_2 adduct, the $[\text{CuOOH}]^{2+}$ core that might be formed upon protonation of the adduct, and $[\text{CuO}]^+$ and $[\text{CuOH}]^{2+}$ cores, all stabilised by the deprotonated histidine brace (details of the mechanism of O-O bond scission by protons and electrons are omitted, but represent important targets of research). It is tempting to suggest that similar species may be involved in methane hydroxylation by pMMOs.

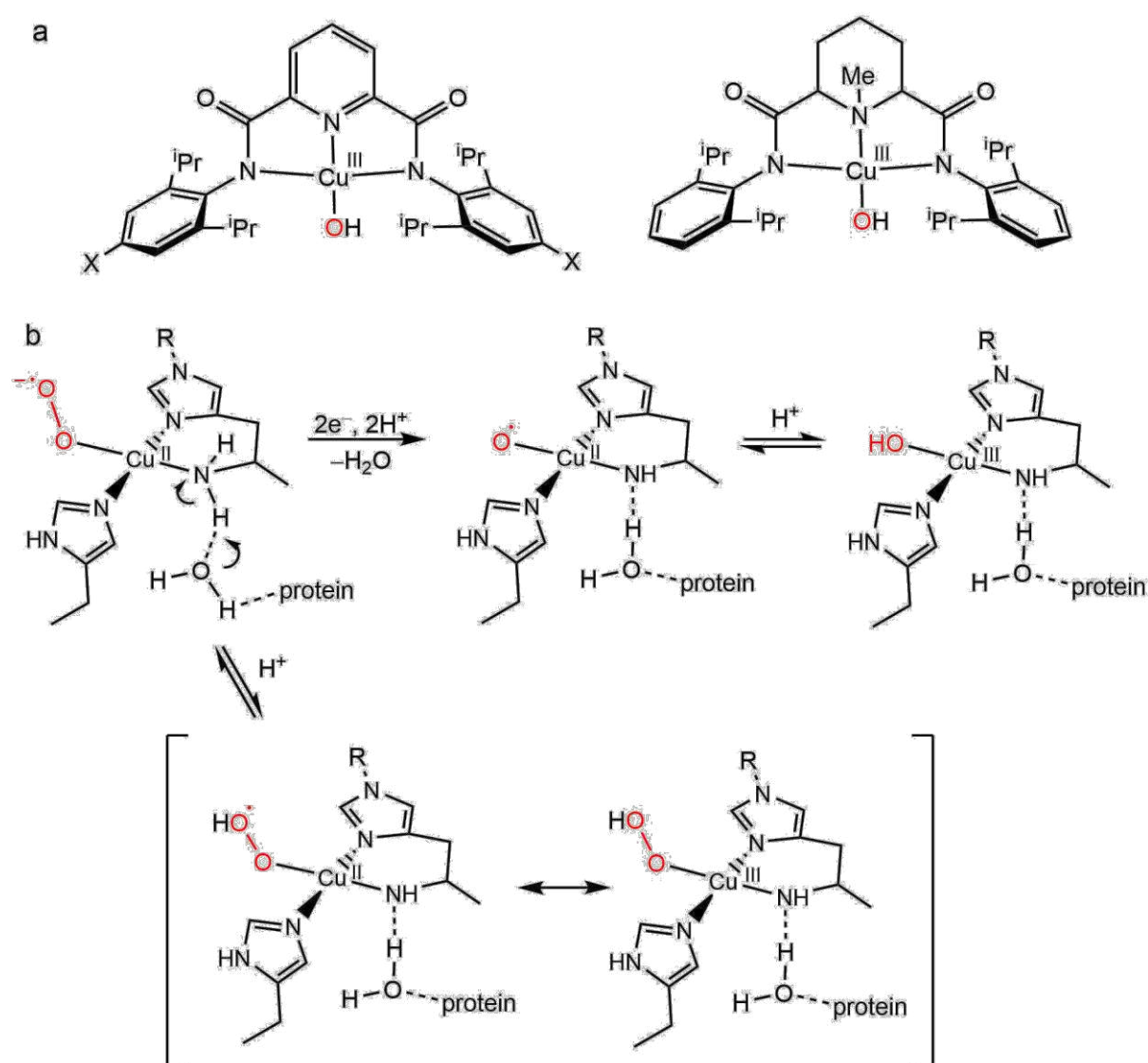


Figure 4. Deprotonation of the amino terminus. a) Complexes with the $[\text{CuOH}]^{2+}$ core ($\text{X} = \text{H}$ or NO_2), b) Proposed intermediates in LPMO highlighting the possible stabilising role of the deprotonated histidine brace ($\text{R} = \text{H}$ or CH_3). Only ligands in the equatorial plane are shown.

Conclusions

The work on LPMOs, pMMOs, and small molecule copper-oxygen complexes now brings into a new light the oxidative power of mononuclear copper species, especially those in which the histidine brace is part of the coordination sphere. The histidine brace appears to be a key common structural element that may facilitate the formation of and stabilise high-valent intermediates, potentially *via* deprotonation of the amino unit. The message from these studies is that monomeric copper sites, especially those which have a histidine brace coordination geometry, are powerful oxidation catalysts, ones that Nature has already recruited for use in the oxidation of recalcitrant biomass and ones that coordination chemists should now see as an attractive and fruitful target for their own studies.

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Author contributions

WBT and PHW wrote the manuscript, GJD analysed LPMO and MMO protein structures and contributed to the writing of the manuscript. LC prepared metrical data for structure analysis and contributed to the writing of the manuscript.

Competing interests

The authors declare no competing interests.

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