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Subtle Microwave-Induced Overheating Effects in an Industrial Demethylation Reaction and Their Direct Use in the Development of an Innovative Microwave Reactor

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ABSTRACT: A systematic study of the conventional and microwave (MW) kinetics of an industrially relevant demethylation reaction is presented. In using industrially relevant reaction conditions the dominant influence of the solvent on the MW energy dissipation is avoided. Below the boiling point, the effect of MWs on the activation energy $E_a$ and $k_0$ is found nonexistent. Interestingly, under reflux conditions, the microwave-heated (MWH) reaction displays very pronounced zero-order kinetics, displaying a much higher reaction rate than observed for the conventionally thermal-heated (CTH) reaction. This is related to a different gas product (methyl bromide, MeBr) removal mechanism, changing from classic nucleation into gaseous bubbles to a facilitated removal through escaping gaseous vapors. Additionally, the use of MWs compensates better for the strong heat losses in this reaction, associated with the boiling of HBr/water and the loss of MeBr, than under CTH. Through modeling, MWH was shown to occur inhomogeneously around gas/liquid interfaces, resulting in localized overheating in the very near vicinity of the bubbles, overall increasing the average heating rate in the bubble vicinity vis-à-vis the bulk of the liquid. Based on these observations and findings, a novel continuous reactor concept is proposed in which the escaping MeBr and the generated HBr/water vapors are the main driving forces for circulation. This reactor concept is generic in that it offers a viable and low cost option for the use of very strong acids and the managed removal/quenching of gaseous byproducts.

INTRODUCTION

The 20th century has been mainly dominated by the use of conventional thermal energy to drive chemical reactions. The use of alternative energy sources, such as microwave (MW) technology, appeared in the late 1970s, and its use could overcome existing bottlenecks in chemical manufacturing and improve the carbon footprint of many reactions. While it is clear that poor instrumentation has led to erroneous reports and set off unrealistic expectations, a systematic and controlled investigation of the influence of MWs on chemical reactions and their kinetics has been lacking. Such background information is however indispensable for possible future implementation of MW technology on an industrial scale. To date, many types of thermal and nonthermal MW effects have been reported; as a guide to past research in these reports, the interested reader is referred to the critical review by De La Hoz and co-workers. Presently, many of the “nonthermal MW effects” have been disproven as being the result of an incorrect temperature measurement. Also, the energy contained in MWs is far too low to break even hydrogen bonds. In the presence of non-MW absorbing solvents the groups of Dudley and Stiegman argue that both substrate and product can temporarily store energy, leading to localized temperature increases at the reactive site, accounting for the observed MW rate enhancements. More recently they also showed that poorly MW-absorbing molecules can be selectively heated by MWs provided association to a nonreactive polar molecule as a good MW absorber, albeit the effect is then less pronounced. In a further advancement the Kappe group used Si−C vessels which were proclaimed to impede the penetration of MWs into the reaction vessel, heating the Si−C material instead, and thus administering in essence conventional heating. That way no difference in reactivity could be observed between conventional and 

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However, very recently the Strouse group showed convincingly that 3 mm Si—C tube walls only retain ~48% of the MWs, pointing at significant MW transmission into, and dissipation inside, the reaction vessel. 

This MW leakage through Si—C tube walls is particularly acute when a strongly absorbing solvent is used. Fan et al. have pioneered the use of MW technology to the hydrothermal depolymerization of cellulose to glucose, demonstrating a distinct influence of the applied MW density and a MW heat input via molecular radiators; notably, they used proton/deuteration exchange techniques to also obtain structural information. Many research groups continue to propose differing activation energies for reactions run in the presence of CTH vis-à-vis MWH. However, the determination of the kinetic parameters ($E_a, k_0$) of a reaction strongly depends on the applied model and thus requires a solid understanding of the reaction mechanism. Also, a sufficient number of data points is needed to ensure confidence, thus requiring a thorough analytical method. In this study we have investigated the MW activation of an industrially (pharmaceutically) relevant demethylation reaction, converting (3-hydroxyphenyl)methylammonium bromide (3HPMA) into (3-hydroxyphenyl)methylammonium bromide (3MPMA) (see Scheme 1)\(^{10,11}\) with a detailed kinetics study. For this purpose, we have used both a SAIREM MiniFlow200SS\(^{12}\) with TM monomode cavity and an Anton Paar Monowave 300, both of which are equipped with fiber-optic temperature measurement. 

The MiniFlow200 uses a solid-state MW generator, which enables precise frequency and microwave power control; in addition, it features forward and reflected power measurement so that an energy balance can be obtained. The absence of this feature in most commonly used MW heating equipment has been demonstrated to result in significant misreadings of the actual MW power transferred to the sample under investigation.\(^{13}\) The experimental study is complemented by a simulation study to obtain additional insight into the interacting physical phenomena: electromagnetics, fluid dynamics, and heat transfer. The demethylation reaction in this study is typically run on multitonne scale and generally employs limited amounts of solvent and reagents, therewith increasing the efficiency of the process and its productivity. This reaction however has received little mechanistic attention, especially under relevant reaction conditions. From a MW point of view, the investigation of a polar reaction with high substrate/product dipoles and continuously changing dielectric properties, as the reaction progresses, presents a great opportunity to gain more knowledge and understanding of how the use of MWs could potentially benefit such chemical reactions.\(^{14,15}\)

### DISCUSSION

To assess the potential influence of different heating methods on the 3MPMA to 3HPMA demethylation reaction, kinetic reaction profiles (as conversion—time plots) were first established for the case of an open vessel CTH covering the 90—118 °C temperature range (Figure 1A). The reaction mixture consisted of 5.038 g of 3MPMA (0.0367 mol) and 12.38 g of 48% HBr (0.0734 mol HBr; twice excess vis-à-vis substrate), and no additional solvent was added. Temperature measurement was performed in a dual way, recording both the oil bath temperature and, by fiber-optic temperature probe (FOTP), the internal reaction temperature. Interestingly, Figure 1A shows that the rate of the reaction becomes equal for oil bath temperatures ≥115 °C, equating by FOTP to an effective internal maximal reaction temperature of 112—112.5 °C. A linear correlation between the target (oil bath) temperatures and the recorded fiber-optic temperatures is observed only up to 113 °C (oil bath) (Figure 1B).

Kinetic analysis of the reaction profiles in Figure 1A shows a complex kinetic behavior which can generally not be explained as a single first- or second-order process for the entire conversion range. More specifically, the beginning and the end of the reaction appear to behave as different consecutive second-order processes, which interchange at the 50–60% conversion level (Figure 1S). In the 90 °C case, the conversion level is below 50% showing thus only the first process. To explain this complex behavior, it was hypothesized that throughout the reaction the number of available protons is reduced by (reversible) protonation of 3HPMA (Scheme 1S). This becomes particularly important in the later stages of the reaction, as close to stoichiometric amounts of reagents are employed in this reaction. As shown in Figure 2, this model (Scheme 1S, eq 15a) fits the experimental kinetic data very well for all temperatures. In a second approach, conversion—time

![Figure 1](image-url)
plots were also established for the MWH reaction. Kinetic analysis of this data shows that the two-step model presented in Scheme 1S (eq 15a) can only be applied up to 111.2 °C (FTOP) (Figure 3A). Furthermore, in the 90−113 °C range the $E_a/k_0$ values of the CTH and MWH reactions vary only slightly (Figures 3B/2B). As shown in Figure 4A, the reaction rate observed at ∼98 °C is independent of the type of heating while at ∼113 °C, MW operation leads to a markedly higher rate of reaction than observed with CTH (Figure 4B). Moreover, Figure 4B shows pronounced zero-order behavior under MW operation at 113.85 °C while under CTH, classic second-order is observed. As the recorded temperature for both heating types (MWH, CTH) is 113 °C, within experimental error, this does demonstrate a pronounced change in the reaction mechanism between the MW and the conventional thermally run reaction.

A multiphysics simulation, primarily using Comsol Multiphysics 5.2 (see also the section on numerical simulation in Supporting Information), was conducted to create additional insight into the electromagnetics, the fluid dynamics, and the heat transfer. To the best of our knowledge, only two studies have previously been published covering the combined simulation of electromagnetic heating and conductive/convective heat transfer in bench-scale MW chemistry systems. These studies both concerned heating in a CEM Discover MW device. More specifically, Robinson et al. studied the MW heating of a variety of solvents in a stirred vial, but their conducted simulation did not include a complete fluid dynamics model. In contrast, Sturm et al. focused on the heating of water in a MW field, including a laminar fluid dynamics model to simulate the free convection under nonstirring conditions. To model the MWH demethylation reaction case adequately, a much more advanced methodology was required, covering all the relevant physical phenomena and stirring. Furthermore, to correctly predict the electromagnetic field inside the cavity and the reactor, knowledge was required of the dielectric properties of the reaction mixture at the relevant temperatures. The medium
properties were measured and determined to be $\varepsilon' = 17$ and $\sigma = 3.3 \, \text{S/m}$. The dissipative ($\sigma$) term is best described as an electrical conductivity due to the high concentration of ions in solution. A visual set up is provided in Figure 5a, and additional details are available from Supporting Information and Table 1S. The fluid agitation in the reactant mixture was simulated by applying a rotating geometrical domain, to account for stirrer bar rotation, in combination with the k-$\varepsilon$ Reynolds-averaged Navier–Stokes model accounting for turbulence and turbulent heat transfer. Figure 5b–e shows the model results for the multiphysics simulation. Additionally, an animated video is available from Supporting Information (video 1S). More specifically Figure 5b shows the MW field (by means of the electric field intensity on logarithmic scale) in and around the cavity applicator and reactor. The simulation shows that the electromagnetic emission and dissipation in the reactor and cavity construction materials are negligible. In addition, it also reconfirms the general nonuniformity of MW fields. The latter feature also expresses itself in the heat generation simulation presented in Figure 5c, where it can be seen that the main zone of heat generation occurs close to the MW antenna of the cavity. However, in Figure 5d it is shown that for the applied stirring speed, the temperature variations in the reactor are too small to have any significant effect on the reaction rate, i.e., less than 1.5% rate variation for the highest temperatures and less than 0.4% below 111 °C. Figure 5e shows the overall electromagnetic dissipation and heat loss to the surroundings versus temperature for both the simulated and the experimental case. It can be seen that the simulation correctly approximates the characteristics of the experimental energy balance. The curves lie close to each other, and both have an accelerating heat loss as the temperature approaches the boiling point. Figure 5f shows a quasi-electrostatic analysis of the MW field around a bubble. The MW field is deformed by the presence of the bubble. Red zones, representing localized overheating, and blue zones, indicative of relatively cooler zones, can be observed. These however do not cancel out: there is a ∼40% average increase in heat generation in a layer of ∼0.1 times the bubble radius. Generally the evaporation of hydrobromic acid into a bubble extracts heat from the reactant liquid adjacent to this bubble due to the expansion of the bubble, which can result in a reduction of vapor pressure and consequently a potential bubble collapse. Though the flow regime in the reactor is turbulent due to stirring, the Kolmogorov length scale is calculated to be 30 to 240 μm. Below this scale, bubbles do not benefit from turbulent convective heat transfer, and their growth potential is limited unless a directly adjacent heat source is present. For the CTH case, only the bubbles contacting the heated reactor wall can grow, so that fewer and larger bubbles are formed. In comparison, for the MWH case, the presence of a locally enhanced volumetric heat source enables many more bubbles to form. This mechanism agrees well with the differences in boiling regime observed during experimentation; illustrative videos of the MWH (video 2S) and CTH (videos 3S and 4S) demethylation reaction at 112 °C are included in Supporting Information.

As the removal of MeBr gas through water/HBr vapor will be governed by the available surface between the reaction mixture and the water vapor bubbles, the occurrence of zero-order kinetics may relate to the available interface becoming insufficient to remove all produced MeBr formed at one given point in time. Scheme 2 represents the proposed mechanism for higher temperature MW operation. The derived rate equation fits the high temperature MW kinetic data very well (see Scheme 2S, eq 28b, and Figure 2S). An additional observation in the MW transmission transient during experiments can be made: Figure 3S shows that fluctuations in the power regulation occur less rapidly with increasing reaction temperature, which indicates a change in medium properties as

$$A + H^+ \xrightarrow{K_1} P + M_L$$
$$M_L + S \xrightarrow{K_2} M_L S$$
$$M_L S \xrightarrow{K_3} M_L + S$$
$$P + H^+ \xrightarrow{K_4} PH^+$$

"A is 3MPMA, P is 3HPMA, $M_L$ is methyl bromide in the liquid phase, $M_L$ is methyl bromide in the gas phase, and $S$ is the surface of the bubbles."

Figure 5. Simulation results of electromagnetic dissipation and heat transfer. (A) Visual setup for measurement of the dielectric properties. (B) MW field in and around the cavity applicator and reactor. (C) Heating rate distribution. (D) Temperature distribution in the reactor. (E) Overall electromagnetic dissipation in both the simulated and the experimental case versus temperature. (F) Heating rate around a bubble.

Scheme 2. Proposed Alternative Model for the Demethylation Reaction under MW Exposure at High Reaction Temperatures"
In summary, we have shown that the main in
crease in reaction order was shown to relate to a deformation of the microwave


