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Kinetics of CH$_2$OO reactions with SO$_2$, NO$_2$, NO, H$_2$O and CH$_3$CHO as a function of pressure

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Abstract

Kinetics of CH$_2$OO Criegee intermediate reactions with SO$_2$, NO$_2$, NO, H$_2$O and CH$_3$CHO and CH$_2$I radical reactions with NO$_2$ are reported as a function of pressure at 295 K. Measurements were made under pseudo-first-order conditions using flash photolysis of CH$_2$I$_2$/O$_2$/N$_2$ gas mixtures in the presence of excess co-reagent combined with monitoring of HCHO reaction products by laser-induced fluorescence (LIF) spectroscopy and, for the reaction with SO$_2$, direct detection of CH$_2$OO by photoionisation mass spectrometry (PIMS). Rate coefficients for CH$_2$OO + SO$_2$ and CH$_2$OO + NO$_2$ are independent of pressure in the ranges studied and are $(3.42 \pm 0.42) \times 10^{-11}$ cm$^3$ s$^{-1}$ (measured between 1.5 and 450 Torr) and $(1.5 \pm 0.5) \times 10^{-12}$ cm$^3$ s$^{-1}$ (measured between 25 and 300 Torr), respectively. The rate coefficient for CH$_2$OO + CH$_3$CHO is pressure dependent, with the yield of HCHO decreasing with increasing pressure. Upper limits of $2 \times 10^{-13}$ cm$^3$ s$^{-1}$ and $9 \times 10^{-17}$ cm$^3$ s$^{-1}$ are placed on the rate coefficients for CH$_2$OO + NO and CH$_2$OO + H$_2$O, respectively. The upper limit for the rate coefficient for CH$_2$OO + H$_2$O is significantly lower than has been reported previously, with consequences for modelling of atmospheric impacts of CH$_2$OO chemistry.

1. Introduction

Criegee intermediates, carbonyl oxide biradicals with the general formula CR$_2$OO, are principally produced in the atmosphere following ozonolysis of unsaturated volatile organic compounds (VOCs) and are key species in the tropospheric oxidation of both biogenic and anthropogenic compounds.$^{1,2}$ The exothermicity of ozonolysis reactions leads to production of vibrationally excited Criegee intermediates with sufficient energy to undergo unimolecular decomposition to products including OH and HO$_2$,$^{3-6}$ representing a significant source of these important oxidising species in certain important environments.$^{7-9}$ However, collisional quenching of the nascent excited Criegee intermediate by N$_2$ or O$_2$, to produce stabilised Criegee intermediates, is competitive with the unimolecular decomposition processes at ambient pressures.$^{1,5}$ and
reactions of stabilised Criegee intermediates have the potential to impact atmospheric budgets of NO\textsubscript{x} (NO\textsubscript{x} = NO + NO\textsubscript{2}), NO\textsubscript{3}, O\textsubscript{3}, HO\textsubscript{x} (HO\textsubscript{x} = OH + HO\textsubscript{2}), SO\textsubscript{2}, H\textsubscript{2}SO\textsubscript{4}, sulfate aerosol and secondary organic aerosol (SOA).\textsuperscript{5,10-17}

Despite their potential importance in atmospheric chemistry, and thus in the assessment and prediction of issues such as air quality and climate change, direct observations of Criegee intermediates have only recently been achieved.\textsuperscript{10-12, 18-20} Kinetics and product yields of Criegee intermediate reactions currently used in atmospheric models are subject to large uncertainties, owing to the reliance of previous investigations on indirect techniques involving measurements of stable species in complex ozonolysis experiments, in which there are several potential sources and sinks of the measured species.\textsuperscript{1,2} Welz \textit{et al.}\textsuperscript{10} reported the first direct measurements of Criegee intermediate kinetics, where the photolysis of CH\textsubscript{2}I\textsubscript{2} in the presence of O\textsubscript{2} was used to generate the CH\textsubscript{2}OO Criegee intermediate at low pressure (4 Torr) and, using synchrotron photoionisation mass spectrometry (PIMS) at the Advanced Light Source (ALS), demonstrated unequivocally that the Criegee intermediate, CH\textsubscript{2}OO, was being monitored:

\begin{align}
\text{CH}_{2}\text{I}_{2} + \text{hv} & \rightarrow \text{CH}_{2}\text{I} + \text{O}_{2} & (R1) \\
\text{CH}_{2}\text{I} + \text{O}_{2} & \rightarrow \text{CH}_{2}\text{OO} + \text{I} & (R2a)
\end{align}

While reactions of CH\textsubscript{2}OO with NO and water vapour were reported to be slow, the reactions of CH\textsubscript{2}OO with SO\textsubscript{2} and NO\textsubscript{2} were shown to be significantly faster than indicated by the indirect methods. Rate coefficients for both CH\textsubscript{2}OO + SO\textsubscript{2} and CH\textsubscript{2}OO + NO\textsubscript{2}, measured at a pressure of 4 Torr and temperature of 298 K, were both approximately 1000 times greater than previously assigned, implying a more significant role of Criegee intermediate chemistry in the atmosphere than expected.

The ability to produce CH\textsubscript{2}OO following photolysis of CH\textsubscript{2}I\textsubscript{2} in the presence of O\textsubscript{2}\textsuperscript{10} has also facilitated spectroscopic investigations of CH\textsubscript{2}OO in the infrared\textsuperscript{19} and ultraviolet,\textsuperscript{20} and has been used to demonstrate the production of NO\textsubscript{3} in the reaction of CH\textsubscript{2}OO with NO\textsubscript{2}.\textsuperscript{21} Subsequent work at the ALS has investigated the reactions of CH\textsubscript{2}OO with acetone, acetaldehyde and hexafluoroacetone at low pressures,\textsuperscript{11} with theoretical investigation\textsuperscript{22} of the reaction between CH\textsubscript{2}OO and acetaldehyde (CH\textsubscript{3}CHO) indicating pressure dependence of the reaction and collisional stabilisation of nascent reaction adducts to produce secondary ozonides (SOZs) at higher pressures which subsequently decompose to generate organic acids.

Taatjes \textit{et al.}\textsuperscript{12} have also recently demonstrated production of the CH\textsubscript{3}CHOO Criegee intermediate following photolysis of CH\textsubscript{3}CHI\textsubscript{2} in the presence of O\textsubscript{2}. The structure of the CH\textsubscript{3}CHOO Criegee intermediate gives rise to the possibility of syn- and anti- conformers, with the conformers sufficiently different in energy, and with a barrier to conversion, leading to the potential for their behaviour as distinct species. Using the synchrotron PIMS technique, Taatjes \textit{et al.}\textsuperscript{12} were not only able to identify both the syn- and anti-CH\textsubscript{3}CHOO conformers, but were also able to assign separate rate coefficients for reactions of the two conformers with SO\textsubscript{2} and water vapour. The anti-conformer was shown to display greater reactivity.
towards both SO$_2$ and H$_2$O compared to the syn-conformer, with rate coefficients for reactions of both syn- and anti- conformers with SO$_2$ greater than previously expected.

Field observations in a boreal forest in Finland have provided further evidence for rapid reactions between Criegee intermediates and SO$_2$, with measurements identifying the presence of oxidising species other than OH which are able to oxidise SO$_2$ to SO$_3$ and ultimately to produce H$_2$SO$_4$. The presence of the unknown oxidising species was shown to be related to emissions of biogenic alkenes, and it was postulated that Criegee intermediates may be responsible, with laboratory measurements of H$_2$SO$_4$ production during alkene ozonolysis reactions in the presence of SO$_2$ and OH scavengers providing further support for the action of Criegee intermediates as atmospheric oxidants of SO$_2$.

Implementation of increased Criegee intermediate + SO$_2$ reaction rates in atmospheric models has been shown to improve model simulations of H$_2$SO$_4$ in forested regions in Finland and Germany, and global modelling has shown that while global production of H$_2$SO$_4$ increases by only 4%, there are increases of up to 100% in the boundary layer in tropical forests. Further modelling work has shown that reactions of Criegee intermediates with SO$_2$ can compete with OH + SO$_2$ in a number of regions, and that Criegee + SO$_2$ reactions may be the dominant removal mechanism for SO$_2$ in certain areas and are major contributors to sulfate aerosol formation on a regional scale. Air quality modelling over the U.S. displayed limited impacts of increased Criegee + SO$_2$ reaction rates on sulfate aerosol production in this region, but the impacts were shown to be highly dependent on the competition between Criegee + SO$_2$ and Criegee + H$_2$O, with a combination of increased Criegee + SO$_2$ and decreased Criegee + H$_2$O reaction rates leading to enhanced sulfate aerosol concentrations. However, such studies have largely been based on the low pressure data for CH$_2$OO + SO$_2$ reported by Welz et al. and there is considerable uncertainty regarding the upper limit for CH$_2$OO + H$_2$O.

Theoretical work has provided support for rapid reactions between Criegee intermediates and SO$_2$, with reactions proceeding via the initial barrierless formation of a cyclic secondary ozonide, and has enabled prediction of potential effects of pressure. For CH$_2$OO + SO$_2$, it has been predicted that the reaction products at atmospheric pressure will be a mixture of HCHO + SO$_3$ (~68%), formyl sulfinic ester (HC(O)OS(O)OH) (~15%) and a singlet bisoxy diradical (CH$_2$(O)O) + SO$_2$ (~17%). In contrast, reactions of larger Criegee intermediates, including CH$_3$CHOO, at ambient pressures are expected to result in production of stabilised secondary ozonide species, with little formation of SO$_3$, and therefore little impact on H$_2$SO$_4$ and sulfate aerosol. Investigation of the reaction products and pressure dependence of Criegee intermediate reactions is thus essential to the accurate determination of their atmospheric impacts.

The yield of CH$_2$OO Criegee intermediates following CH$_3$I photolysis in O$_2$ was studied by Huang et al., and in our previous work, as a function of pressure. Both investigations indicate that the initial reaction between CH$_3$I radicals and O$_2$ (R2) produce a chemically activated species, CH$_3$IO$_2^+$, which decomposes at
low pressures to produce CH\(_2\)OO + I (R2a), but is also collisionally stabilised at higher pressures to produce the CH\(_3\)IO\(_2\) peroxy radical (R2b).

\[
\text{CH}_3\text{I}_2 + \text{hv} \rightarrow \text{CH}_3\text{I} + \text{I} \quad \text{(R1)}
\]

\[
\text{CH}_3\text{I} + \text{O}_2 \rightarrow \text{CH}_3\text{IO}_2^\# \quad \text{(R2)}
\]

\[
\text{CH}_3\text{IO}_2^\# \rightarrow \text{CH}_2\text{OO} + \text{I} \quad \text{(R2a)}
\]

\[
\text{CH}_3\text{IO}_2^\# + \text{M} \rightarrow \text{CH}_3\text{IO}_2 + \text{M} \quad \text{(R2b)}
\]

Our previous work\(^{26}\) indicates a yield of \(~18\%\) CH\(_2\)OO following photolysis of CH\(_2\)I\(_2\) in air at 760 Torr, with recent results from Huang \textit{et al.}\(^{27}\) in reasonable agreement. This result has potential significance for modelling of atmospheric chemistry in iodine-rich regions,\(^{28-31}\) and also indicates potential for pressure dependent studies of CH\(_2\)OO kinetics using photolysis of CH\(_2\)I\(_2\) in \text{O}_2.

In this work, we report kinetics of CH\(_2\)OO reactions with SO\(_2\), NO\(_2\), NO, H\(_2\)O and CH\(_3\)CHO at pressures between 25 and 450 Torr at a temperature of 295 K, using photolysis of CH\(_2\)I\(_2\)/\text{O}_2/\text{N}_2 mixtures under pseudo-first-order conditions combined with monitoring of the HCHO reaction products by laser-induced fluorescence (LIF) spectroscopy, and, for the CH\(_2\)OO + SO\(_2\) reaction at \(~1.5\) Torr, direct monitoring of CH\(_2\)OO by photoionisation mass spectrometry (PIMS). We also report kinetics of the CH\(_2\)I + NO\(_2\) reaction at pressures between 25 and 300 Torr at 295 K.

2. Experimental

2.1 Laser-Induced Fluorescence Experiments

Apparatus and experimental procedures for the laser-induced fluorescence (LIF) experiments have been described elsewhere in detail,\(^{26, 32}\) therefore only a brief description is given here. Kinetics of CH\(_2\)OO reactions were studied by monitoring of HCHO reaction products by LIF spectroscopy. Radicals were generated by the laser flash photolysis of CH\(_2\)I\(_2\)/\text{O}_2/\text{N}_2 gas mixtures (R1-R2) with the addition of excess co-reagent (NO\(_2\), NO, SO\(_2\), H\(_2\)O or CH\(_3\)CHO) to ensure pseudo-first-order conditions. Experiments to investigate CH\(_2\)I + NO\(_2\) kinetics were performed in the absence of \text{O}_2, while those to investigate CH\(_2\)OO + NO\(_2\) were performed using a limited range of NO\(_2\) concentrations in order to avoid production of HCHO through the reaction of CH\(_2\)I with NO\(_2\) (see Section 3.1), whilst maintaining pseudo-first-order conditions.

CH\(_2\)I\(_2\) (Sigma-Aldrich, 99 %) was used as a dilute gas in \text{N}_2 either by filling a glass bulb containing liquid CH\(_2\)I\(_2\) with \text{N}_2 or by bubbling a slow flow of \text{N}_2 through liquid CH\(_2\)I\(_2\). Reagents (NO, NO\(_2\), SO\(_2\), CH\(_3\)CHO) were prepared at known concentrations in \text{N}_2 and stored in glass bulbs. NO (BOC Special Gases, 99.5 %) was purified prior to use by a series of freeze-pump-thaw cycles. CH\(_3\)I\(_2\), CH\(_3\)CHO (Sigma-Aldrich, 99.5 %)
NO (Sigma-Aldrich, 99.5%), SO₂ (Sigma-Aldrich, 99.9%), N₂ (BOC, 99.99%) and O₂ (BOC, 99.999%) were used as supplied. Water vapour was added to the gas mixture by bubbling a known flow of N₂ gas through a bubbler containing deionised water at a known temperature. Gases were mixed in a gas manifold and passed into a six-way cross reaction cell at known flow rates (determined by calibrated mass flow controllers). The pressure in the reaction cell was monitored by a capacitance manometer (MKS Instruments, 626A) and controlled by throttling the exit valve to the reaction cell. The total gas flow rate through the reaction cell was adjusted with total pressure to maintain an approximately constant gas residence time in the cell (~0.1 s). All experiments were performed at \( T = (295 \pm 2) \) K unless stated otherwise.

For experiments using NO₂, NO, CH₃CHO or H₂O as co-reagents, initiation of chemistry within the cell was achieved using an excimer laser (KrF, Tui ExciStar M) operating at \( \lambda = 248 \) nm with typical laser fluence in the range 30 – 80 mJ cm\(^{-2}\). Experiments in which SO₂ was present as the co-reagent were performed at a photolysis wavelength of 355 nm (typical fluence ~ 150 mJ cm\(^{-2}\)), generated by frequency tripling the output of a Nd:YAG laser (Spectron Laser Systems) to avoid potential multi-photon photolysis of SO₂ at shorter wavelengths.³³-³⁵

Production of HCHO was monitored by laser-induced fluorescence (LIF) of HCHO at \( \lambda \sim 353.1 \) nm.³⁶ Approximately 2 to 4 mJ pulse\(^{-1}\) of laser light at ~ 353.1 nm was generated by a dye laser (Lambda Physik, FL3002) operating on DMQ/dioxirane dye and pumped by a 308 nm excimer laser generating ~ 50 mJ pulse\(^{-1}\) (XeCl, Lambda Physik LPX100). The output of the dye laser was passed through the reaction cell in an orthogonal axis to the 248 nm / 355 nm photolysis laser output, with HCHO fluorescence detected in the visible region of the spectrum by a channel photomultiplier (CPM, Perkin-Elmer C1943P) orthogonal to both the photolysis laser and the LIF excitation laser beams. A Perspex filter was used to prevent scattered laser light from the photolysis laser and the LIF excitation laser reaching the CPM. The HCHO fluorescence signal was monitored as a function of time following photolysis of CH₂I₂ by varying the time delay between firing the photolysis laser and the LIF excitation laser through use of a delay generator (SRS DG535). Results from between 5 and 20 photolysis shots were typically averaged prior to analysis.

### 2.2 Photoionisation Mass Spectrometry Experiments

Photoionisation mass spectrometry (PIMS) experiments were performed in this work to determine the kinetics of CH₂OO + SO₂ at low pressure (~1.5 Torr) and 295 K by direct monitoring of CH₂OO in reactions performed under pseudo-first-order conditions. The PIMS apparatus has been described previously in detail³², ³⁷, ³⁸ and only a brief description is given here. Gas mixtures of CH₂I₂/O₂/N₂ and CH₂I₂/O₂/N₂/SO₂ were prepared in a gas handling line, with reagents and reagent preparation as described above for the LIF experiments, and introduced to the steel reaction flow tube (10.5 mm internal diameter,
70 cm in length) via calibrated mass flow controllers. The pressure in the reaction flow tube was monitored by a capacitance manometer (MKS Instruments, 626A) and controlled by throttling the exit valve to the flow tube.

Chemistry was initiated by a pulsed excimer laser (Lambda Physik, Compex 205) at a wavelength of 248 nm, with typical fluence of ~50 mJ cm\(^{-2}\), through reactions R1 and R2. A representative sample from the reaction mixture effused into a high vacuum chamber (< 10\(^{-5}\) Torr, maintained by diffusion and turbo pumps) via a 1 mm pinhole situated in the sidewall of the reaction flow tube. Components of the gas mixture were photoionised using 118 nm vacuum ultraviolet (VUV) laser light (typically 10\(^{11}\) photons pulse\(^{-1}\)), generated by frequency tripling of the third harmonic of a Nd:YAG laser (Continuum Powerlite, 8010) in a Xe gas cell, and passed across the effusing gas flow within 2-3 mm of the sampling pinhole. VUV light of 118 nm (equivalent to 10.5 eV) is sufficiently energetic to ionise CH\(_2\)OO (threshold = 10.02 eV), but is below the threshold required to ionise other isomers at \(m/z = 46\) (dioxirane, threshold = 10.82 eV; formic acid, threshold = 11.33 eV).\(^{10}\) Ions were sampled by the time of flight mass spectrometer (TOF-MS, Kore Technology Ltd.), and detected by an electron multiplier. The ion signals were amplified and boxcar averaged on an oscilloscope and then stored on the control computer. The ion signals were monitored as a function of time following photolysis of CH\(_2\)I\(_2\) by varying the time delay between the excimer laser and the Nd:YAG laser, used to generate the VUV radiation, through use of a delay generator (SRS DG35). These kinetic traces consisted of typically 200 time points, with typically between 10 and 25 shot averaging per time point.

### 3. Results and Discussion

#### 3.1 Photolysis of CH\(_2\)I\(_2\)/O\(_2\)/N\(_2\) mixtures

Figure 1 shows the HCHO fluorescence signal following photolysis of CH\(_2\)I\(_2\)/O\(_2\)/N\(_2\) mixtures (\textit{i.e.} in the absence of any additional co-reagent), resulting in production of HCHO through reactions R1-R6.\(^{26,32}\)

\[
\begin{align*}
\text{CH}_2\text{I}_2 + \text{hv} & \rightarrow \text{CH}_2\text{I} + \text{O}_2 \quad \text{(R1)} \\
\text{CH}_2\text{I} + \text{O}_2 & \rightarrow \text{CH}_2\text{OO} + \text{I} \quad \text{(R2a)} \\
\text{CH}_2\text{I} + \text{O}_2 + \text{M} & \rightarrow \text{CH}_2\text{IO}_2 + \text{M} \quad \text{(R2b)} \\
\text{CH}_2\text{OO} + \text{I} & \rightarrow \text{HCHO} + \text{IO} \quad \text{(R3)} \\
\text{CH}_2\text{IO}_2 + \text{I} & \rightarrow \text{CH}_2\text{IO} + \text{IO} \quad \text{(R4)} \\
\text{CH}_2\text{IO}_2 + \text{CH}_2\text{IO}_2 & \rightarrow 2 \text{CH}_2\text{IO} + \text{O}_2 \quad \text{(R5)} \\
\text{CH}_2\text{IO} & \rightarrow \text{HCHO} + \text{I} \quad \text{(R6)}
\end{align*}
\]
Previous work in this laboratory\textsuperscript{26} has shown that the yields of \(\text{CH}_2\text{OO}\) and \(\text{CH}_2\text{IO}_2\) from R2 are dependent on pressure, owing to initial formation of the excited species \(\text{CH}_2\text{IO}_2^*\), which can either decompose to produce the \(\text{CH}_2\text{OO}\) Criegee intermediate and iodine atoms (R2a) or can be collisionally stabilised to produce the peroxy radical \(\text{CH}_2\text{OO}_2\) (R2b). Since subsequent reactions of both \(\text{CH}_2\text{OO}\) and \(\text{CH}_2\text{IO}_2\) in the absence of any additional co-reagent result in production of HCHO, there is no change in the total HCHO yield as a function of pressure following photolysis of \(\text{CH}_2\text{I}_2/\text{O}_2/\text{N}_2\) mixtures.

Production of HCHO in reactions R1-R6 can be approximated by Equation 1:\textsuperscript{26, 32}

\[
S_{\text{HCHO},t} = S_0 \left[ \exp(-k_{\text{loss}} t) \right] + \frac{S_1 k'_g}{k'_g - k_{\text{loss}}} \left[ \exp(-k_{\text{loss}} t) - \exp(-k'_g t) \right] \quad \text{(Equation 1)}
\]

where \(S_{\text{HCHO},t}\) is the HCHO signal at time \(t\), \(S_0\) is the height of the HCHO signal at time zero, \(S_1\) is the maximum HCHO signal, \(k'_g\) is the pseudo-first-order rate coefficient for HCHO growth, and \(k_{\text{loss}}\) is the rate coefficient representing the slow loss of HCHO from the detection region \textit{via} diffusion. Although the HCHO growth through reactions R1-R6 is not strictly first-order, our previous work\textsuperscript{26} demonstrates that Equation 1 can faithfully reproduce the HCHO growth kinetics. In the presence of excess co-reagent (e.g. \(\text{SO}_2, \text{NO}_2\)) the kinetics of HCHO production from \(\text{CH}_2\text{OO}\) are under pseudo-first-order conditions. Figure 1 shows the fits to HCHO production in the absence and presence of additional co-reagent, indicating the fidelity of the fit to the analytical equation.

In the absence of any additional co-reagent, the first-order rate coefficient approximating the production of HCHO, \(k'_g\), was found to vary from \(~300\) s\(^{-1}\) to \(~3500\) s\(^{-1}\), depending on the concentration of \(\text{CH}_2\text{I}_2\), and thus of I atoms, in the system, in keeping with the work of Welz \textit{et al.}\textsuperscript{10} and Taatjes \textit{et al.}\textsuperscript{11} Some initial HCHO production was observed owing to multi-photon photolysis of \(\text{CH}_2\text{I}_2\) and the subsequent rapid reaction of \(^3\text{CH}_2\) with \(\text{O}_2\), with \(S_0\) typically no greater than 5 – 10 \% of \(S_1\).\textsuperscript{39-43}

### 3.2 \(\text{CH}_2\text{OO} + \text{SO}_2\)

The reaction of \(\text{CH}_2\text{OO}\) with \(\text{SO}_2\) (R7) was investigated in separate experiments using the PIMS method to monitor \(\text{CH}_2\text{OO}\) and the LIF method to monitor HCHO production.

\[
\text{CH}_2\text{OO} + \text{SO}_2 \rightarrow \text{HCHO} + \text{SO}_3 \quad \text{(R7)}
\]

Experiments using the PIMS method were performed at a total pressure of 1.5 Torr. Figure 2 shows a typical decay for \(\text{CH}_2\text{OO}\) observed in the presence of excess \(\text{SO}_2\), with the pseudo-first-order rate coefficient for \(\text{CH}_2\text{OO}\) decay found by least-squares fitting to Equation 2:
$S_{\text{CH}2\text{OO},j} = \frac{S_{\text{max}} k_{\text{sampling}}}{k_{\text{sampling}} - k} \left[ \exp(-k' t) - \exp(-k_{\text{sampling}} t) \right]$ \hspace{1cm} (Equation 2)

where $S_{\text{CH}2\text{OO},j}$ is the CH$_2$OO ion signal at time $t$, $S_{\text{max}}$ is the maximum CH$_2$OO ion signal, $k'$ is the pseudo-first-order rate coefficient for CH$_2$OO decay, and $k_{\text{sampling}}$ is the rate coefficient representing the transport of molecules in the reactor to the ionisation region (~30,000 s$^{-1}$, described in detail by Baeza-Romero et al.$^{38}$).

The bimolecular rate coefficient for CH$_2$OO + SO$_2$ ($k_7$) determined using the PIMS method at 1.5 Torr was $(3.6 \pm 0.5) \times 10^{-11}$ cm$^3$ s$^{-1}$ (Figure 3), similar to the value of $(3.9 \pm 0.7) \times 10^{-11}$ cm$^3$ s$^{-1}$ at 4 Torr reported by Welz et al.$^{10}$ and several orders of magnitude greater than the values typically used in atmospheric models.

The LIF experiments monitoring HCHO production from CH$_2$OO + SO$_2$ were performed over the pressure range 50 – 450 Torr, with SO$_2$ concentrations in the range $2.4 \times 10^{14}$ to $1.6 \times 10^{15}$ cm$^{-3}$. The HCHO growth (Figure 4) was observed to display biexponential behaviour, with no decrease in the total HCHO yield compared to experiments performed in the absence of any co-reagent, indicating complete titration of both CH$_2$OO and CH$_2$IO$_2$ to HCHO. Kinetic parameters were determined by fitting to Equation 3:

$$
S_{\text{HCHO},j} = S_0 \left[ \exp(-k_{\text{loss}} t) \right] + \frac{S_1 f k'_{g1}}{k'_{g1} - k_{\text{loss}}} \left[ \exp(-k_{\text{loss}} t) - \exp(-k'_{g1} t) \right] + \frac{S_1 (1-f) k'_{g2}}{k'_{g2} - k_{\text{loss}}} \left[ \exp(-k_{\text{loss}} t) - \exp(-k'_{g2} t) \right]
$$

where $S_{\text{HCHO},j}$ is the HCHO signal at time $t$, $S_0$ is the height of the HCHO signal at time zero, $S_1$ is the maximum HCHO signal, $k'_{g1}$ is the pseudo-first-order rate coefficient for the fast HCHO growth, $k'_{g2}$ is the pseudo-first-order rate coefficient for the slower HCHO growth, $f$ is the fractional contribution of the fast growth process to the total HCHO yield (hence $(1-f)$ is the fractional contribution of the slower growth process to the total HCHO yield), and $k_{\text{loss}}$ is the rate coefficient representing the slow loss of HCHO from the detection region via diffusion. For the SO$_2$ experiments (conducted using a photolysis wavelength of 355 nm) there was no contribution from $S_0$ (i.e. $S_0 = 0$).

The initial fast growth of HCHO displayed a linear dependence on [SO$_2$], while the slower growth was independent of [SO$_2$] and at a similar rate to the observed HCHO production in the absence of any additional co-reagent. The yields of HCHO from the faster growth process were consistent with production from CH$_2$OO + SO$_2$, while those from the slower process were consistent with production from reactions of CH$_2$IO$_2$ (i.e. reactions R4-R6). We thus determine $k_7$ from linear fits of $k'_{g1}$ (Equation 3) against [SO$_2$]. The validity of describing the system using Equation 3 is discussed in our previous work.$^{26}$
Figure 5 and Table 1 show the values of \( k_7 \) as a function of pressure. No significant dependence of \( k_7 \) on pressure was observed, with an average value of \( (3.42 \pm 0.42) \times 10^{11} \) cm\(^3\) s\(^{-1}\) for all experiments (PIMS and LIF) described in this work (all errors are 1σ unless stated otherwise). Moreover, there is no significant change in the HCHO yield from the reaction of CH\(_2\)OO with SO\(_2\) as a function of pressure, indicating there is little stabilisation of reaction products. These results are consistent with the low pressure results obtained by Welz et al.\(^{10}\) and theoretical work by Vereecken et al.\(^{13}\), and support arguments for an increased role of CH\(_2\)OO + SO\(_2\) in the atmosphere. Taatjes et al.\(^{12}\) have also shown that the reaction of the C\(_2\) Criegee intermediate, CH\(_3\)CHOO, with SO\(_2\) at a pressure of 4 Torr is also significantly faster than previously expected, potentially indicating an increased role for CH\(_3\)CHOO + SO\(_2\) in the atmosphere. However, theoretical calculations predict that reactions of larger Criegee intermediates will exhibit pressure dependence,\(^{13}\) and that production of SO\(_3\) in reactions of larger Criegee intermediates at atmospheric pressures is unlikely owing to stabilisation of SO\(_2\)-Criegee intermediate complexes to produce secondary ozonide species, thus reducing the impacts of SO\(_2\) + Criegee intermediate reactions on H\(_2\)SO\(_4\) and sulfate aerosol production.\(^{13}\) Field observations and laboratory studies by Mauldin et al.\(^{23}\) indicate that larger Criegee intermediates, such as those produced in the ozonolysis of monoterpenes, do impact on atmospheric concentrations of H\(_2\)SO\(_4\) through oxidation of SO\(_2\), but that the impacts may not be as great as those reported for CH\(_2\)OO, potentially owing to stabilisation of reaction products. Further work is thus required to investigate the effects of pressure on the reactions of larger Criegee intermediates. Moreover, modelled impacts of increases in the rates of Criegee intermediate reactions with SO\(_2\) are highly dependent on the competition with rates of Criegee intermediate reactions with water vapour. We thus investigate CH\(_2\)OO + H\(_2\)O in Section 3.6.

### 3.3 CH\(_2\)I + NO\(_2\)

Production of HCHO following photolysis of CH\(_2\)I/NO\(_2\)/N\(_2\) mixtures was examined as a function of pressure to facilitate assessment of the competition between CH\(_2\)I + O\(_2\) (R2) and CH\(_2\)I + NO\(_2\) (R8) in CH\(_2\)OO + NO\(_2\) experiments (Section 3.4).

\[
\text{CH}_2\text{I} + \text{NO}_2 \rightarrow \text{HCHO} + \text{products} \quad \text{(R8)}
\]

The production of HCHO could be described by Equation 1 (above), where \( k'_g = k_g[\text{NO}_2] \), with concentrations of NO\(_2\) between \( 1 \times 10^{14} \) and \( 9 \times 10^{14} \) cm\(^{-3}\). Pseudo-first-order rate coefficients (\( k'_g \)) were in the range ~5000 to 45,000 s\(^{-1}\), and typically large compared to the rate coefficients describing HCHO production in the absence of any additional co-reagent (Section 3.1). The bimolecular rate coefficient \( k_8 \) was determined from plots of \( k'_g \) against [NO\(_2\)] at each pressure (Figure S1), and was found to increase with increasing pressure (Figure S2 and Table S1), with a corresponding decrease in the HCHO yield as the pressure was increased (Figure S3).
A previous investigation of CH\textsubscript{2}I + NO\textsubscript{2} at pressures of 2 to 5 Torr gave a value of $k_8 = (2.2 \pm 0.1) \times 10^{-11}$ cm\textsuperscript{3} s\textsuperscript{-1}.\textsuperscript{44} Results of this work show $k_8$ to be $2.56 \pm 0.17 \times 10^{-11}$ cm\textsuperscript{3} s\textsuperscript{-1} at 50 Torr, increasing to $(5.07 \pm 0.28) \times 10^{-11}$ cm\textsuperscript{3} s\textsuperscript{-1} at 300 Torr.

The rate coefficient for reaction of CH\textsubscript{2}I radicals with O\textsubscript{2} (R2), has been shown previously to be $\sim 1.6 \times 10^{-12}$ cm\textsuperscript{3} s\textsuperscript{-1}.\textsuperscript{45, 46} Experiments to investigate HCHO production in the reaction of CH\textsubscript{2}OO (produced by CH\textsubscript{2}I + O\textsubscript{2}) with NO\textsubscript{2} must therefore be conducted at sufficiently high [O\textsubscript{2}] to avoid complications owing to HCHO production from CH\textsubscript{2}I + NO\textsubscript{2}.

### 3.4 CH\textsubscript{2}OO + NO\textsubscript{2}

Experiments to investigate CH\textsubscript{2}OO + NO\textsubscript{2} (R9) kinetics were performed with sufficient NO\textsubscript{2} concentrations ($1.0 \times 10^{14}$ to $1.4 \times 10^{15}$ cm\textsuperscript{-3}) to ensure pseudo-first-order conditions for CH\textsubscript{2}OO loss whilst also ensuring that $k_3[O_2] > k_8[NO_2]$ at all times to avoid potential complications owing to HCHO production through CH\textsubscript{2}I + NO\textsubscript{2}.

\[
\begin{align*}
\text{CH}_2\text{I} + \text{hv} & \rightarrow \text{CH}_2\text{I} + \text{O}_2 & \text{(R1)} \\
\text{CH}_2\text{I} + \text{O}_2 & \rightarrow \text{CH}_2\text{OO} + \text{I} & \text{(R2a)} \\
\text{CH}_2\text{I} + \text{NO}_2 & \rightarrow \text{HCHO} + \text{products} & \text{(R8)} \\
\text{CH}_2\text{OO} + \text{NO}_2 & \rightarrow \text{HCHO} + \text{NO}_3 & \text{(R9)}
\end{align*}
\]

Figure 1 shows the evolution of the HCHO signal following photolysis of CH\textsubscript{2}I/O\textsubscript{2}/N\textsubscript{2}/NO\textsubscript{2} mixtures. Experiments in which NO\textsubscript{2} was used as a co-reagent resulted in a decrease in the total HCHO yield when compared to experiments performed in the absence of any co-reagent. We attribute this to the formation of the peroxy nitrate species CH\textsubscript{2}IO\textsubscript{2}NO\textsubscript{2} which inhibits formation of HCHO through reactions R4-R6.

Experiments performed at 273 K to increase the lifetime of CH\textsubscript{2}IO\textsubscript{2}NO\textsubscript{2} with respect to dissociation to CH\textsubscript{2}IO\textsubscript{2}NO\textsubscript{2} did not result in any significant decrease in the HCHO yield compared to equivalent experiments at 295 K, indicating that the CH\textsubscript{2}IO\textsubscript{2}NO\textsubscript{2} lifetime at 295 K is sufficiently long to minimise production of HCHO from CH\textsubscript{2}IO\textsubscript{2}. Thus, while there is a small contribution to the HCHO signal owing to rapid chemistry following multi-photon photolysis of CH\textsubscript{2}I\textsubscript{2}, the growth of HCHO observed following photolysis of CH\textsubscript{2}I\textsubscript{2}/O\textsubscript{2}/N\textsubscript{2}/NO\textsubscript{2} mixtures can be attributed to CH\textsubscript{2}OO + NO\textsubscript{2} (R9) exclusively.

The pseudo-first-order rate coefficient for the reaction of CH\textsubscript{2}OO with NO\textsubscript{2} was determined by least-squares fitting to Equation 1, with $k'_9 = k_9[\text{NO}_2]$. The bimolecular rate coefficient for CH\textsubscript{2}OO + NO\textsubscript{2} ($k_9$) was subsequently determined from plots of $k'_9$ against [NO\textsubscript{2}], as shown in Figure 6. Fits to experimental
data using the numerical integration package Kintecus\textsuperscript{47} to determine $k_9$, detailed in the supplementary information, gave results within 10% of those obtained using the analytical expression (Equation 1).

Values for $k_9$ as a function of pressure are shown in Figure 7 and Table 2. No significant dependence of $k_9$ on total pressure was observed over the pressure range investigated (25 to 300 Torr), with an average value of $k_9 = (1.5 \pm 0.5) \times 10^{-12}$ cm$^3$ s$^{-1}$. Errors in $k_9$ include the 1$\sigma$ errors in the fits to the bimolecular plots at each pressure and an error of $\pm 10\%$ to account for any differences between fits using the analytical expression and those obtained by numerical integration (see supplementary information).

Yields of HCHO in the presence of NO$_2$, determined relative to experiments performed in the absence of NO$_2$ (i.e. production through reactions R3-R6), were consistent with the yields of CH$_2$OO determined in our previous work\textsuperscript{26} (Figure 8). This result demonstrates that $\sim$100% of CH$_2$OO is titrated to HCHO by CH$_2$OO + NO$_2$, indicating a lack of pressure dependence in $k_9$, and that there is insignificant HCHO production from CH$_2$IO$_2$ in the presence of NO$_2$. Recent measurements by Ouyang \textit{et al.}\textsuperscript{21} have demonstrated the production of NO$_3$ at atmospheric pressure from the reaction of CH$_3$OO with NO$_2$, thus also suggesting little stabilisation of reaction products to a secondary ozonide species in this system.

No significant difference in $k_9$ or in yields of HCHO were observed between experiments performed in O$_2$ bath gas and N$_2$ bath gas (results shown in Table 2), providing further evidence for similar quenching of the nascent excited CH$_2$IO$_2$\# species (produced in R2) by O$_2$ and N$_2$, as discussed in our previous work.\textsuperscript{26}

Results for $k_9$ obtained in this work, while lower than those reported by Welz \textit{et al.}\textsuperscript{10}, are on the same order of magnitude, and demonstrate a significantly faster reaction between CH$_2$OO and NO$_2$ than suggested by previous indirect measurements.\textsuperscript{1}

### 3.5 CH$_2$OO + NO

Production of HCHO following photolysis of CH$_2$I$_2$/O$_2$/N$_2$ mixtures in the presence of excess NO (3.6 $\times$ 10$^{14}$ to 1.7 $\times$ 10$^{15}$ cm$^{-3}$) exhibits biexponential growth, as shown in Figure 9, similar to experiments with SO$_2$. Again, no decrease in the total HCHO yield compared to experiments performed in the absence of any co-reagent, indicating complete titration of both CH$_2$OO and CH$_3$IO$_2$ to HCHO. Kinetic parameters for the processes contributing to HCHO production were obtained by fitting to Equation 3 (above).

The rate coefficient describing the fast HCHO growth process, $k'_{g1}$, was observed to increase linearly with increasing [NO], with the slope of a plot of $k'_{g1}$ against [NO] giving a bimolecular rate coefficient of $(1.07 \pm 0.06) \times 10^{-11}$ cm$^3$ s$^{-1}$ at 250 Torr (Figure S4). The rate coefficient describing the slower HCHO growth, $k'_{g2}$, was found to be independent of [NO], and similar to the rate coefficient for HCHO production obtained in the absence of NO. Reactions of peroxy radicals (RO$_2$) with NO are well established, and are typically on the order of $10^{-12}$ to $10^{-11}$ cm$^3$ s$^{-1}$,\textsuperscript{48,49} with a rate coefficient for CH$_3$O$_2$ + NO of $7.2 \times 10^{-12}$ cm$^3$ s$^{-1}$,\textsuperscript{49} while
Welz et al.\textsuperscript{10} reported an upper limit of $6 \times 10^{-14}$ cm$^3$ s$^{-1}$ for the rate coefficient for CH$_2$OO with NO. Thus, in contrast to the experiments with SO$_2$, we attribute the fast HCHO growth to the rapid decomposition of CH$_2$IO (R6), produced in the reaction of CH$_2$IO$_2$ with NO (R10) and assign $k_{10} = (1.07 \pm 0.06) \times 10^{-11}$ cm$^3$ s$^{-1}$ at 250 Torr.

$$\text{CH}_2\text{IO}_2 + \text{NO} \rightarrow \text{CH}_2\text{IO} + \text{NO}_2 \quad \text{(R10)}$$

$$\text{CH}_2\text{IO} \rightarrow \text{HCHO} + \text{I} \quad \text{(R6)}$$

The slower HCHO growth thus contains contributions from CH$_2$OO + I (R3) and potentially CH$_2$OO + NO (R11). In the absence of NO, production of HCHO was observed with a pseudo-first-order rate coefficient of $1860 \pm 100$ s$^{-1}$ (Equation 1). On addition of up to $1.7 \times 10^{15}$ cm$^3$ NO, the average value for the rate coefficient describing the slow HCHO growth ($k_{g2}$ in Equation 3) was $1800 \pm 340$ s$^{-1}$. Any potential influence of NO on the observed rates of HCHO production is assumed to be within the error of the experiment, and we thus place an upper limit of $2 \times 10^{13}$ cm$^3$ s$^{-1}$ on the rate coefficient for reaction of CH$_2$OO + NO ($k_{11}$).

$$\text{CH}_2\text{OO} + \text{NO} \rightarrow \text{HCHO} + \text{NO}_2 \quad \text{(R11)}$$

The upper limit for $k_{11}$ determined here is higher than that reported by Welz et al. ($k_{11} < 6 \times 10^{-14}$ cm$^3$ s$^{-1}$), owing to increased uncertainties associated with the biexponential fit, relatively low concentrations of NO, and higher concentrations of CH$_2$I$_2$ used in these experiments compared to those performed by Welz et al., which lead to increased iodine atom concentrations in this work and thus increased rates of HCHO production through CH$_2$OO + I (R3). In subsequent experiments (notably those used to investigate the kinetics of CH$_2$OO + H$_2$O) lower CH$_2$I$_2$ concentrations were used by changing the delivery method for CH$_2$I$_2$. There are also additional uncertainties in the rate coefficients for reactions with NO owing to the potential for production of NO$_2$ in the gas lines leading to the reaction cell through oxidation of NO by O$_2$ (the gas mixture has a residence time of $\sim$1 s in the gas lines leading from the mixing line to the reaction cell), leading to the potential for contributions to the observed HCHO growth from reactions involving NO$_2$.

### 3.6 CH$_2$OO + H$_2$O

Welz et al. did not observe any change in the rate of CH$_2$OO decay on addition of water vapour to the system, and reported an upper limit of $4 \times 10^{15}$ cm$^3$ s$^{-1}$ for the rate coefficient for reaction of CH$_2$OO with H$_2$O (R12):

$$\text{CH}_2\text{OO} + \text{H}_2\text{O} \rightarrow \text{HCHO} + \text{H}_2\text{O}_2 \quad \text{(R12)}$$
Similarly to the results of Welz et al., the addition of water vapour to the LIF experiments in this work did not result in any significant change to the rate of HCHO production. The total HCHO yield was also unaffected by the presence of water vapour, indicating complete titration of CH$_2$OO and CH$_2$IO to HCHO through reactions R3-R6. Figure 10 shows the HCHO fluorescence signals following photolysis of CH$_2$I$_2$/O$_2$/N$_2$ in the absence and presence of water vapour. While the HCHO signal is reduced in the presence of water vapour, there is no change in the kinetics and the reduction in signal is attributed to increased fluorescence quenching by water vapour.

At 200 Torr the pseudo-first-order rate coefficient for HCHO production was determined to be $41 \pm 15$ s$^{-1}$ by fitting to Equation 1, and was lower than the typical values reported in Section 3.1 as a result of lower concentrations of CH$_2$I$_2$ to reduce the rate of HCHO production through radical-radical reactions in the absence of water vapour. On addition of up to $1.7 \times 10^{17}$ cm$^{-3}$ water vapour to the system, a value of $52 \pm 13$ s$^{-1}$ was obtained, with no obvious dependence on the concentration of water vapour added. Owing to the higher total pressures used in this work, enabling the addition of a higher number density of water vapour to the system compared to the low pressure experiments by Welz et al., we are able to place an upper limit of $9 \times 10^{-17}$ cm$^3$ s$^{-1}$ on $k_{12}$ at 295 K by assuming any influence of water vapour is within the error of the experiment. Ouyang et al.$^{21}$ have reported a value for $k_{12}$ of $(2.5 \pm 1) \times 10^{-17}$ cm$^3$ s$^{-1}$ at 760 Torr, determined in a relative rate experiment monitoring NO$_3$ production and using the absolute value for $k_9$ (CH$_2$OO + NO$_2$) reported by Welz et al.$^{10}$ $(7 \times 10^{-12}$ cm$^3$ s$^{-1}$). Using the relative rate coefficient ratio reported by Ouyang et al., with the value for $k_9$ determined in this work $(1.5 \times 10^{-12}$ cm$^3$ s$^{-1}$), a value of $k_{12} = 5.4 \times 10^{-18}$ cm$^3$ s$^{-1}$ can be obtained.

Modelling studies investigating the impacts of CH$_2$OO chemistry on the atmospheric oxidation of SO$_2$ may therefore be underestimating the effects of increasing the rate coefficient for CH$_2$OO + SO$_2$ owing to overestimation of the competition with CH$_2$OO + H$_2$O, resulting in more significant impacts on atmospheric production of H$_2$SO$_4$ and sulfate aerosol than indicated thus far. However, Taatjes et al.$^{12}$ have shown that the anti-CH$_3$CHO Criegee intermediate does react with water vapour ($k = (1.0 \pm 0.4) \times 10^{-14}$ cm$^3$ s$^{-1}$), and the lack of reaction between CH$_2$OO and water vapour may not be representative of all Criegee intermediates. Modelling of Criegee chemistry in forested regions in Finland and Germany has indicated that concentration of the CH$_2$OO Criegee intermediate is only ~20-33 % of the concentrations of larger Criegee intermediates derived from monoterpenes,$^{14}$ with global modelling indicating that the production rate of CH$_2$OO comprises ~40 % of the total global production rate of all Criegee intermediates.$^{15}$ The chemistry of larger Criegee intermediates warrants further attention.

**3.7 CH$_2$OO + CH$_3$CHO**
The reactions of Criegee intermediates with carbonyl compounds are of interest not only for their potential atmospheric relevance, but also to facilitate the use of carbonyl compounds as scavengers of Criegee intermediates in alkene ozonolysis experiments, enabling the determination of product yields of ozonolysis reactions.

Horie et al.\textsuperscript{50} studied the relative rates of CH\textsubscript{2}OO reactions with CH\textsubscript{3}CHO (R13) and CF\textsubscript{3}COCF\textsubscript{3} (R14) at 730 Torr in synthetic air using FT-IR spectroscopy to monitor the decay of CF\textsubscript{3}COCF\textsubscript{3} and the production of the secondary ozonide propene ozonide (methyl-1,2,4-trioxolane) from the reaction with CH\textsubscript{3}CHO, and found the reaction with CF\textsubscript{3}COCF\textsubscript{3} to be 13 times faster than that with CH\textsubscript{3}CHO.

\[
\begin{align*}
\text{CH}_2\text{OO} + \text{CH}_3\text{CHO} & \rightarrow \text{products} \quad \text{(R13)} \\
\text{CH}_2\text{OO} + \text{CF}_3\text{COCF}_3 & \rightarrow \text{products} \quad \text{(R14)}
\end{align*}
\]

Secondary ozonide products were observed by Horie et al. for both R13 and R14 at 730 Torr, while photoionisation mass spectrometry experiments by Taatjes et al.\textsuperscript{11} at 4 Torr observed a secondary ozonide product for R14 but not for R13. Absolute rate coefficients for CH\textsubscript{2}OO + CH\textsubscript{3}CHO and CH\textsubscript{2}OO + CF\textsubscript{3}COCF\textsubscript{3} were measured by Taatjes et al.\textsuperscript{11} at 4 Torr in He by direct monitoring of CH\textsubscript{2}OO, with results indicating the reaction with CF\textsubscript{3}COCF\textsubscript{3} to be \sim 32 times faster than that with CH\textsubscript{3}CHO and \( k_{13} = (9.4 \pm 0.7) \times 10^{-13} \text{ cm}^3 \text{s}^{-1} \) at 4 Torr. As discussed by Taatjes et al.\textsuperscript{11}, the differences between the results of Horie et al. and Taatjes et al. may arise from differences in the fall-off behaviour of the two reactions, indicating pressure dependence of one or both of the reactions over the range of pressures investigated. Differences in product observations between the two studies also suggest pressure dependence in \( k_{13} \). In the low pressure experiments, Taatjes et al. do not observe formation of secondary ozonide products. At 730 Torr, propene ozonide was observed as the major product of R13, indicating collisional stabilisation of the nascent secondary ozonide at high pressures. Recent theoretical work\textsuperscript{22} has investigated the potential energy surface for the reaction of CH\textsubscript{2}OO with CH\textsubscript{3}CHO, and supports the observed pressure dependence of the reaction. Reaction products are predicted to be collisionally stabilised to a secondary ozonide (SOZ) species, with significant production of the SOZ at atmospheric pressure (760 Torr) and the SOZ dominating the reaction products at pressures above 1000 Torr.

Pressure dependent kinetics are expected to be typical for reactions of larger Criegee intermediates with atmospherically relevant species, including SO\textsubscript{2}, and investigation of the CH\textsubscript{2}OO + CH\textsubscript{3}CHO system may therefore provide insight to the behaviour of other Criegee intermediates.

In this work, we investigate HCHO production from CH\textsubscript{2}OO + CH\textsubscript{3}CHO (R13) at total pressures between 25 and 300 Torr and concentrations of CH\textsubscript{3}CHO in the range 2 \times 10^{14} to 1 \times 10^{15} \text{ cm}^{-3}. Production of HCHO displayed single exponential growth, and the HCHO fluorescence signal was fitted to Equation 1 (Figure 11). Figure 12 shows the bimolecular plot used to determine \( k_{13} \) at 25 Torr, giving \( k_{13} = (1.48 \pm 0.04) \times 10^{12} \text{ cm}^3 \text{s}^{-1} \) at 25 Torr. The HCHO yield from R13 (corrected for any HCHO production from
CH$_2$O$_2$ in reactions R4-R6 using the results of our previous work) was observed to decrease with increasing pressure, indicating stabilisation of the CH$_2$OO + CH$_3$CHO reaction product at higher pressures (R13b) and pressure dependence in $k_{13}$.

\[
\text{CH}_2\text{OO} + \text{CH}_3\text{CHO} \rightarrow \text{CH}_2\text{OO-CH}_3\text{CHO}^\# \\
\text{CH}_2\text{OO-CH}_3\text{CHO}^\# \rightarrow \text{HCHO} + \text{CH}_3\text{C(O)OH} \quad \text{(R13a)} \\
\text{CH}_2\text{OO-CH}_3\text{CHO}^\# + M \rightarrow \text{propene ozonide + M} \quad \text{(R13b)}
\]

Figure 13 shows the Stern-Volmer plot for HCHO yields from R13, giving an intercept of 1.19 ± 0.39 and slope ($k_{13b}/k_{13a}$) of (1.09 ± 0.08) $\times$ 10$^{-18}$ cm$^3$. Using an intercept of 1, at 4 Torr we estimate a yield of HCHO of 88%, with a yield of 4% at 730 Torr, reconciling the results of Taatjes et al.$^{11}$ and Horie et al.$^{50}$ and in agreement with theoretical work of Jalan et al.$^{22}$

Owing to the decrease in HCHO yield with increasing pressure, assignment of the kinetics of R13 at pressures above 25 Torr is challenging. Using the results of Taatjes et al.$^{11}$ at 4 Torr ($k_{13} = (9.5 \pm 0.7) \times 10^{-13}$ cm$^3$ s$^{-1}$), together with those determined here at 25 Torr ($k_{13} = (1.48 \pm 0.04) \times 10^{-12}$ cm$^3$ s$^{-1}$), 50 Torr ($\sim 2.2 \times 10^{-12}$ cm$^3$ s$^{-1}$) and the determination of $k_{13b}/k_{13a}$ from the Stern-Volmer plot ((1.09 ± 0.08) $\times$ 10$^{-18}$ cm$^3$), we estimate a low pressure limit ($k_{13,0}$) of $\sim 1.6 \times 10^{-20}$ cm$^6$ s$^{-1}$ and a high pressure limit ($k_{13,\text{inf}}$) of $\sim 1.7 \times 10^{-12}$ cm$^3$ s$^{-1}$ (see supplementary information).

**Conclusions**

Reactions of the CH$_2$OO Criegee intermediate with NO$_2$, NO, SO$_2$, H$_2$O and CH$_3$CHO have been investigated over a range of pressures. The reactions of CH$_2$OO with NO$_2$, SO$_2$ and CH$_3$CHO are rapid, in agreement with recent measurements by Welz et al.$^{10}$ and Taatjes et al.$^{11}$ but in contrast to recommendations for atmospheric modelling based on indirect measurements. Rate coefficients for reactions of CH$_2$OO with NO$_2$ and SO$_2$ are essentially independent of pressure over the pressure ranges studied in this work. The rate coefficient for CH$_2$OO + CH$_3$CHO is pressure dependent, with stabilisation to form the secondary ozonide reaction products at high pressures.

We observe no evidence for reactions of CH$_2$OO with NO or H$_2$O under the conditions employed in this work, and place upper limits on rate coefficients for these reactions of $2 \times 10^{-13}$ cm$^3$ s$^{-1}$ and $9 \times 10^{-17}$ cm$^3$ s$^{-1}$, respectively. The upper limit for the rate coefficient for CH$_2$OO + H$_2$O is significantly lower than has been reported previously. Earlier assessments$^{2,14,15,17}$ of the impacts of increased reaction rates for CH$_2$OO + SO$_2$ and CH$_2$OO + NO$_2$ will therefore be lower limits owing to overestimation of the impacts of CH$_2$OO + H$_2$O.
Acknowledgements

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### Tables

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Table 1: Bimolecular rate coefficients for CH$_2$OO + SO$_2$ ($k_7$) as a function of pressure. Errors are 1σ. $^a$Data at 1.5 Torr are from the PIMS experiments.

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<td>This work</td>
</tr>
<tr>
<td>100$^a$</td>
<td>1.38 ± 0.33</td>
<td>This work</td>
</tr>
<tr>
<td>150$^a$</td>
<td>1.19 ± 0.30</td>
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</tr>
<tr>
<td>200$^a$</td>
<td>2.00 ± 0.56</td>
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</tr>
<tr>
<td>250$^a$</td>
<td>0.96 ± 0.29</td>
<td>This work</td>
</tr>
<tr>
<td>300$^a$</td>
<td>2.53 ± 0.47</td>
<td>This work</td>
</tr>
</tbody>
</table>

Table 2: Bimolecular rate coefficients for CH$_2$OO + NO$_2$ ($k_9$) as a function of pressure. Errors include the 1σ in the fits to the bimolecular plots and an error of ± 10 % to account for any differences between the fits using the analytical expression and those obtained by numerical integration. $^a$Measured using N$_2$ as the bath gas; $^b$Measured using O$_2$ as the bath gas.
Figure 1: HCHO fluorescence signals at 200 Torr following photolysis of CH$_2$I$_2$ in the presence of O$_2$ in the absence of any co-reagent (black open squares) and in the presence of NO$_2$ (red open circles). The fits to Equation 1 are shown by the solid lines, and give $k'_g = (460 \pm 30)$ s$^{-1}$ in the absence of any additional co-reagent and $k'_g = (1490 \pm 50)$ s$^{-1}$ in the presence of NO$_2$. The ratio of $S_1$ (Equation 1) in the presence of NO$_2$ to that in the absence of NO$_2$ is 0.37.
Figure 2: CH$_2$OO ion signals at 1.5 Torr following photolysis of CH$_2$I/O$_2$/N$_2$ in the presence of SO$_2$, with the fit to Equation 3 (solid red line). For these data, $k' = (3310 \pm 450)$ s$^{-1}$. 
Figure 3: a) Pseudo-first-order rate coefficients ($k'$) at 1.5 Torr, derived from fits to Equation 3, for the decay of the CH$_2$OO ion signal ($m/z = 46$, ionised using VUV radiation at 118 nm) following photolysis of CH$_2$I$_2$/O$_2$/N$_2$ in the presence of SO$_2$. Error bars are 1σ. The fit to the data (shown in red) gives the bimolecular rate coefficient for CHOO + SO$_2$ ($k_7$); b) Pseudo-first-order rate coefficients ($k'$) for the rapid HCHO production at 250 Torr following photolysis of CH$_2$I$_2$/O$_2$/N$_2$ in the presence of SO$_2$ derived from fits to Equation 2. Error bars are 1σ. The fit to the data (shown in red) gives the bimolecular rate coefficient for CHOO + SO$_2$ ($k_7$).
Figure 4: HCHO fluorescence signals at 250 Torr following photolysis of CH$_2$I$_2$/O$_2$/N$_2$ in the presence of SO$_2$, with the fit to Equation 2 (solid red lines). The inset panel shows the evolution of the signal to longer times. For these data, $k'_{g1} = (45500 \pm 2240)$ s$^{-1}$; $k'_{g2} = (3580 \pm 280)$ s$^{-1}$; $k_{loss} = (40 \pm 9)$ s$^{-1}$; $f = (0.49 \pm 0.01)$; $S_0 = (0.43 \pm 0.01)$. 
Figure 5: Bimolecular rate coefficients for CH$_2$OO + SO$_2$ ($k_7$) as a function of pressure. Error bars are 1σ. The plot includes results from the PIMS experiments (at 1.5 Torr) and the LIF experiments (pressures ≥ 50 Torr). The data point shown by the red open circle is that determined by Welz et al.$^{10}$
Figure 6: Pseudo-first-order rate coefficients ($k'$) for HCHO production at 50 Torr, derived from fits to Equation 1, following photolysis of CH$_2$I$_2$/O$_2$/N$_2$ in the presence of NO$_2$. Error bars are 1σ. The fit to the data (shown in red) gives the bimolecular rate coefficient for CH$_2$OO + NO$_2$ ($k_0$).
Figure 7: Bimolecular rate coefficients for \( \text{CH}_2\text{OO} + \text{NO}_2 \) (\( k_9 \)) as a function of pressure. Error bars are 1σ. The data point shown by the red open circle is that determined by Welz et al.\(^\text{10}\).
Figure 8: Stern-Volmer plot showing (inverse) yields of CH$_2$OO as a function of pressure from the reaction of CH$_2$I with O$_2$. Results from our previous work are shown for experiments monitoring iodine atom production in the system (black squares), and monitoring of HCHO production in experiments with SO$_2$ (blue triangles) and NO (red circles), with the best fit line (red). Yields of HCHO from the reaction of CH$_2$OO with NO$_2$ (this work, green diamonds), determined relative to the HCHO yields in the absence of NO$_2$ (i.e. through reactions R3-R6), suggest that there is 100% titration of CH$_2$OO to HCHO in the presence of NO$_2$ at all pressures (i.e. there is no stabilisation of reaction products), and that there is little production of HCHO from CH$_2$IO$_2$ in the system. The fit to our previous work (comprising data from the I atom, NO and SO$_2$ experiments) gives an intercept of 1.10 ± 0.23 and a slope of $(1.90 ± 0.22) \times 10^{-19}$ cm$^3$. The NO$_2$ experiments give an intercept of 1.05 ± 0.12 and a slope of $(1.70 ± 0.18) \times 10^{-19}$ cm$^3$. 
Figure 9: HCHO fluorescence signals at 250 Torr following photolysis of CH$_2$I$_2$/O$_2$/N$_2$ in the presence of NO, with the fit to Equation 2 (solid red lines). The inset panel shows the evolution of the signal to longer times. For these data, $k'_{g1} = (24800 \pm 1400)$ s$^{-1}$; $k'_{g2} = (2660 \pm 320)$ s$^{-1}$; $k_{\text{loss}} = (10 \pm 2)$ s$^{-1}$; $f = (0.70 \pm 0.02)$; $S_0 = (1.33 \pm 0.01)$. 

\[ S \text{ signal / arbitrary units} = \frac{S_0}{1 + k'_{g1} t_{g1} + k'_{g2} t_{g2} + k_{\text{loss}} t_{\text{loss}} + f \cdot k_{\text{loss}} t_{\text{loss}} } \]
Figure 10: HCHO fluorescence signals at 200 Torr following photolysis of CH$_2$I$_2$/O$_2$/N$_2$ in the absence (black open circles) and presence of water vapour (red open triangles), with the fits to Equation 1 (solid lines). The differences in the amplitude of the signal result from the quenching of the fluorescence signal by H$_2$O. For these data, $k' = (41 \pm 15)$ s$^{-1}$ in the absence of water vapour and $k' = (52 \pm 13)$ s$^{-1}$ in the presence of water vapour.
Figure 11: HCHO fluorescence signals at 25 Torr following photolysis of CH$_2$I$_2$/O$_2$/N$_2$ in the presence of CH$_3$CHO, with the fit to Equation 1 (solid red line). For these data, $k' = (2040 \pm 120)$ s$^{-1}$. 
Figure 12: Pseudo-first-order rate coefficients \((k')\) for HCHO production at 25 Torr, derived from fits to Equation 1, following photolysis of CH\(_2\)I\(_2\)/O\(_2\)/N\(_2\) in the presence of CH\(_3\)CHO. Error bars are 1σ. The fit to the data (shown in red) gives the bimolecular rate coefficient for CH\(_2\)OO + CH\(_3\)CHO \((k_{13})\).
Figure 13: Stern-Volmer analysis for HCHO yields from CH$_2$OO + CH$_3$CHO (R13) (corrected for HCHO production from CH$_3$IO$_2$ chemistry) as a function of total pressure, with the fit to the data (red). Error bars are 1σ.
References


